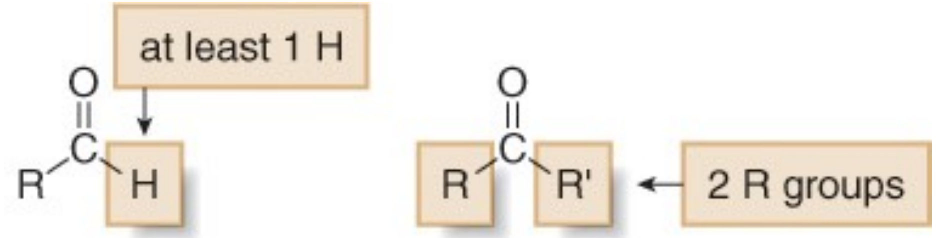
Carbonyl Compounds

Introduction

Introduction

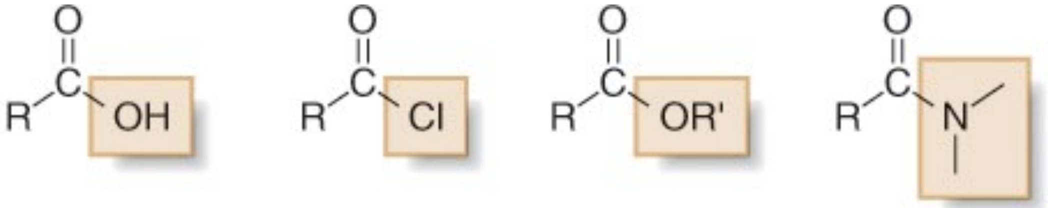
Two broad classes of compounds contain the carbonyl group:

[1] Compounds that have only carbon and hydrogen atoms bonded to the carbonyl



aldehyde ketone

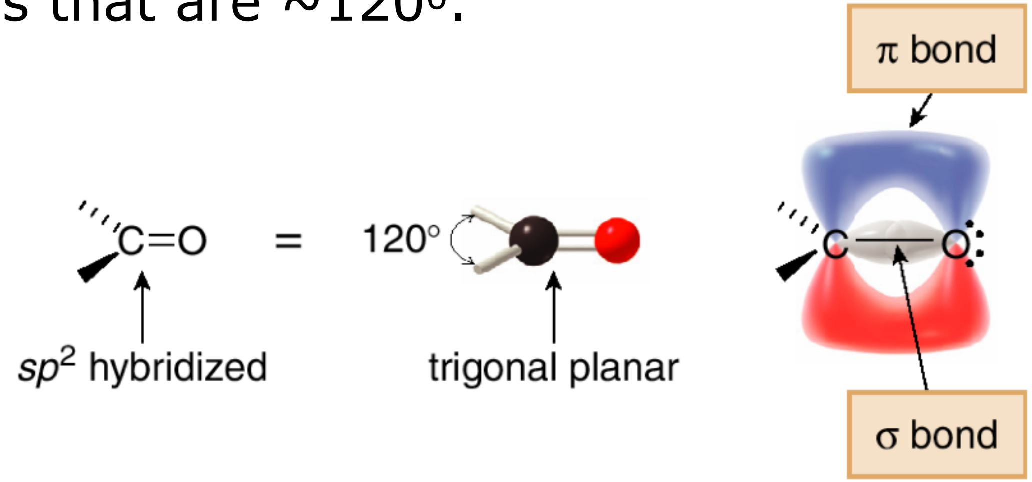
[2] Compounds that contain an eteroatom (N, O, S, Cl) bonded to the carbonyl



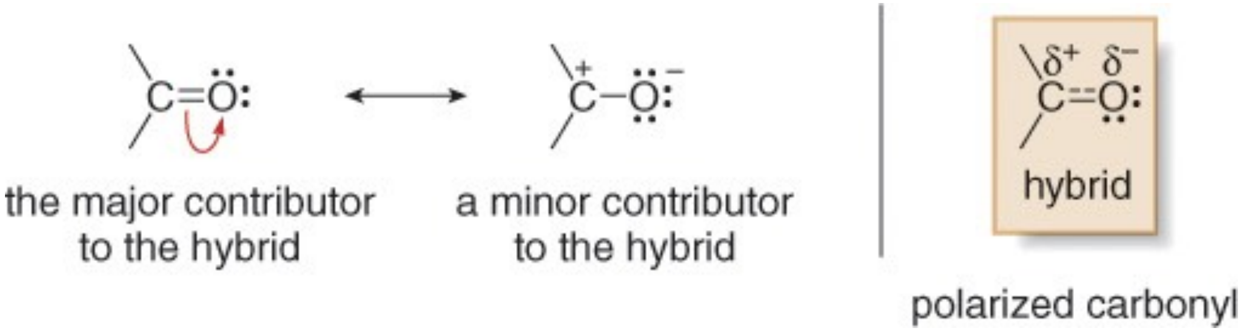
carboxylic acid acid chloride ester amide

Introduction

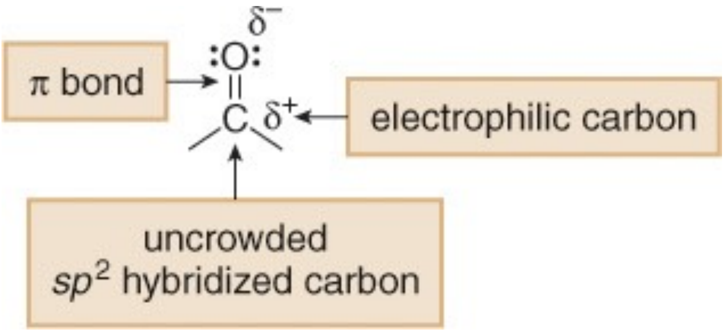
- Carbonyl carbons are *sp*^{2} hybridized, trigonal planar, and have bond angles that are ~120^0.



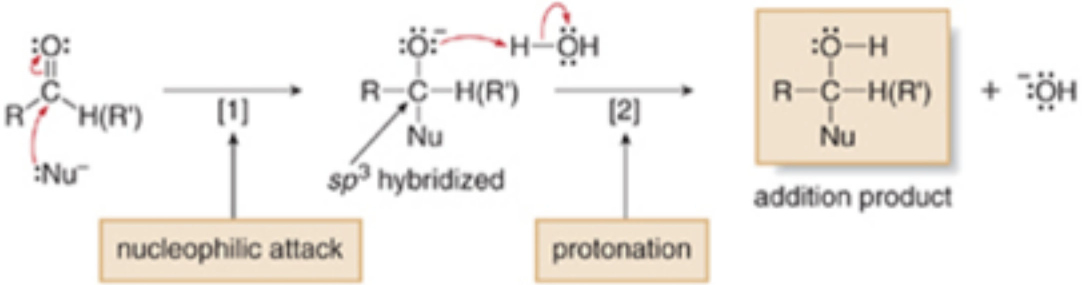
- The electronegative oxygen atom in the carbonyl group means that the bond is polarized, making the carbonyl carbon electron deficient.



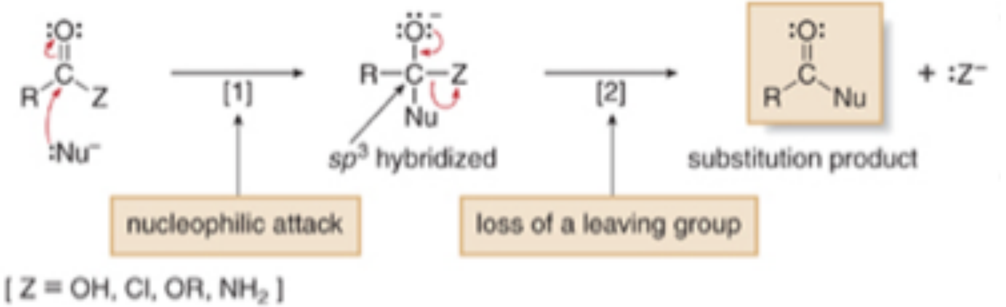
General Reactions of Carbonyl Compounds



Aldehydes and ketones



Acyl derivatives



Aldehydes and Ketones

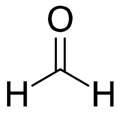
Chapter 19

Organic Chemistry, *8th Edition*

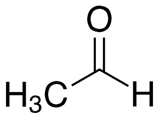
John McMurry

Nomenclature of Aldehydes

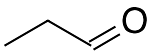
- Find the longest chain containing the CHO group, and change the –*e* ending of the parent alkane to the suffix *–al*. If the CHO group is bonded to a ring, name the ring and add the suffix *–carbaldehyde*.
- A common name for an aldehyde is formed by taking the common parent name and adding the suffix *–aldehyde*.



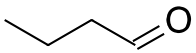
methanal
(formaldehyde)



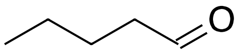
ethanal
(acetaldehyde)



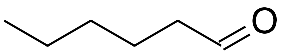
propanal
(propionald.)



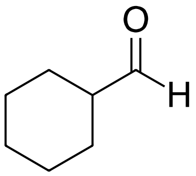
butanal
(butyraldehyde)



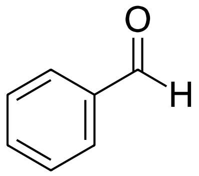
pentanal
(valeraldehyde)



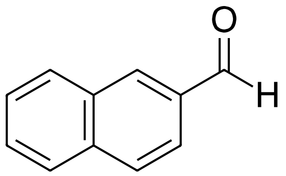
hexanal
(caproic aldehyde)



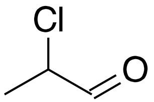
cyclohexanecarbaldehyde



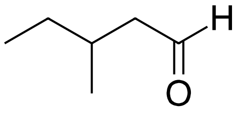
benezenecarbaldehyde
(benzaldehyde)



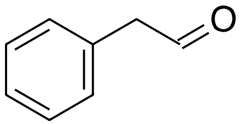
2-naphthalenecarbaldehyde



2-chloropropanal

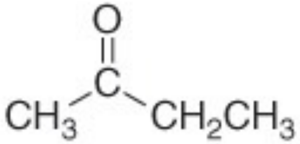


3-methylpentanal

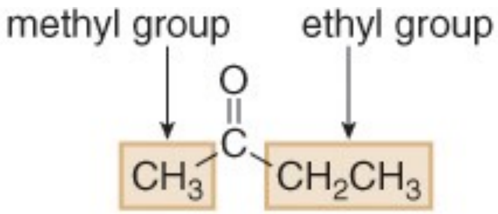


phenylethanal

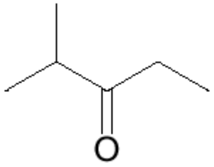
Nomenclature of Ketones



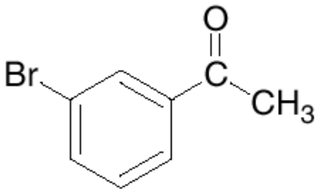
IUPAC name: **2-butanone**



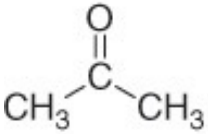
Common name: **ethyl methyl ketone**



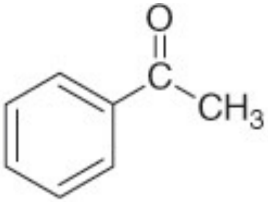
IUPAC name: 2-methyl-3-pentanone
Common name: ethyl isopropyl ketone



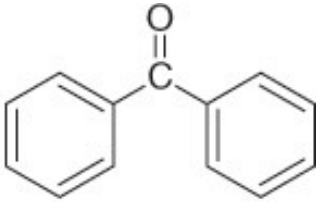
m-bromoacetophenone
or
3-bromoacetophenone



acetone

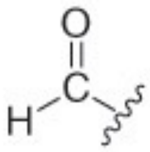


acetophenone

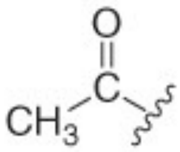


benzophenone

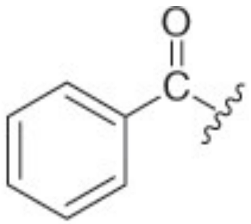
Nomenclature of Aldehydes and Ketones



formyl group

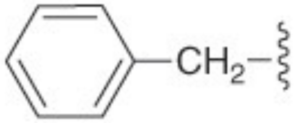


acetyl group



benzoyl group

Do not confuse a **benzyl** group with a **benzoyl** group.



benzyl group

Physical Properties

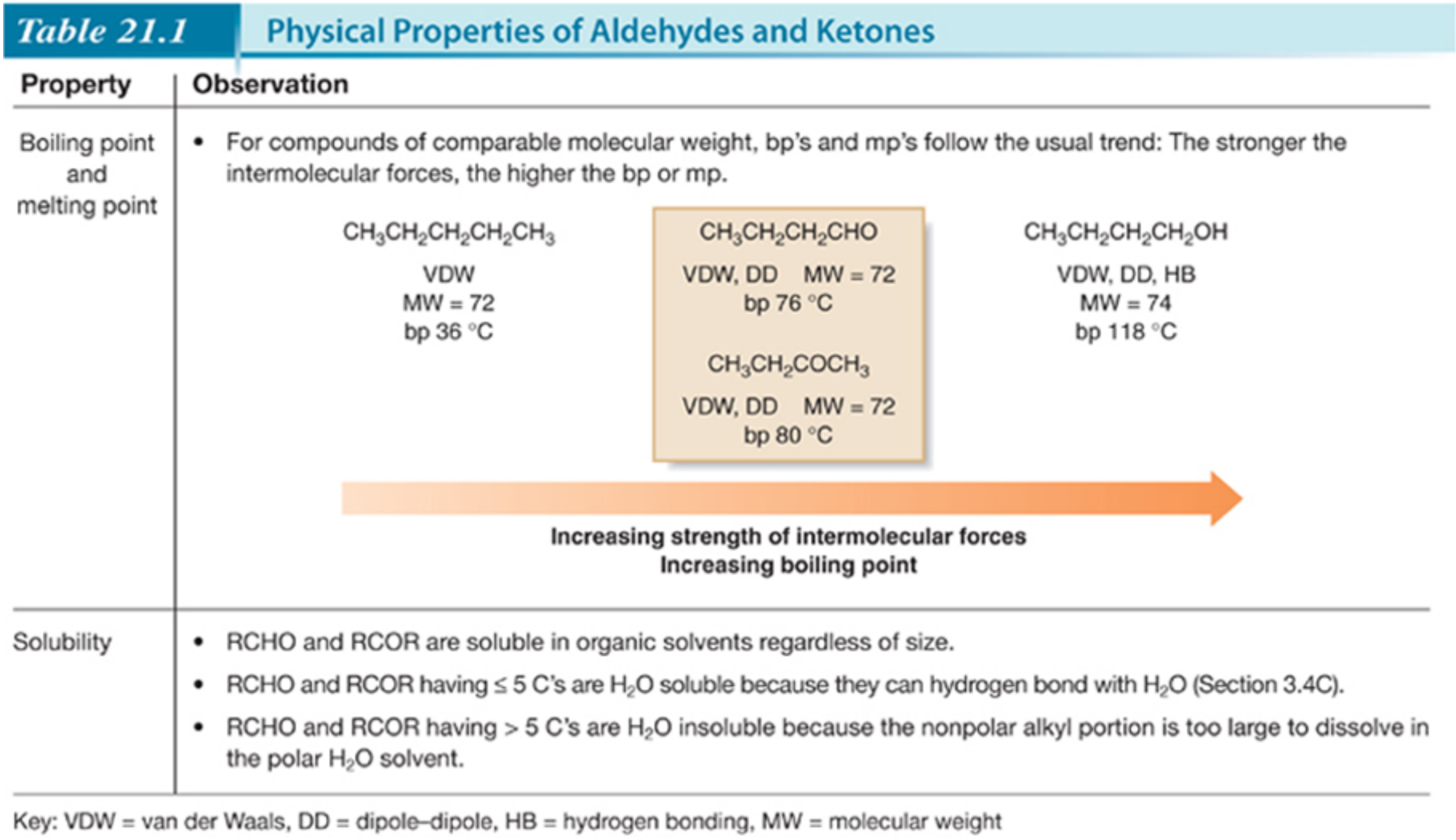
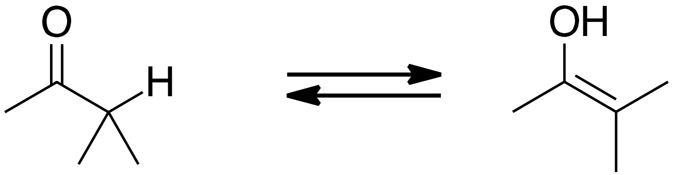


Table 21.1 Physical Properties of Aldehydes and Ketones

Property	Observation
Boiling point and melting point	• For compounds of comparable molecular weight, bp's and mp's follow the usual trend: The stronger the intermolecular forces, the higher the bp or mp. $CH_3CH_2CH_2CH_2CH_3$ VDW MW = 72 bp 36 °C $CH_3CH_2CH_2CHO$ VDW, DD MW = 72 bp 76 °C $CH_3CH_2COCH_3$ VDW, DD MW = 72 bp 80 °C $CH_3CH_2CH_2CH_2OH$ VDW, DD, HB MW = 74 bp 118 °C **Increasing strength of intermolecular forces** **Increasing boiling point**
Solubility	• RCHO and RCOR are soluble in organic solvents regardless of size. • RCHO and RCOR having ≤ 5 C's are H_2O soluble because they can hydrogen bond with H_2O (Section 3.4C). • RCHO and RCOR having > 5 C's are H_2O insoluble because the nonpolar alkyl portion is too large to dissolve in the polar H_2O solvent.

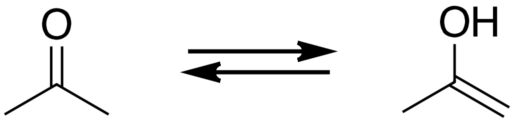
Key: VDW = van der Waals, DD = dipole–dipole, HB = hydrogen bonding, MW = molecular weight

Keto-Enol Tautomerism



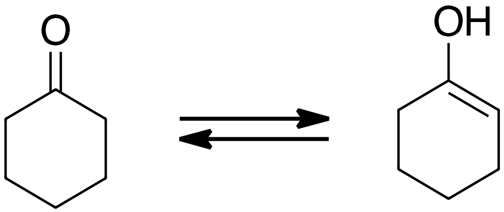
Keto tautomer

Enol tautomer



99,9999999%

0,0000001%

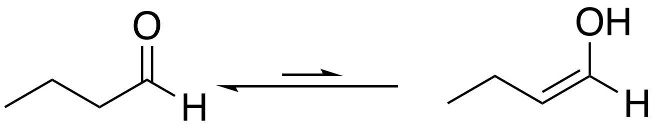


99,9999%

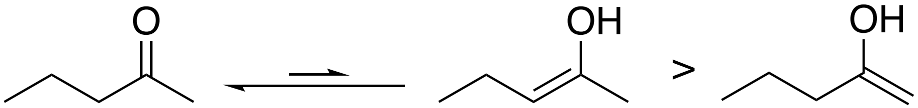
0,0001%

Keto-Enol Tautomerism

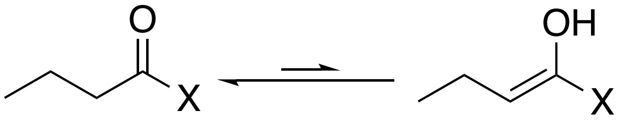
Aldehydes: 1 enol



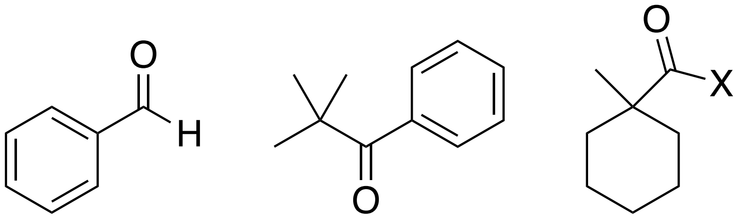
Ketones: 2 enols



Acid derivatives:
1 enol

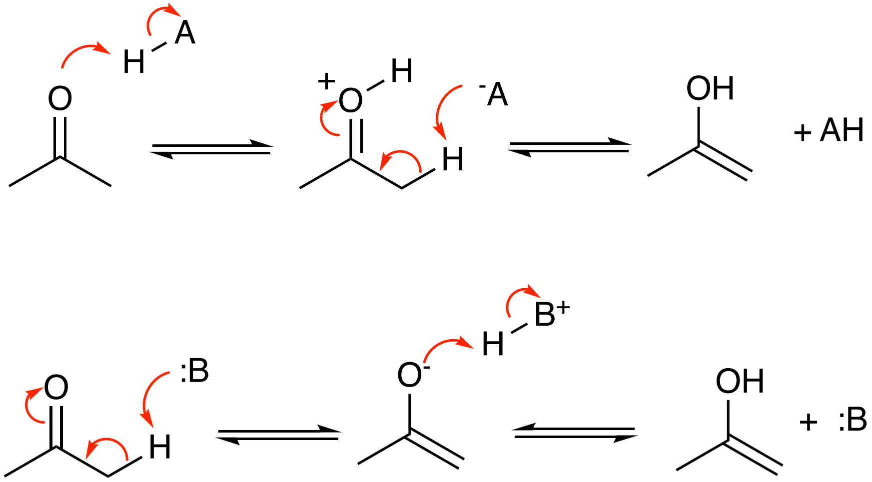


Non-enolizable



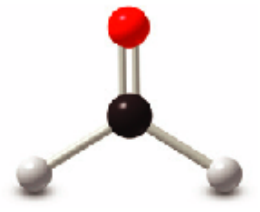
Keto-Enol Tautomerism

Enolization is catalyzed by both acids and bases



The catalyst accelerates the equilibrium; it does not influence its position

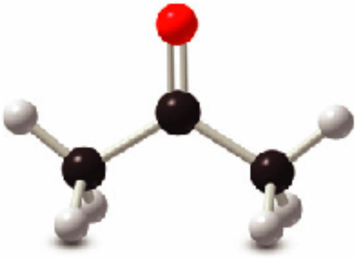
Interesting Aldehydes and Ketones



formaldehyde

$CH_2{=}O$

Billions of pounds of formaldehyde are produced annually from the oxidation of methanol. It is sold as a 37% solution called formalin which is used as a disinfectant, antiseptic, and preservative for biological specimens. It is a product of incomplete combustion of coal, and is partly responsible for the irritation caused by smoggy air.



acetone

$(CH_3)_2C{=}O$

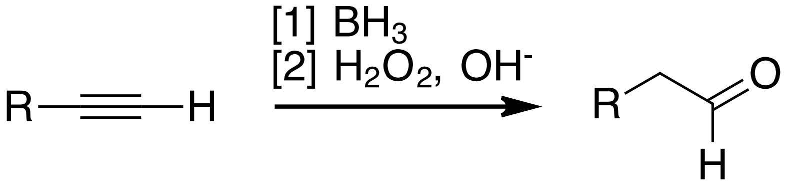
Acetone is an industrial solvent. It is also produced in vivo during breakdown of fatty acids. Diabetics often have unusually high levels of acetone in their blood streams.

Interesting Aldehydes and Ketones

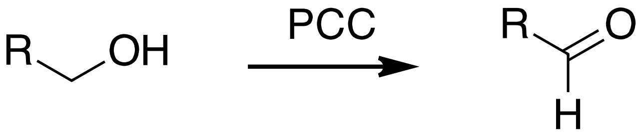
Many aldehydes and ketones with characteristic odors occur in nature.

Preparation of Aldehydes

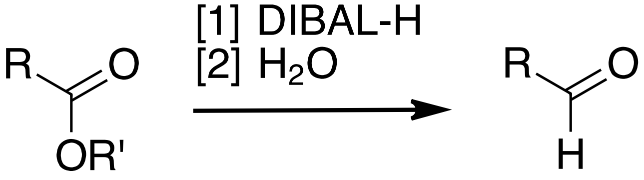
Hydration of an alkyne (hydroboration-oxidation)



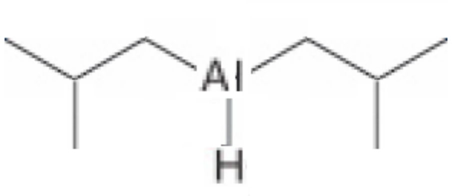
Oxidation of 1ry alcohols



Reduction of esters and acyl chlorides



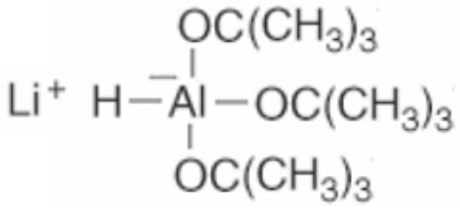
Preparation of Aldehydes and Ketones



$i\text{-}Bu_2AlH$

DIBAL-H

Diisobutilaluminium hydride



$$Li^+ \; H-\overset{-}{Al}(OC(CH_3)_3)_3$$

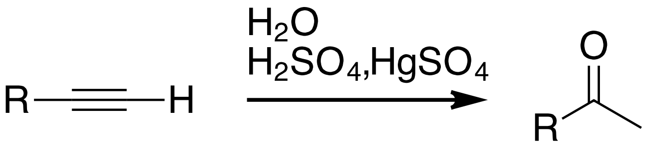
$Li^+ \overset{-}{Al}H[O\text{-}t\text{-}Bu]_3$

Lithium tri-*tert*-butoxyaluminium hydride

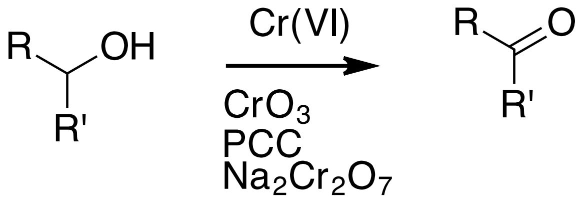
Sterically hindered - Less reactive than $LiAlH_4$

Preparation of Ketones

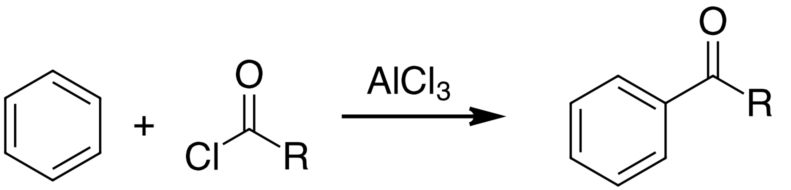
Hydration of alkynes



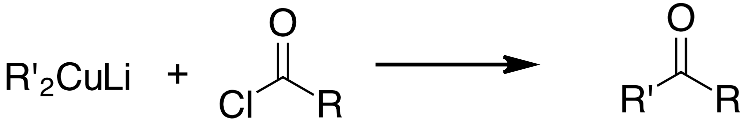
Oxidation of 2^{ry} alcohols



Friedel-Crafts acylation

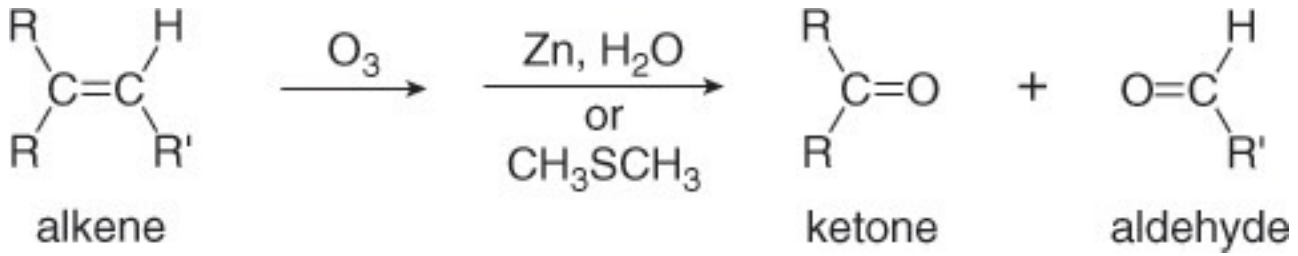


Acylation of organocuprates



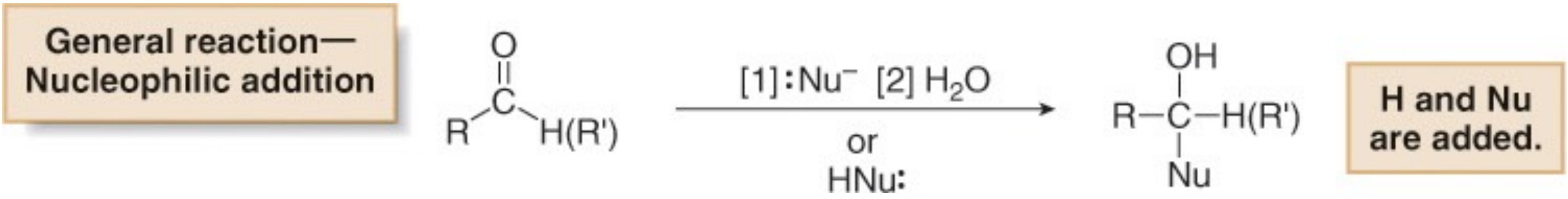
Preparation of Aldehydes and Ketones

Aldehydes and ketones are also both obtained as products of the oxidative cleavage of alkenes.



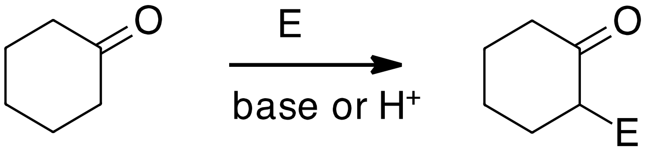
Reactions of Aldehydes and Ketones—General

[1] Nucleophilic addition

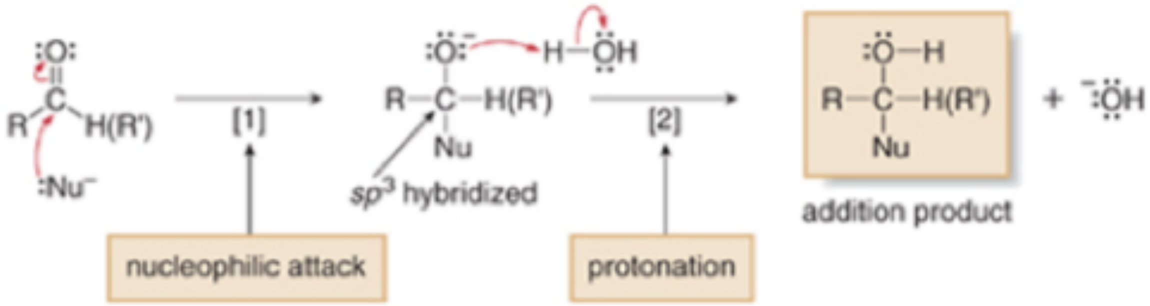


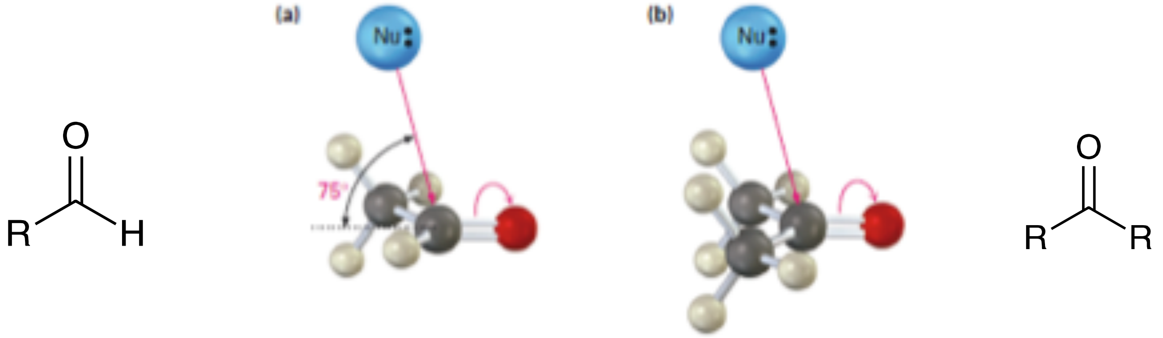
[2] Oxidation

[3] Reaction at the α carbon



Nucleophilic Addition

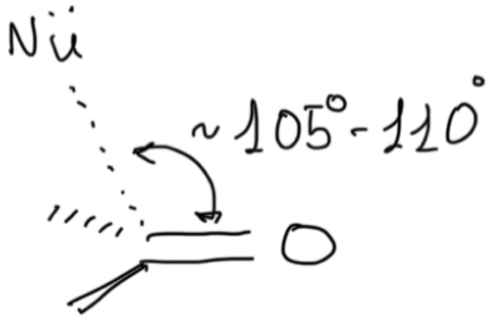




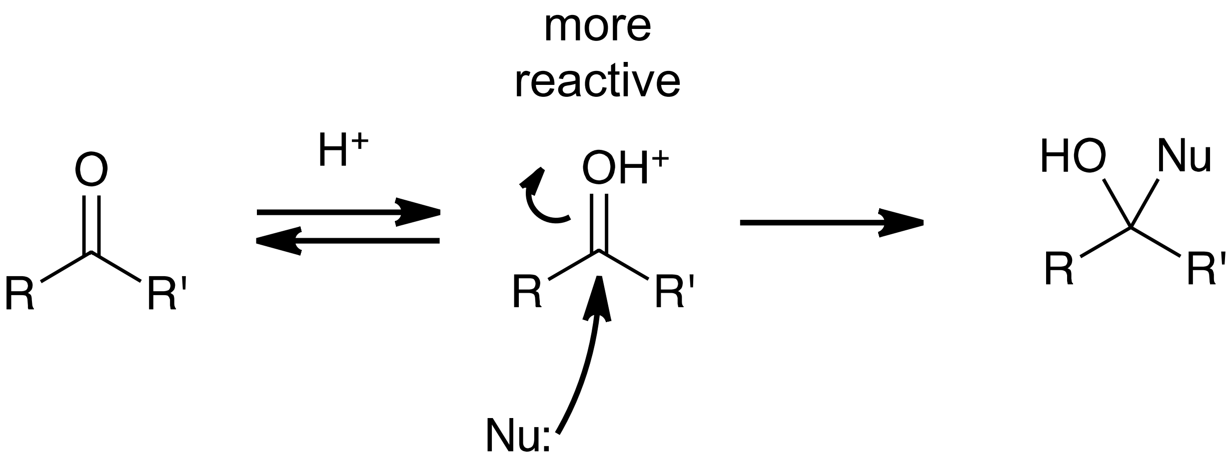
aldehyde
less crowded
more reactive

ketone
more crowded
less reactive

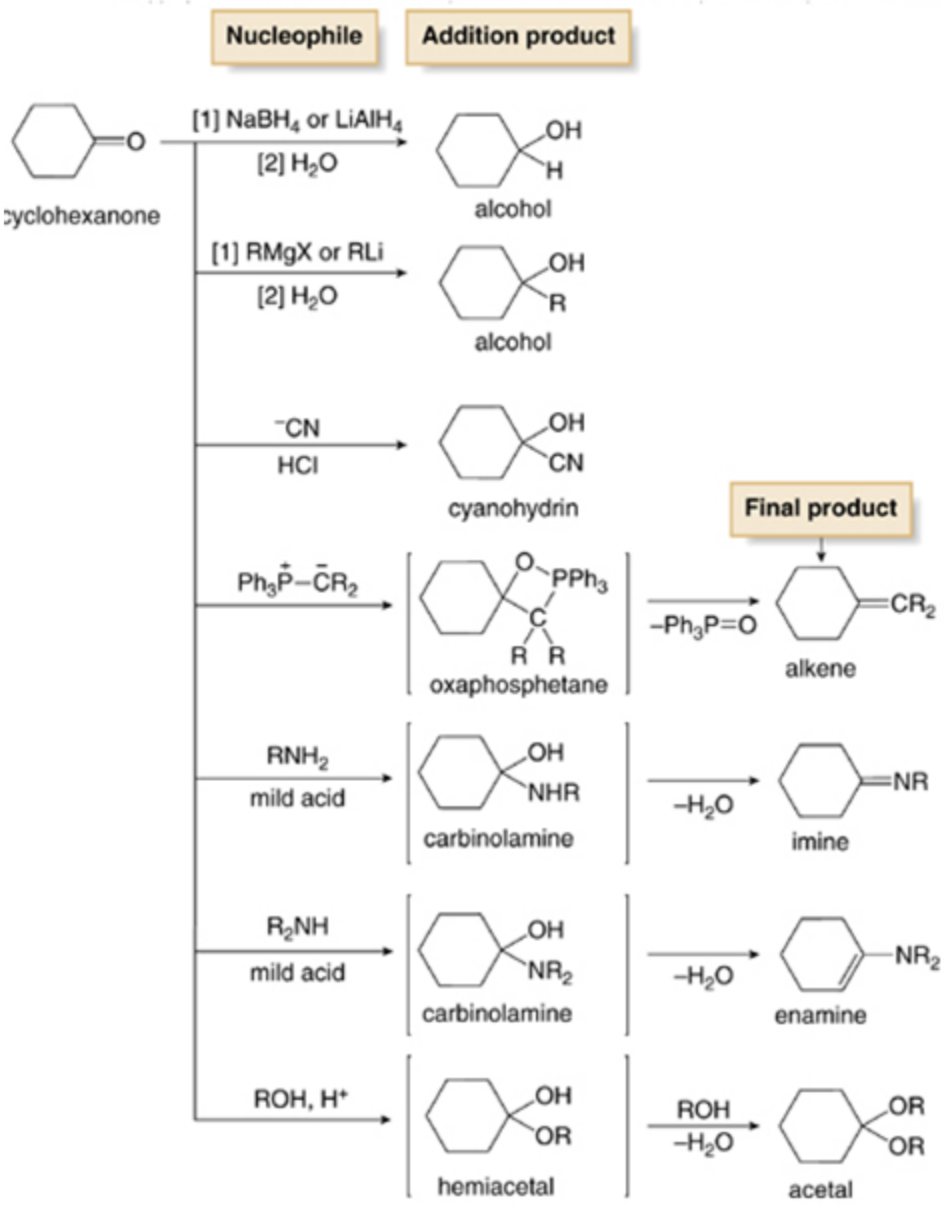
Nucleophilic Addition: Geometry



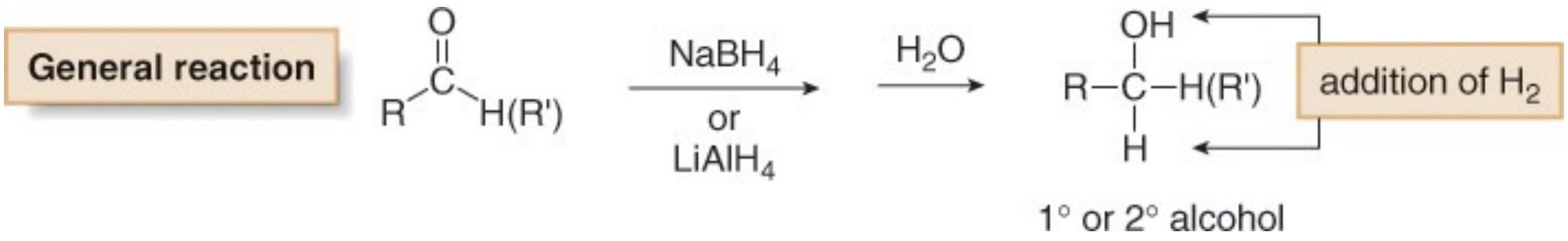
Acid Catalysis

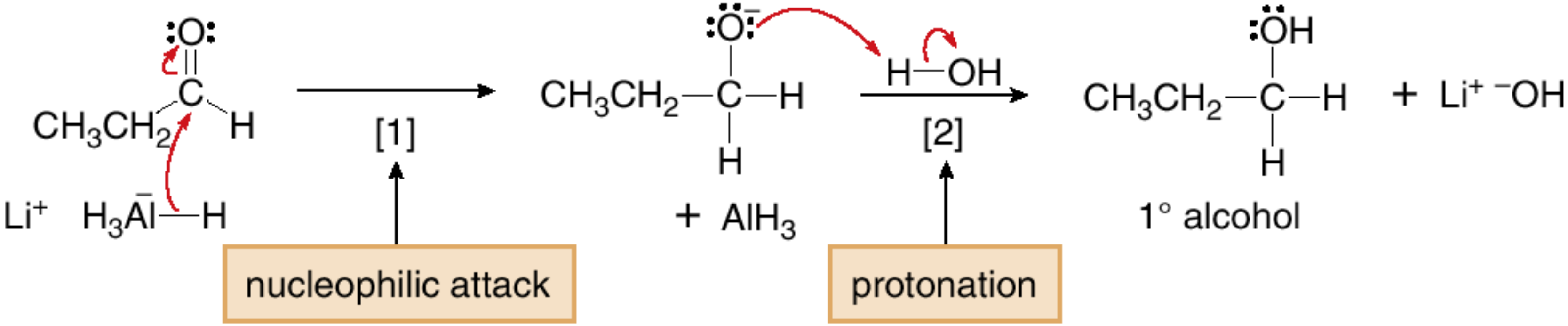


Nucleophilic Addition

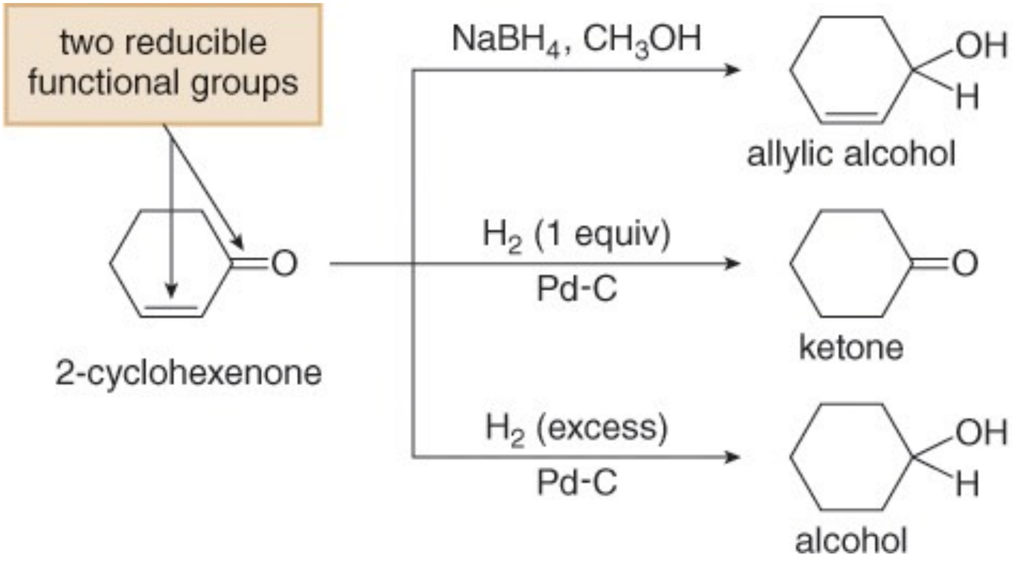


Nucleophilic Addition of H^- and R^-
A Review





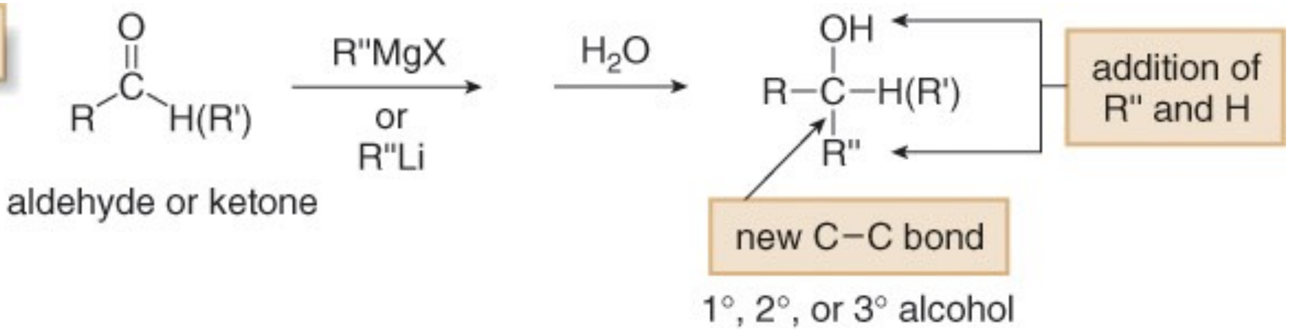
Nucleophilic Addition of H^- and R^-
A Review



- $NaBH_4$ reduces the C=O selectively to form an allylic alcohol.
- One equivalent of H_2 reduces the C=C selectively to form a ketone.
- Excess H_2 reduces both π bonds to form an alcohol.

Nucleophilic Addition of H¯ and R¯
A Review

General reaction

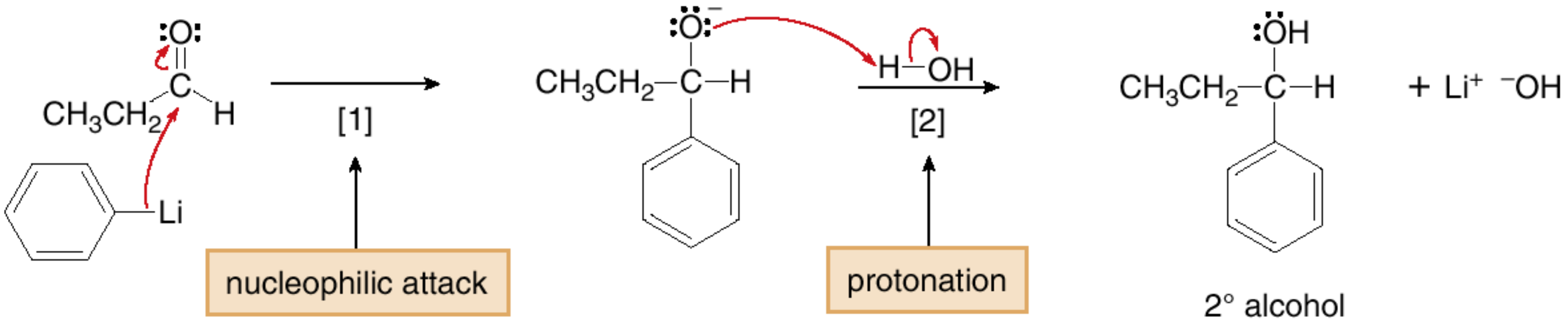


aldehyde or ketone

addition of R'' and H

new C–C bond

1°, 2°, or 3° alcohol



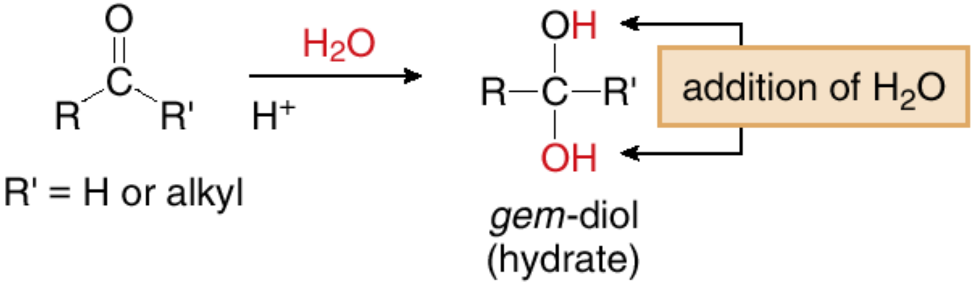
nucleophilic attack

protonation

2° alcohol

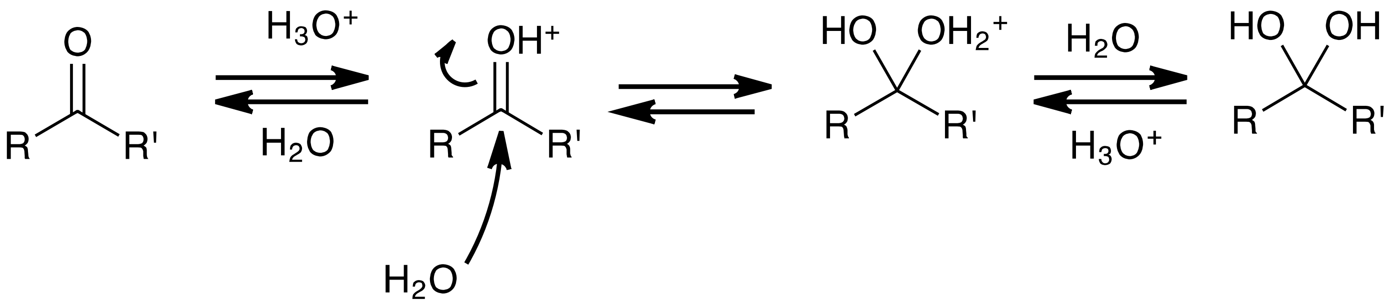
Addition of H_2O — Hydration

Nucleophilic addition of H_2O

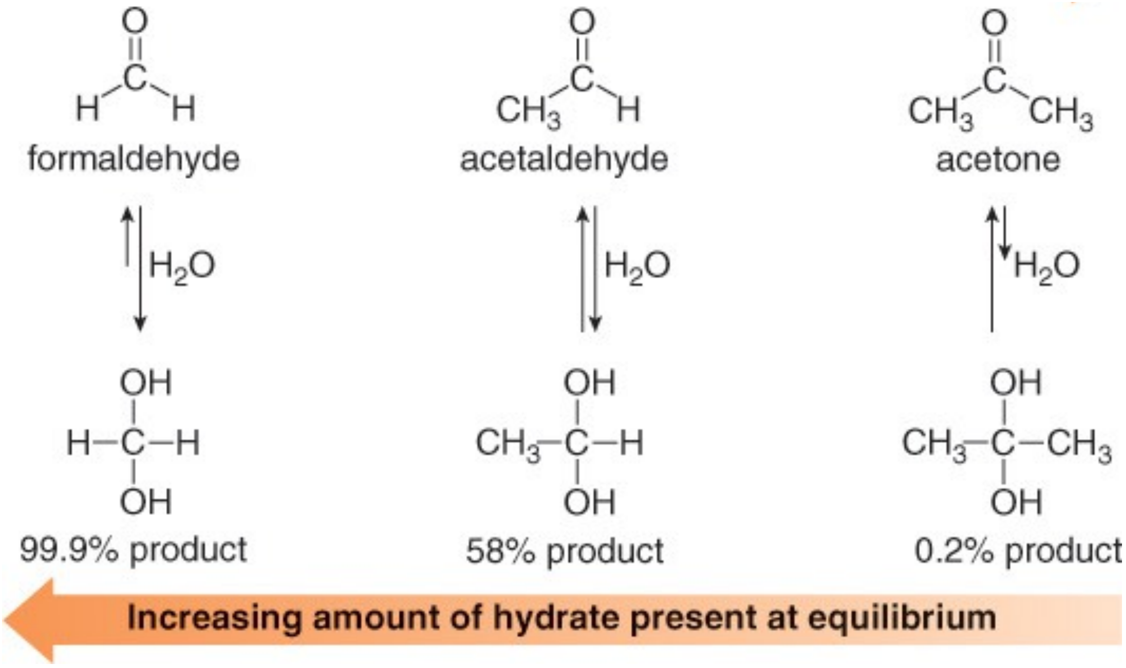


R' = H or alkyl

gem-diol
(hydrate)

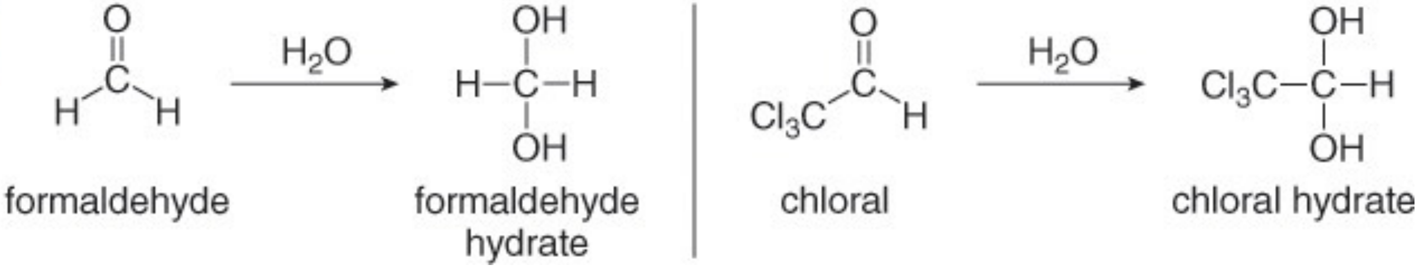


Addition of H_2O — Hydration



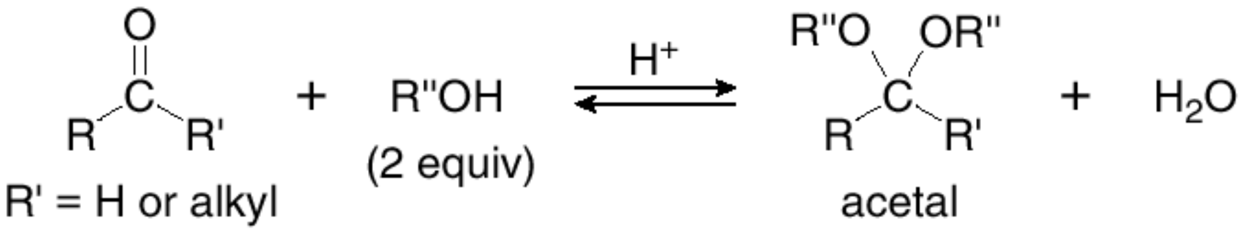
Gem-diol product yields are good only when unhindered aldehydes or aldehydes with nearby electron withdrawing groups are used.

Examples

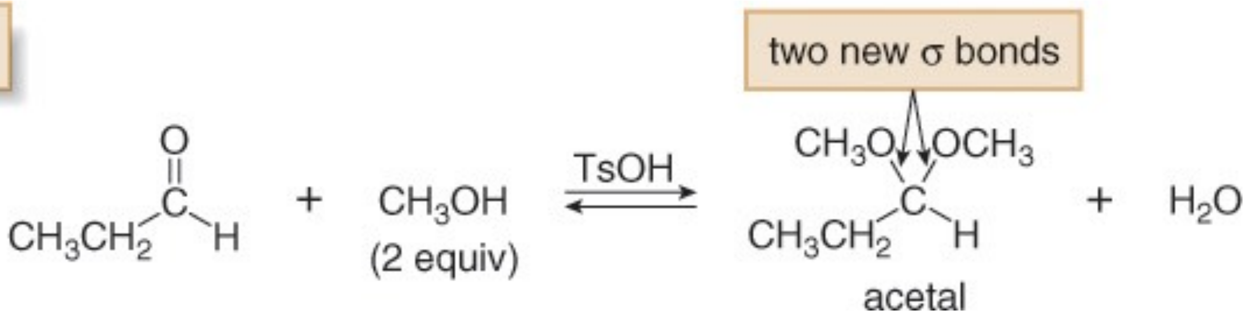


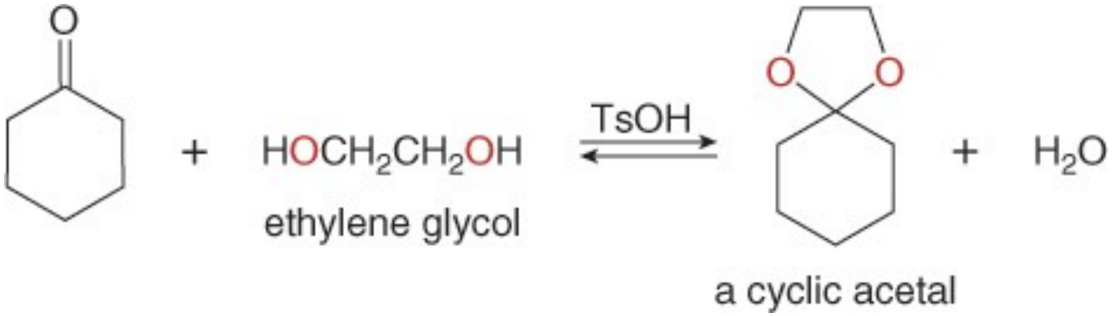
Addition of Alcohols — Acetal Formation

Acetal formation

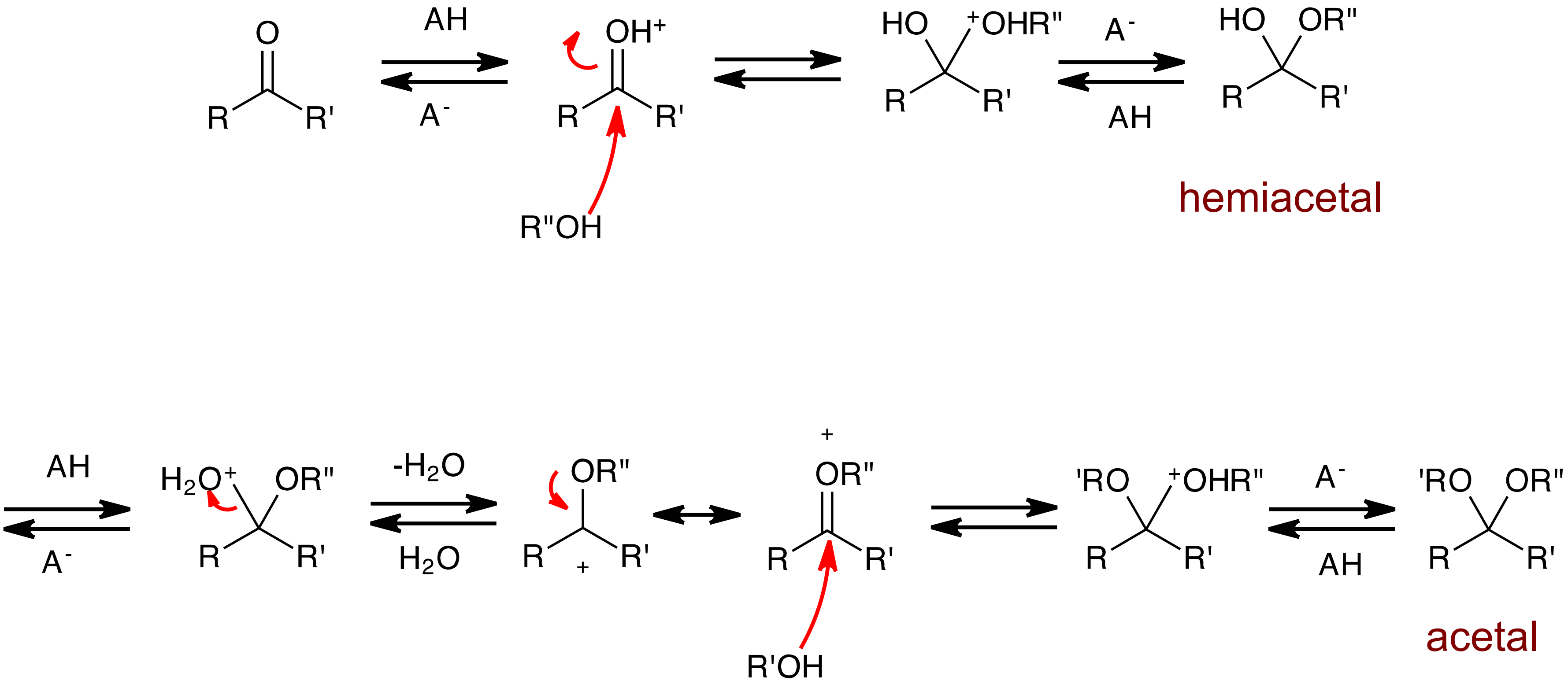


Example





Addition of Alcohols — Acetal Formation



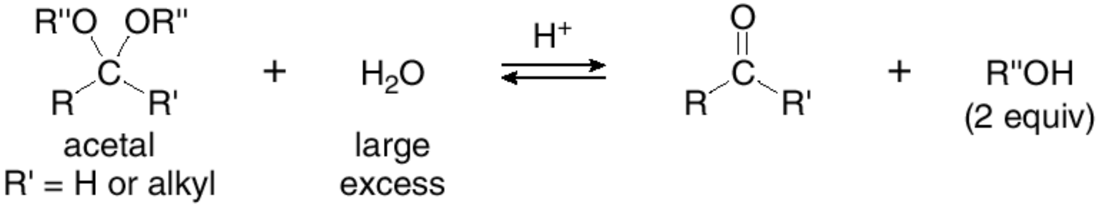
hemiacetal

resonance stabilized carbocation

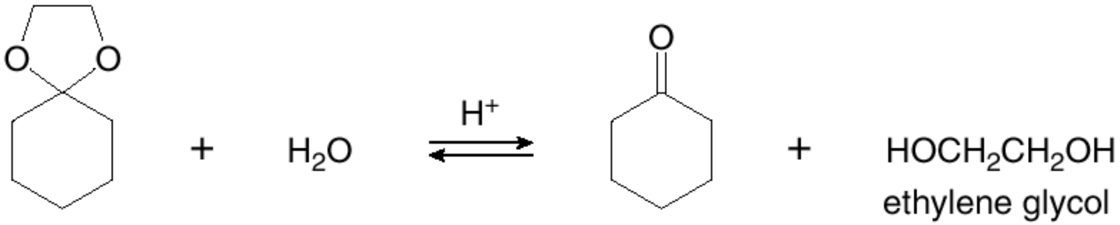
acetal

Acetal Hydrolysis

Acetal hydrolysis

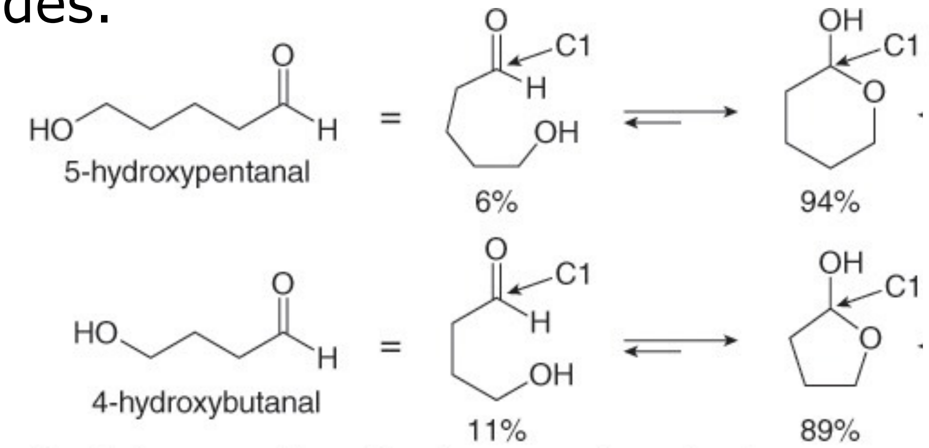


Example



Cyclic Hemiacetals and Acetals

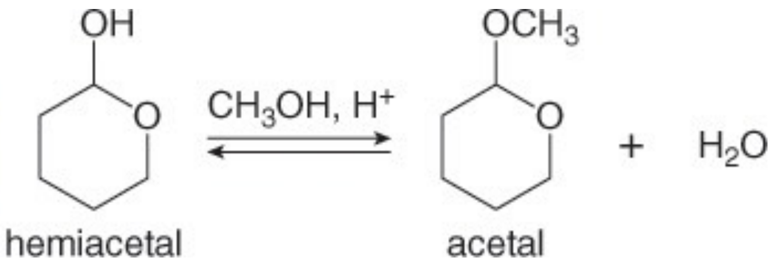
Cyclic hemiacetals are formed by intramolecular cyclization of hydroxy aldehydes.



[Equilibrium proportions of each compound are given.]

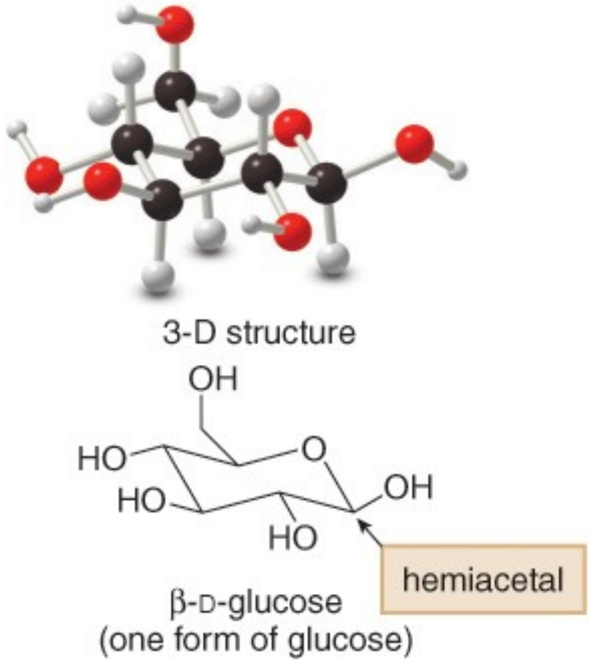
Cyclic hemiacetals can be converted to acetals by treatment with an alcohol and acid.

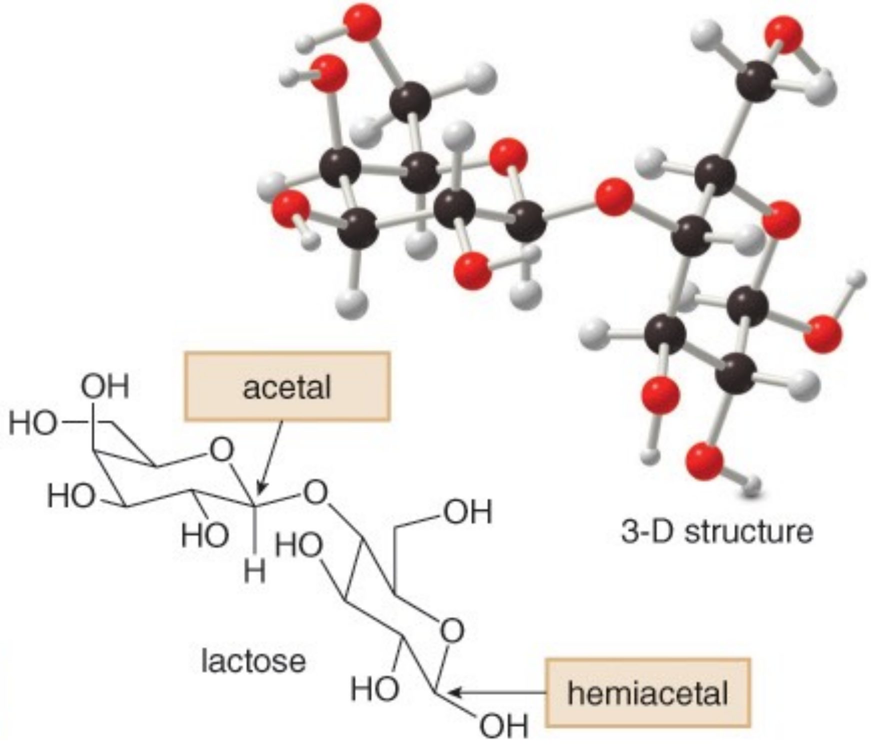
Converting a hemiacetal to an acetal



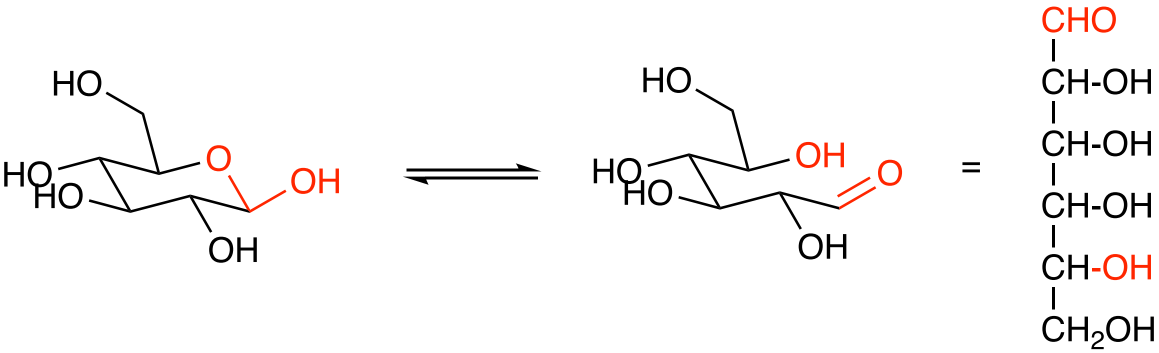
Introduction to Carbohydrates

- Carbohydrates, commonly referred to as sugars and starches, are polyhydroxy aldehydes and ketones, or compounds that can be hydrolyzed to them.
- Many carbohydrates contain cyclic acetals or hemiacetals. Examples include glucose and lactose.



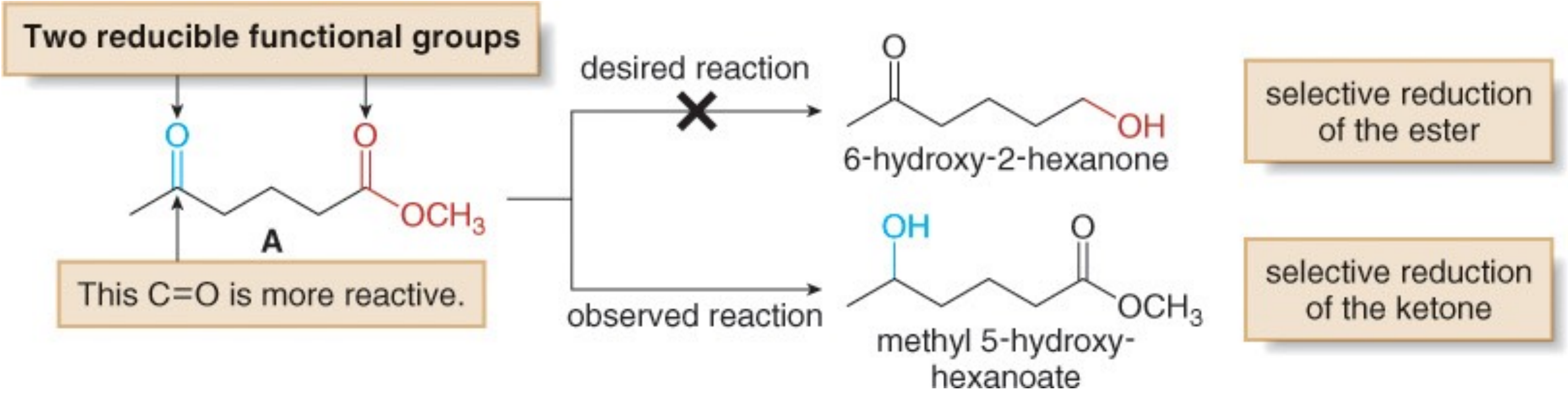


Equilibrium Between Hemiacetal and Open Chain Forms of Glucose



CHO
|
CH-OH
|
CH-OH
|
CH-OH
|
CH-OH
|
CH_2OH

Acetals as Protecting Groups



Two reducible functional groups

This C=O is more reactive.

desired reaction

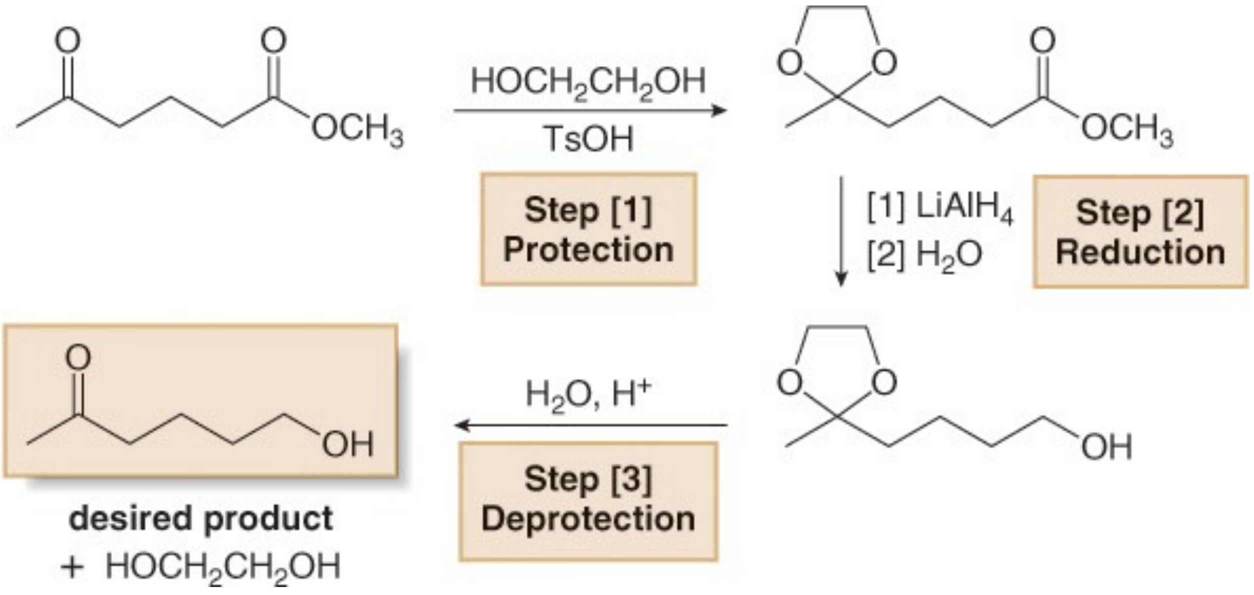
6-hydroxy-2-hexanone

selective reduction of the ester

observed reaction

methyl 5-hydroxyhexanoate

selective reduction of the ketone



$HOCH_2CH_2OH$, TsOH

Step [1] Protection

[1] $LiAlH_4$
[2] H_2O

Step [2] Reduction

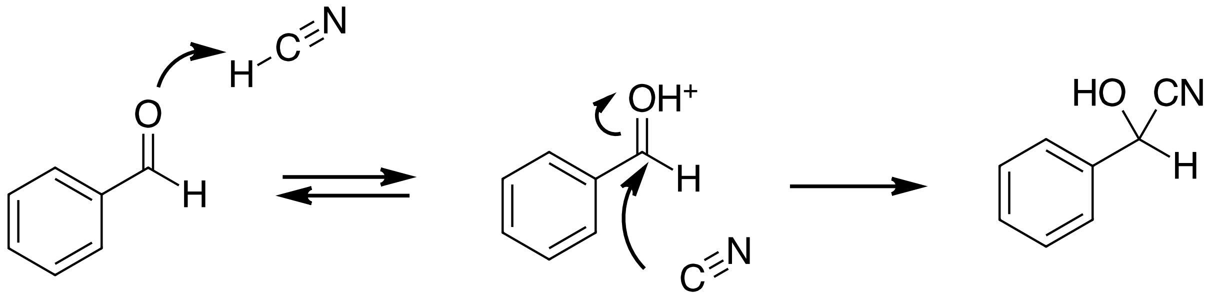
H_2O, H^+

Step [3] Deprotection

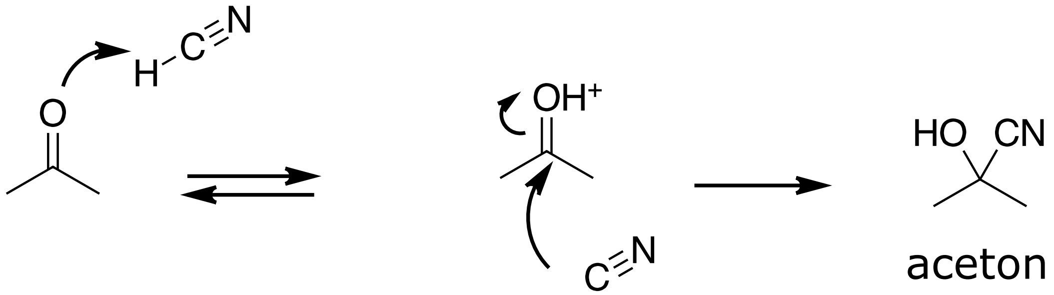
desired product
+ $HOCH_2CH_2OH$

Nucleophilic Addition of CN⁻

- Treatment of an aldehyde or ketone with HCN gives a cyanohydrin.



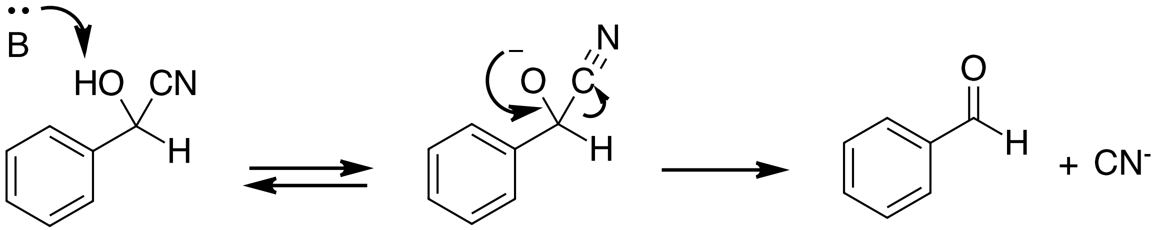
benzaldehyde cyanohydrin



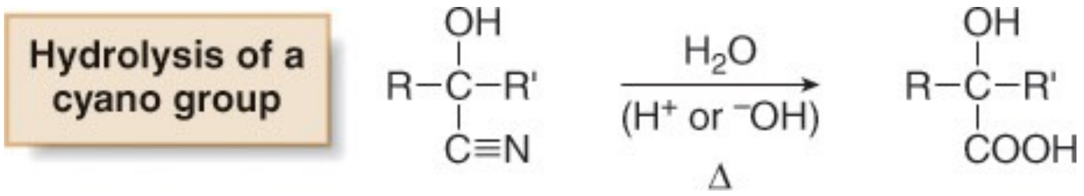
aceton cyanohydrin

Nucleophilic Addition of CN$^-$

- Cyanohydrins can be reconverted to carbonyl compounds by treatment with base. This process is just the reverse of the addition of HCN: deprotonation followed by elimination of $^-$CN.

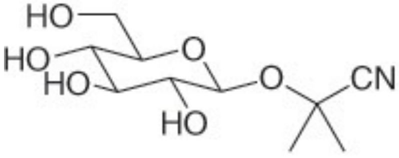


- The cyano group of a cyanohydrin is readily hydrolyzed to a carboxy group by heating with aqueous acid or base.

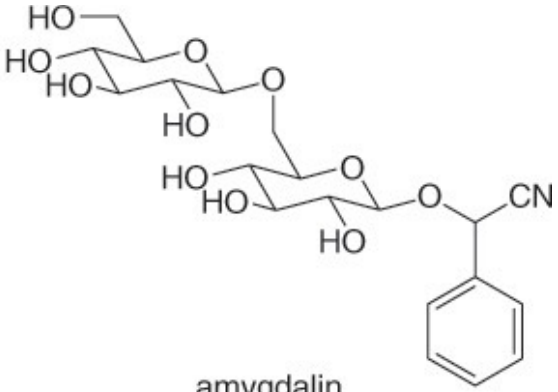


Nucleophilic Addition of CN^-

- Linamarin and Amygdalin are two naturally occurring cyanohydrin derivatives.



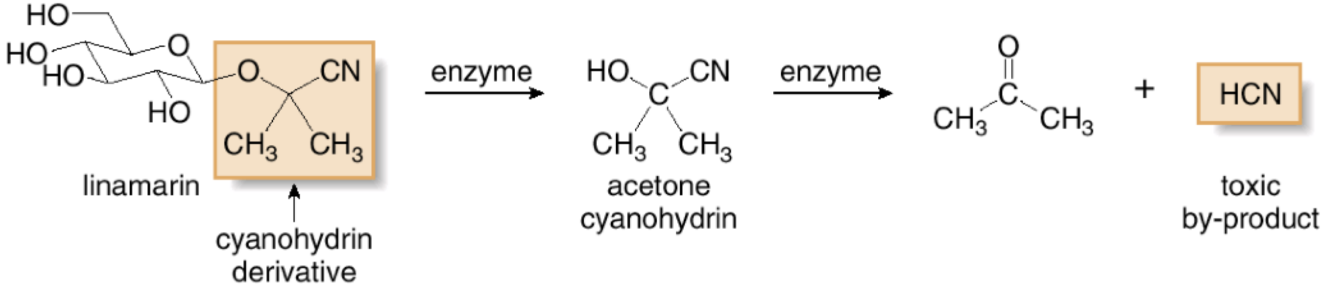
linamarin
(found in cassava root)



amygdalin
(commonly called laetrile)

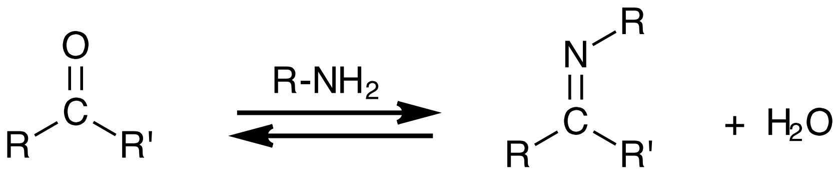
- Both compounds are toxic because they are metabolized to cyanohydrins, which are hydrolyzed to carbonyl compounds and HCN gas.

The breakdown of linamarin to HCN



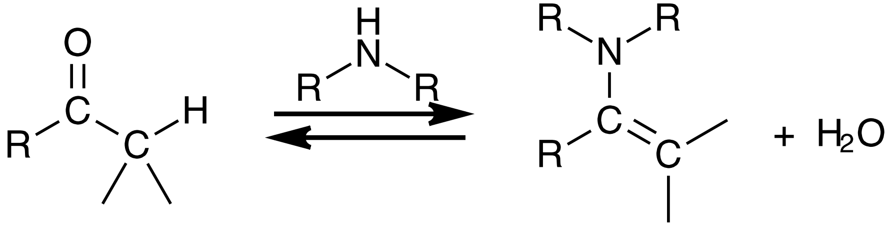
Addition of Amines

- Treatment of an aldehyde or a ketone with a 1ry amine affords an imine (also called a Schiff base).



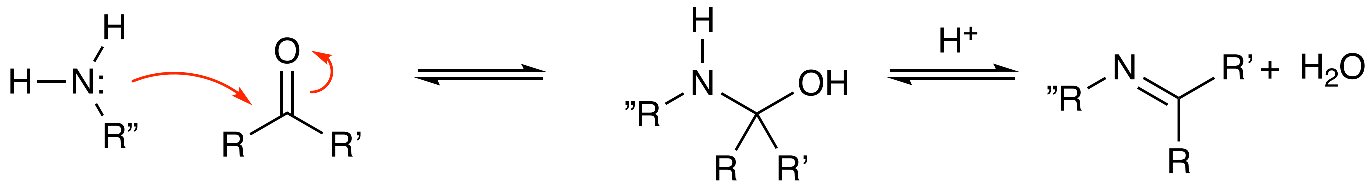
The N atom of an imine is *sp*^{2} hybridized, making the C—N—R bond angle 120° (not 180°).

- Treatment of an aldehyde or a ketone with a 2ry amine affords an enamine.



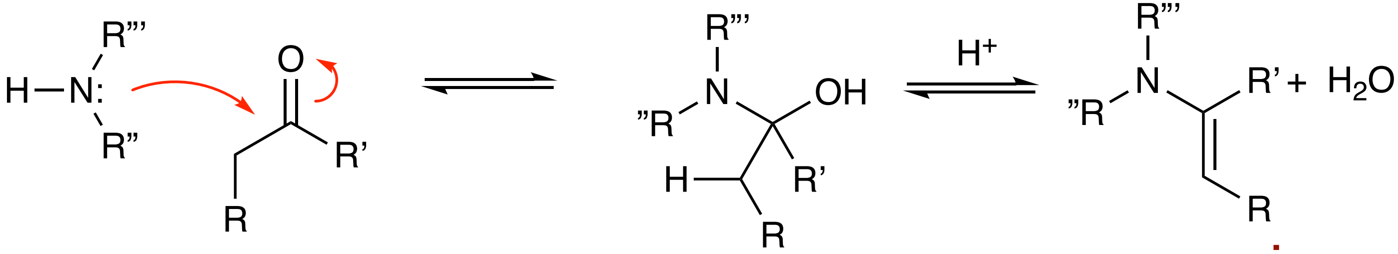
Imines and Enamines: Simplified Mechanism

Primary amines:



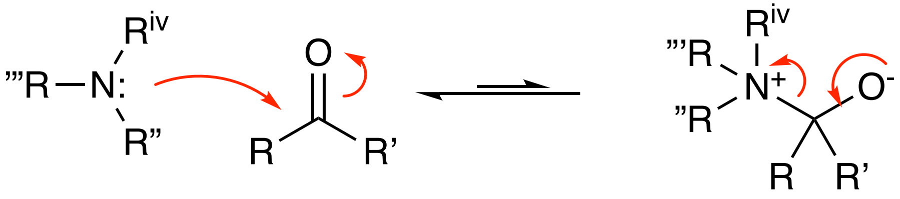
imine

Secondary amines:



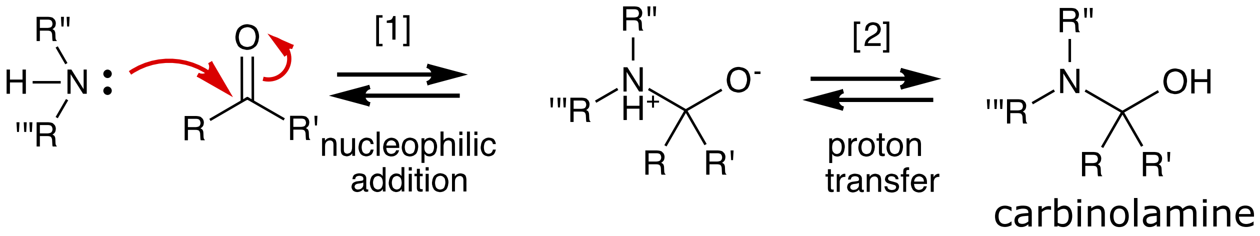
enamine

Tertiary amines:

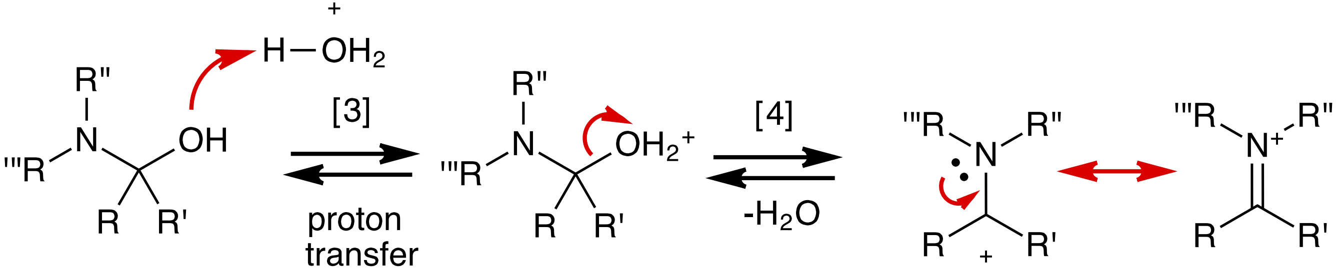


Imines and Enamines: Complete Mechanism

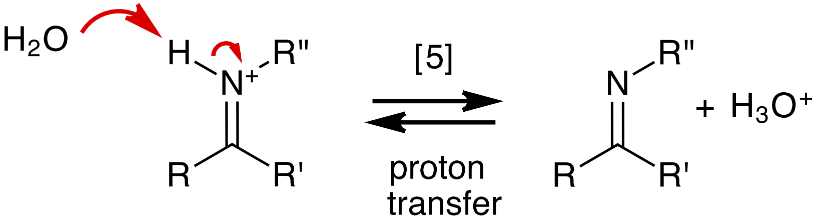
1. Amine addition



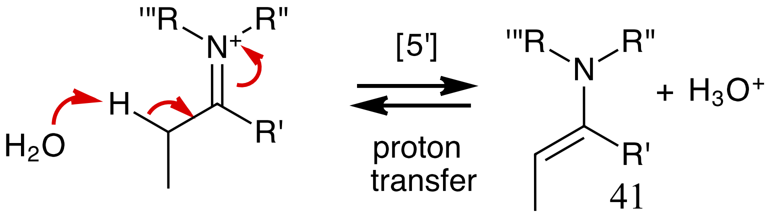
2. Elimination of water (E1)



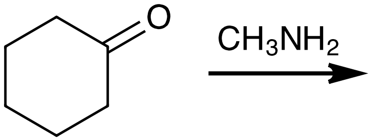
a: NH_3, $R''NH_2$

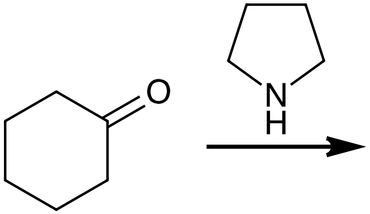


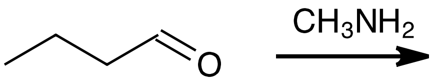
b: $R''R'''NH$

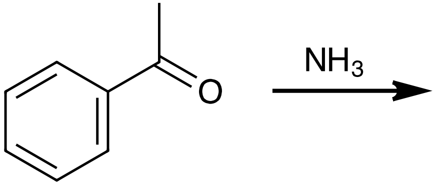


Addition of Amines





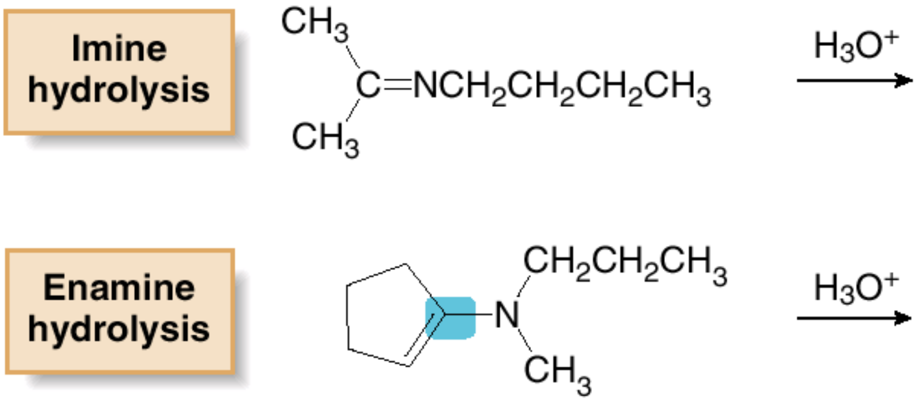




Imine and Enamine Hydrolysis

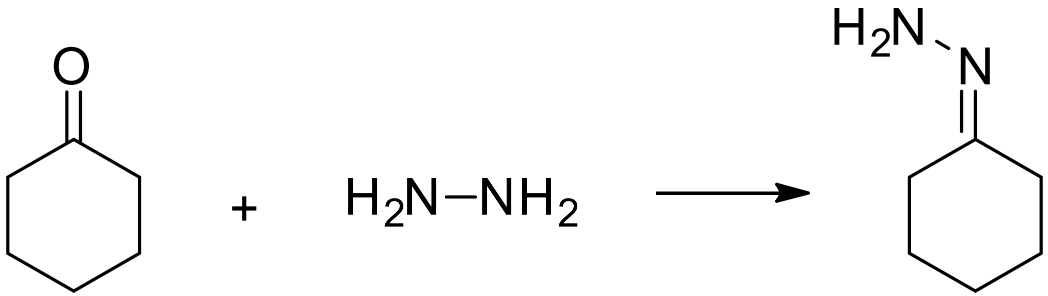
- Because imines and enamines are formed by a reversible set of reactions, both can be converted back to carbonyl compounds by hydrolysis with mild acid.
- The mechanism of hydrolysis is the exact reverse of the mechanism written for formation of imines and enamines.

- **Hydrolysis of imines and enamines forms aldehydes and ketones.**



$$(CH_3)_2C=NCH_2CH_2CH_2CH_3 \xrightarrow{H_3O^+}$$

Other Amines



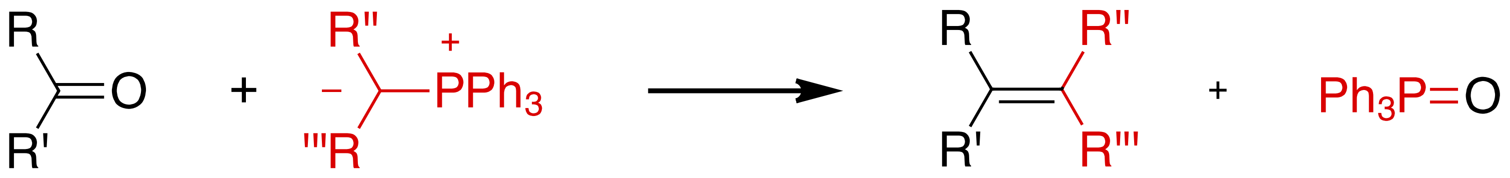
hydrazine

hydrazone

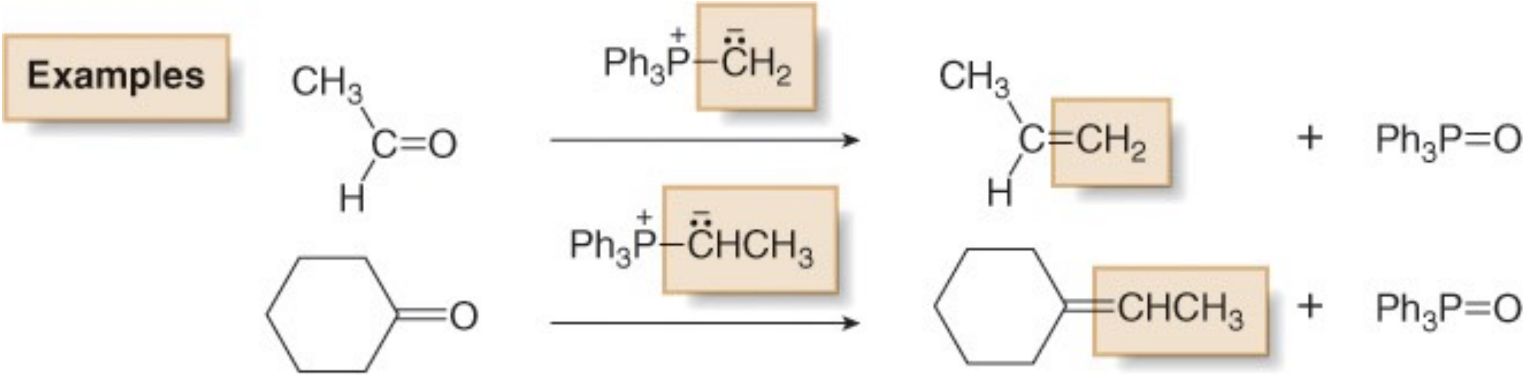
hydroxylamine

oxime

The Wittig Reaction

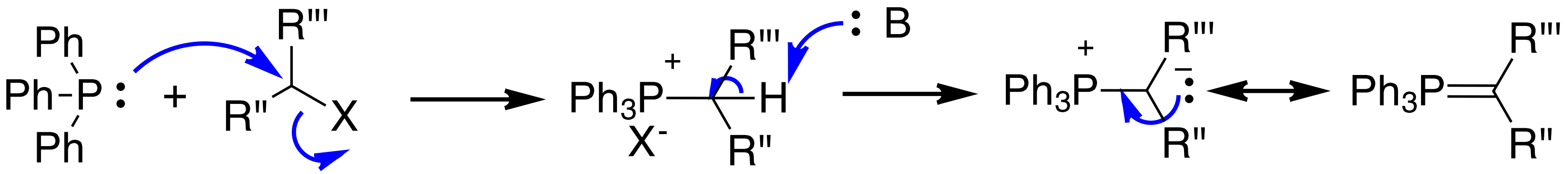


phosphonium ylids
(Wittig reagents)



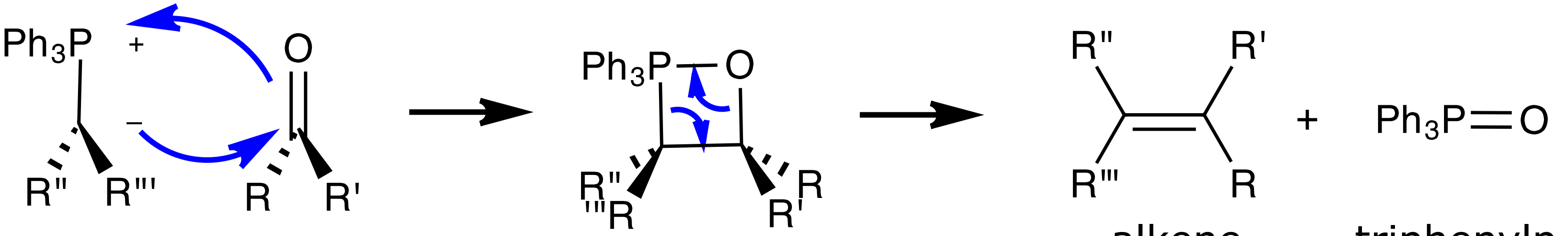
The Wittig Reaction

Preparation of phosphorus ylides (phosphoranes)   B: nBuLi, $NaNH_2$, NaH



triphenylphosphine   phosphonium salt   phosphonium ylide

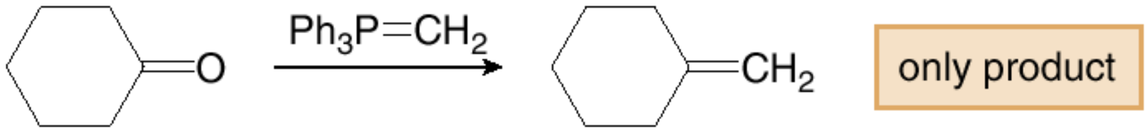
Reaction of phosphorus ylides with carbonyl compounds



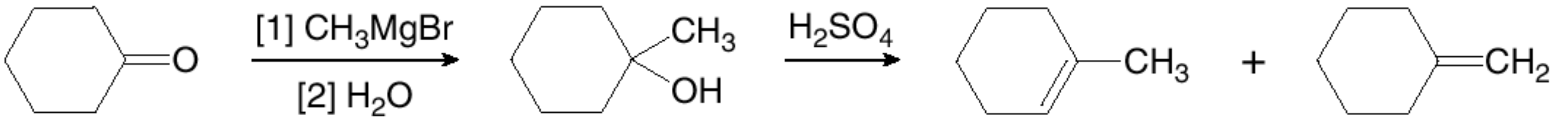
alkene   triphenylphosphine oxide

The Wittig Reaction

Advantage: the Wittig reaction always gives a single constitutional isomer.



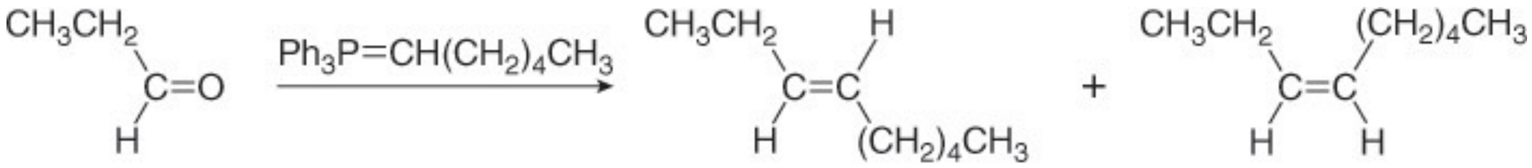
cyclohexanone **B**



B

cyclohexanone 3° alcohol trisubstituted C=C **major product** disubstituted C=C **minor product**

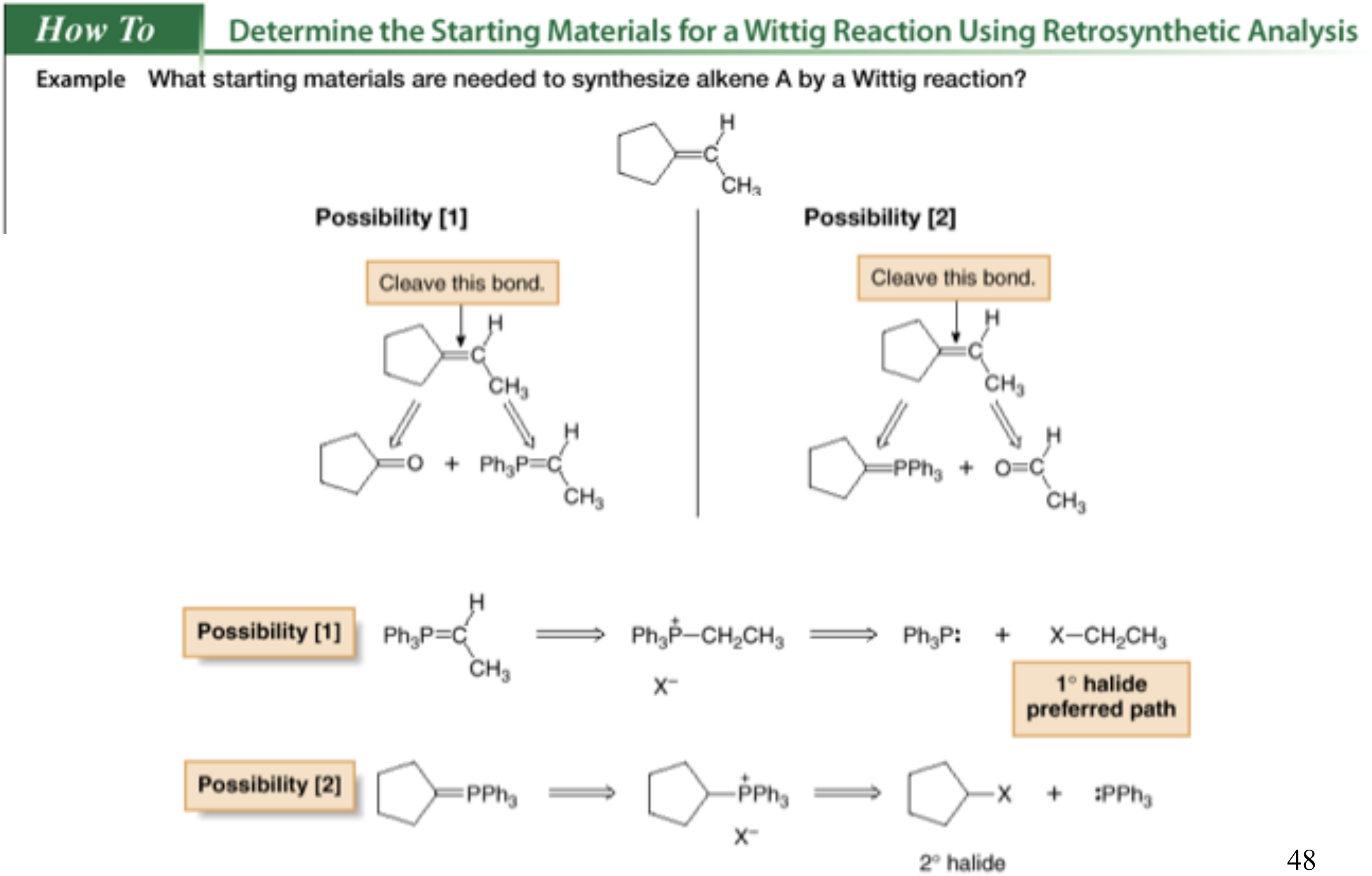
Limitation: a mixture of stereoisomers is sometimes formed.



E isomer
59%

Z isomer
41%

The Wittig Reaction



How To Determine the Starting Materials for a Wittig Reaction Using Retrosynthetic Analysis

Example What starting materials are needed to synthesize alkene A by a Wittig reaction?

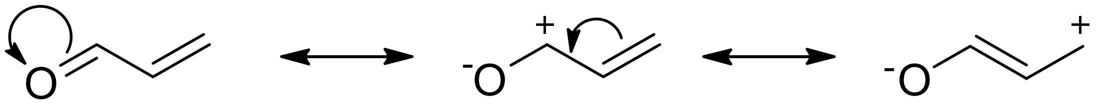
Possibility [1]

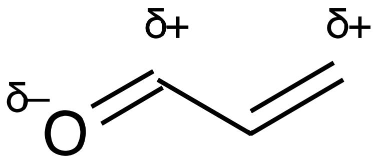
Cleave this bond.

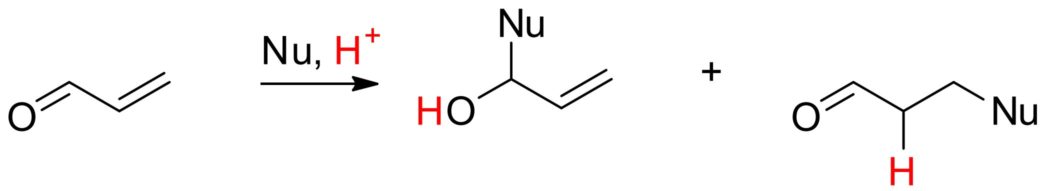
Possibility [2]

Cleave this bond.

α,β-Unsaturated Carbonyl Compounds





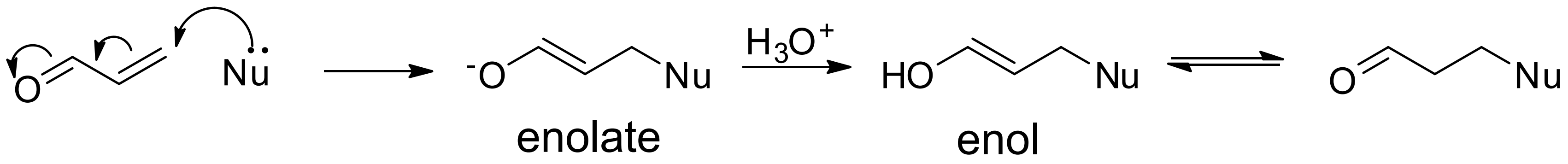


1,2-addition

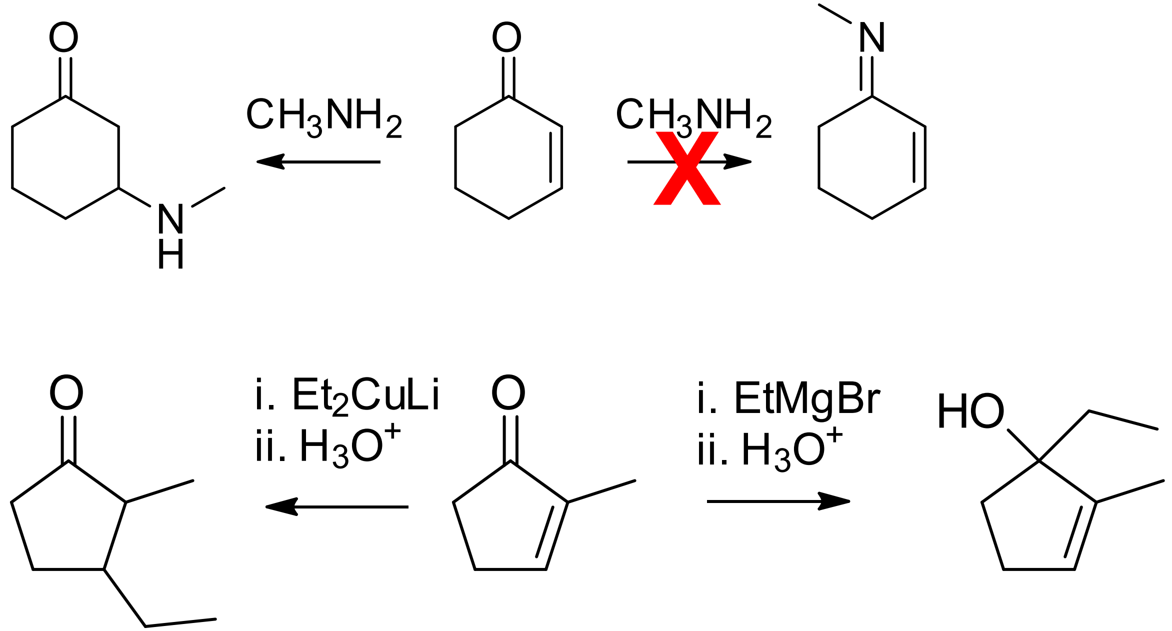
1,4-addition
(conjugate addition)

Conjugate Addition

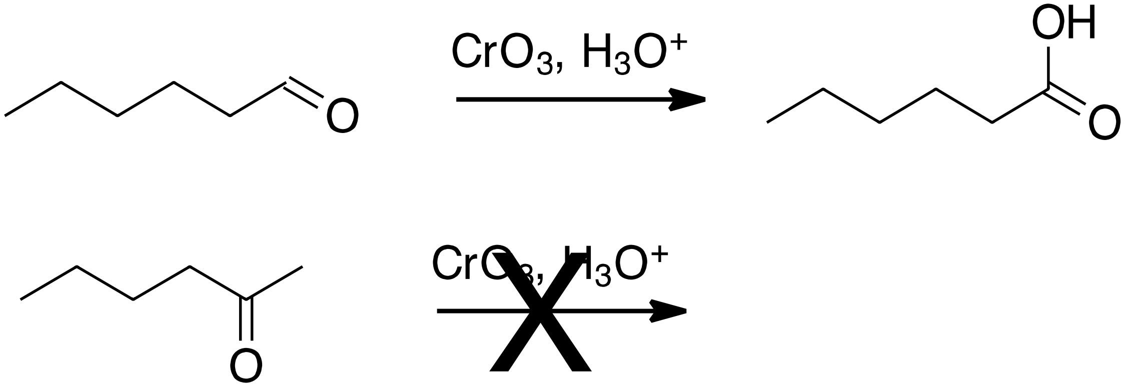
general mechanism:



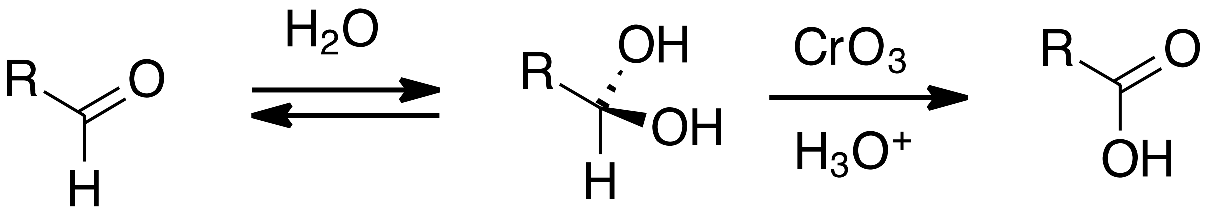
examples:



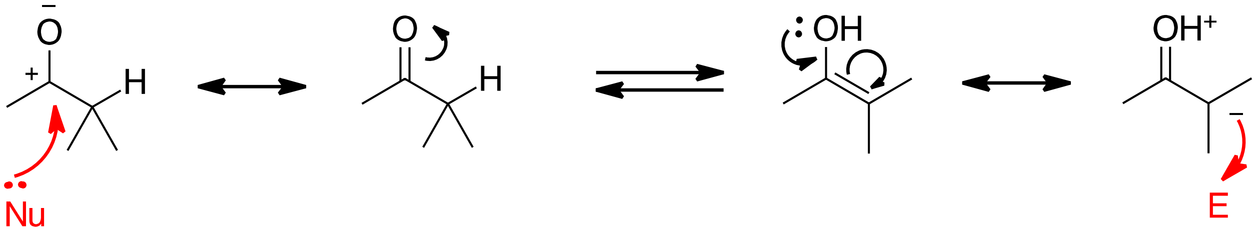
Oxidation



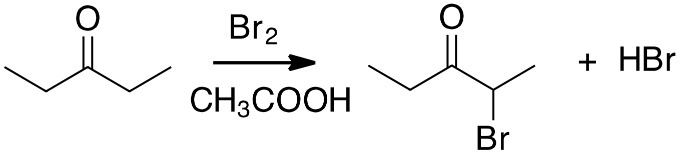
mechanism



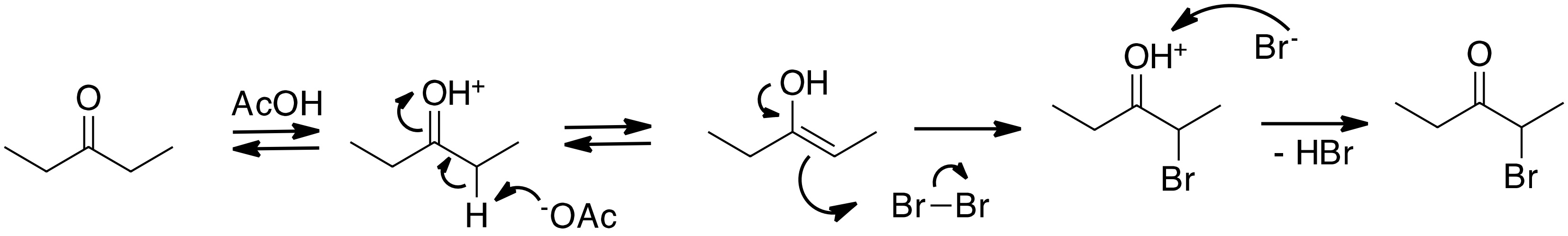
Keto-Enol Tautomerism. Reaction at the α-Carbon



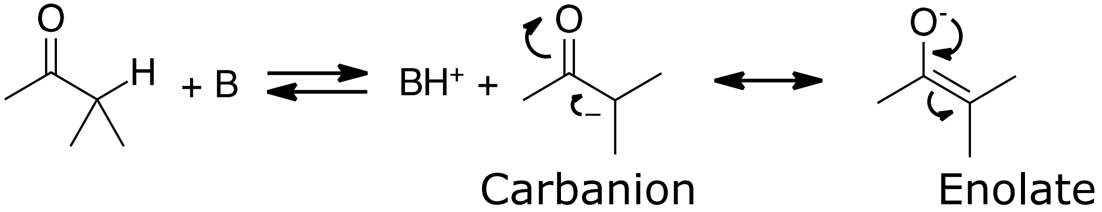
α-halogenation:



mechanism:



Enolates. Reaction at the α-Carbon



pKa ≈ 20

Example: 2-methylcyclohexanone