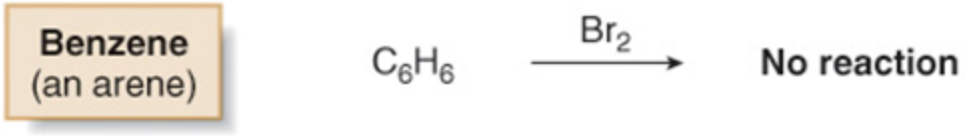
Benzene and Aromatic Compounds

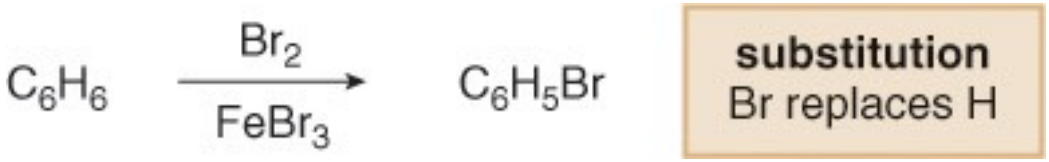
Chapter 15
Organic Chemistry, *8th Edition*
John McMurry

Background

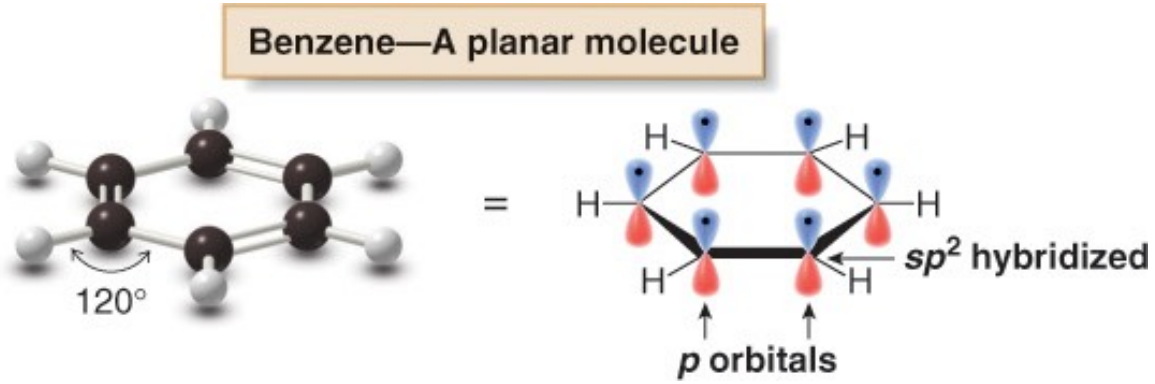
- Benzene (C_6H_6) is the simplest aromatic hydrocarbon (or arene).
- Four degrees of unsaturation.
- It is planar.
- All C—C bond lengths are equal.
- Whereas unsaturated hydrocarbons such as alkenes, alkynes and dienes readily undergo addition reactions, benzene does not.

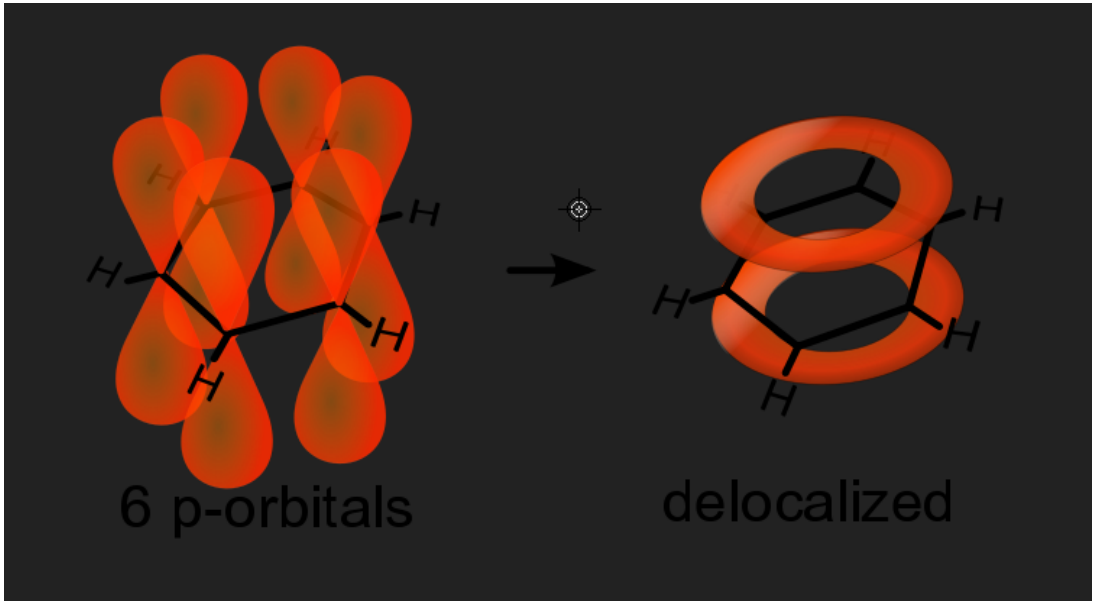


- Benzene reacts with bromine only in the presence of $FeBr_3$ (a Lewis acid), and the reaction is a substitution, not an addition.



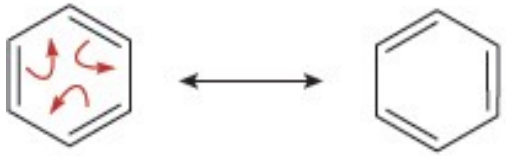
The Structure of Benzene: MO





The Structure of Benzene: Resonance

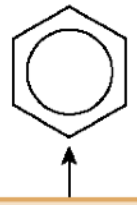
The true structure of benzene is a resonance hybrid of the two Lewis structures.



The hybrid

The electrons in the π bonds are **delocalized** around the ring.

Some texts draw benzene as a hexagon with an inner circle:



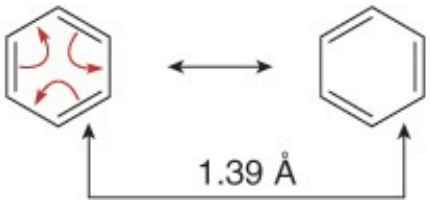
The circle represents the **six π electrons,** distributed over the six atoms of the ring.

CH_3-CH_3

1.53 Å

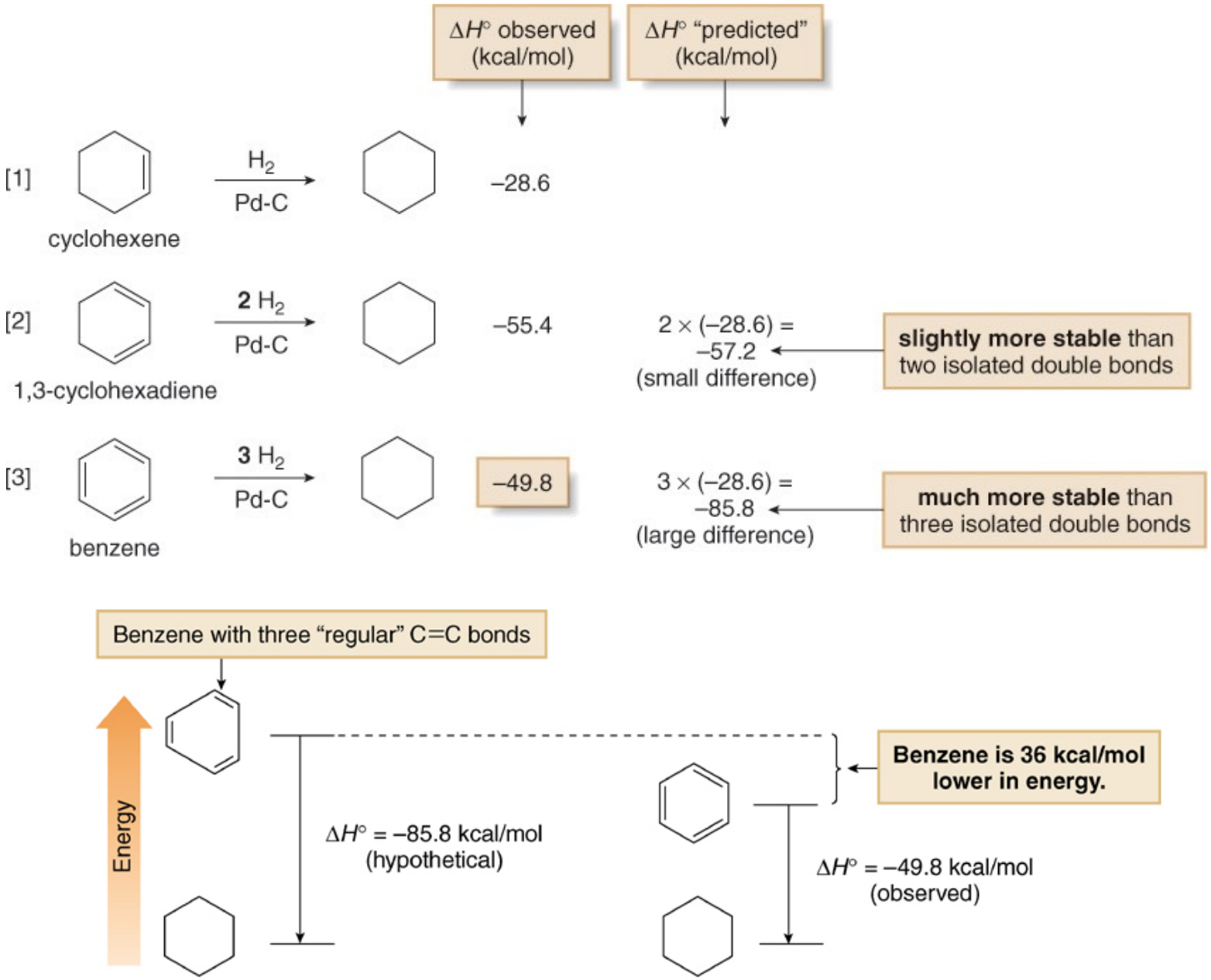
$CH_2=CH_2$

1.34 Å



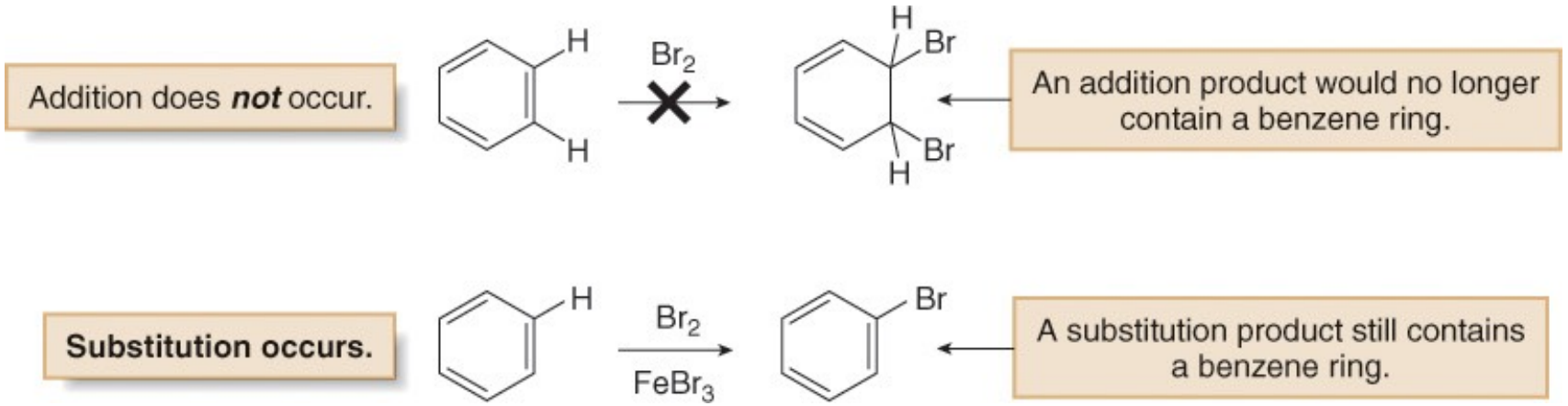
The C–C bonds in benzene are equal and intermediate in length.

Aromaticity – Resonance Energy



Stability of Benzene - Aromaticity

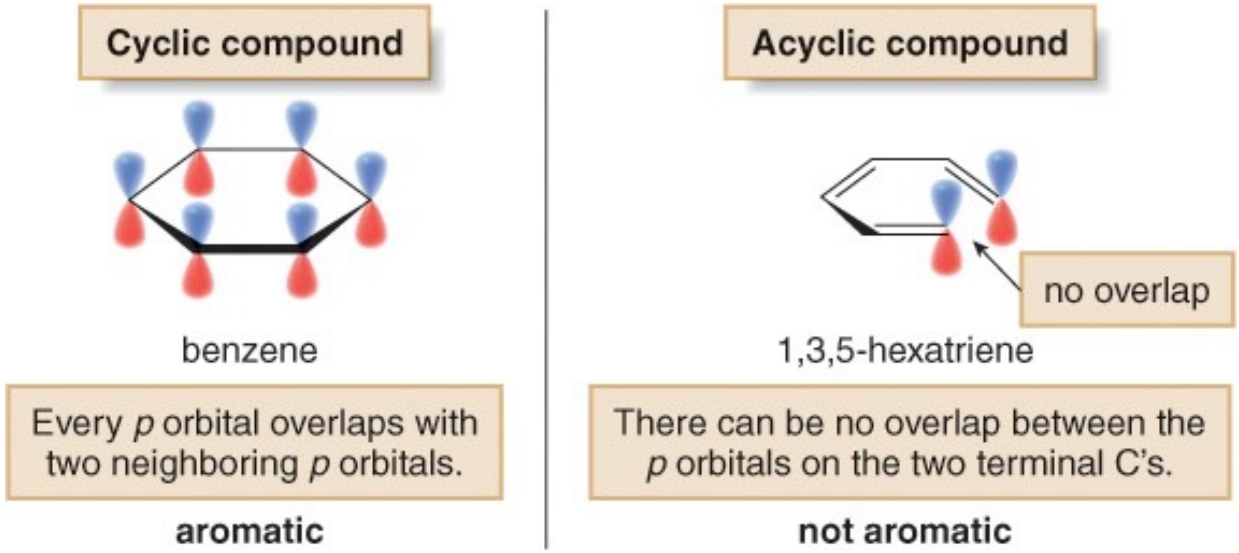
Benzene does not undergo addition reactions typical of other highly unsaturated compounds, including conjugated dienes.



The Criteria for Aromaticity

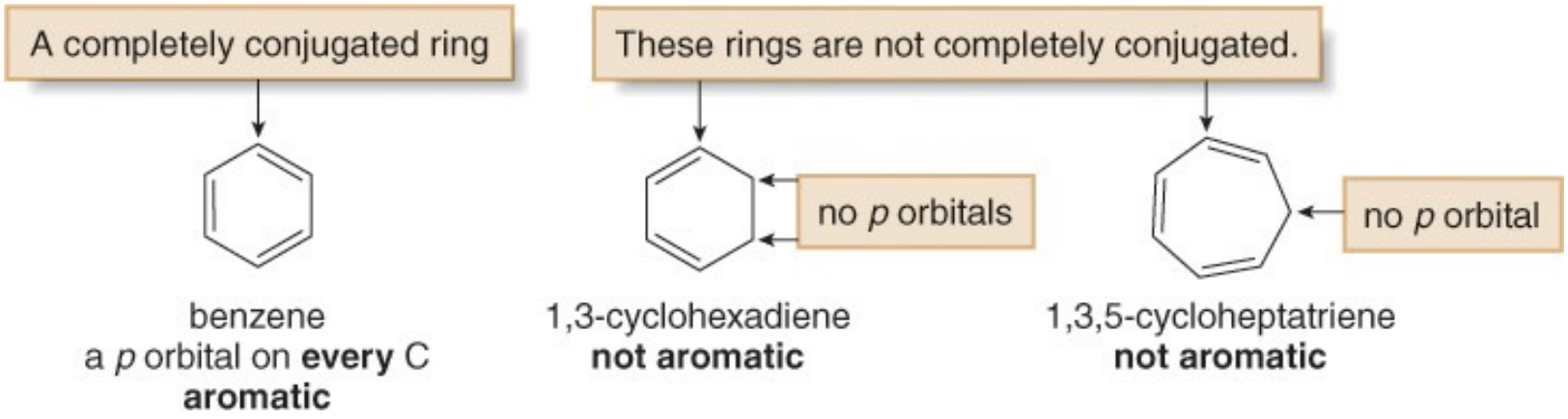
Four structural criteria must be satisfied for a compound to be aromatic.

[1] A molecule must be cyclic.



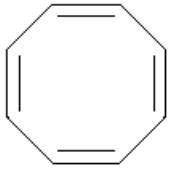
The Criteria for Aromaticity

[2] A molecule must be completely conjugated (all atoms sp^2).

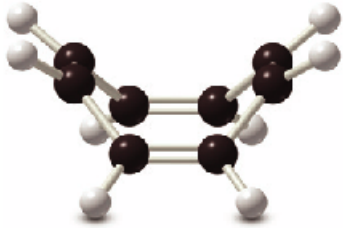


The Criteria for Aromaticity

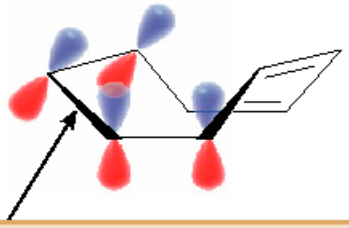
[3] A molecule must be planar.



cyclooctatetraene
not aromatic



a tub-shaped,
eight-membered ring



Adjacent *p* orbitals cannot overlap.
Electrons cannot delocalize.

The Criteria for Aromaticity—Hückel's Rule

[4] A molecule must satisfy Hückel's rule.

- An aromatic compound must contain $4n + 2$ π electrons (n = 0, 1, 2, and so forth).
- Cyclic, planar, and completely conjugated compounds that contain $4n$ π electrons are especially unstable, and are said to be *antiaromatic.*

Benzene
An aromatic compound

$4n + 2 = 4(1) + 2 =$
6 π electrons
aromatic

Cyclobutadiene
An antiaromatic compound

$4n = 4(1) =$
4 π electrons
antiaromatic

Table 17.2 The Number of π Electrons That Satisfy Hückel's Rule

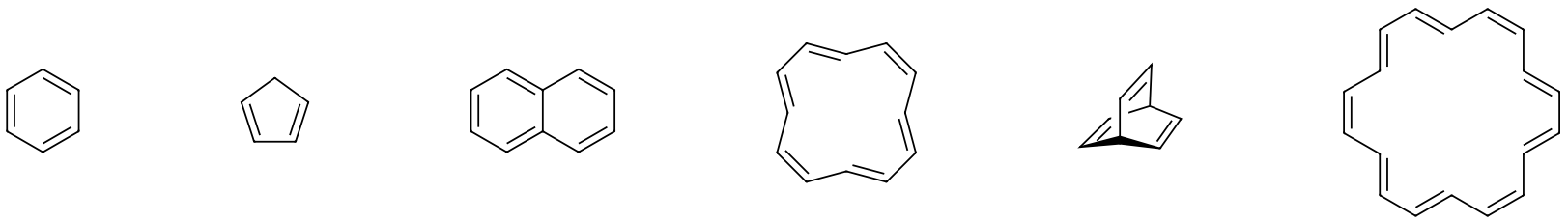
n	$4n + 2$
0	2
1	6
2	10
3	14
4, etc.	18

The Criteria for Aromaticity—Hückel's Rule

1. Aromatic—A cyclic, planar, completely conjugated compound with $4n + 2$ π electrons.

3. Antiaromatic—A cyclic, planar, completely conjugated compound with $4n$ π electrons.

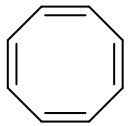
5. Not aromatic (nonaromatic)—A compound that lacks one (or more) of the following requirements for aromaticity: being cyclic, planar, and completely conjugated.



aromatic nonaromatic aromatic antiaromatic nonaromatic aromatic

Annulenes

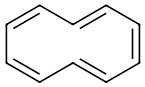
Cyclooctatetraene
8 π electrons



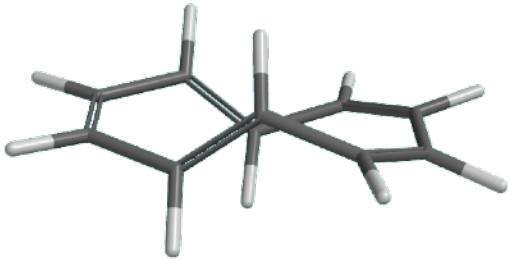
antiaromatic

nonaromatic

10-Annulene
10 π electrons



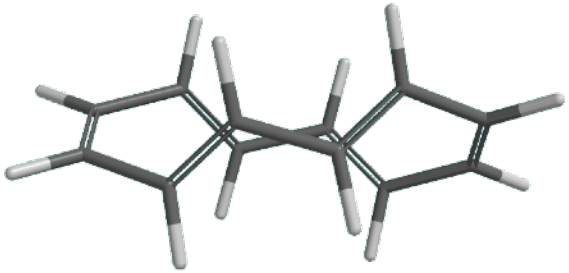
aromatic



nonaromatic

12-Annulene
12 π electrons

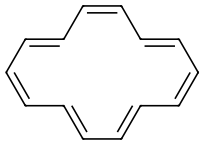
antiaromatic



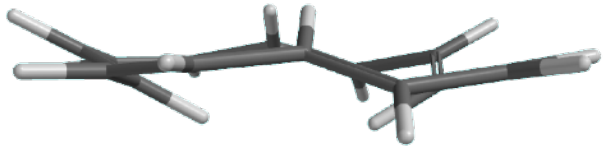
nonaromatic

Annulenes

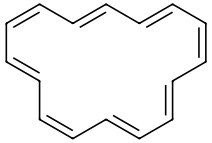
14-Annulene
14 π electrons



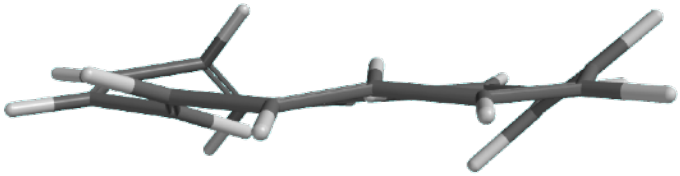
aromatic



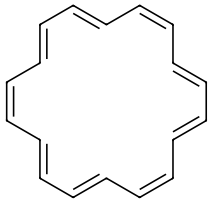
16-Annulene
16 π electrons



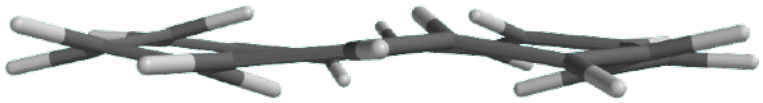
antiaromatic



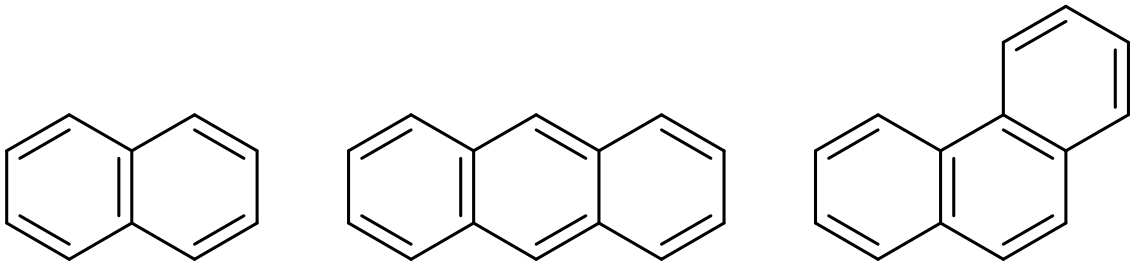
18-Annulene
18 π electrons



aromatic



Polycyclic Aromatic Hydrocarbons (PAH)

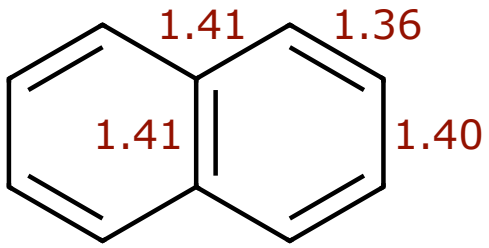


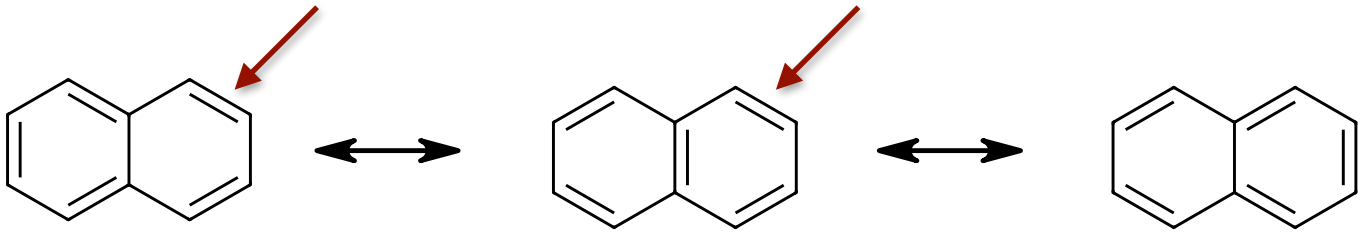
10 π 14 π

naphthalene anthracene phenanthrene

resonance energy	61	83	91
per ring	30.5	27.7	30.3

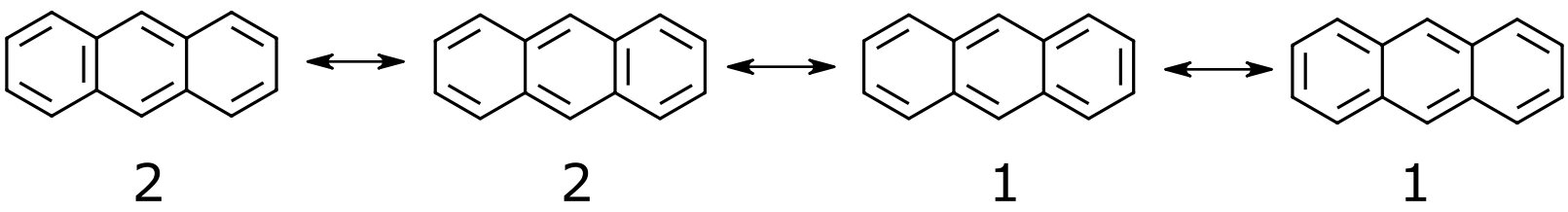
Polycyclic Aromatic Hydrocarbons (PAH)

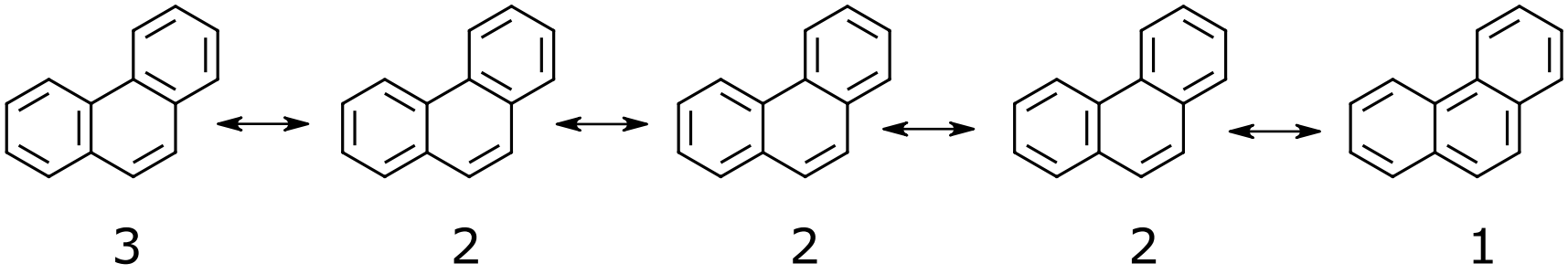




In 2 resonance structures there is a double bond between C1 and C2

Polycyclic Aromatic Hydrocarbons (PAH)

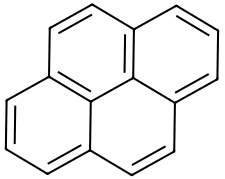




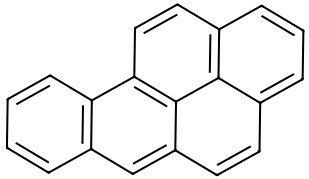
In phenanthrene there are more resonance structures containing 6-electron benzene rings

Polycyclic Aromatic Hydrocarbons (PAH)

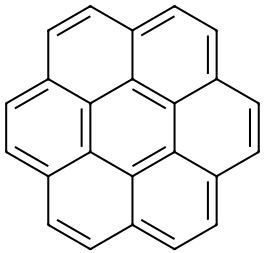
the Hückel rule does not apply to large fused systems: the following are all aromatic



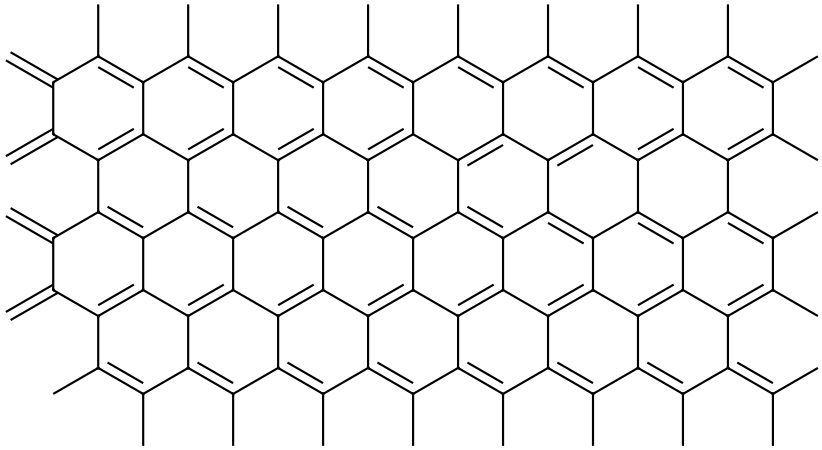
pyrene
16 π



benzopyrene
20 π

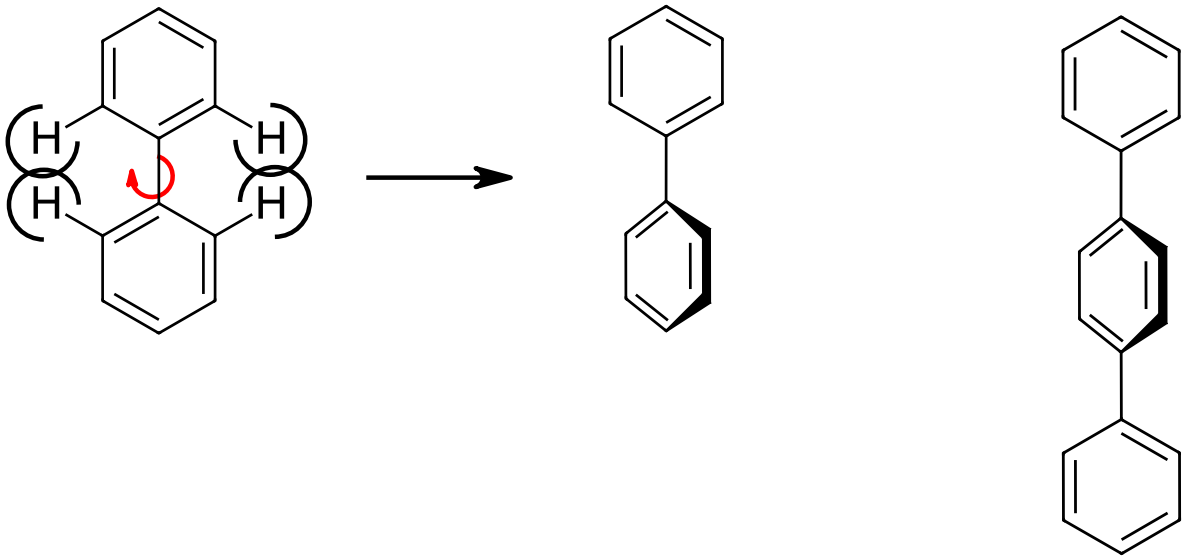


coronene
24 π



graphene

Diphenyl and Terphenyl



- In diphenyl and terphenyl the rings are twisted around the central single bond to avoid torsional strain.

- The aromatic rings are independent.

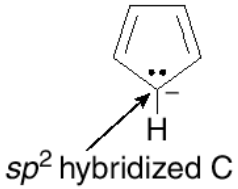
Other Aromatic Compounds

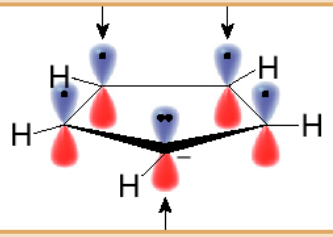
cyclopentadiene
not aromatic
$pK_a = 15$

cyclopentadienyl anion
aromatic
a stabilized conjugate base

The cyclopentadienyl anion

The ring is completely conjugated with **6 π electrons**.





The lone pair resides in a *p* orbital.

- **The cyclopentadienyl anion is aromatic because it is cyclic, planar, completely conjugated, and has six π electrons.**

Other Aromatic Compounds

cyclopentadienyl anion
- **6** π electrons
- contains $4n + 2$ π electrons

aromatic

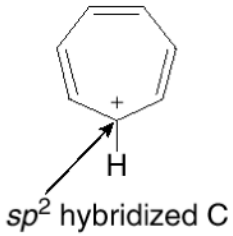
cyclopentadienyl cation
- **4** π electrons
- contains $4n$ π electrons

antiaromatic

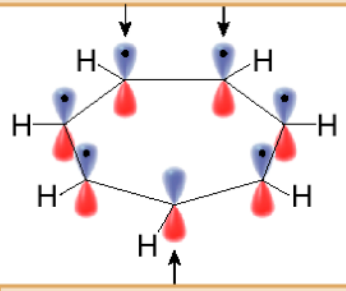
cyclopentadienyl radical
- **5** π electrons
- does not contain either $4n$ or $4n + 2$ π electrons

nonaromatic

The tropylium cation

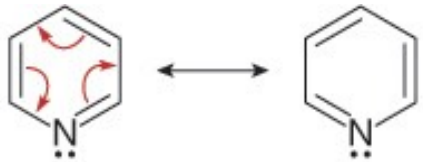


The ring is completely conjugated with **6 π electrons**.

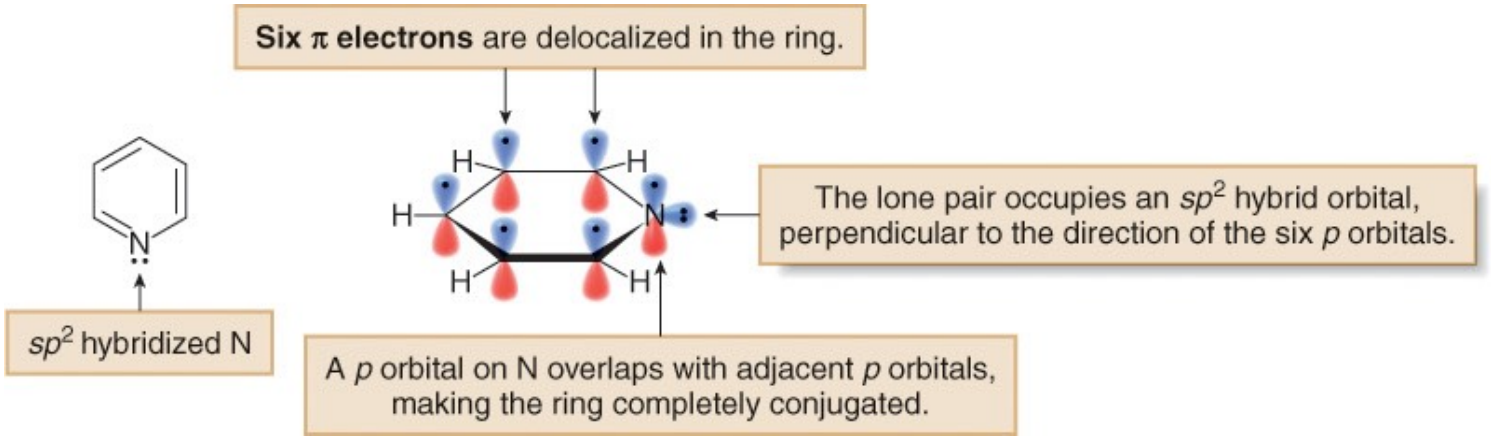


- The tropylium cation is aromatic because it is cyclic, planar, completely conjugated, and has six π electrons delocalized over the seven atoms of the ring.

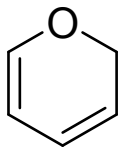
Aromatic Heterocycles: Pyridine



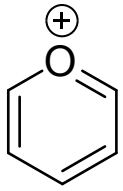
two resonance structures for pyridine
6 π electrons



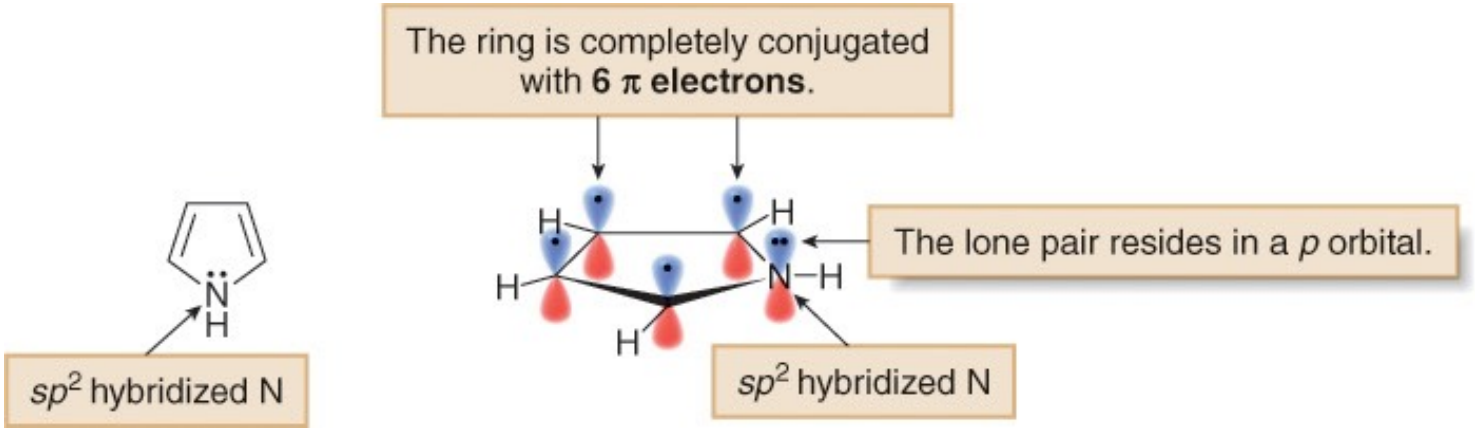
2H-pyran
4 π electrons
nonaromatic

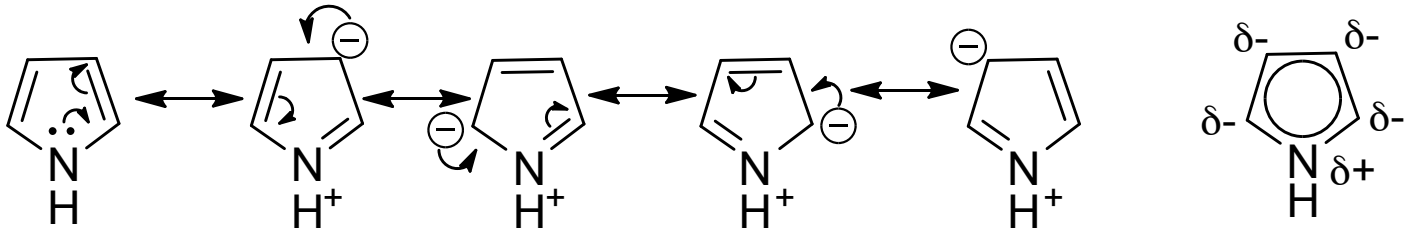


pyrilium ion
6 π electrons
aromatic

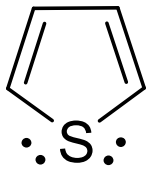


Aromatic Heterocycles: Pyrrole

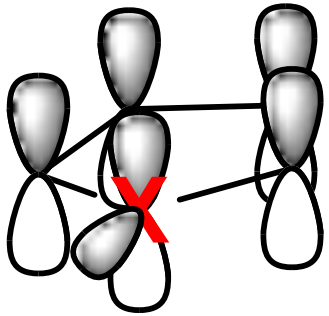




furan

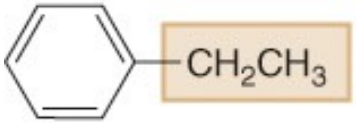


thiophen



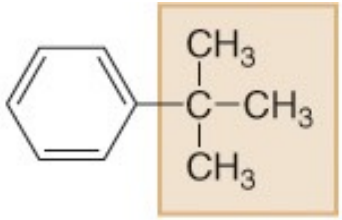
Nomenclature: 1 Substituent

Systematic:



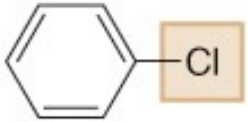
ethyl group

ethylbenzene



tert-butyl group

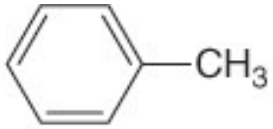
tert-butylbenzene



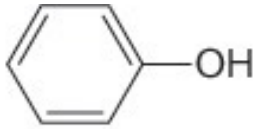
chloro group

chlorobenzene

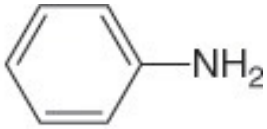
Common:



toluene
(methylbenzene)



phenol
(hydroxybenzene)

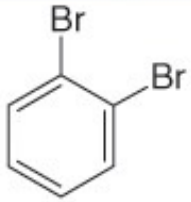


aniline
(aminobenzene)

Nomenclature: 2 Substituents

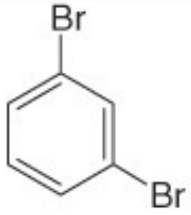
Identical:

1,2-disubstituted benzene
ortho isomer



1,2-dibromobenzene
o-dibromobenzene

1,3-disubstituted benzene
meta isomer



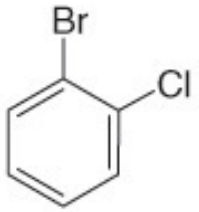
1,3-dibromobenzene
m-dibromobenzene

1,4-disubstituted benzene
para isomer

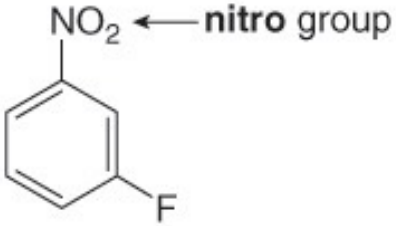
1,4-dibromobenzene
p-dibromobenzene

Different:

Alphabetize two different substituent names:

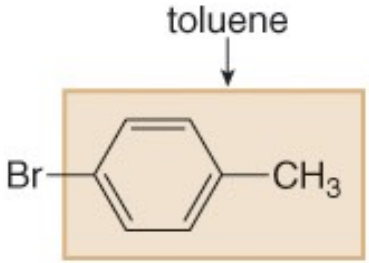


***o*-bromochloro-benzene**

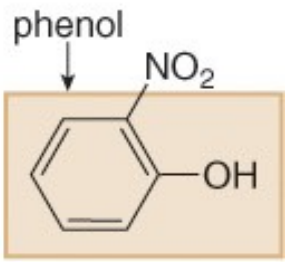


***m*-fluoronitro-benzene**

Use a common root name:



***p*-bromo**toluene

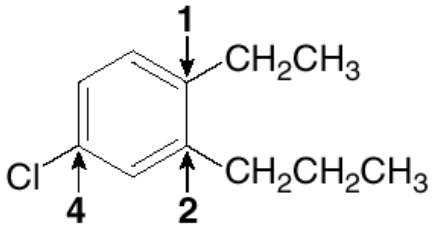


***o*-nitro**phenol

Nomenclature: 3 or More Substituents

Examples of naming polysubstituted benzenes

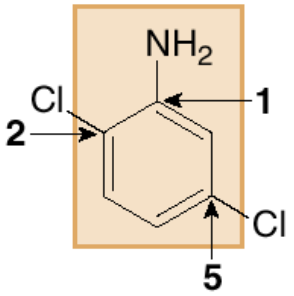
[1]



- Assign the lowest set of numbers.
- Alphabetize the names of all the substituents.

4-chloro-1-ethyl-2-propylbenzene

[2]

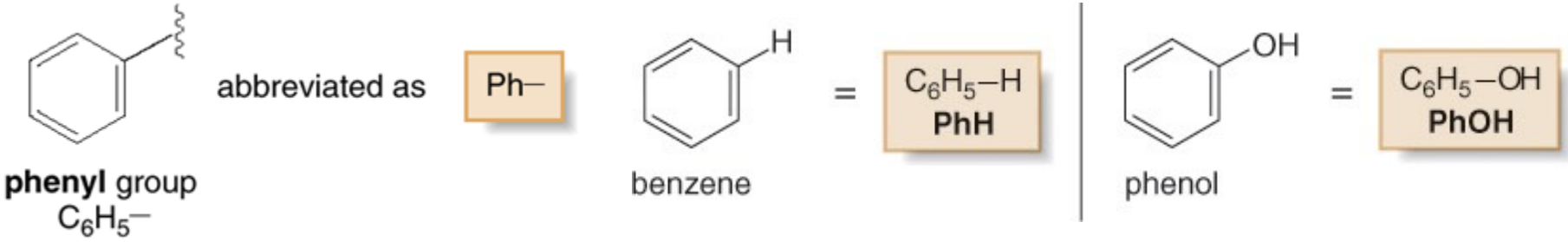


- Name the molecule as a derivative of the common root **aniline.**
- Designate the position of the NH_2 group as "1," and then assign the lowest possible set of numbers to the other substituents.

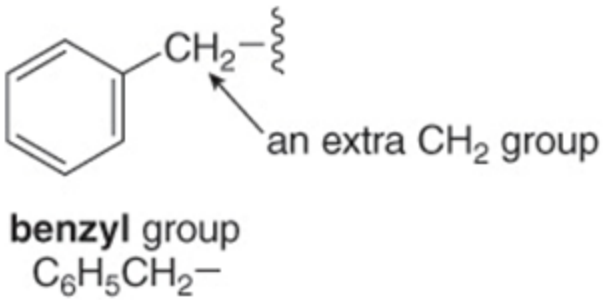
2,5-dichloroaniline

Nomenclature

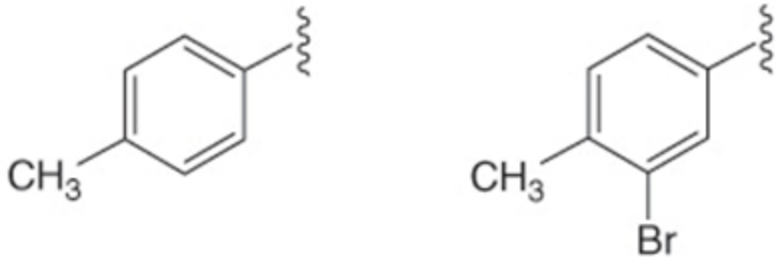
A benzene substituent is called a phenyl group, and it can be abbreviated in a structure as "Ph-".



The benzyl group:



Aryl groups:

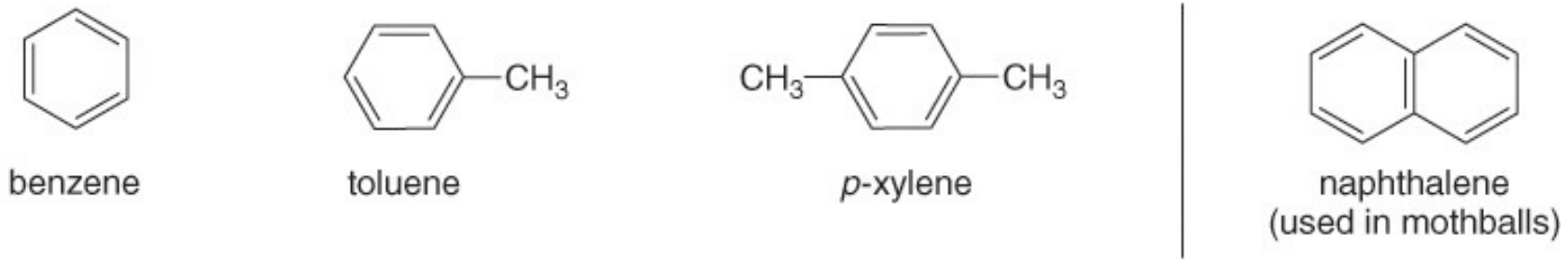


Interesting Aromatic Compounds

- Benzene and toluene, are obtained from petroleum refining and are useful starting materials for synthetic polymers.

- Compounds containing two or more benzene rings that share carbon—carbon bonds are called polycyclic aromatic hydrocarbons (PAHs). Naphthalene, the simplest PAH, is the active ingredient in mothballs.

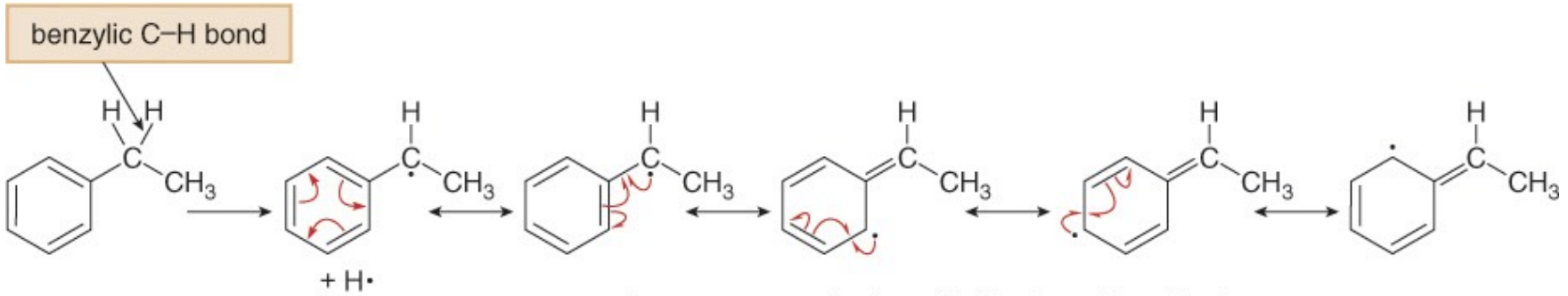
The components of the gasoline additive BTX



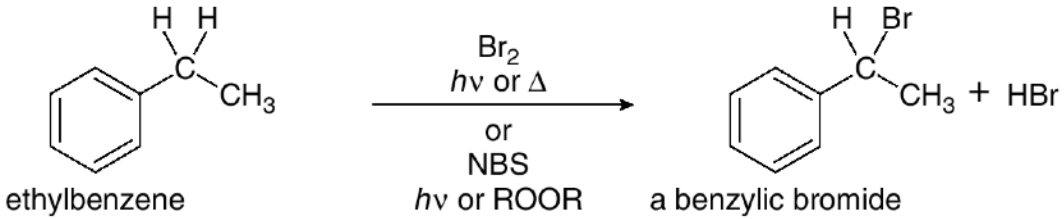
benzene | toluene | *p*-xylene | naphthalene (used in mothballs)

Side Chain Reactivity: Radical Halogenation

Benzylic C—H bonds are weaker than most other sp^3 hybridized C—H bonds, because homolysis forms a resonance-stabilized benzylic radical.



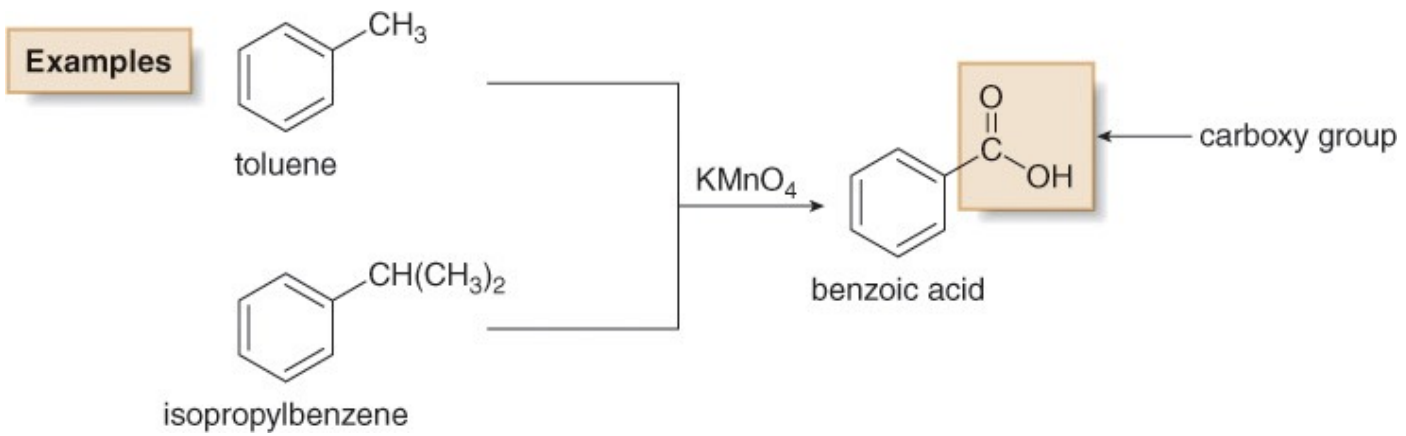
five resonance structures for the benzylic radical

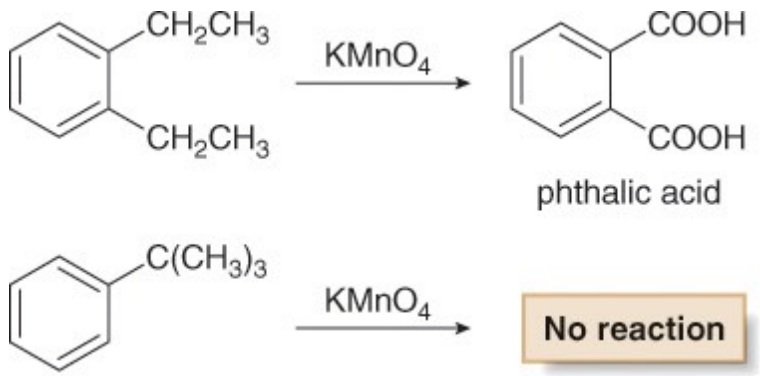


ethylbenzene — a benzylic bromide

radical conditions

Side Chain Reactivity: Oxidation





Side Chain Reactivity: Reduction