TÉCNICO LISBOA

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Metal-Catalyzed Polymerization

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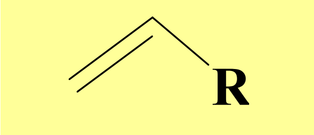
Metal-Catalyzed Polymerization

(Coordination Polymerization)

- *Polymerization of Olefins (Insertion)* ✓ ***Prof. Barbara Milani***
- ***Polymerization of Dienes (Insertion)***
- ***Polymerization of Alkynes***
- ***Ring Opening Metathesis Polymerization (ROMP)***
- ***Classical Anionic Polymerization***
- ***Ring Opening Polymerization (ROP)***
- ***Metal-mediated Radical Polymerization***

POLYMERIZATION OF VINYL MONOMERS



TYPES OF POLYMERIZATION

- **RADICAL** — initiators: ROOR, ROOH, R-N≡N-R
- **CATIONIC** — initiators: Brönsted acids, Lewis acids, stable cations
- **ANIONIC** — initiators: alkyl or aryl lithium or sodium compounds, sodium
- **COORDINATION** — catalysts: Ziegler-Natta, metallocenes, post-metallocenic (the only method that homo- and copolymerizes propylene and α-olefins)

GENERAL MECHANISM

Chain Growth

- **INITIATION**
- **PROPAGATION**

→ common to all types

- **CHAIN TRANSFER**
- **TERMINATION**

→ they can be suppressed when the polymerization is "controlled"

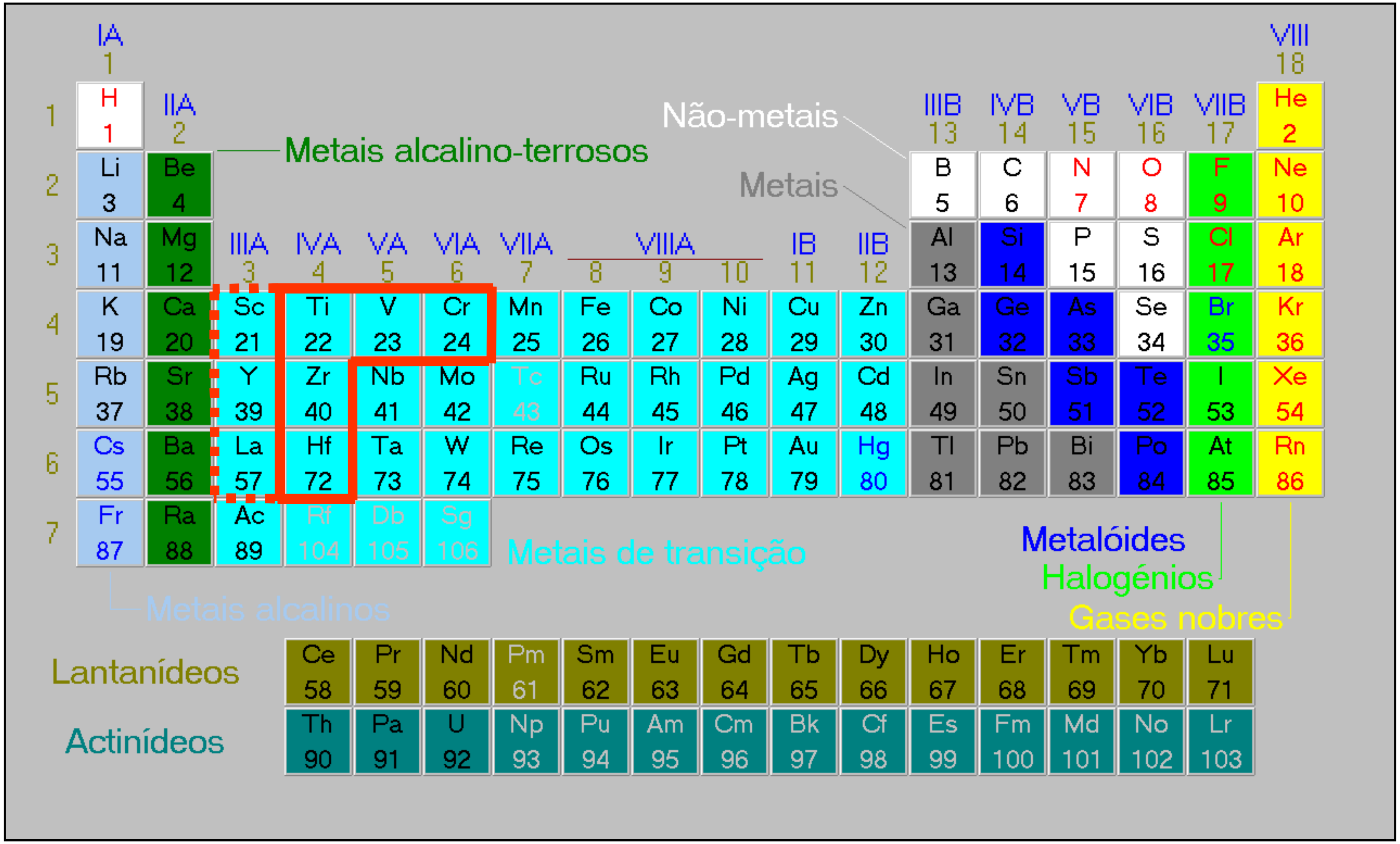
↓

"Living" polymerization

ZIEGLER-NATTA CATALYSTS

DEFINITION:

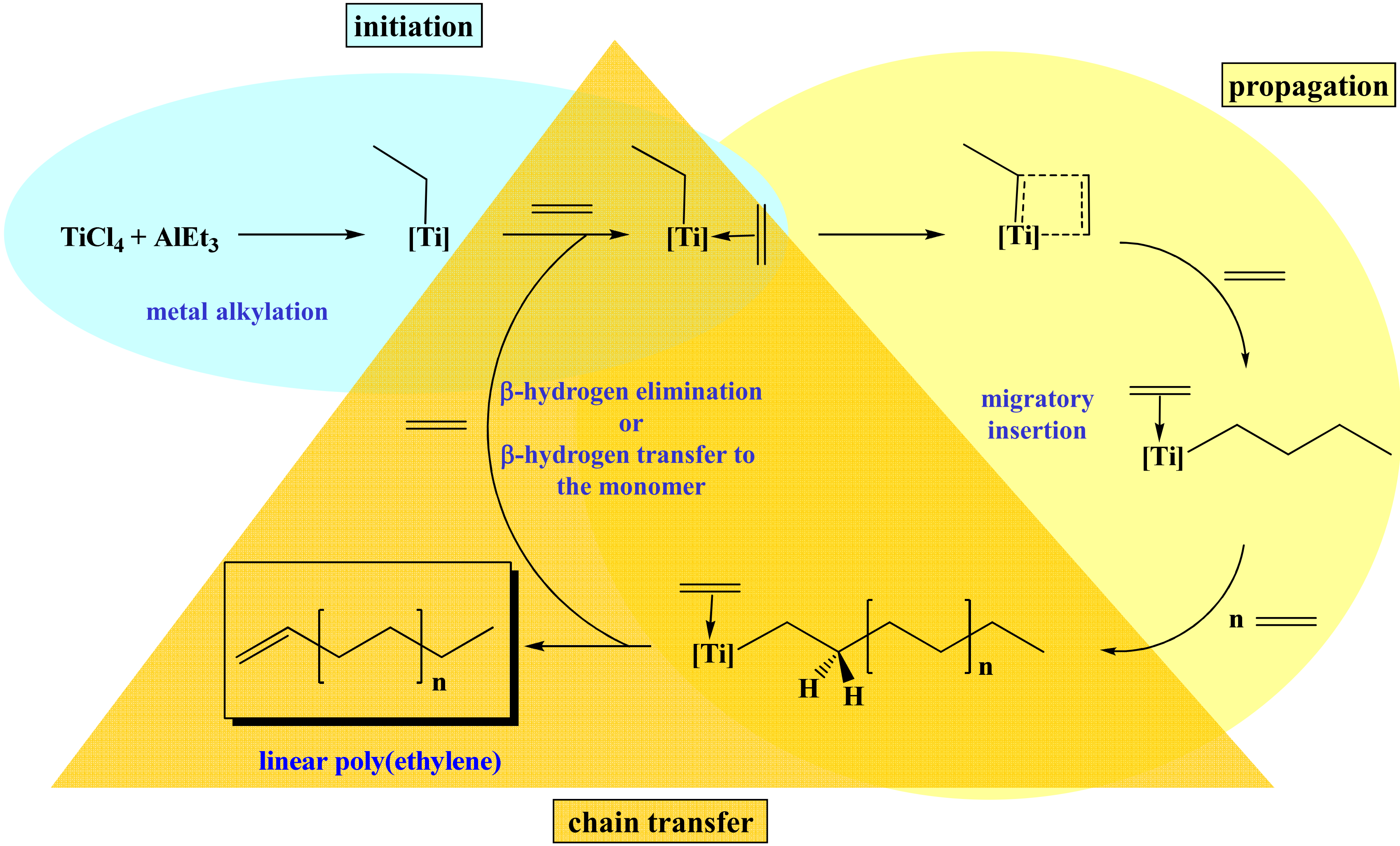
Transition Metal Compound (groups 4 - 10)	+	Metal Alkyl, Aryl or Hydride (groups 1, 2, 13 and 14)



IA
1
IIA
2
Metais alcalino-terrosos
Não-metais
Metais
IIIA IVA VA VIA VIIA VIIIA IB IIB
3 4 5 6 7 8 9 10 11 12
IIIB IVB VB VIB VIIB VIII
13 14 15 16 17 18
H 1
Li 3 Be 4
Na 11 Mg 12
K 19 Ca 20 Sc 21 Ti 22 V 23 Cr 24 Mn 25 Fe 26 Co 27 Ni 28 Cu 29 Zn 30
Rb 37 Sr 38 Y 39 Zr 40 Nb 41 Mo 42 Tc 43 Ru 44 Rh 45 Pd 46 Ag 47 Cd 48
Cs 55 Ba 56 La 57 Hf 72 Ta 73 W 74 Re 75 Os 76 Ir 77 Pt 78 Au 79 Hg 80
Fr 87 Ra 88 Ac 89 Rf 104 Db 105 Sg 106
B 5 C 6 N 7 O 8 F 9 Ne 10 He 2
Al 13 Si 14 P 15 S 16 Cl 17 Ar 18
Ga 31 Ge 32 As 33 Se 34 Br 35 Kr 36
In 49 Sn 50 Sb 51 Te 52 I 53 Xe 54
Tl 81 Pb 82 Bi 83 Po 84 At 85 Rn 86
Metais de transição
Metalóides
Halogénios
Gases nobres
Metais alcalinos
Lantanídeos
Ce 58 Pr 59 Nd 60 Pm 61 Sm 62 Eu 63 Gd 64 Tb 65 Dy 66 Ho 67 Er 68 Tm 69 Yb 70 Lu 71
Actinídeos
Th 90 Pa 91 U 92 Np 93 Pu 94 Am 95 Cm 96 Bk 97 Cf 98 Es 99 Fm 100 Md 101 No 102 Lr 103

By analogy with homogeneous metallocene catalysts

Mechanism of Olefin Polymerization with Ziegler-Natta Catalysts



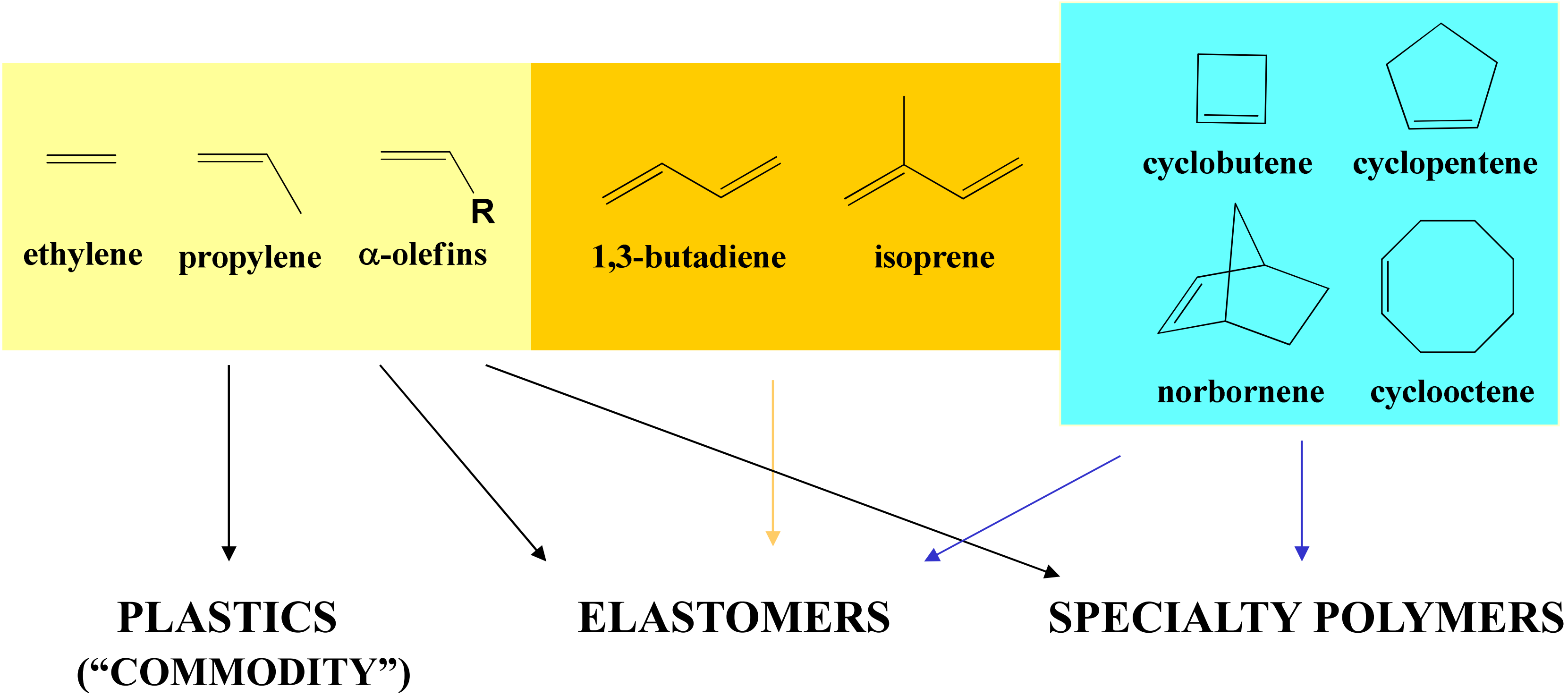
[Ti] ≡ heterogeneous metal site (it can also be a metallocene or a post-metallocene)

POLYOLEFINS

Olefin = unsaturated hydrocarbon = Alkene

Typical Olefin Monomers:

Aliphatic monoolefins (C_nH_{2n}): **Aliphatic diolefins (C_nH_{2n-2}):** **Cycloolefins (C_nH_{2n-2}):**



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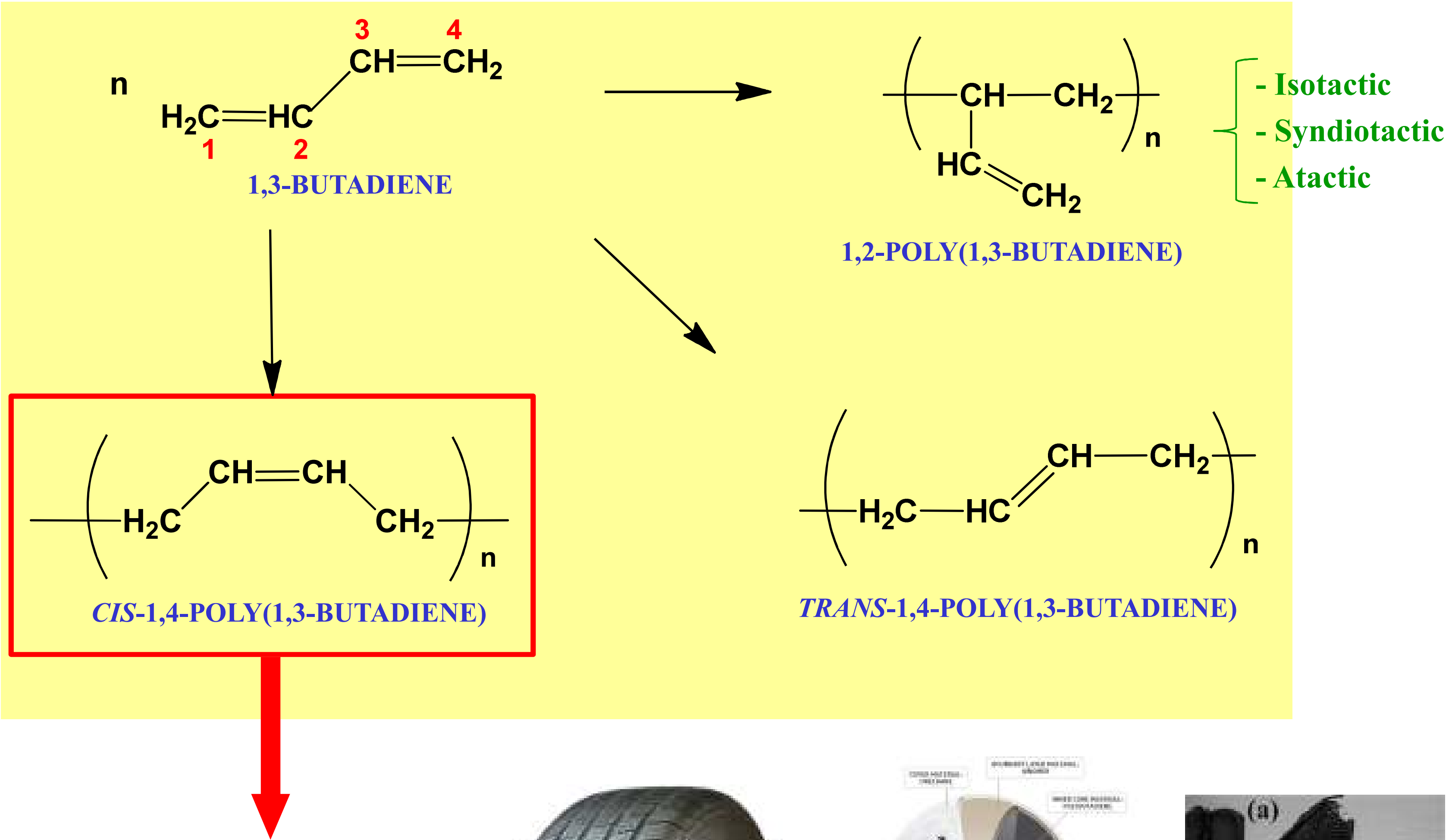
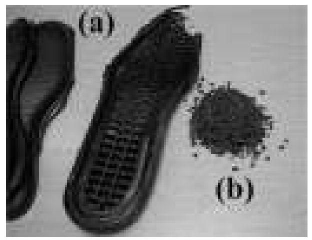
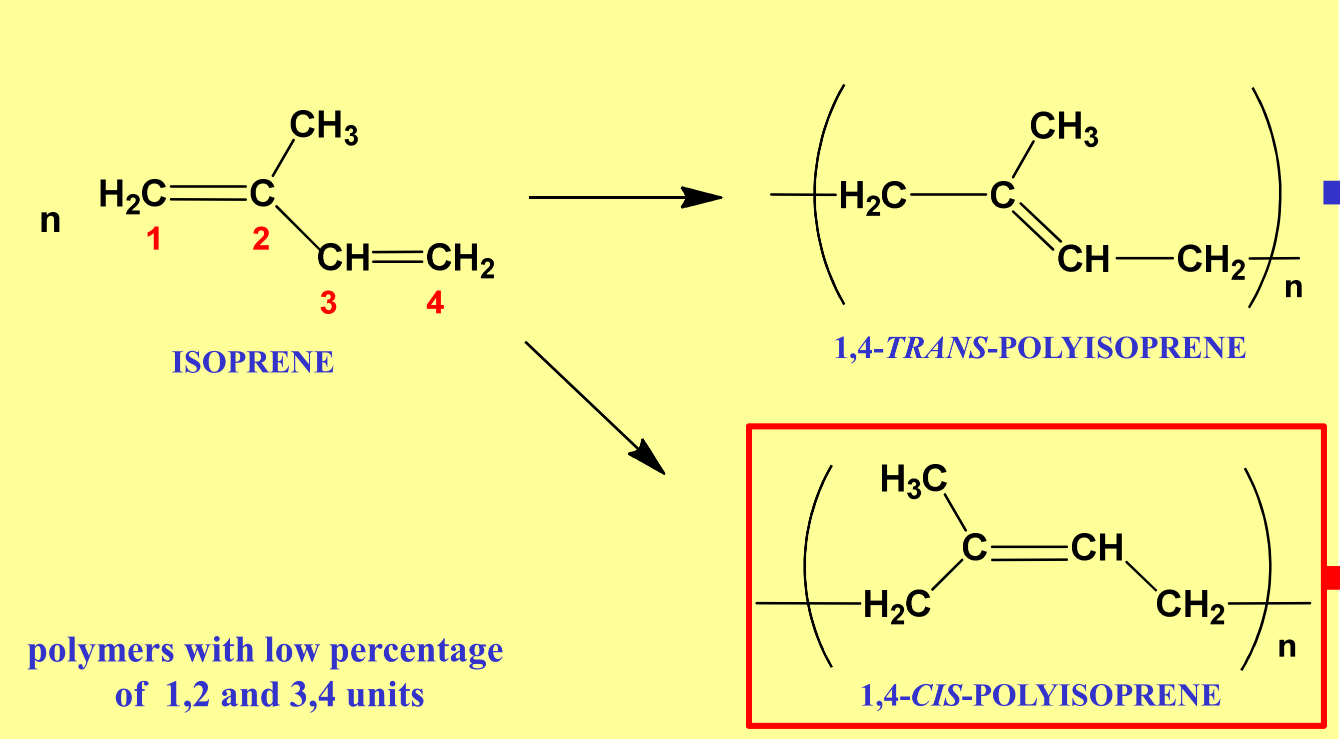
POLYMERIZATION OF DIENES



SYNTHETIC RUBBER





From natural sources:
GUTTA-PERCHA

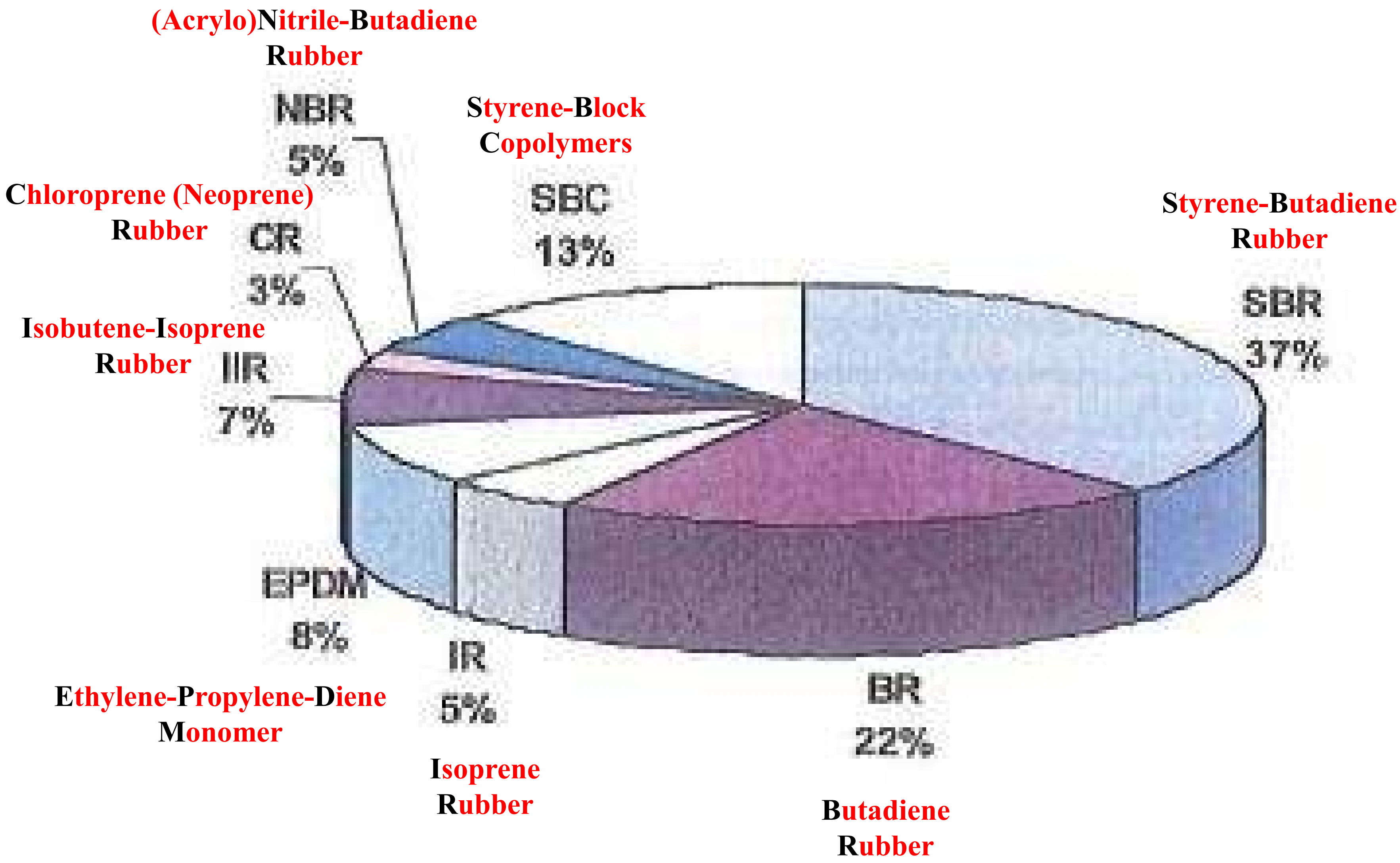
Hard material:
- **Insulator**
- **Old golf balls**

From natural sources:
NATURAL RUBBER

- ***Hevea Brasiliensis***
- ***Guayule***

SYNTHETIC NATURAL RUBBER

SYNTHETIC RUBBER MARKET



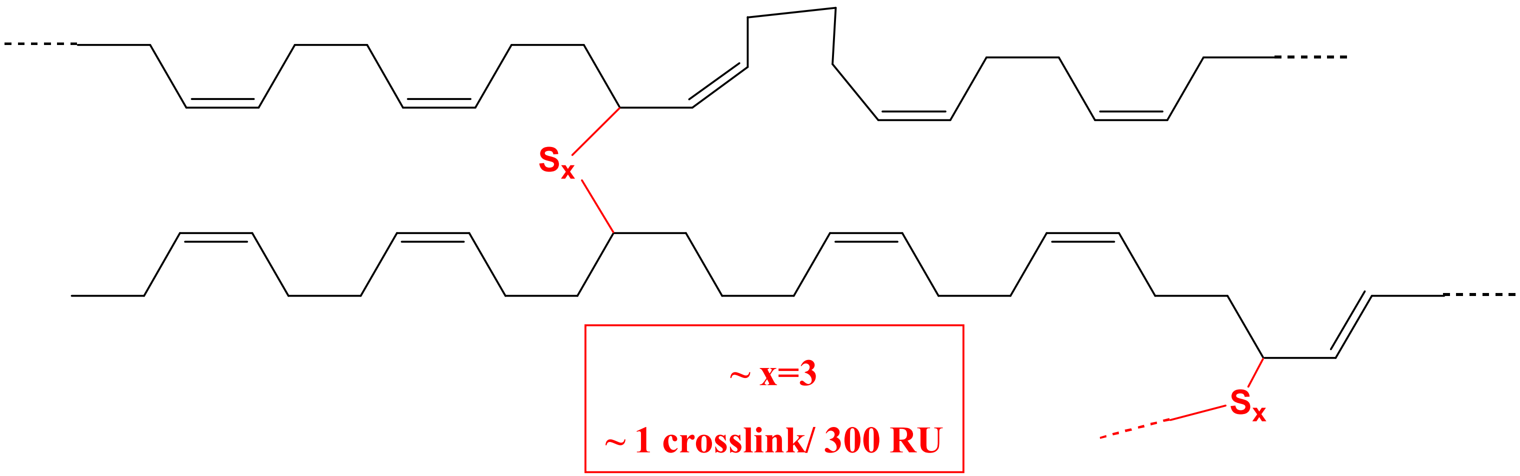
STEREOREGULAR ELASTOMERS

1,4-*CIS*-POLYBUTADIENE	1,4-*CIS*-POLYISOPRENE
~3.2 Mton	~0.75 Mton

Production depends on the NATURAL RUBBER market

~10.3 Mton

S_8 ↓ **Vulcanization**



MACROMOLECULAR NETWORK

PolyButadiene Rubber World Producers

Company	Annual Capacity (thousand of metric tons)	% of World Capacity
LANXESS	488	15,4
Sinopec	**390**	**12,3**
Goodyear	265	8,4
Korea Kumho Petrochemicals	222	7,0
UBE Industries	173	5,5
PetroChina	**160**	**5,1**
Polimeri Europa	**160**	**5,1**
Firestone Polymers	150	4,7
Voronezhsynthiezkauchuk	141	4,5
others	1010	32,0
Total	3159	100

North America	Central and South America	Western Europe	Central and Western Europe	Asia*	Middle East/ Africa/Oceania	*Total*
755	93	355	421	**1520**	45	3159

*China, Taiwan, India, Japan, Korea, Thailandia

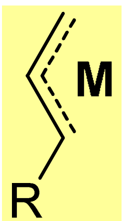
STEREOSPECIFIC CATALYSTS - POLYBUTADIENE

• ZIEGLER-NATTA CATALYSTS

High content (>90%) in:

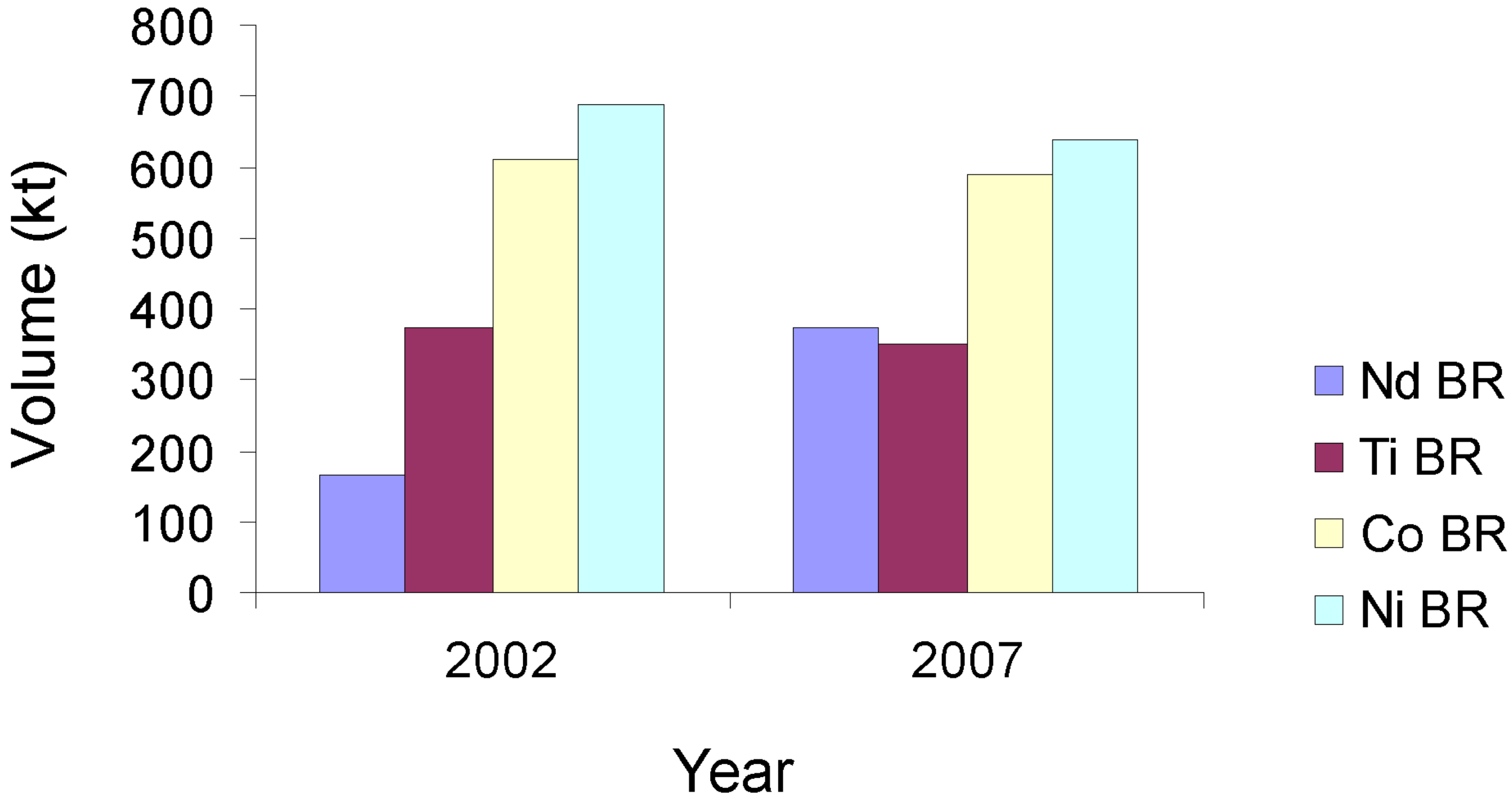
1,4-cis	1,4-*trans*	1,2 (syndiotactic)
TiI_4 + Al(*i*Bu)$_3$ (1:4-5), 30 ºC	γ-$TiCl_3$ + $AlEt_3$	Ti(OR)$_4$ + $AlEt_3$ (1:7), 15 ºC
$CoCl_2$ + $Al_2Cl_3Et_3$ (1:1000), 5 ºC	VCl_3 + $AlEt_3$ (1:2), 15 ºC	V(acac)$_3$ + $AlEt_3$ (1:6-10), 15°C
Co(acac)$_2$ + $AlEt_2Cl$ + H_2O (branched polymer)	VCl_4 + $AlEt_3$ (1:1.8), 15 ºC	Cr(C_6H_5CN) + $AlEt_3$ (1:2) Cr(C_6H_5CN) + $AlEt_3$ (1:10) (isot.)
Ni(octanoate)$_2$ + $AlEt_3$ + BF_3·OEt_2 (1:17:15), 50 ºC	V(acac)$_3$ + $AlEt_2Cl$ + Cl_3CCO_2H, 80 ºC	Co(acac)$_3$ + $AlEt_3$ (1:50), 16°C
U(OR)$_4$ + $AlEt_2Cl$	$VOCl_3$ + $AlEt_3$	Co(acac)$_3$ + $AlEt_3$ + H_2O + CS_2
Nd(neodecanoate)$_3$ + AlR_2Cl + AlR_3 , 60 ºC	V(acac)$_3$ + MAO (1:1000)	

• ALLYL METAL CATALYSTS (without Alkyl Aluminium cocatalyst)

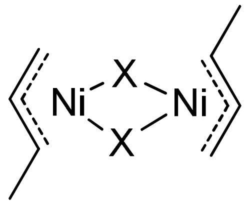


M= Cr, Co, Nb, W, Rh, U, Ni

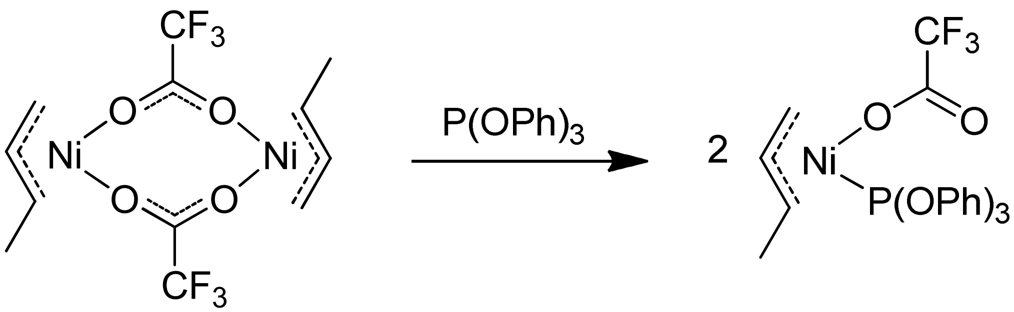
Catalysts for high *cis*-1,4 Polybutadiene



EXAMPLES OF ALLYL NICKEL STEREOSPECIFIC CATALYSTS



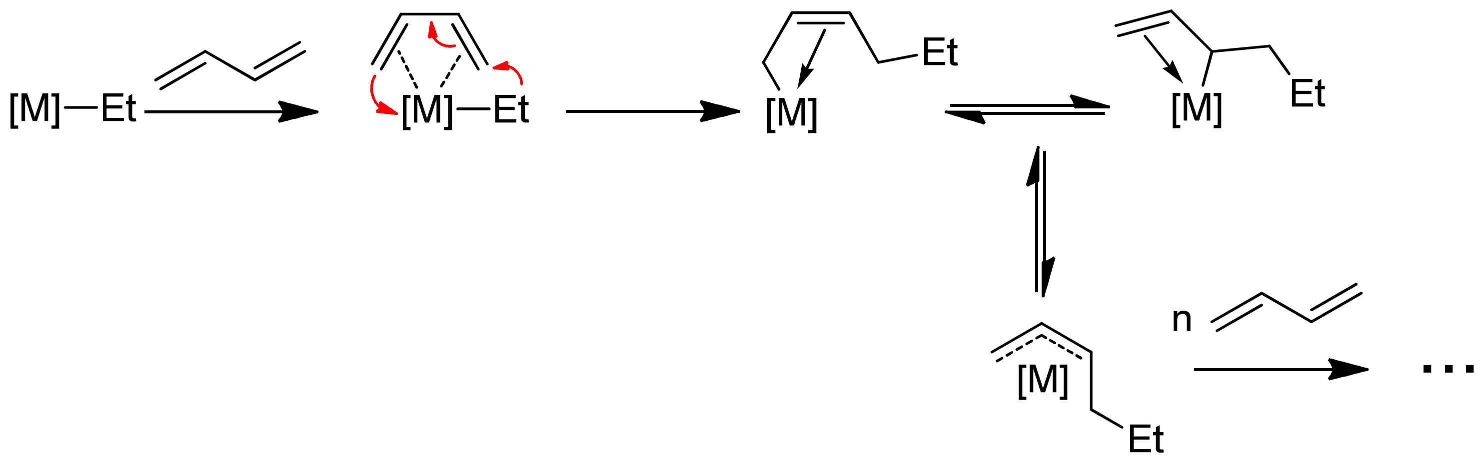
X	1,4-*cis*	1,4-*trans*	1,2
Cl	92	6	2
Br	72	25	3
I	0	97	3



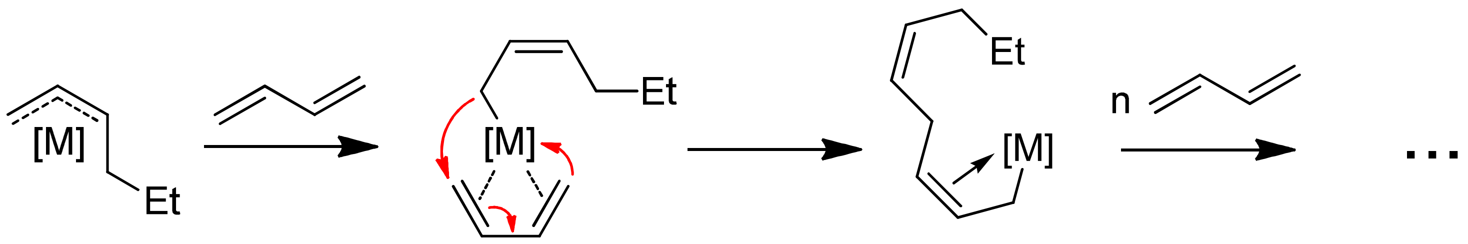
1,4-*cis*	97	0
1,4-*trans*	2	96
1,2	1	4

STEREOREGULATION MECHANISM

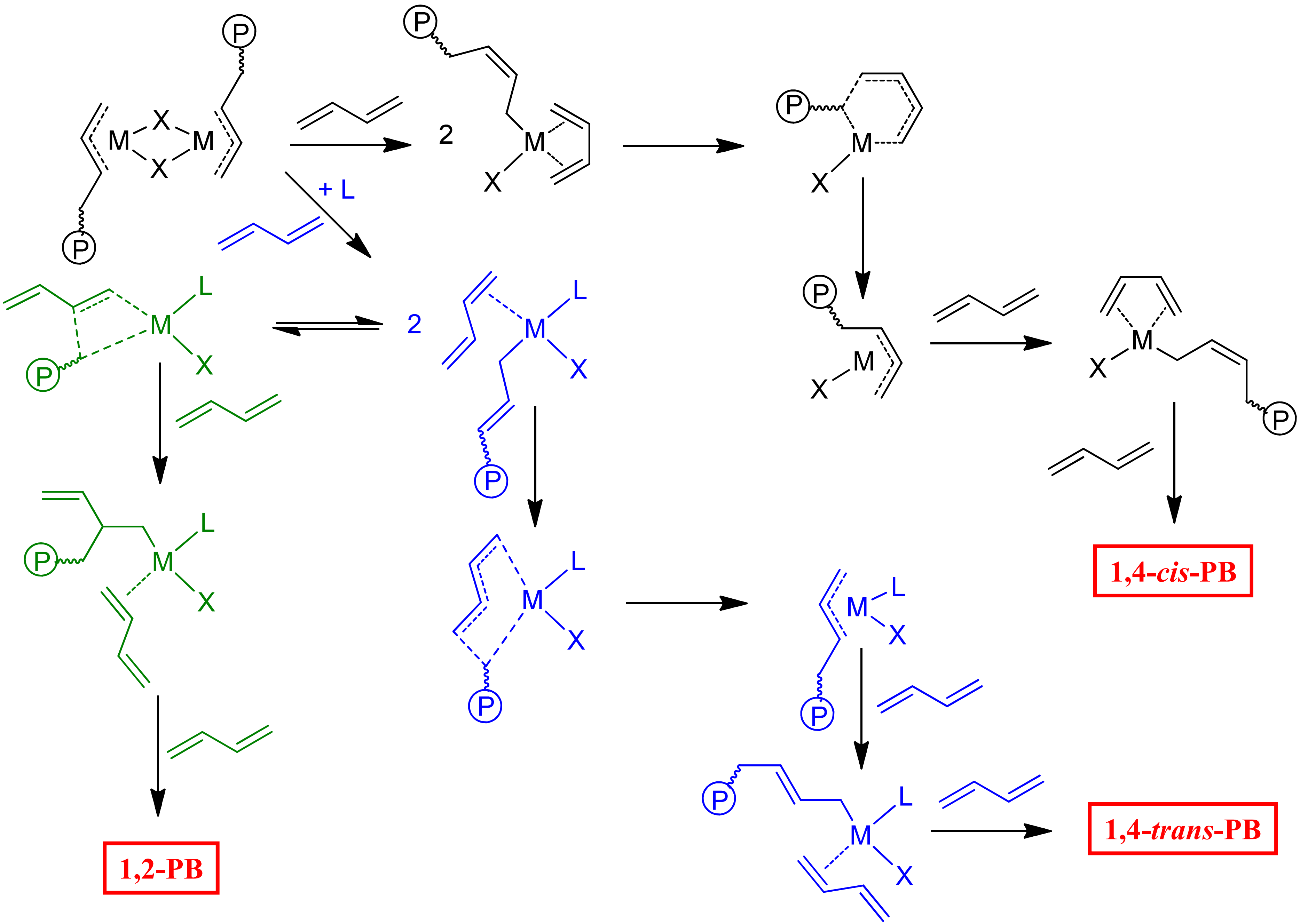
• INITIATION (Z-N CATALYSTS)



• PROPAGATION



STEREOREGULATION MECHANISM



STEREOSPECIFIC CATALYSTS - POLYISOPRENE

- **ZIEGLER-NATTA CATALYSTS**

High content (>90%) in:		
1,4-cis	**1,4-*trans***	**3,4**
$TiCl_4 + AlEt_3$ (Al/Ti > 1)	$TiCl_4 + AlEt_3$ (Al/Ti < 1)	$Ti(OR)_4 + AlEt_3$
	α-$TiCl_3 + AlR_3$	
	$VCl_3 + AlEt_3$	

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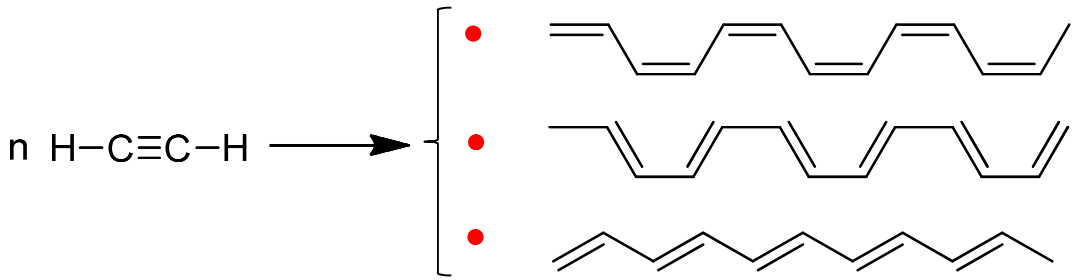
POLYMERIZATION OF ALKYNES

(POLYMERIZATION OF ACETYLENE)



- cis-transoid
- trans-cisoid
- trans-transoid

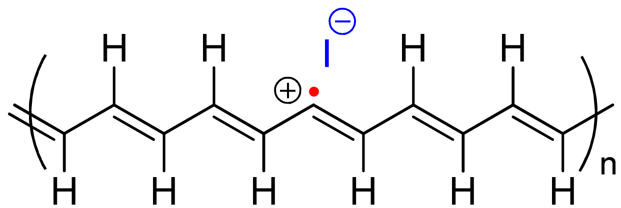
cis
insulator
($\sigma = 4\times10^{-9}\ \Omega^{-1}\ cm^{-1}$)

trans
semiconductor
($\sigma = 9\times10^{-5}\ \Omega^{-1}\ cm^{-1}$)

doped with oxidants (1977)
(ex: I_2)

conductor
$\sigma = 1.5\times10^{3}\ \Omega^{-1}\ cm^{-1}$

Au ($\sigma = 2.4\times10^{4}$)
Cu ($\sigma = 7.2\times10^{4}$)



CONDUCTION
(?)

- Intermolecular
- Interchain
- Intergrain

POLYACETYLENE

- **Insoluble (in organic solvents)**
- **Unstable in air (double bonds oxidation)**
- **Bad mechanical properties (poorly processable)**

the objective was to make conducting films

● SYNTHESIS - CATALYSTS

- $Ti(OBu)_4 + AlEt_3$ (Natta, 1958)

- $Ti(OBu)_4 + AlEt_3$ (Al/Ti ~ 4) (Shirakawa, 1974) **good films when $[Ti]<10^{-3}$ M**

- $Ti(OBu)_4$ + LiBu (Li/Ti ~ 2) **high *trans* %**

- $MoCl_5 + SnPh_4$

- $WCl_6 + SnPh_4$

- $NiX_2(PR_3)_2$ (X= Cl, Br, I) **high *trans* %**

POLYACETYLENE

- **BASF METHOD** (best commercial polyacetylene)

- $Ti(OBu)_4 + AlEt_3$ (Al/Ti ~ 4)
- $Ti(OBu)_4$ + LiBu (Li/Ti ~ 2)

SOLVENT:

very viscous silicone

↓

Polymers with lower content of *sp*$_3$ defects

↓

Better morphological properties (better oriented fibres)

Films produced on HDPE or PP supports

↓ Stretching 7× the original length

Highly oriented transparent films

↓ Doping with I_2

Ti/Al catalyst: $\sigma = 2\times10^4\ \Omega^{-1}\ cm^{-1}$ (20 μm film)

$\sigma = 8\times10^3\ \Omega^{-1}$ cm-1 (0.1 μm film)

Ti/Li catalyst: $\sigma \sim 10^5\ \Omega^{-1}\ cm^{-1}$

BASF abandoned this process with the appearance of new conducting polymers (more stable and processable)

POLYMERIZATION OF SUBSTITUTED ACETYLENES

H−C≡C−R **or** R−C≡C−R

POLYMERS:
- **More soluble (in organic solvents)**
- **More stable in air**
- **The majority are electric insulators**

• CATALYSTS

- **$MoCl_5 + SnPh_4$**
- **$WCl_6 + SnPh_4$**

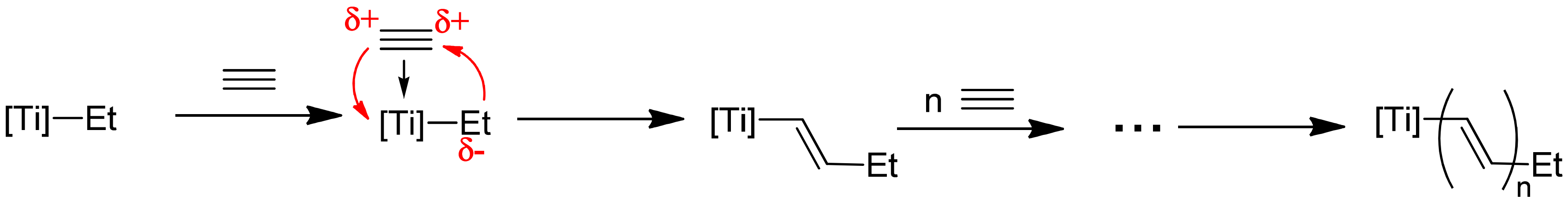
} **Small R groups**

- **$NbCl_5 + SnPh_4$**
- **$TaCl_5 + SnPh_4$**

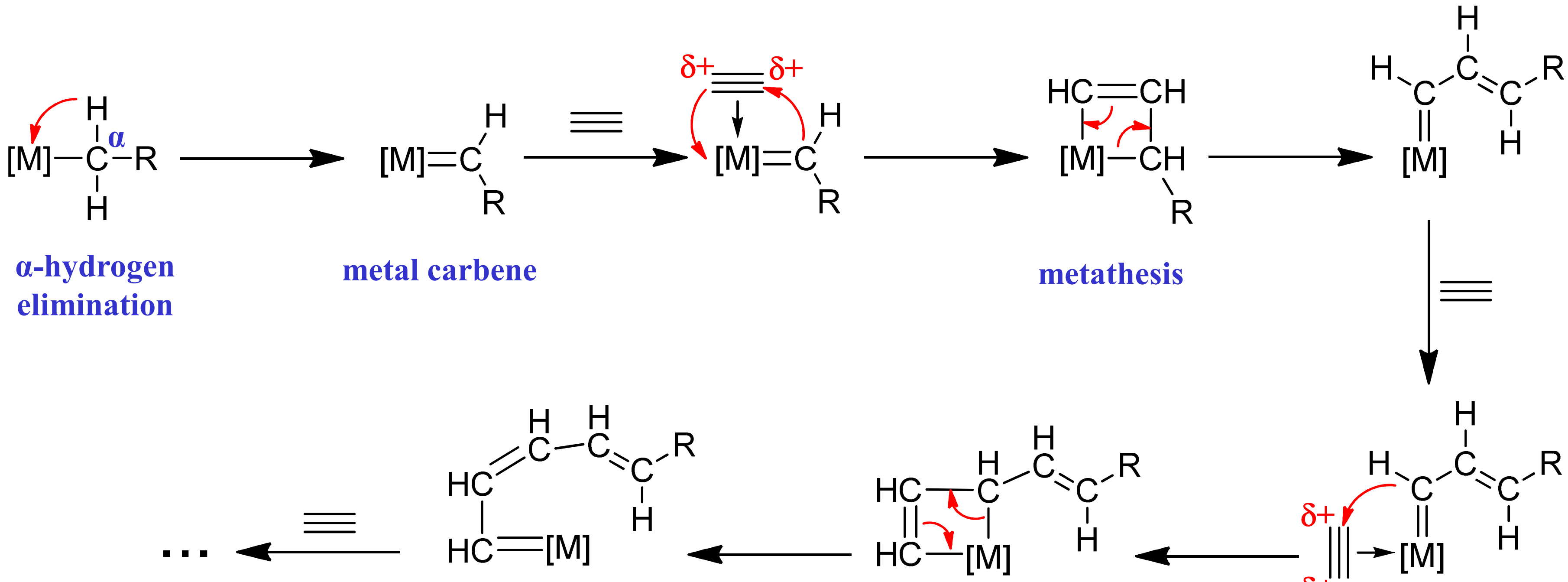
} **Bulky R groups (ex: $-SiR_3$)**

MECHANISM

• Ti ZIEGLER-NATTA CATALYSTS (insertion mechanism)



• Mo, W, Nb, Ta (Groups 5 and 6) ZIEGLER-NATTA CATALYSTS (metathesis mechanism ??)



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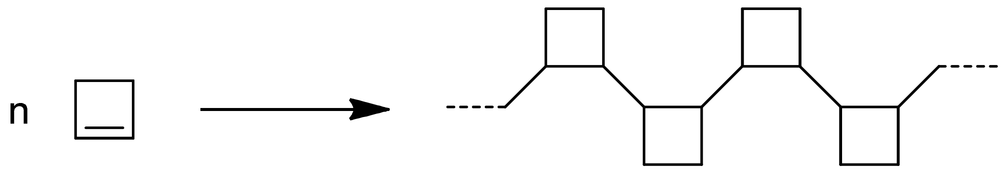
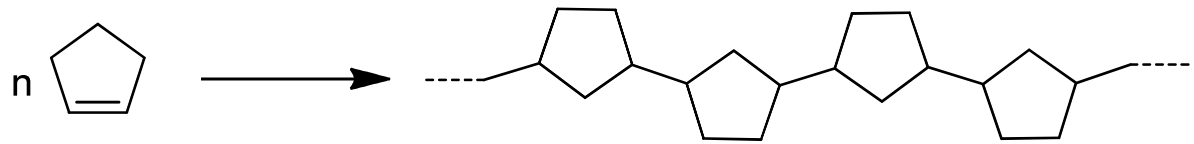
RING-OPENING METATHESIS POLYMERIZATION (ROMP)

MONOMERS: Cycloolefins and Cycloalkynes

When Ziegler-Natta or metallocene catalysts based on Group 4 metals (Ti, Zr, Hf) or post-metallocene catalysts are used the polymerization occurs by Insertion (or Vinyl-addition) Polymerization:



cyclobutene   poly(cyclobutene)



cyclopentene   poly(cyclopentene)

- **The polymers do not have double bonds in the main chain**
- **Very rigid polymers (high melting temperatures)**
- **May copolymerize with linear α-olefins (metallocene catalysts) to give amorphous copolymers**

BUT…

RING-OPENING METATHESIS POLYMERIZATION (ROMP)

When <u>Ziegler-Natta</u> catalysts based on <u>Group 6 metal</u> (Mo, W) or <u>metal carbene</u> catalysts are used the polymerization occurs by Ring-Opening Methathesis Polymerization:

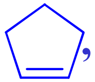
cyclopentene poly(1-pentenylene)

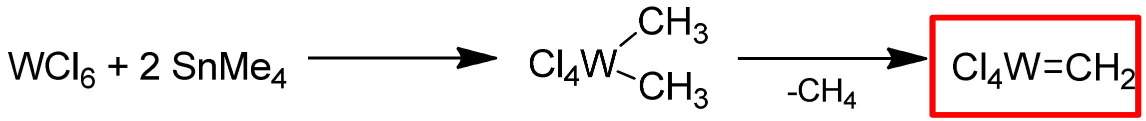
norbornene poly(cyclopentenylenevinylene)

- **The monomer double bond is retained in the polymer main chain**
- **The polymerization has a living character**
- **Easy block-copolymerization**
- **Ring tension favours ROMP (thermodynamically)**

ROMP CATALYSTS

• ZIEGLER–NATTA CATALYSTS

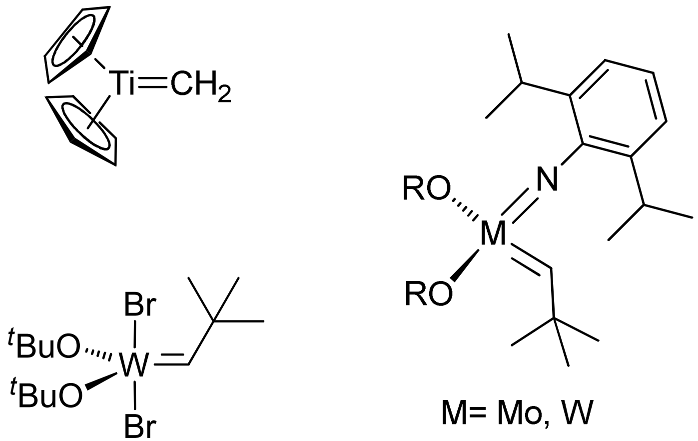
- $WCl_6 + SnMe_4$
- $WCl_6 + ZnMe_2$
- MoO_3 / γ-Al_2O_3
- $MoCl_5 + AlEt_3$ (, Natta, 1964)



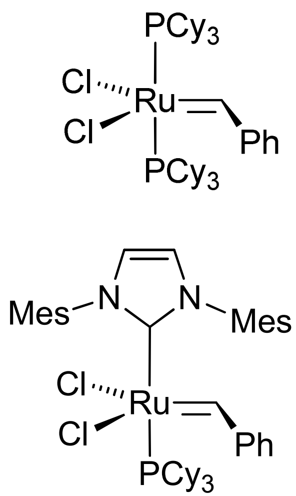
α-hydrogen elimination

carbene

• WELL-DEFINED CARBENE CATALYSTS



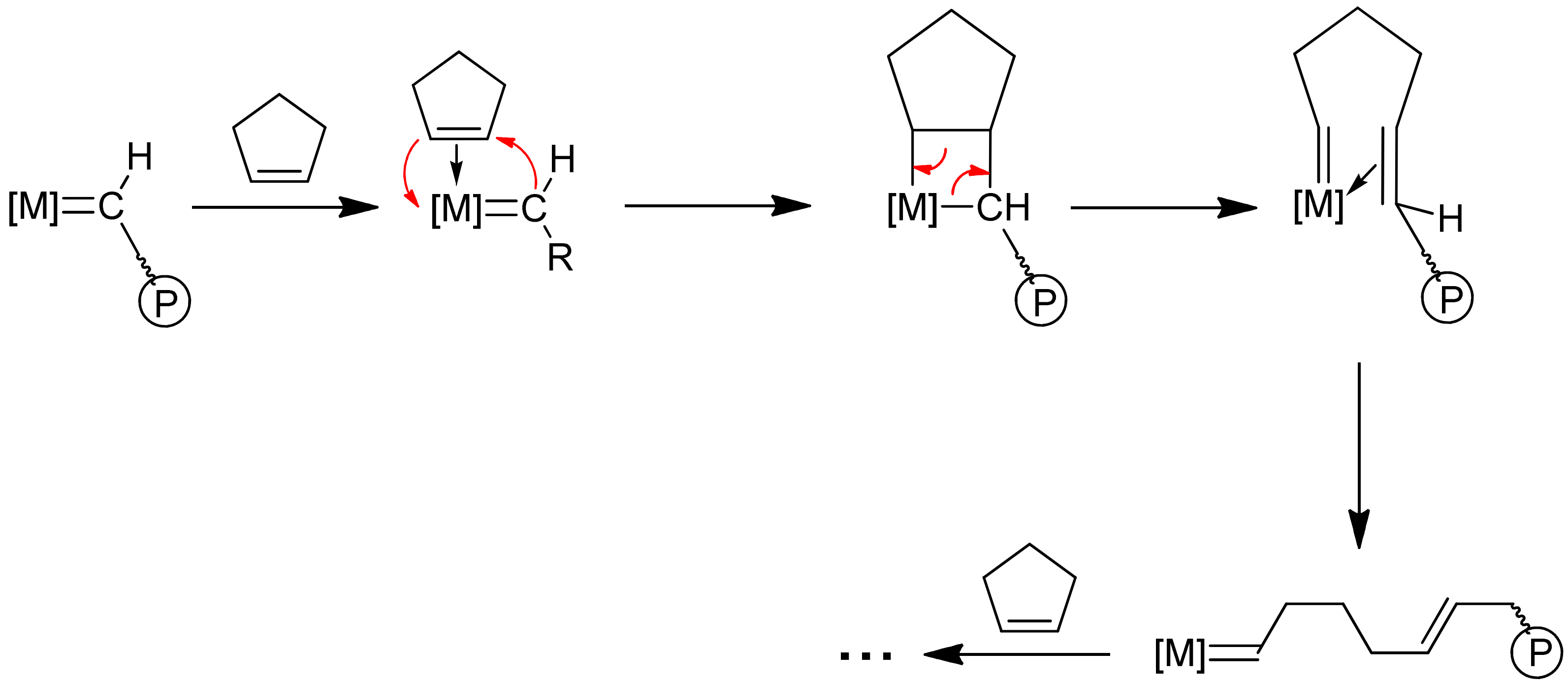
Schrock type catalysts



Grubbs type catalysts

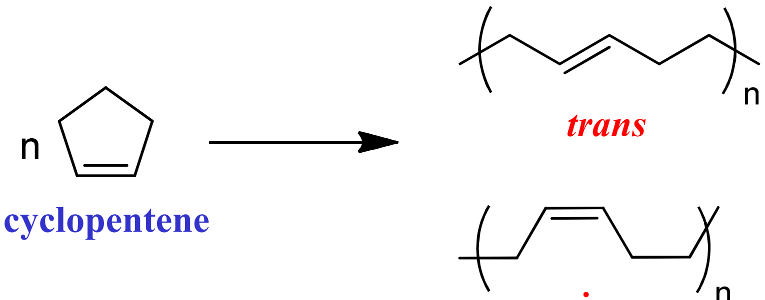
MECHANISM

• PROPAGATION



Living polymerization

ASSORTED EXAMPLES OF ROMP

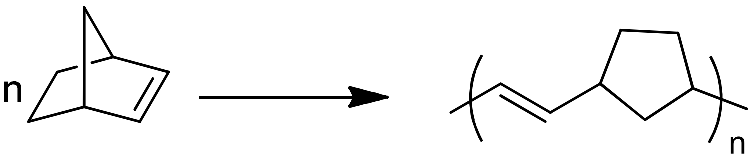


cyclopentene

trans — good elastomer

cis — good elastomer
- Tires

poly(1-pentenylene)

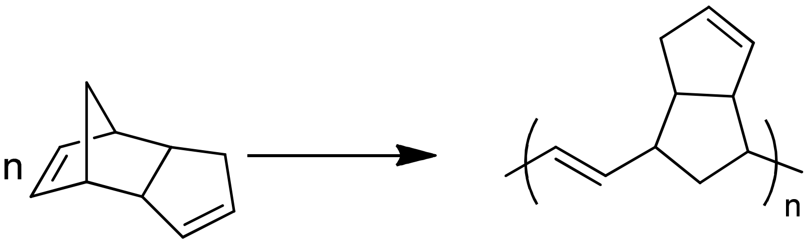


norbornene

poly(cyclopentenylenevinylene)

Norsorex®

- Oil superabsorbent (400% elongation)
- Cleaning up oil spills
- Acoustic insulator
- Gaskets
- Anti-vibration material
- Shock absorption material

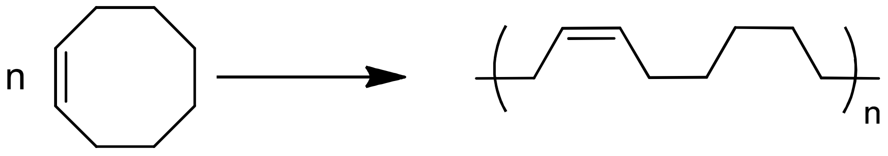


endo-dicyclopentadiene

poly(dicyclopentadiene)

Metton®

- Commercial engineering plastic for moulding

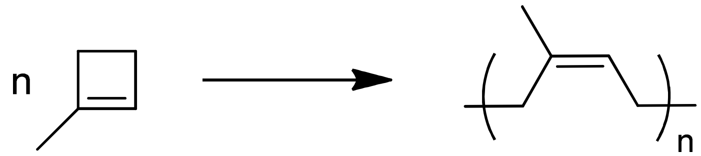


cyclooctene

poly(1-octenylene)

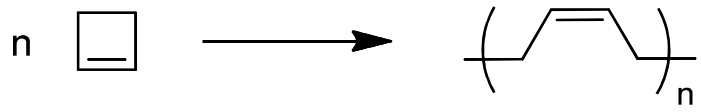
Vestenamer®

- **Minor component in elastomer blends with SBR for gaskets, brake hoses and printing rollers**



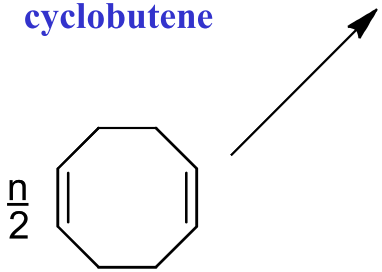
1-methyl-cyclobutene

1,4-*cis*-polyisoprene

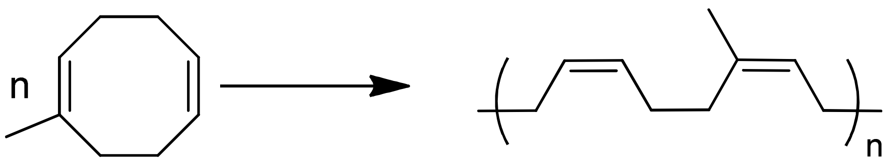


cyclobutene

1,4-*cis*-polybutadiene

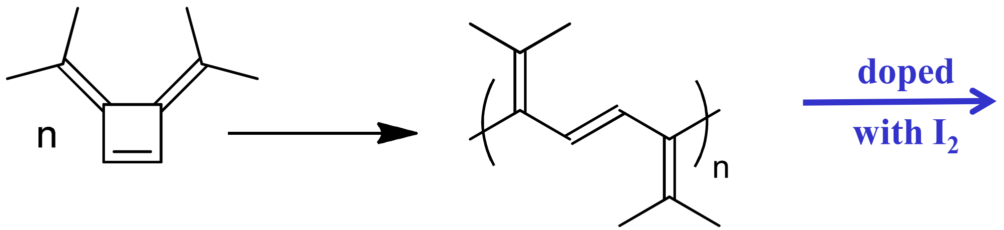


1,5-cyclooctadiene



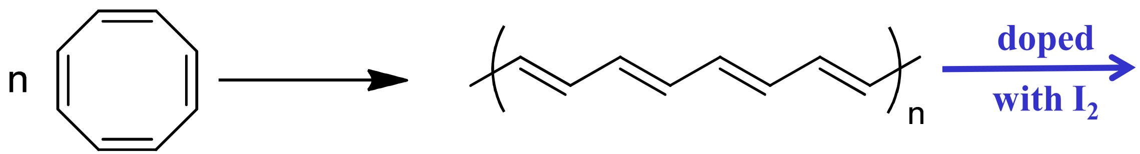
1-methyl-1,5-cyclooctadiene

poly(1,4-*cis*-butadiene)-*alt*-poly(1,4-*cis*-isoprene)



Colourless, soluble, oxidizable

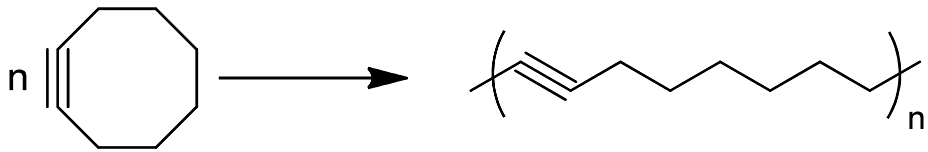
$\sigma = 2\times10^{2}\ \Omega^{-1}\ cm^{-1}$



1,3,5,7-cyclooctatetraene

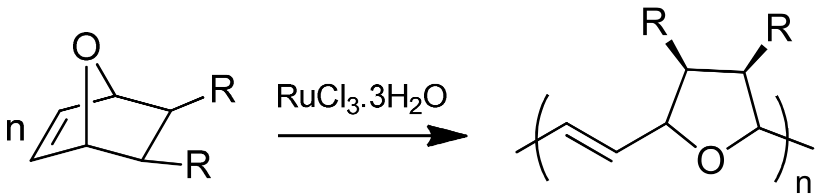
polyacetylene

$\sigma = 10^{3}\ \Omega^{-1}\ cm^{-1}$



cyclooctyne

poly(1-octynylene)



7-oxa-2,3-dimethyl norbornene

$M_n = 10^6$

in aqueous solution !!

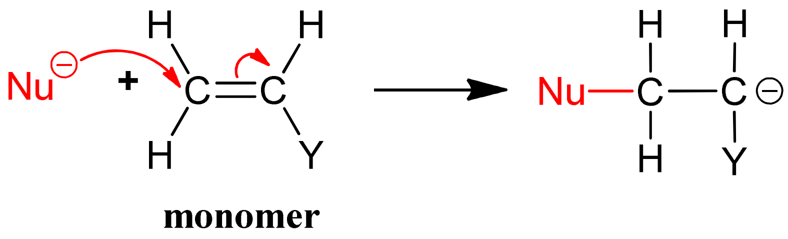
In general, the <u>coordination polymerization</u> experimental conditions impose strict absence of H_2O and O_2

Metal-Catalyzed Polymerization

(Coordination Polymerization)

- Polymerization of Olefins (Insertion) ✓ **Prof. Barbara Milani**
- Polymerization of Dienes (Insertion)
- Polymerization of Alkynes
- Ring Opening Metathesis Polymerization (ROMP)
- **Classical Anionic Polymerization**
- Ring Opening Polymerization (ROP)
- Metal-mediated Radical Polymerization

ANIONIC POLYMERIZATION

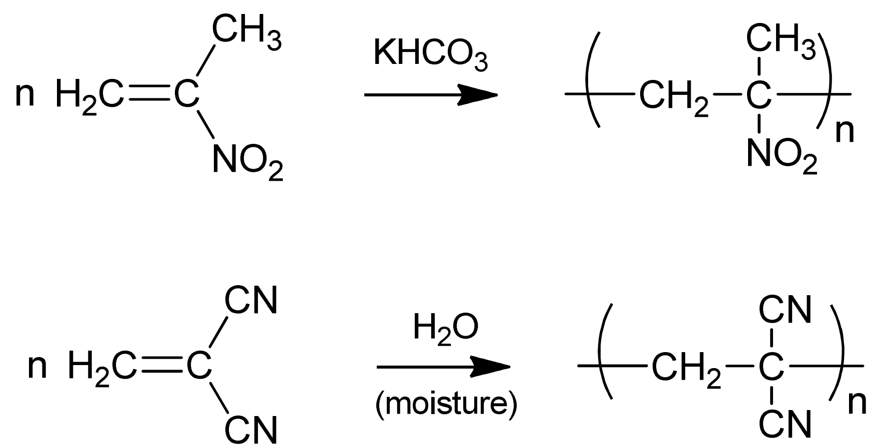


MONOMERS

Most reactive monomers: those containing Y substituents that stabilize carbanions (electron withdrawing by induction and/or resonance). Exs: nitro, cyano, carboxyl, vinyl, phenyl

The more electron attractor group Y is, the less need for strong bases in the initiation:

Exs:



ANIONIC INITIATORS

- ADDITION OF A NEGATIVE ION TO THE MONOMER
- ELECTRON TRANSFER TO THE MONOMER

• INITIATION BY ADDITION OF A NEGATIVE ION TO THE MONOMER

- **Organolithium compounds**

  Exs: $Li\text{-}CH_3$ **(LiMe)**, $Li\text{-}CH_2\text{-}CH_2\text{-}CH_2\text{-}CH_3$ **(LiBu)** **(soluble in inert solvents)**

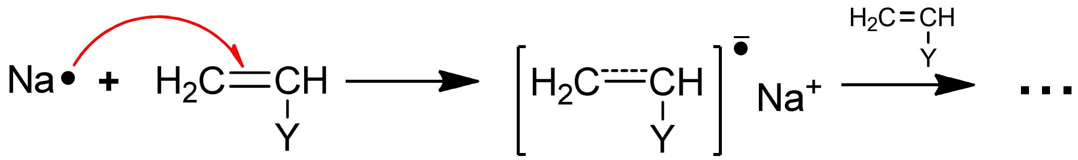
- **Organometallic compounds of the higher alkali metals (Na, K, Rb,…)**
  - higher ionic character than those based on Li
  - less soluble (generally heterogeneous)

- **Organometallic compounds of the alkaline earth metals (Ca, Ba)**
- **Grignard Reagents (RMgX)**

LESS USED (the last two)

• **INITIATION BY ELECTRON TRANSFER TO THE MONOMER**

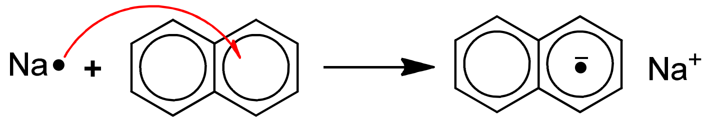
- Alkali Metals
- in solution (of NH_3 or certain ethers)
- in suspension (in inert solvents –"sands")
- supported (in alumina)



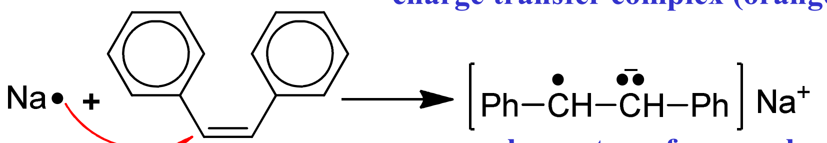
radical anion

- Alkali Metal Complexes (<u>soluble in inert solvents</u>)

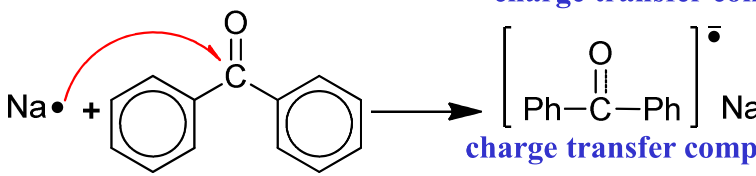
Exs:



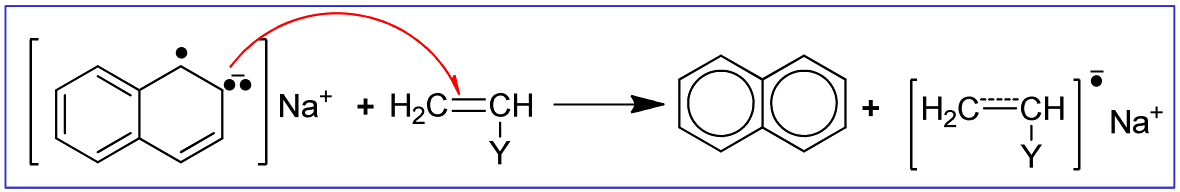
charge transfer complex (orange)



charge transfer complex



charge transfer complex (blue)



MECHANISM AND KINETICS

- **INITIATION BY ANIONIC SPECIES**

Exs: Li-CH_3 (LiMe), Li-CH_2-CH_2-CH_2-CH_3 (LiBu)

$$I^{\ominus} + M \xrightarrow{k_i} P_1^{\ominus}$$

$$P_i^{\ominus} + M \xrightarrow{k_p} P_{i+1}^{\ominus}$$

Normally, $r_i >> r_p$ ➡ chains grow all at the same time ➡ $[P_i^{\ominus}]_o = [I]_o$

$$r_p = -\frac{d[M]}{dt} = k_p\,[I]_o[M]$$

$$[M] = [M]_o\, e^{-k_p[I]_o t}$$

$$\overline{DP}_n = \frac{\overline{M}_n}{M_{RU}}$$

$$\overline{DP}_n = \bar{x} = \frac{[M]_o - [M]}{[I]_o} = \frac{p[M]_o}{[I]_o}$$

$\overline{DP}$ = degree of polymerization

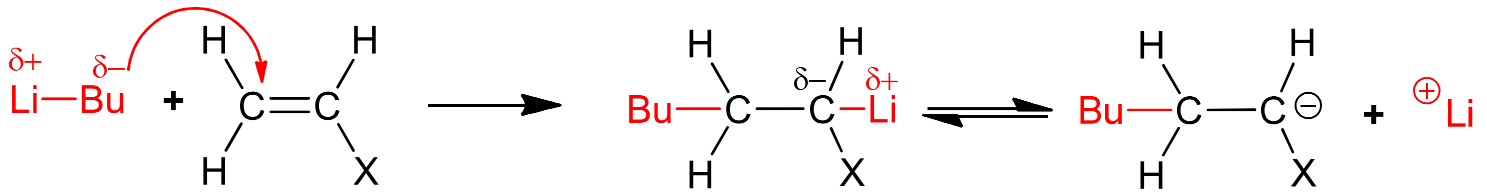
If the monomer is totally consumed (conversion = 100%):

$$\overline{DP}_n = \frac{[M]_o}{[I]_o}$$

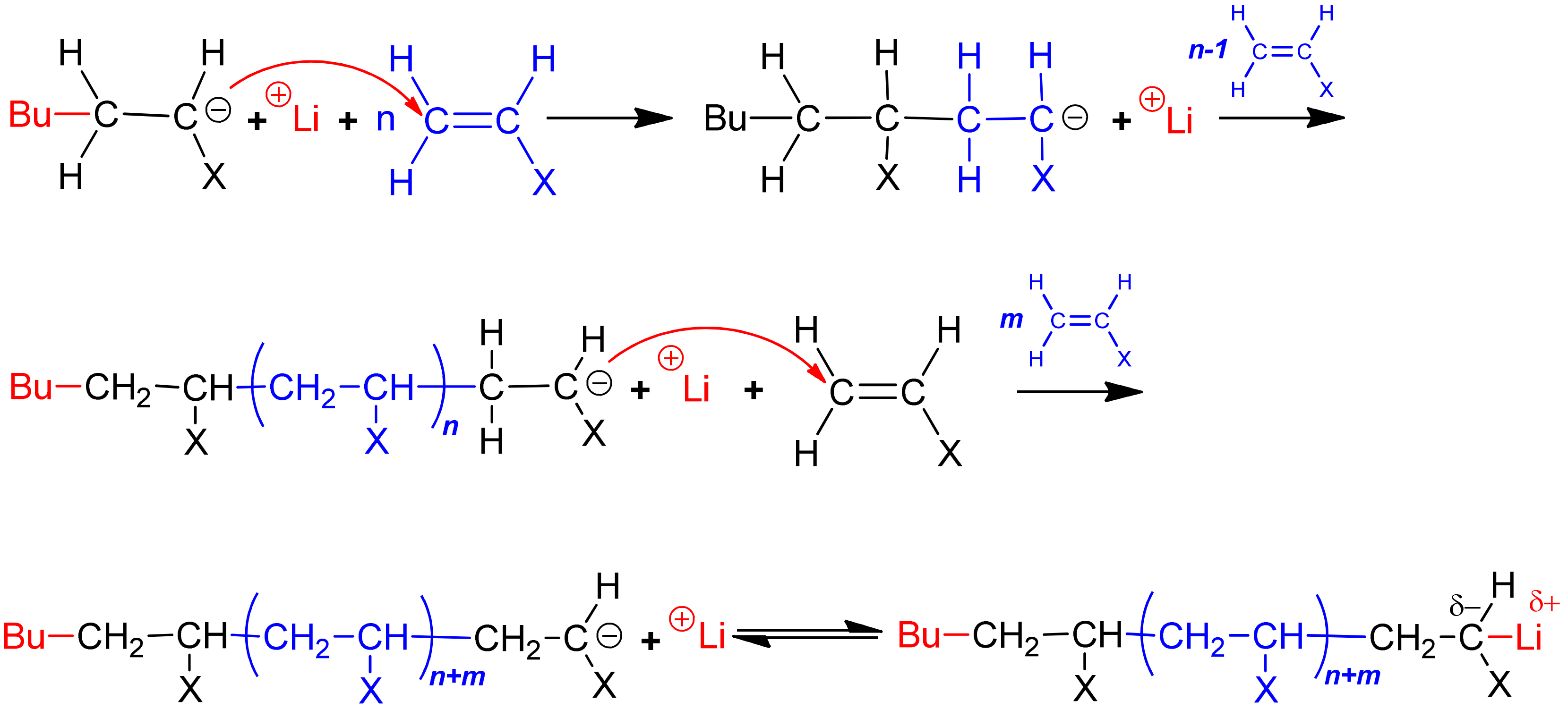
$$\frac{\overline{M}_w}{\overline{M}_n} = 1 + \frac{1}{\overline{DP}_n}$$

Poisson distribution

- **Initiation**



- **Propagation**



living polymer

“dormant” species

- **If the there are no transfer agents in the reaction medium (including impurities in the solvent):**

↓

LIVING POLYMERIZATION

because there is no chain termination

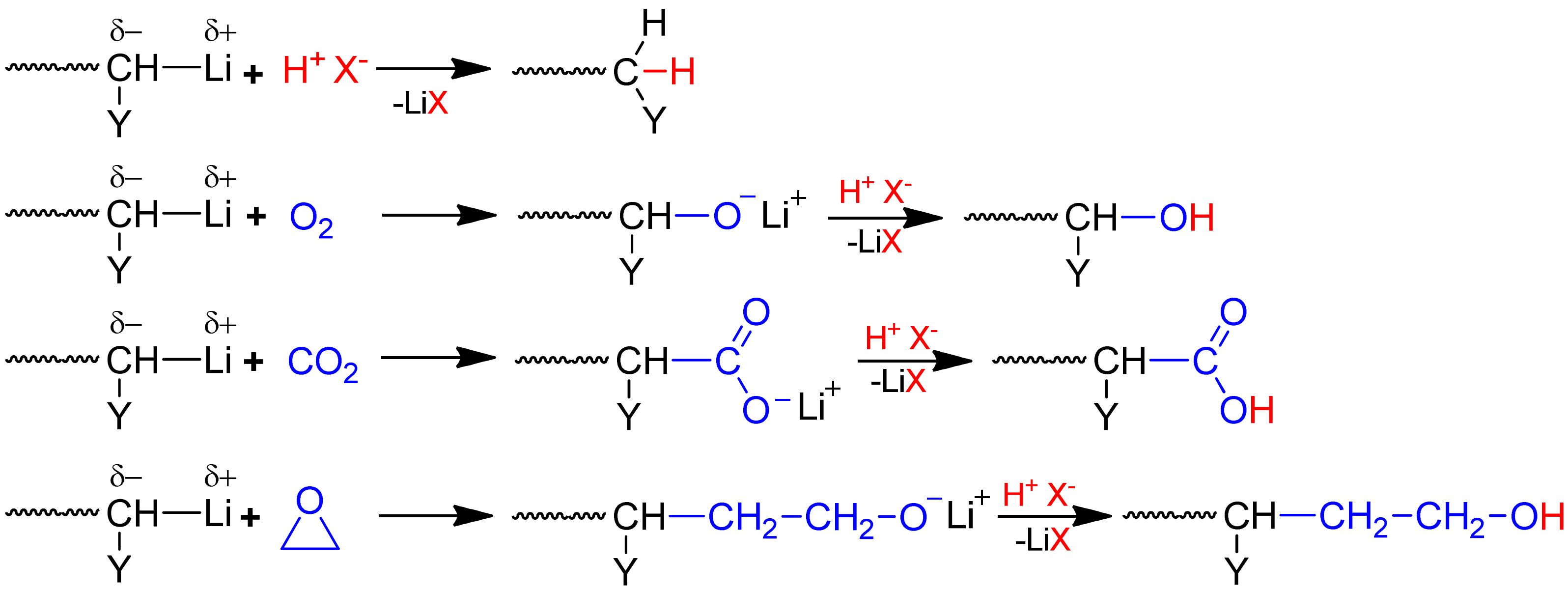
↓

Polymers have a narrow molecular weight distribution i.e. $M_n/M_w \approx 1$

(e.g. GPC/SEC standards)

- **If there is no termination, the chain end is living and can be used for:**

  - **chain end functionalization**
  - **block copolymerization**

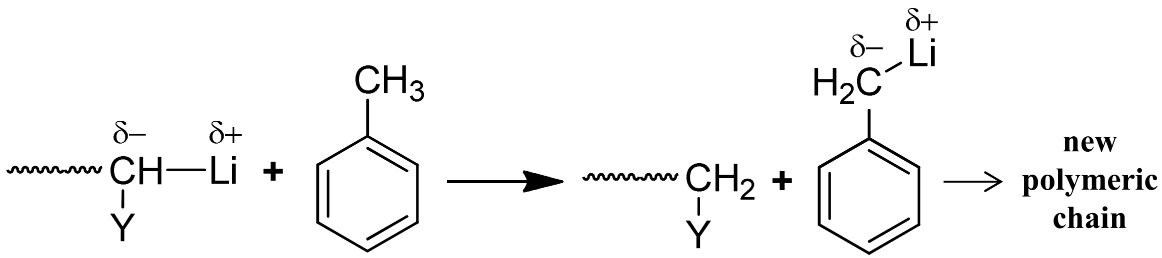


- **Easy control of the molecular weight:**

Routes

- **Stoichiometry** ($\overline{DP}_n = \frac{[M]_o}{[I]_o}$)
- **Termination at time *t* with addition of a terminating agent**
- **Addition of a chain transfer agent** → **to decrease M_n**

Ex:



- The rate of propagation can be influenced by the <u>degree of association</u> between anion and cation, which depends strongly on the SOLVENT:

  - **<u>Non-coordinating and weakly polar solvents</u> (*e.g.* hydrocarbons)**

    ↓

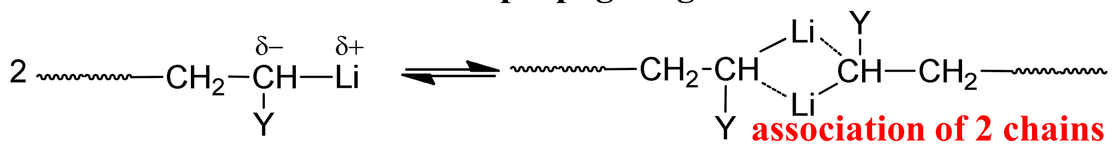
    Higher degree of association of the initiator → Lower polymerization rate (r_p)

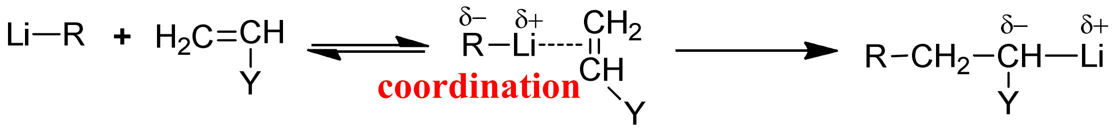
    - Association of the initiator:

$$4\ LiMe \rightleftharpoons (LiMe)_4$$

$$6\ LiBu \rightleftharpoons (LiBu)_6$$

- Association of the propagating chain:

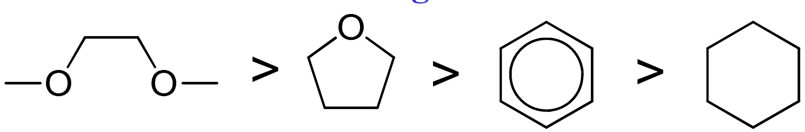




  - **<u>More coordinating and/or more polar solvents</u> (*e.g.* ethers)**

    ↓

    Lower degree of association and higher solvation → Higher polymerization rate (r_p)



DME THF Benzene Cyclohexane

Ex:

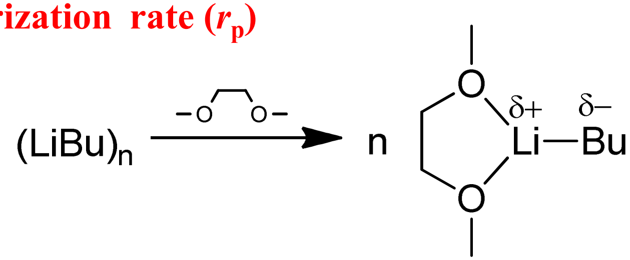


TABLE 7.4. Representative Anionic Propagation Rate Constants, k_p, for Polystyrene[a]

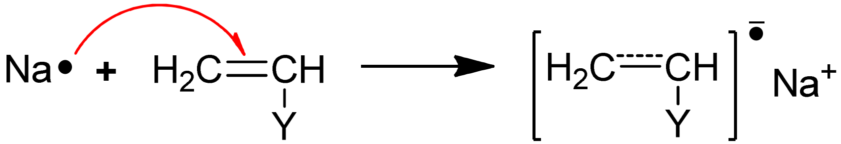
Counterion	*Solvent*	k_p *(L/mol s)*[b]
Na^+	Tetrahydrofuran	80
Na^+	1,2-Dimethoxyethane	3600
Li^+	Tetrahydrofuran	160
Li^+	Benzene	10^{-3}–10^{-1} [c]
Li^+	Cyclohexane	$(5\text{-}100)\times 10^{-5}$ [c]

[a] Data from Morton.[30]

[b] At 25°C unless otherwise noted.

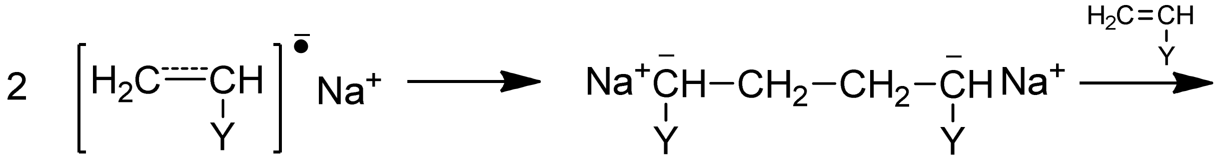
[c] Variable temperature.

• INITIATION BY ELECTRON TRANSFER



radical anion

Recombination of 2 radical anions:



. . . propagation

Dianionic chain
with 2 propagation centres

↓

TELECHELIC POLYMERIZATION

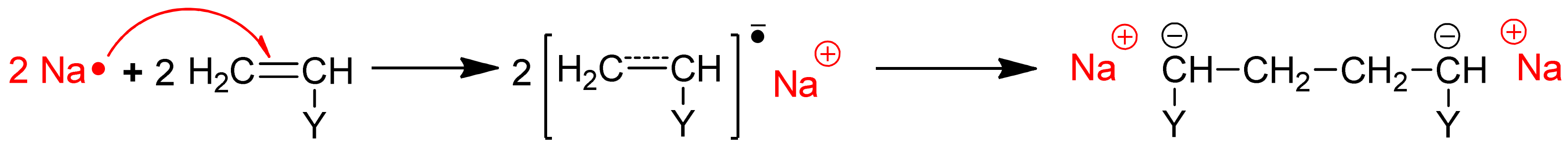
↓

$$\overline{DP}_n = \overline{x} = 2\,\frac{[M]_o - [M]}{[I]_o}$$

↓

average kinetic chain length

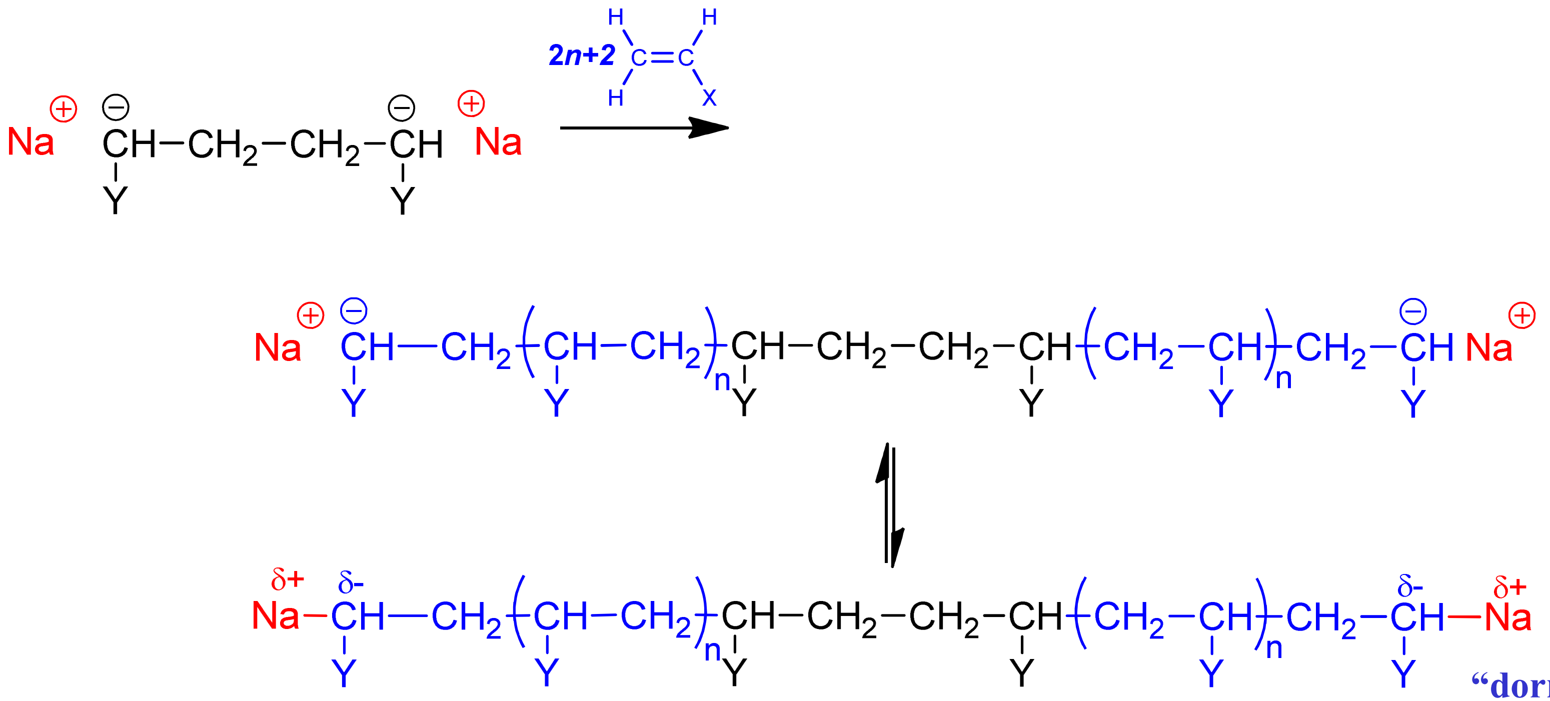
• **Initiation**



radical anion

radical recombination

• **Propagation**

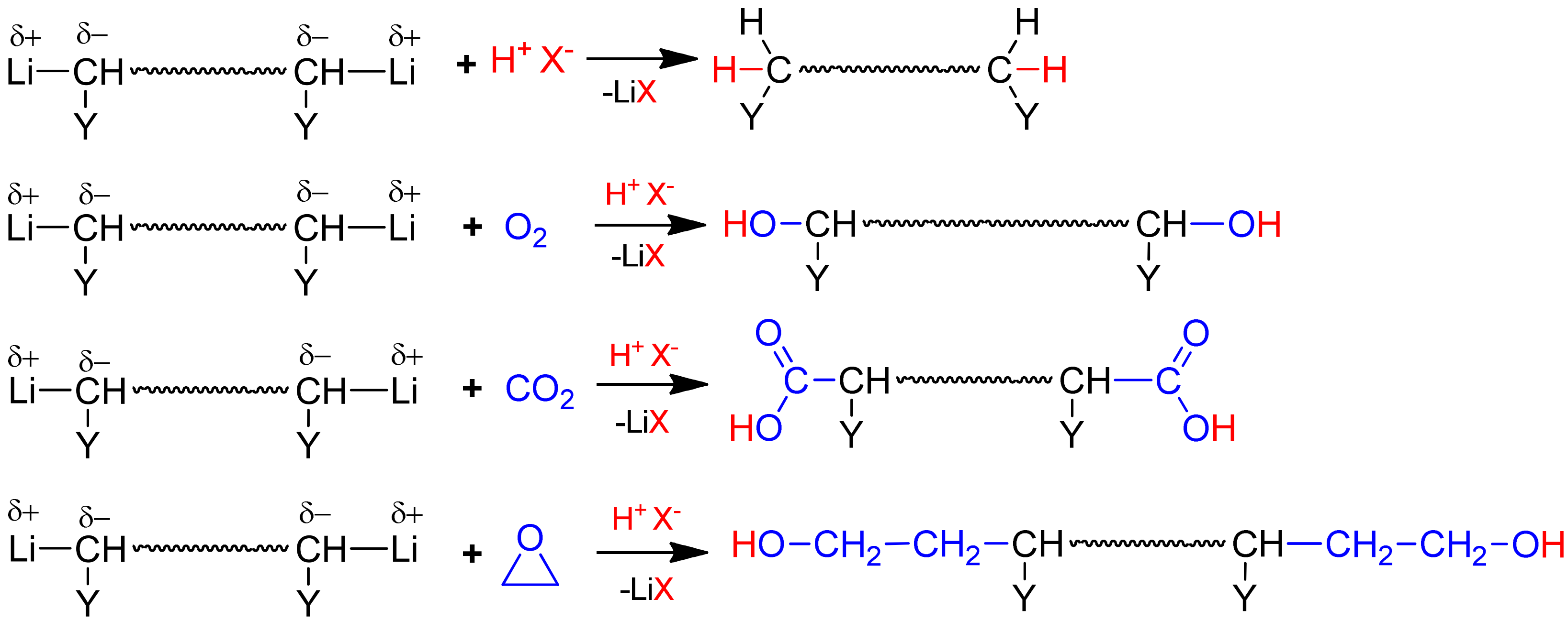


"dormant" species

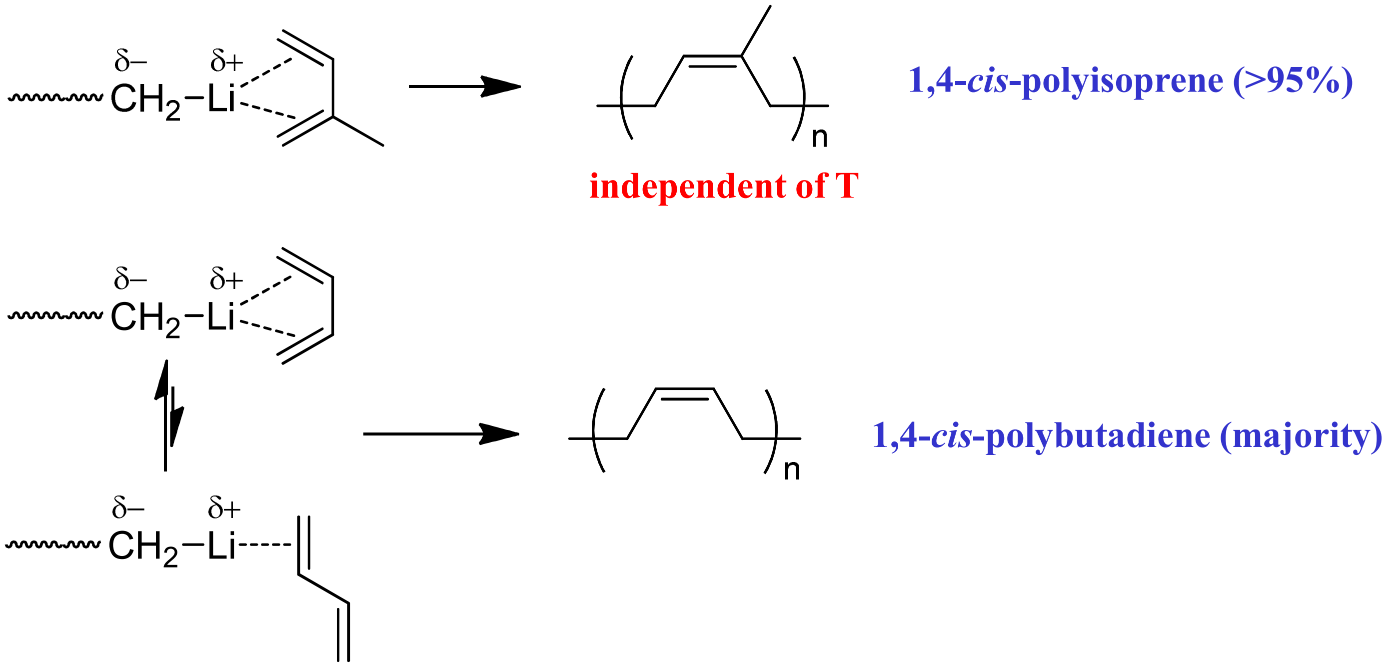
living telechelic polymer

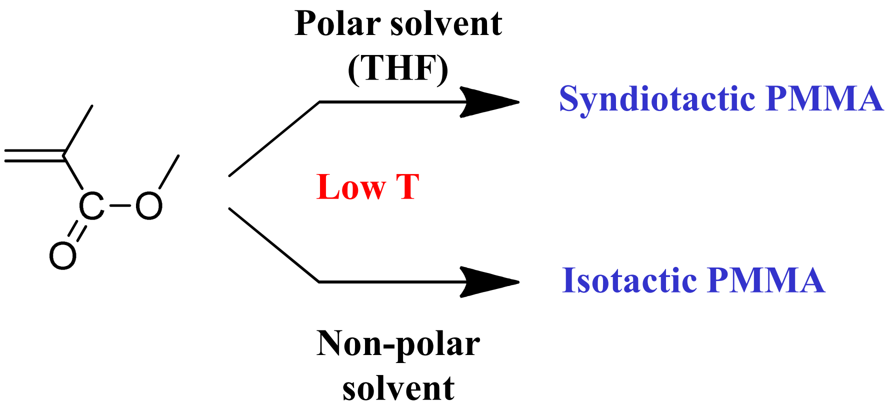
Telechelic chain growth (chain growth at both ends)

- **Functionalization of both chain ends:**



STEREOCHEMISTRY OF PROPAGATION

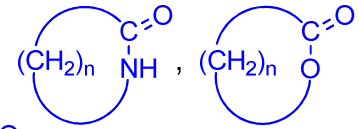


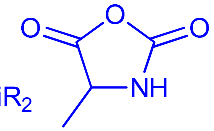


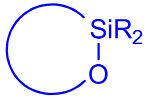
The Li+ counter-cation always assists the chain growth by bonding covalently to the chain end, protecting it and avoiding chain termination or chain transfer reactions

ANIONIC RING OPENING POLYMERIZATION

4 Types of Monomers

- **EPOXIDES** 
- **CYCLIC LACTAMS AND LACTONES** 

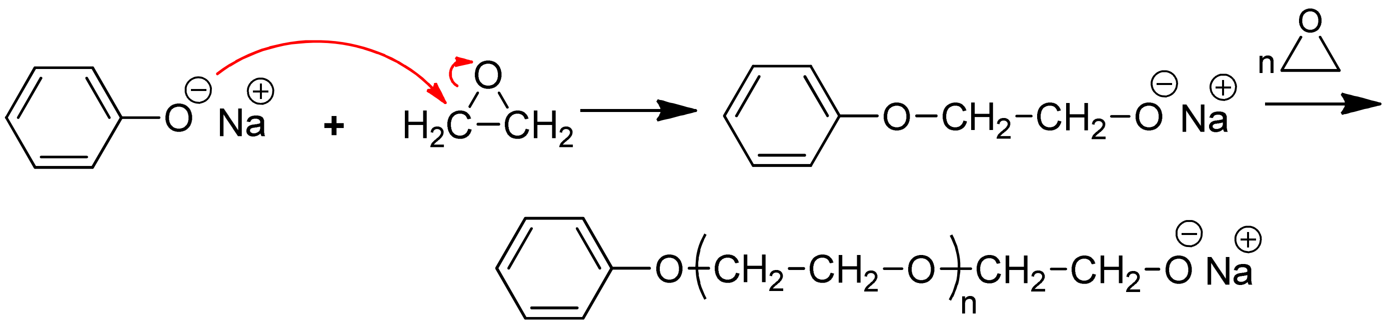
- ***N*-CARBOXYANHYDRIDES** 

- **CYCLIC SILOXANES** 

Some examples have already been given in the Step-Growth Polymerization

• EPOXIDES

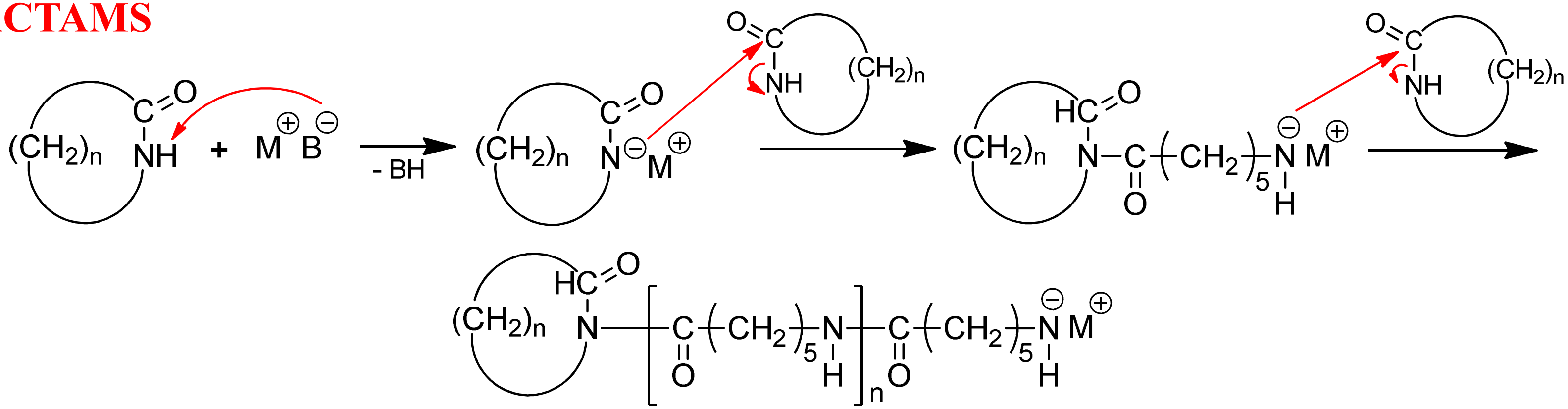
Ex:



Poly(ethylene oxide) (PEO)

• LACTAMS

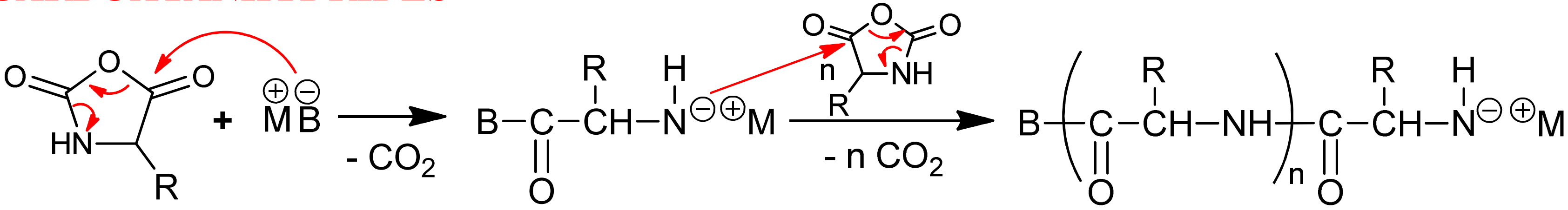
Ex:



Nylon 6

• *N*-CARBOXYANHYDRIDES

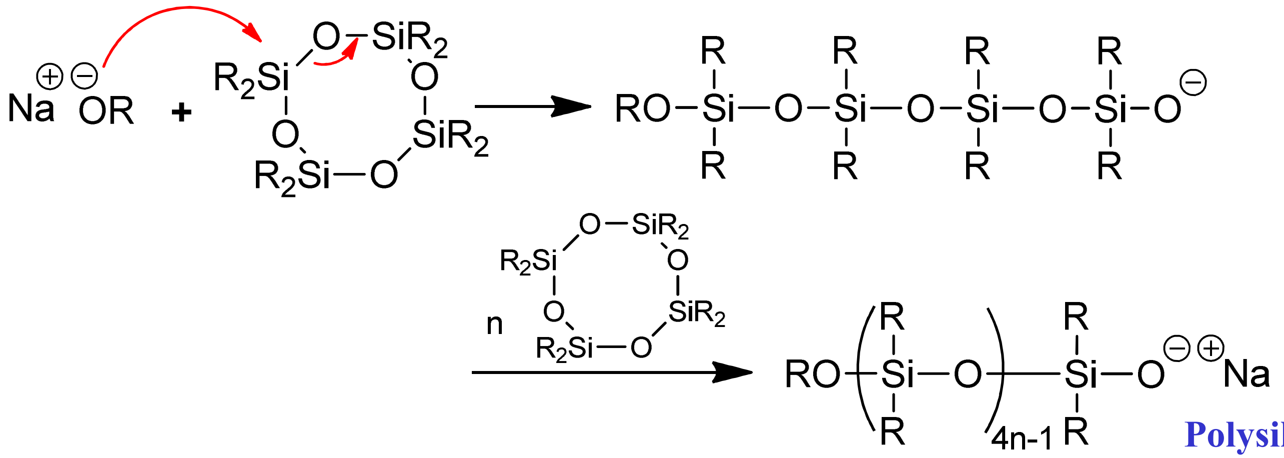
Ex:



Polypeptide (M_n ~10^6)

• CYCLIC SILOXANES

Ex:



Polysiloxane

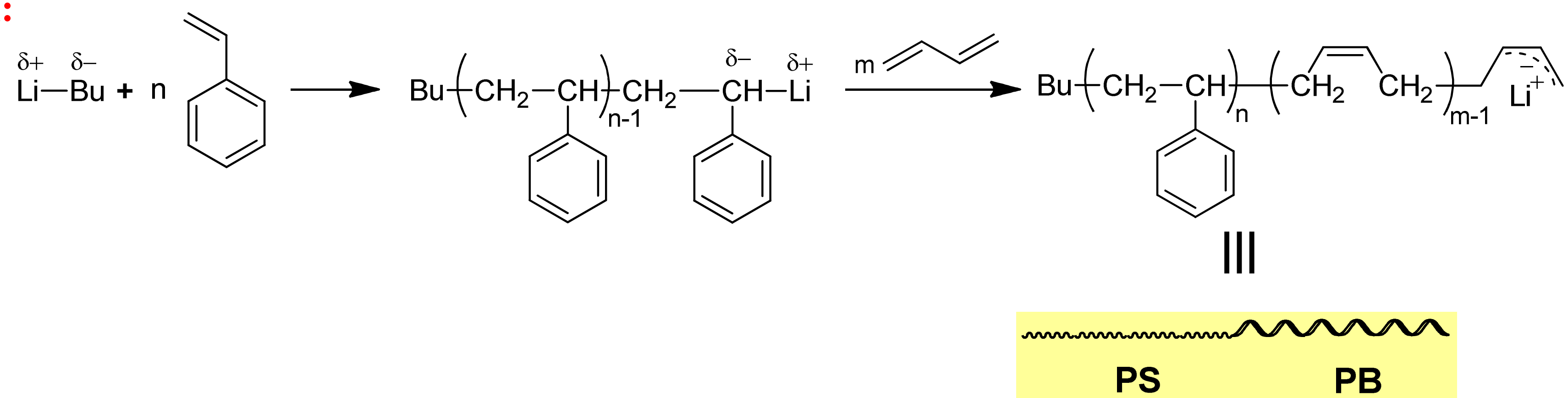
ANIONIC COPOLYMERIZATION

• BLOCK COPOLYMERIZATION (COMONOMERS ADDED SEQUENTIALLY)

Owing to the living nature of anionic polymerization (absence of chain transfer and chain termination reactions) and easy control of molecular weight, this technique is very much used in block copolymerization

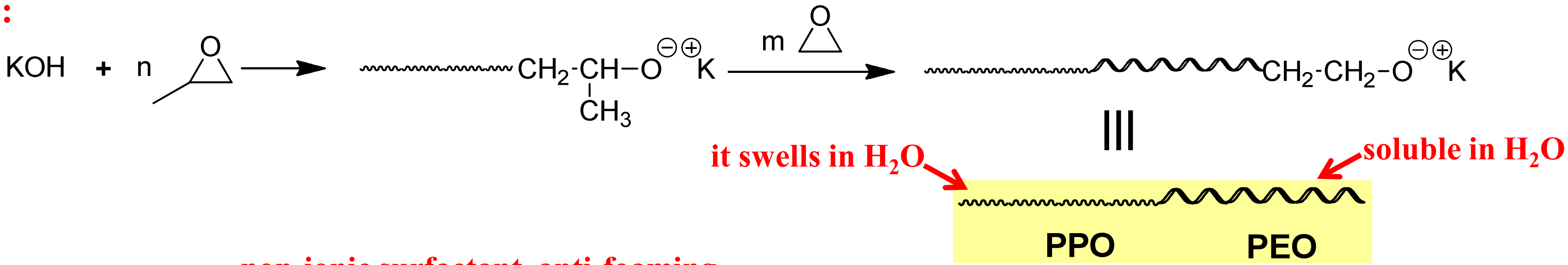
• From mononegative chains

Ex:



Polystyrene-*block*-1,4-*cis*-Polybutadiene

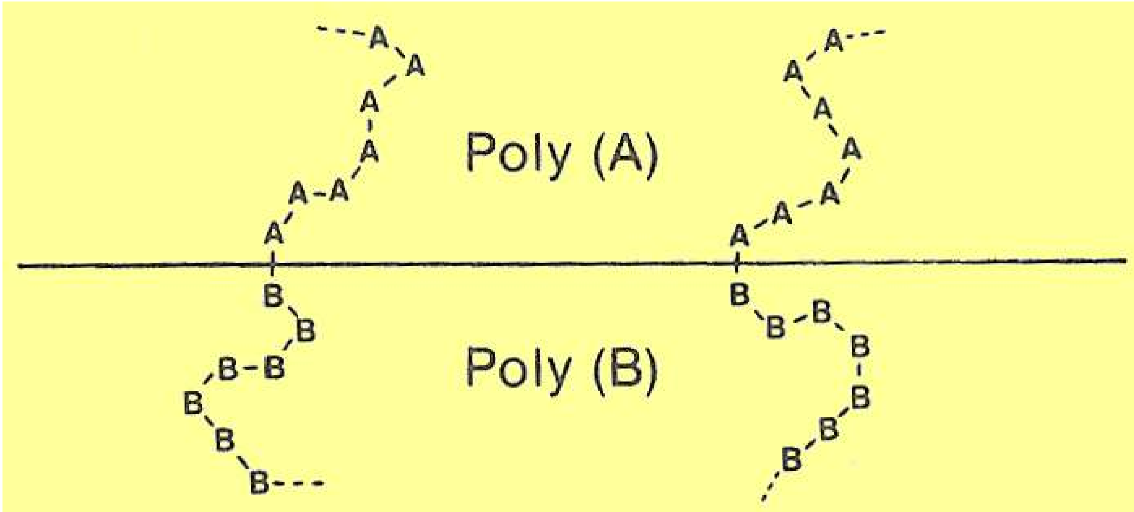
Ex:



non-ionic surfactant, anti-foaming, thickener, wetting agent (Pluronic®)

Poly(propylene oxide)-*block*-Poly(ethylene oxide)

COMPATIBILIZERS OF IMMISCIBLE POLYMER BLENDS



AB BLOCK COPOLYMERS ARE COMPATIBILIZERS

Example: commercial SB (polystyrene-*b*-polybutadiene)

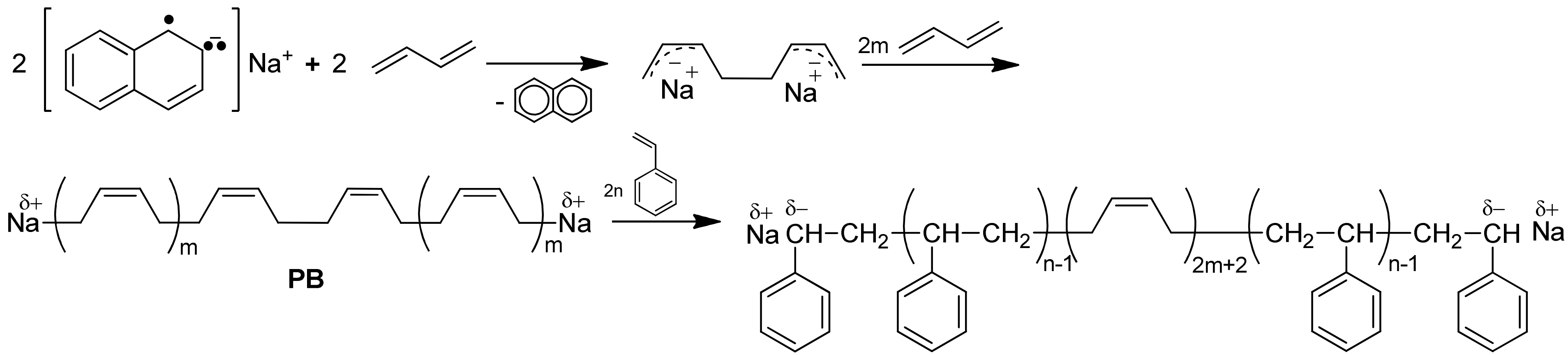
Blends of polybutadiene and polystyrene are immiscible

AB copolymers improve the adhesion between phases and compatibilize them

Sequential polymerization of different monomers (with total monomer consumption) is possible due to the living nature of the polymeric chain end. The preparation of the following types of block copolymers can be performed:

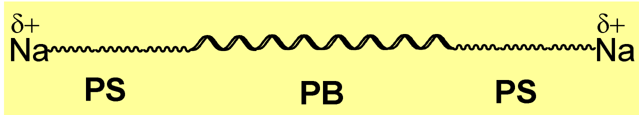
~~AAAAAABBBBBB~~~~	**DIBLOCK AB TYPE**
~~AAAAAABBBB~~~~BBBBCCCCCC~~~~	**TRIBLOCK ABC TYPE**
~~AAAAAABBBB~~~~BBBBAAAAAA~~~~	**TRIBLOCK ABA TYPE**
—[~~~~AAAAAABBBBBB~~~~]—$_n$	**MULTIBLOCK [AB] TYPE**

• From dinegative chains

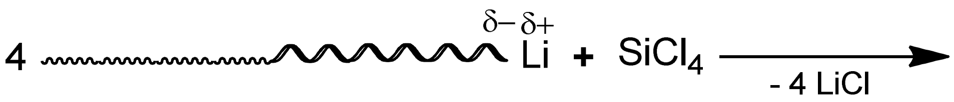


Triblock SBS copolymer

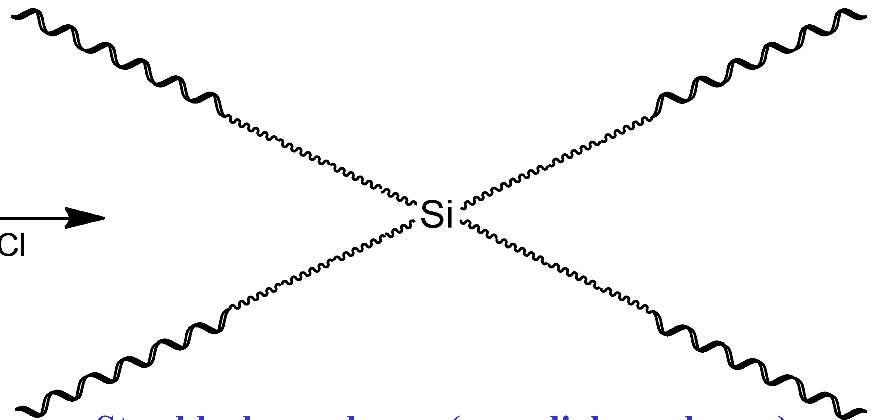
The ABA copolymers are thus obtained in 2 steps whereas from mononegative initiators is obtained in 3 steps



• Star-block copolymers



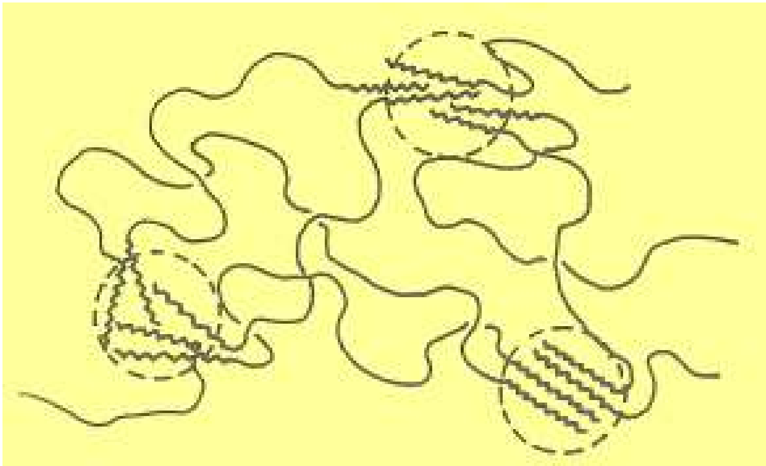
Living AB block copolymer



Star-block copolymer (or radial copolymer)

When melted, star-block copolymers exhibit lower viscosities, even when M_n are high

THERMOPLASTIC ELASTOMERS (TPE)



ABA BLOCK COPOLYMERS ARE THERMOPLASTIC ELASTOMERS
(A= rigid polymer; B= flexible polymer)

<u>Example</u>: commercial SBS (polystyrene-*b*-polybutadiene-*b*-polystyrene)

- PB blocks – *ca.* M_n = 50000 - 70000

- PS blocks – *ca.* M_n = 10000 - 15000

Semicrystalline blocks at the ends (S) tend to aggregate in microdomains, whereas amorphous central blocks (B) form the matrix

Aggregation ≡ elastic behaviour ≡ physical crosslinks

• **NORMAL COPOLYMERIZATION (COMONOMERS MIXED IN THE FEED)**

Relatively few reactivity ratios have been determined for anionic "normal" copolymerization

TABLE 7.5. Representative Anionic Reactivity Ratios (r)[a]

Monomer 1	*Monomer 2*	*Initiator*[b]	*Solvent*[c]	*Temperature*[d] *(°C)*	r_1	r_2
Styrene	Methyl methacrylate	Na	NH_3		0.12	6.4
		n-BuLi	None		e	e
	Butadiene	n-BuLi	None	25	0.04	11.2
		n-BuLi	Hexane	25	0.03	12.5
		n-BuLi	Hexane	50	0.04	11.8
		n-BuLi	THF	25	4.0	0.3
		n-BuLi	THF	−78	11.0	0.4
		EtNa	Benzene		0.96	1.6
	Isoprene	n-BuLi	Cyclohexane	40	0.046	16.6
	Acrylonitrile	RLi	None		0.12	12.5
	Vinyl acetate	Na	NH_3		0.01	0.01
Butadiene	Isoprene	n-BuLi	Hexane	50	3.38	0.47
Methyl methacrylate	Acrylonitrile	$NaNH_2$	NH_3		0.25	7.9
		RLi	None		0.34	6.7
	Vinyl acetate	$NaNH_2$	NH_3		3.2	0.4

[a] Data from Morton.[30]
[b] Bu = butyl, Et = ethyl, R = alkyl.
[c] THF = tetrahydrofuran.
[d] Temperature not specified in some instances.
[e] No detectable styrene in polymer.

POLYMER ARCHITECTURES ACCESSIBLE BY LIVING POLYMERIZATION

Monofunctionalized polymers

- Dispersing agents
- Synthesis of macromonomers

Difunctionalized polymers

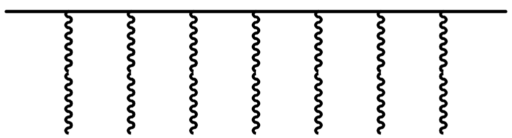
- Synthesis of elastomers
- Crosslinking agents

AB block copolymers

- Dispersing agents
- Compatibilizers of polymer blends

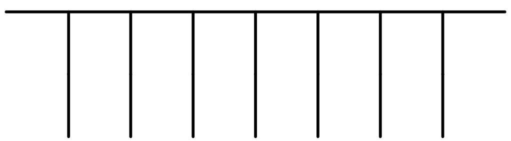
ABA block copolymers

- Thermoplastic elastomers (TPE)



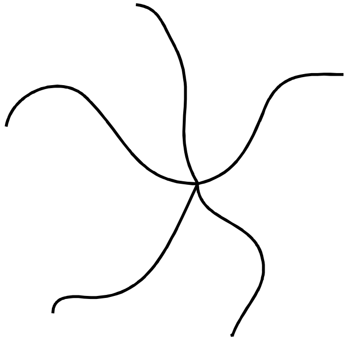
Graft copolymers

- Elastomers
- Adhesives



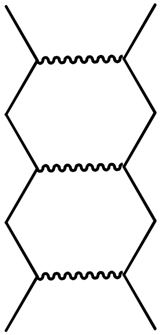
Comb polymers

- Elastomers
- Adhesives



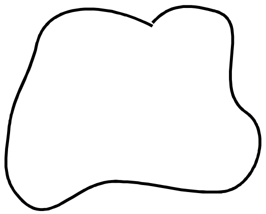
Star polymers/ copolymers

- Rheology control



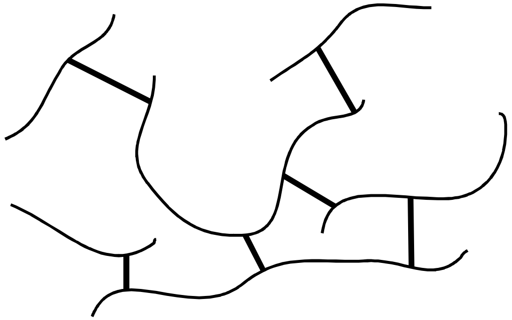
Ladder polymers/ copolymers

- High temperature plastics
- Membranes
- Elastomers



Cyclic polymers

- Rheology control



Amphiphilic networks

- Biocompatible polymers

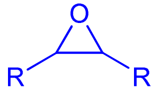
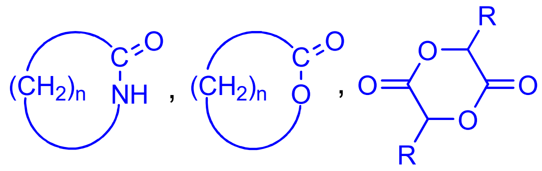
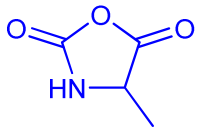
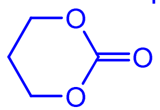
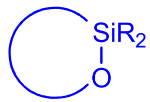
Metal-Catalyzed Polymerization

(Coordination Polymerization)

- *Polymerization of Olefins (Insertion)* ✓ *Prof. Barbara Milani*
- *Polymerization of Dienes (Insertion)*
- *Polymerization of Alkynes*
- *Ring Opening Metathesis Polymerization (ROMP)*
- *Classical Anionic Polymerization*
- ***Ring Opening Polymerization (ROP)***
- *Metal-mediated Radical Polymerization*

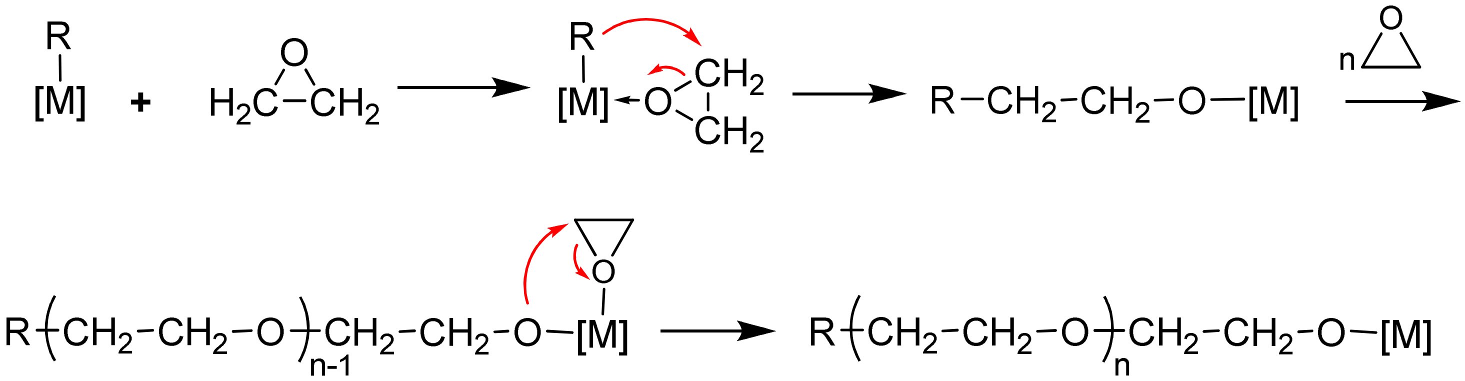
RING-OPENING POLYMERIZATION (ROP)

Types of Monomers (heterocyclic rings)

- **EPOXIDES** 
- **CYCLIC LACTAMS, LACTONES AND DILACTONES** 
- ***N*-CARBOXYANHYDRIDES** 
- **CYCLIC CARBONATES** 
- **CYCLIC SILOXANES** 

• EPOXIDES

Ex:



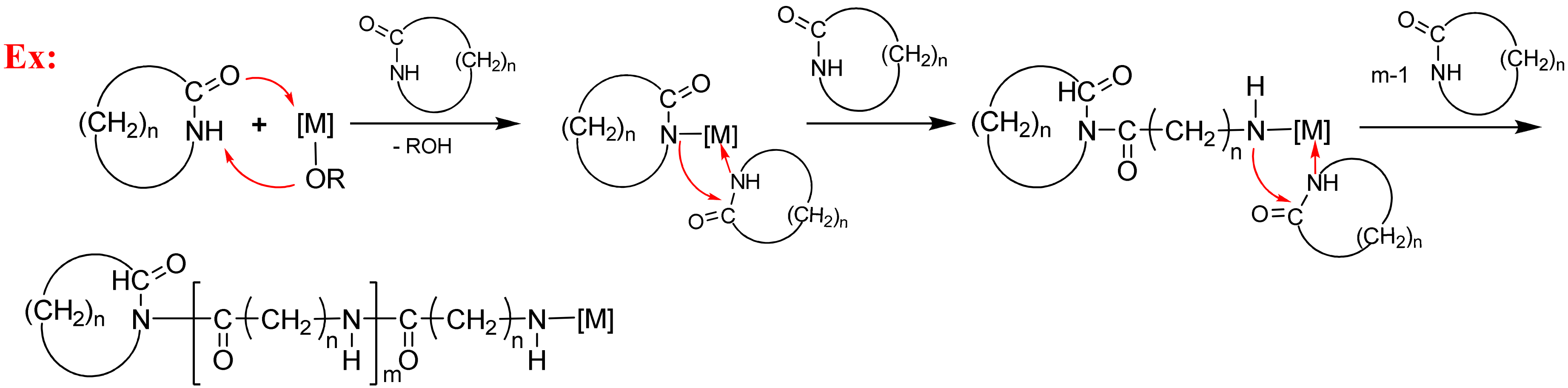
<u>Polyethers</u> [Poly(ethylene oxide) (PEO)]

Main catalytic systems used for the coordination ROP of epoxides

Monomer	*Catalysts*
Methyloxirane	$FeCl_3$/POx, $ZnEt_2/H_2O$, $AlEt_3/H_2O$/pyridine, and others
Phenyloxirane	$ZnEt_2(H_2O)$
(Haloalkyl)oxiranes (e.g., ECH)	$FeCl_3$/POx, $AlEt_3/H_2O$(/pyridine)
Oxiranes substituted with acetal groups	$ZnEt_2$/MeOH, $ZnEt_2$/cyclohexanol
Oxiranes substituted with ester groups	$AlEt_3/H_2O$/acetylacetone
Oxiranes substituted with organosilane or organosiloxane	$ZnEt_2/H_2O$
Oxiranes substituted with nitrile	$Al(i\text{-}Bu)_3/H_2O$/acetylacetone
2,3-Dimethyloxirane	$Al(i\text{-}Bu)_3/H_2O$, $ZnEt_2/H_2O$
bis(Chloromethyl)oxirane	$Al(i\text{-}Bu)_3/H_2O$
1,2-Epoxycyclohexane	$ZnEt_2$, $(EtZnOMe)_4$, $Al(i\text{-}Bu)_3/H_2O$, $AlEt_3/H_2O$/acetylacetone, and others
Others (ethyl, *tert*-butyl, neopentyl, allyl amines, sulfones, ether, amides)	$ZnEt_2/H_2O$

• LACTAMS

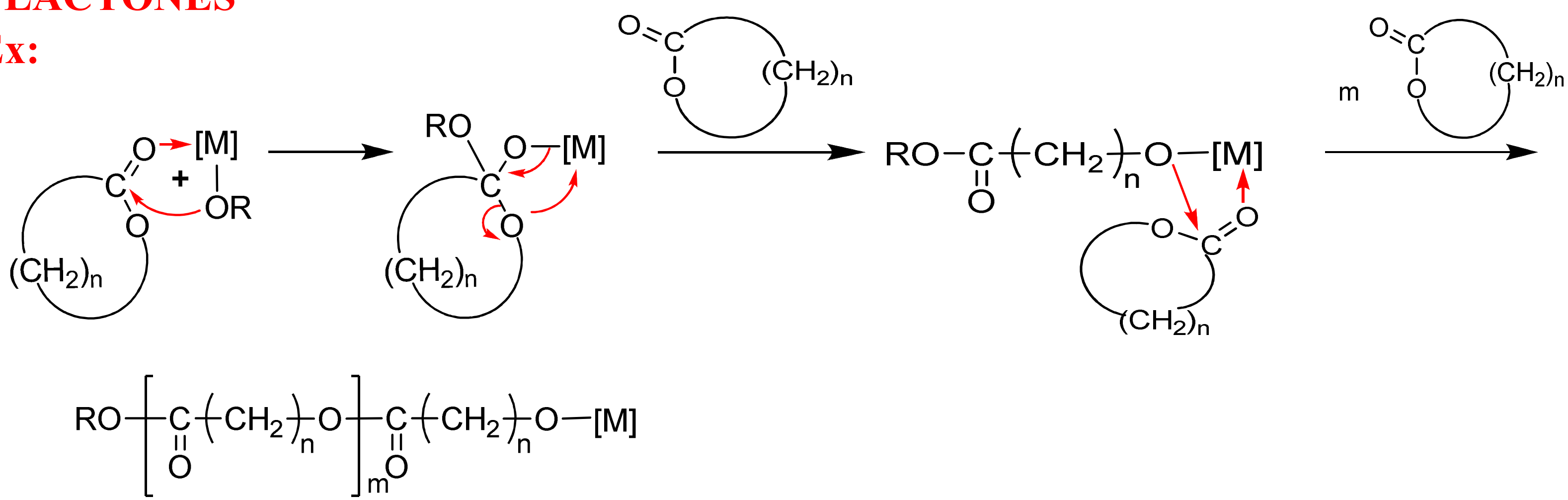
Ex:



Polyamides [nylon 6 or poly(ε-caprolactam) (n=5)]

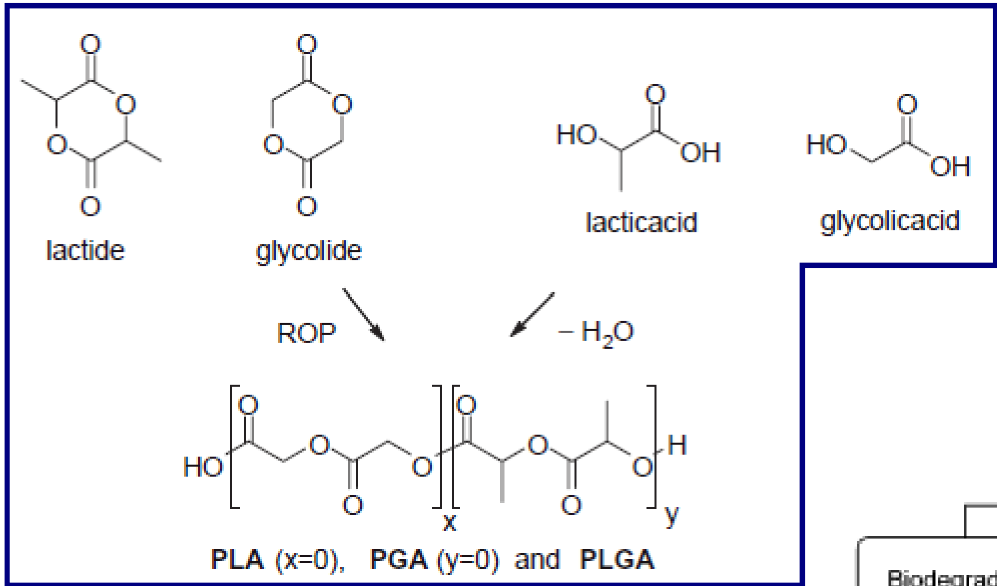
• LACTONES

Ex:



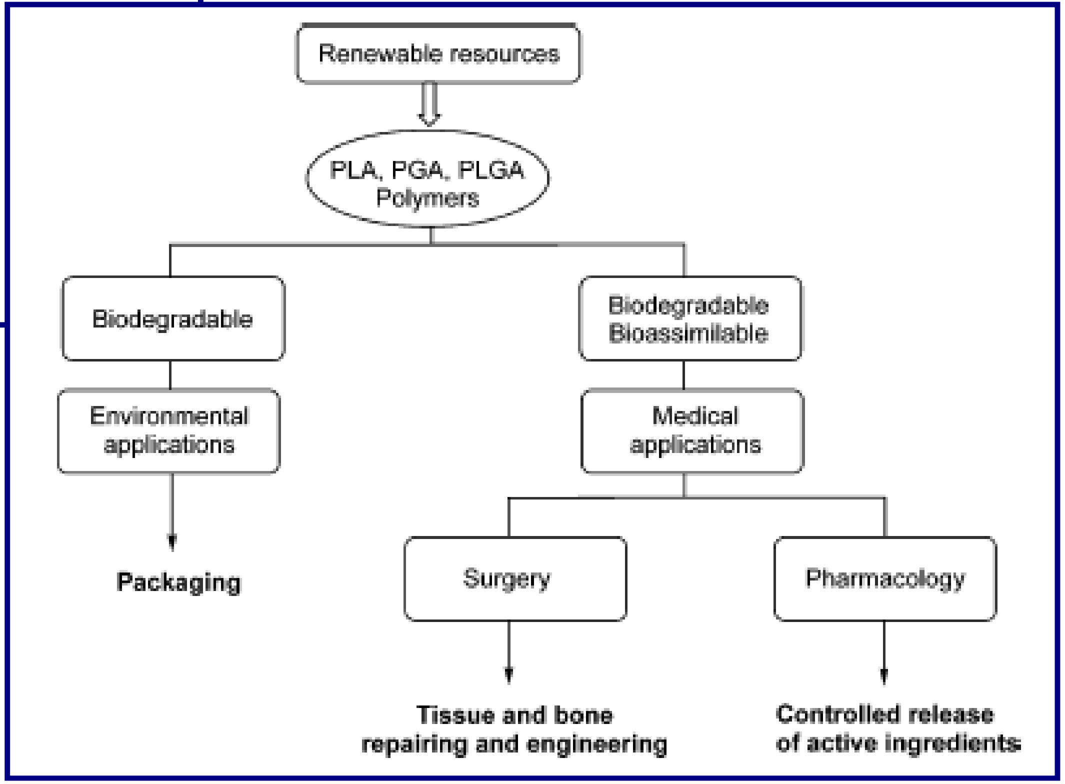
Polyesters [poly(ε-caprolactone)] (n=5)

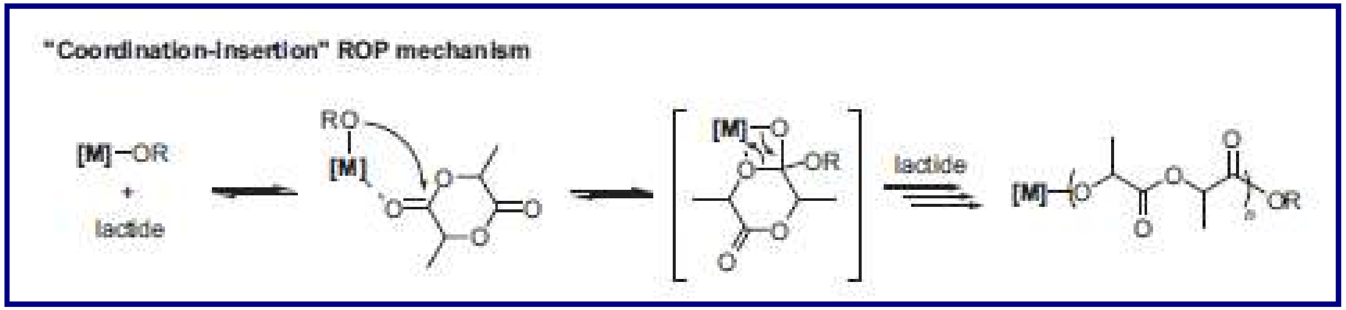
- **DILACTONES**



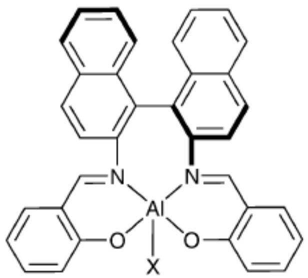
Monomers, polymers and copolymers

Polyesters

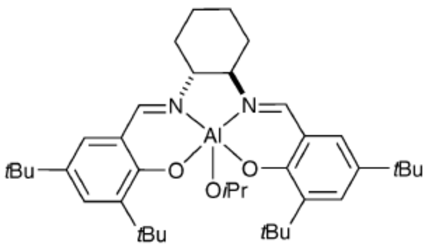




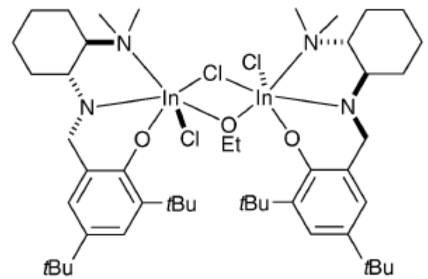
Catalysts used for coordination ROP of lactide



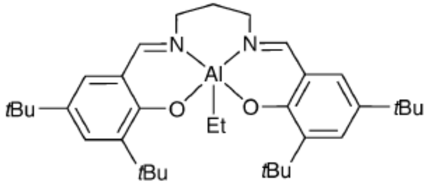
1a: X = O*i*Pr
1b: X = OMe



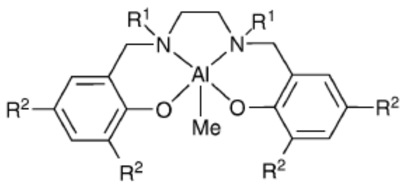
2



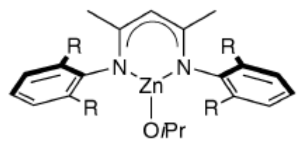
5



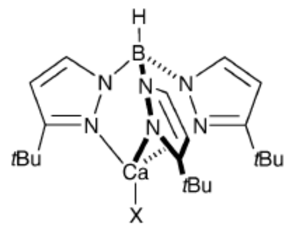
3



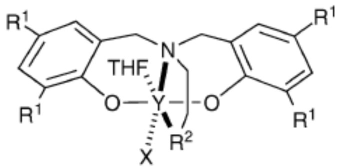
4a: R^1 = Me, R^2 = H
4b: R^1 = CH_2Ph, R^2 = H
4c: R^1 = CH_2Ph, R^2 = Cl
4d: R^1 = Me, R^2 = Me



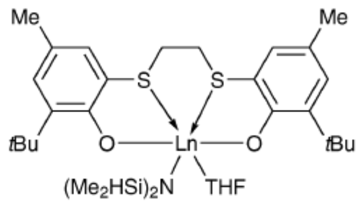
6a: R = Et
6b: R = *n*Pr
6c: R = *i*Pr



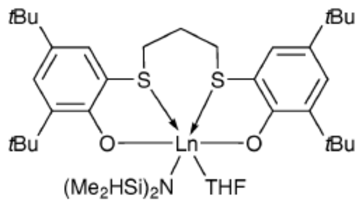
7a: X = $N(SiMe_3)_2$
7b: X = OC_6H_3-2,6-*i*Pr_2



9a: R^1 = CMe_3; R^2 = OMe; X = $N(SiHMe_2)_2$
9b: R^1 = $CPhMe_2$; R^2 = OMe; X = $N(SiHMe_2)_2$
9c: R^1 = $CPhMe_2$; R^2 = OMe; X = O*i*Pr
9d: R^1 = Me; R^2 = OMe; X = $N(SiMe_3)_2$
9e: R^1 = CMe_3; R^2 = NMe_2; X = CH_2SiMe_3
9f: R^1 = CMe_3; R^2 = NEt_2; X = CH_2SiMe_3

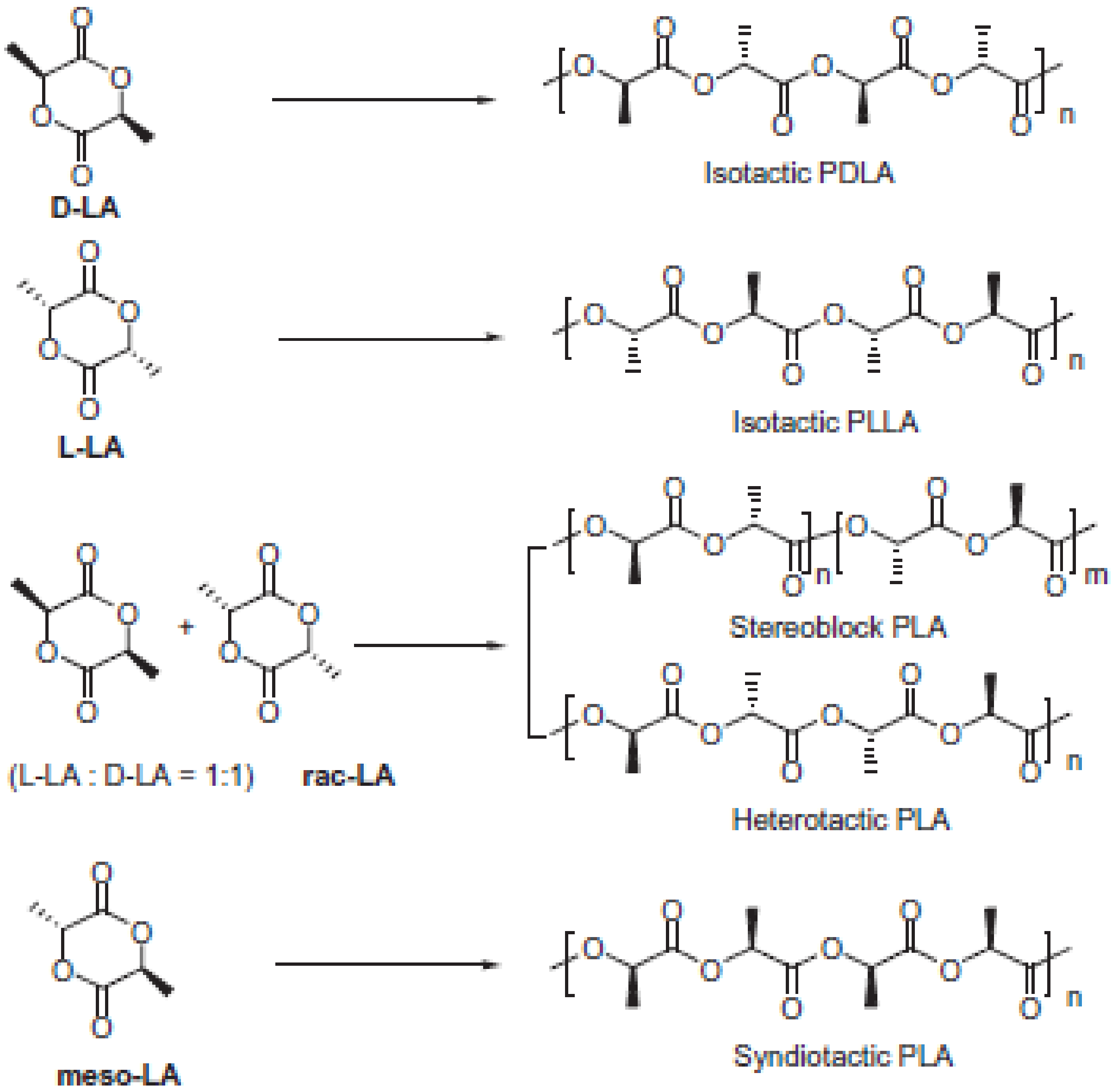


10a: Ln = Sc
10b: Ln = Lu
10c: Ln = Y



11a: Ln = Sc
11b: Ln = Lu
11c: Ln = Y

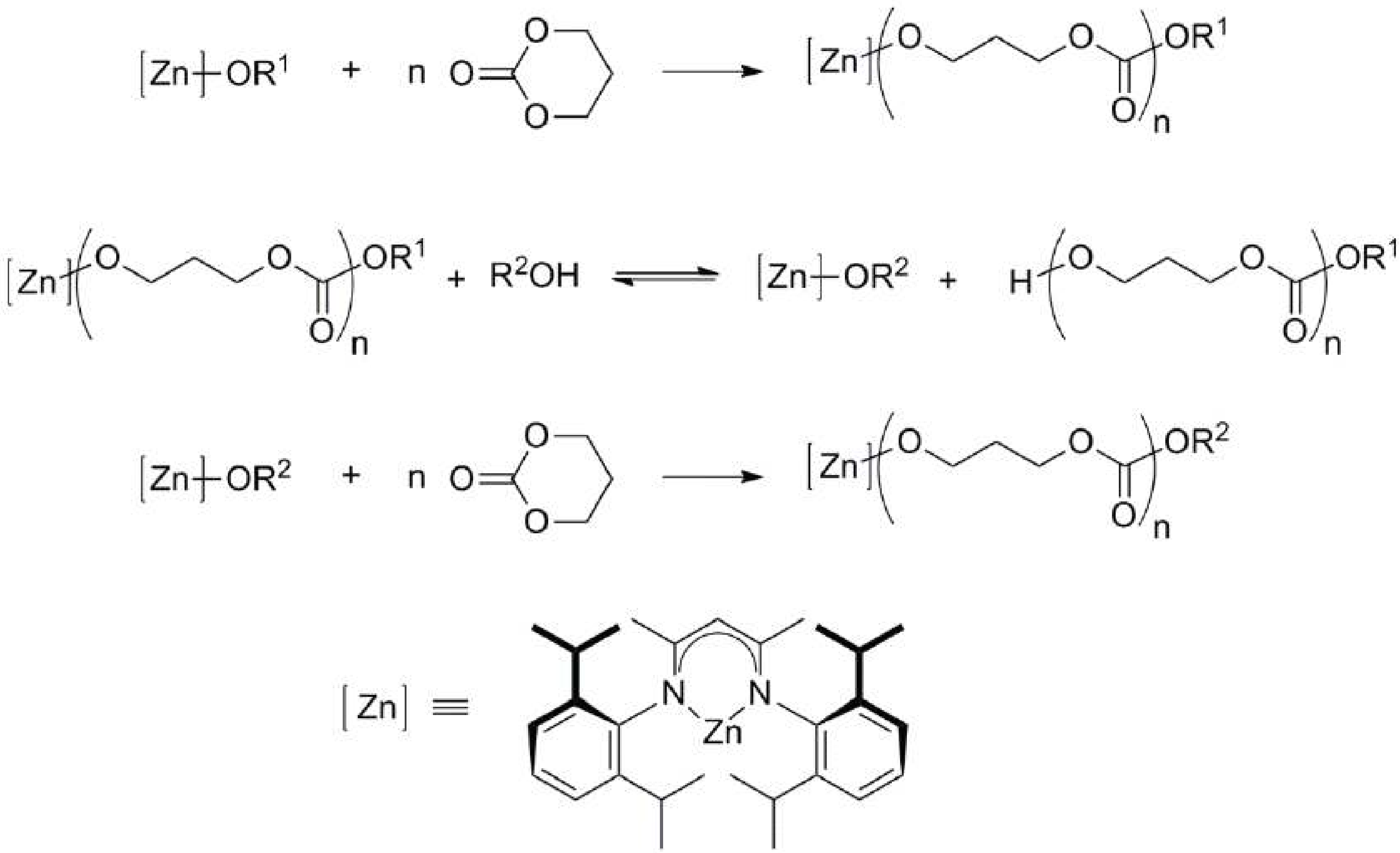
Synthesis of stereoregular PLAs by ROP



Crystalline
T_m **= 170-180 ºC**

Amorphous

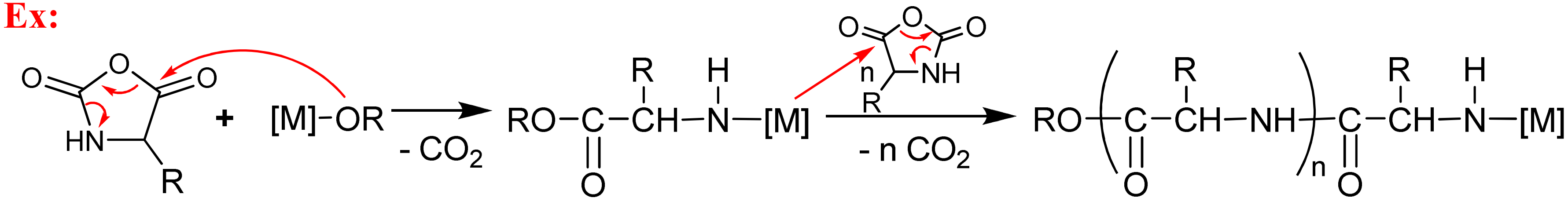
Synthesis of polycarbonates by ROP



[Zn] mediated living ROP of trimethylene carbonate

• *N*-CARBOXYANHYDRIDES

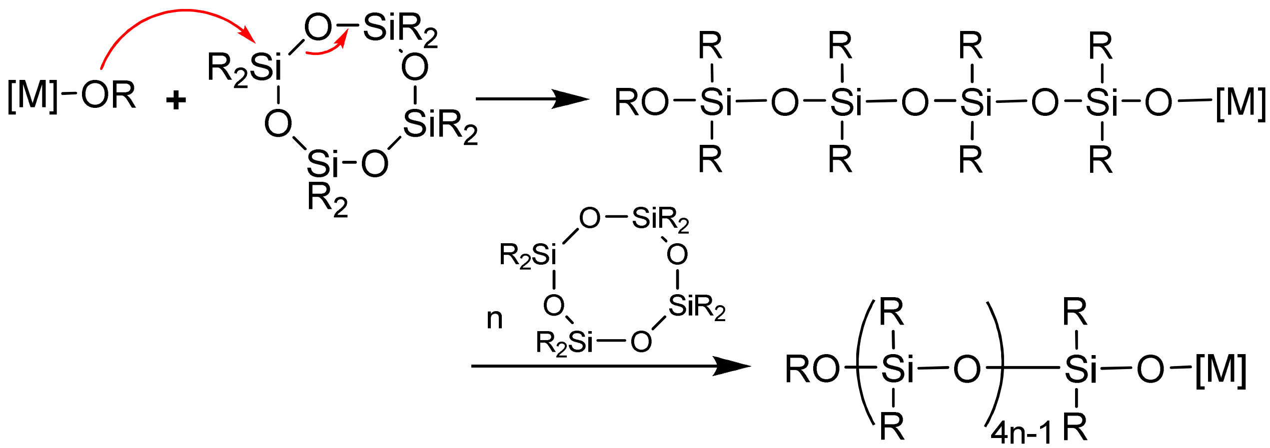
Ex:



Polypeptide (M_n ~10^6)

• CYCLIC SILOXANES

Ex:



Polysiloxane

Metal-Catalyzed Polymerization

(Coordination Polymerization)

- Polymerization of Olefins (Insertion) ✓ **Prof. Barbara Milani**

- Polymerization of Dienes (Insertion)

- Polymerization of Alkynes

- Ring Opening Metathesis Polymerization (ROMP)

- Classical Anionic Polymerization

- Ring Opening Polymerization (ROP)

- **Metal-mediated Radical Polymerization**

FREE RADICAL POLYMERIZATION

It involves:
- **Monomer**
  +
- **Radical Initiator**

STEPS

- **INITIATION**
  addition of the radical initiator to the monomer
- **PROPAGATION**
  radical chain growth by sequential addition of monomers
- **CHAIN TERMINATION**
  "death" of the radical propagating species by reaction with other radical species
- **CHAIN TRANSFER**
  growing chain reacts with a neutral molecule and abstracts one of its atoms, the latter becoming a new radical

---

INITIATION

$$I_2 \longrightarrow I\bullet + I\bullet$$

$$I\bullet + M \longrightarrow P_1\bullet$$

PROPAGATION

$$P_i\bullet + M \longrightarrow P_{i+1}\bullet$$

TERMINATION

$$P_i\bullet + P_j\bullet \longrightarrow P_{i+j}$$

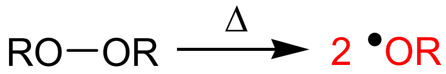
$$P_i\bullet + P_j\bullet \longrightarrow P_i + P_j$$

CHAIN TRANSFER

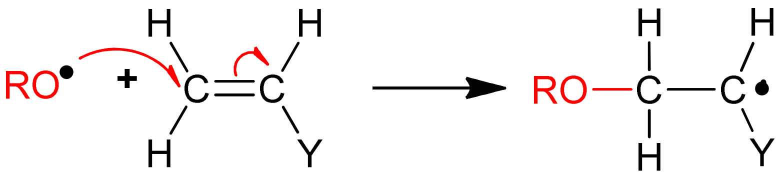
$$P_i\bullet + X \longrightarrow P_i + X\bullet$$

MECHANISM OF CLASSICAL ‘FREE RADICAL POLYMERIZATION’ (GENERAL)

• INITIATION



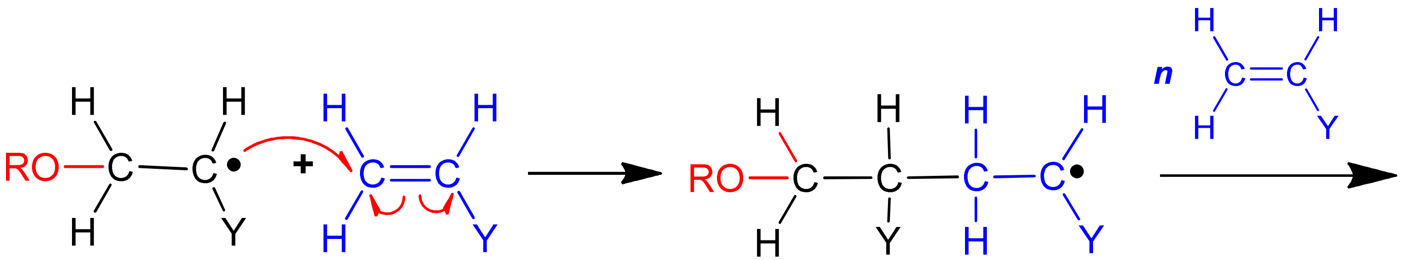
Initiator decomposition

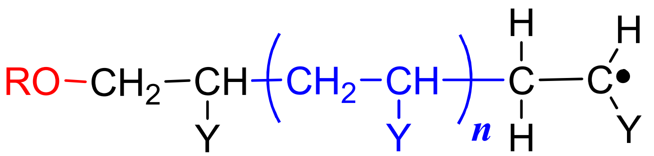


monomer

Addition of initiator radical to monomer

• PROPAGATION





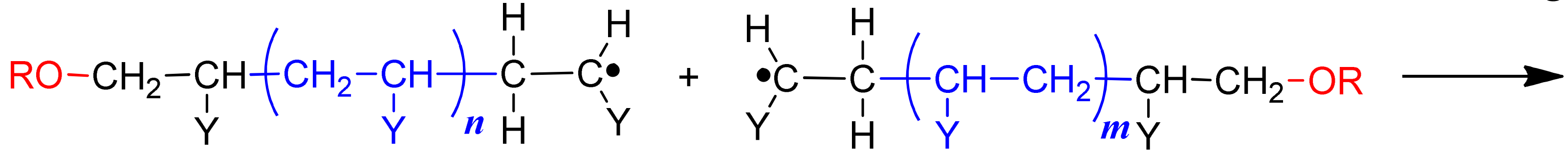
mer

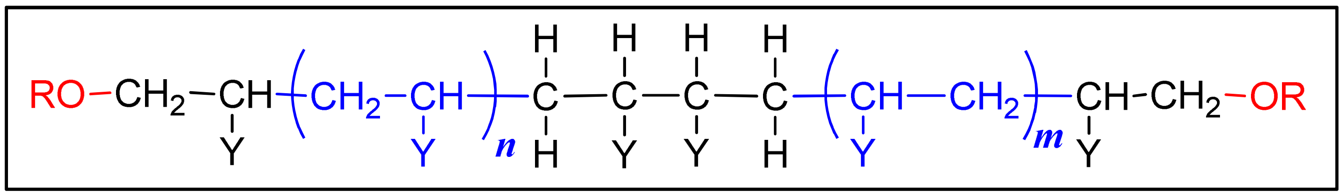
Sequential addition of monomers to radical growing chain → CHAIN GROWTH

• TERMINATION

• Recombination (or Coupling) (*low* temperatures)

Recombination of 2 radical growing chains

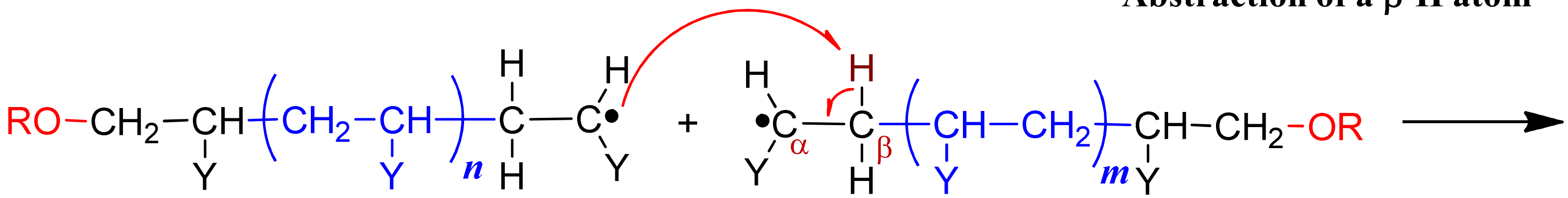


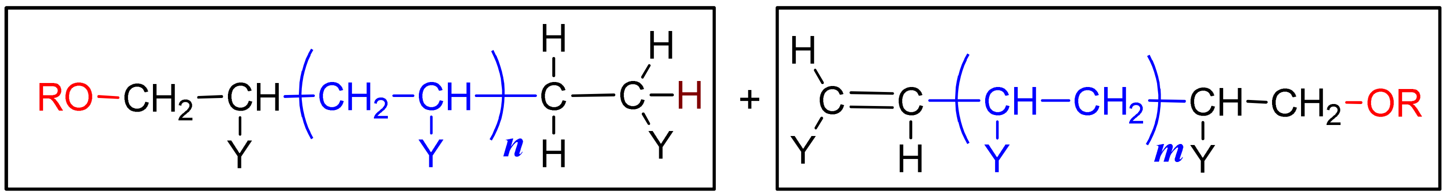


1 neutral chain
(molecular weight doubles)

• Disproportionation (*high* temperatures)

Abstraction of a β-H atom



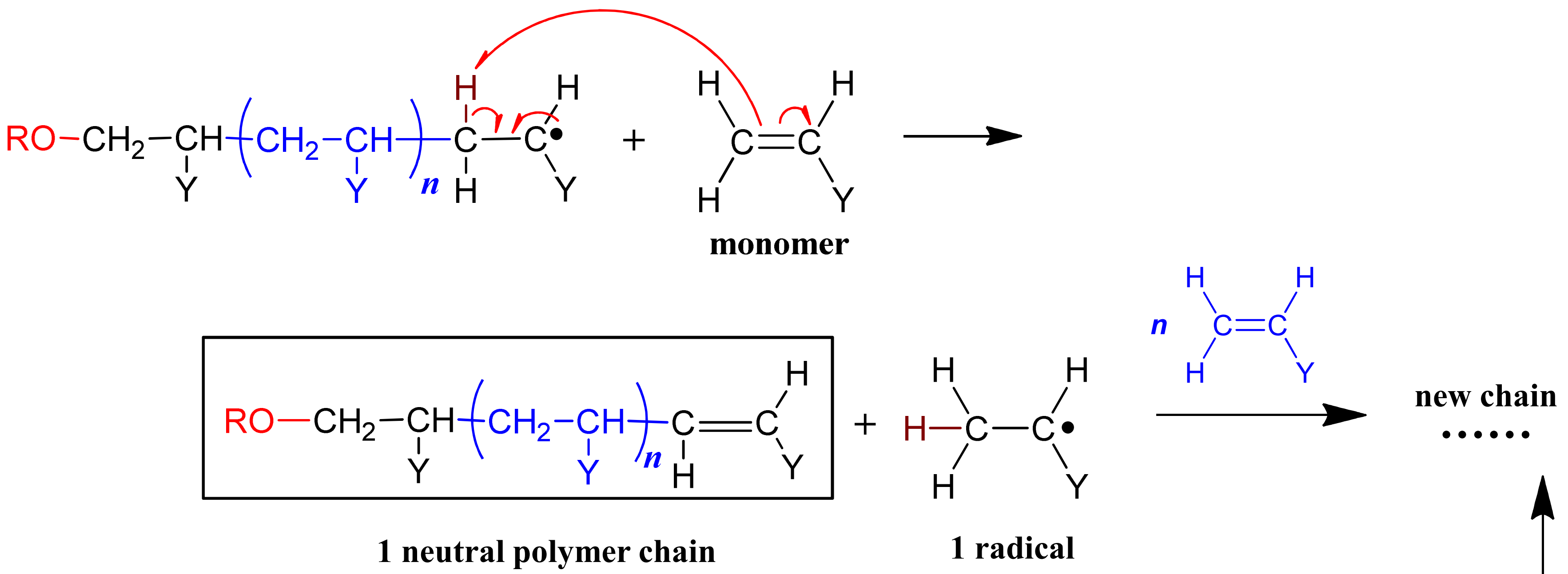


2 neutral polymer chains

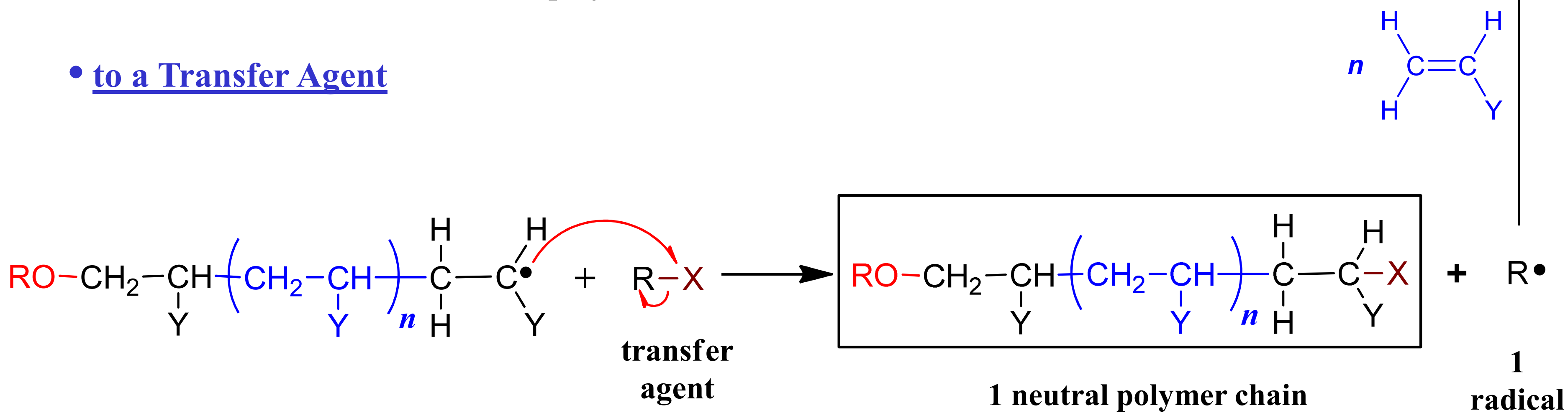
• CHAIN TRANSFER

Step responsible for the decrease in molecular weight
and for the broadening of the molecular weight distribution

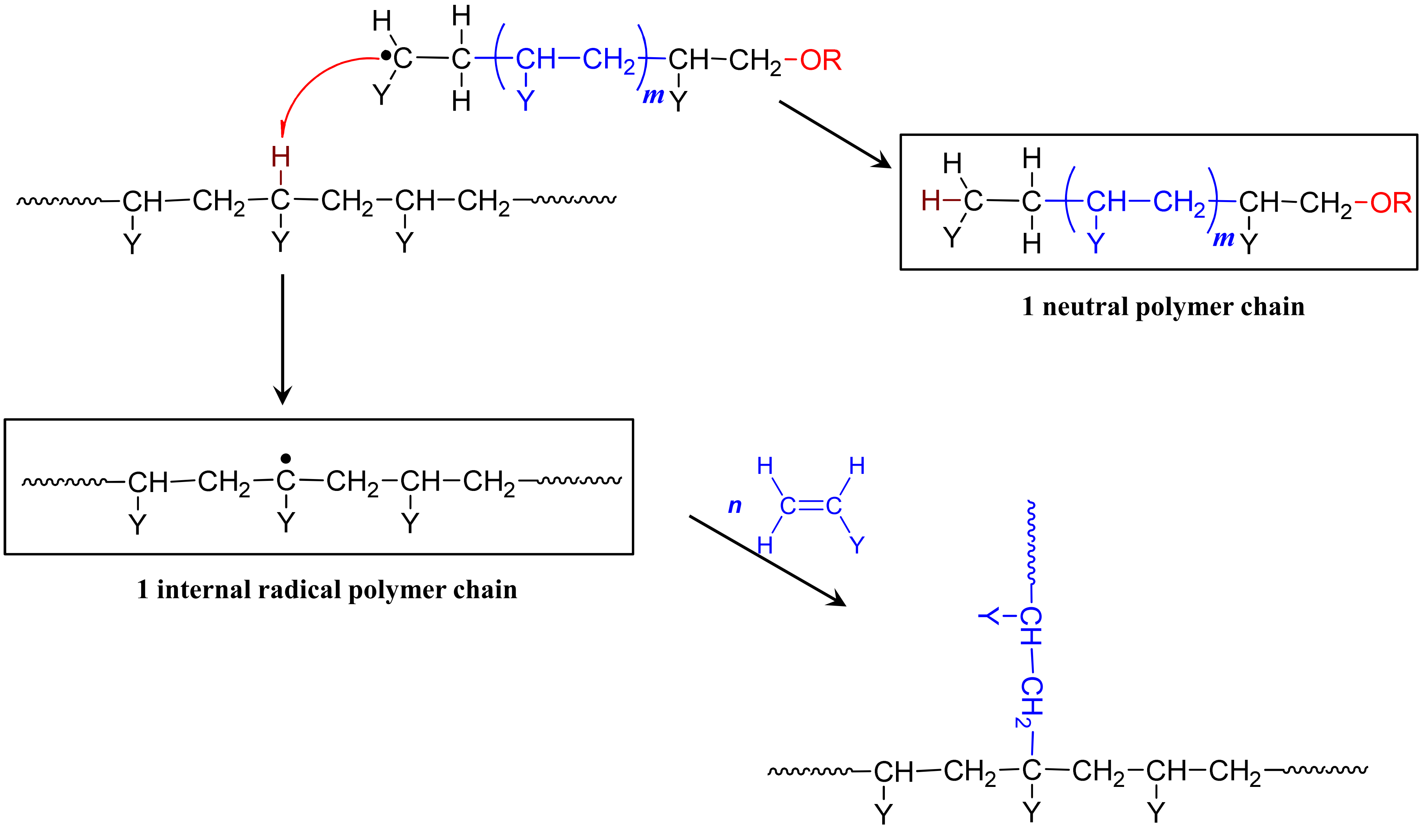
• to the Monomer



• to a Transfer Agent



• **to the Polymer (*very high* temperatures)**



BRANCHING

CONTROLLED RADICAL POLYMERIZATION (CRP)

Absence of TERMINATION or CHAIN TRANSFER

⬇

Living Polymers (Living Polymerization) → Living Chain Ends

⬇

Molecular Weight Control → Narrow Molecular Weight Distribution

⬇

Block Copolymers

Several Techniques of CRP (examples)

- ATRP
- OMRP
- NMP
- RAFT

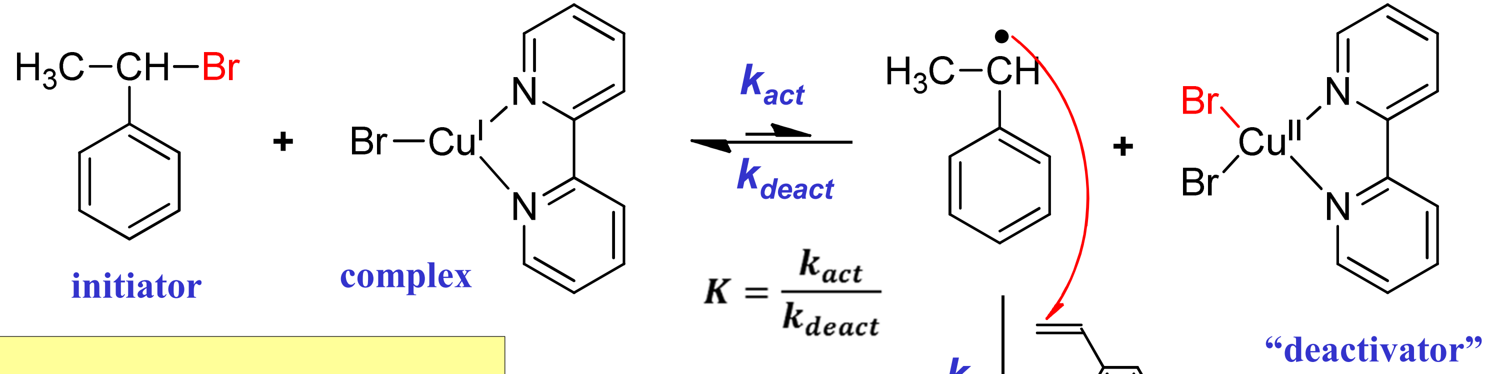
(OMRP, NMP, RAFT: SFRP)

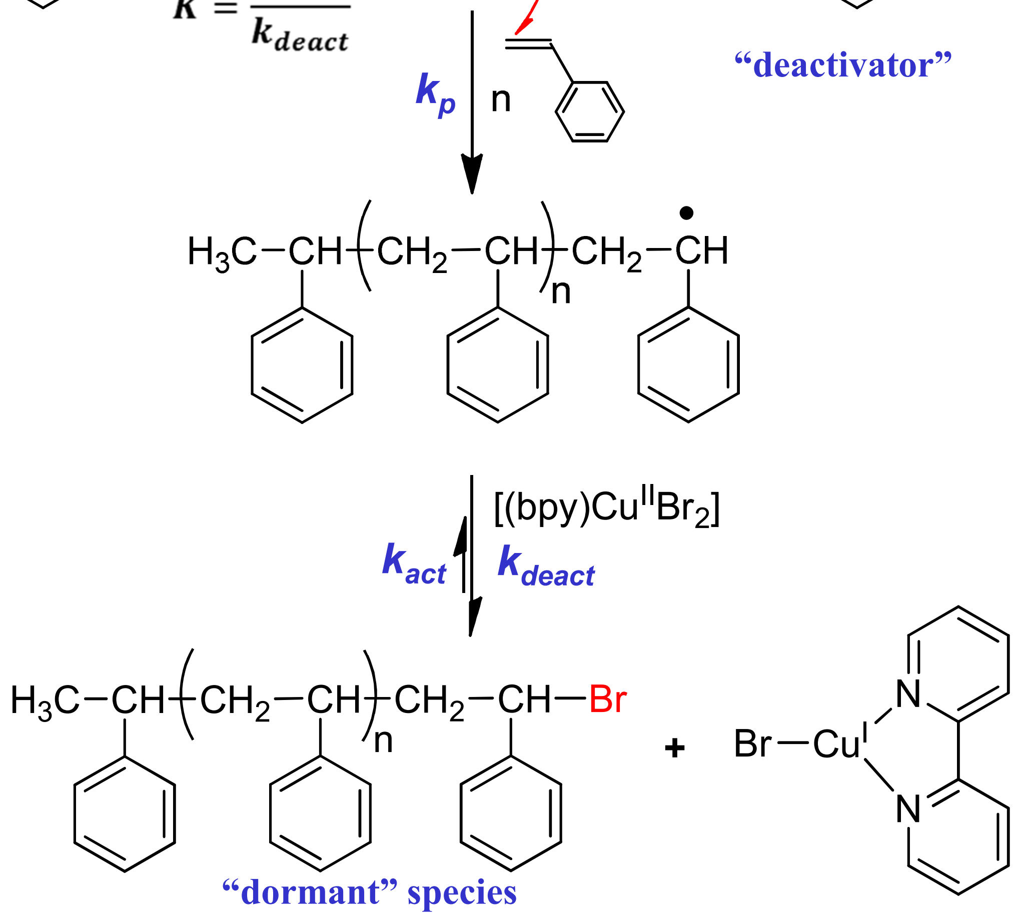
SFRP = Stable Free Radical Polymerization

ATOM-TRANSFER RADICAL POLYMERIZATION (ATRP)

Initiator: Organic halide undergoing a reversible redox catalyzed by a transition-metal complex

Ex:





$$r_p = \frac{k_p K[\mathrm{I}][\mathrm{Cu}^+]}{[\mathrm{Cu}^{2+}]}[\mathrm{M}]$$

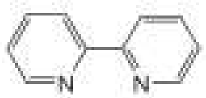
$$\ln\frac{[\mathrm{M}]_o}{[\mathrm{M}]} = \frac{k_p K[\mathrm{I}][\mathrm{Cu}^+]}{[\mathrm{Cu}^{2+}]}t$$

$$\overline{DP}_n = \frac{[\mathrm{M}]_o - [\mathrm{M}]}{[\mathrm{I}]_o} = \frac{p[\mathrm{M}]_o}{[\mathrm{I}]_o}$$

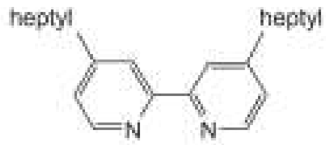
$$\frac{\bar{M}_w}{\bar{M}_n} = 1 + \frac{1}{\overline{DP}_n}$$

Poisson distribution

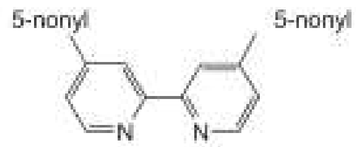
• COMMON LIGANDS (L) OF ATRP



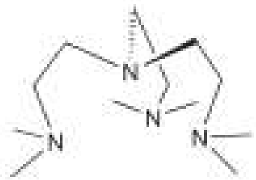
bipy



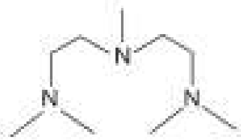
dHbipy



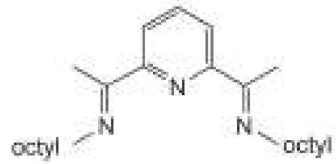
dNbipy



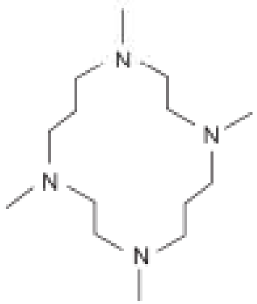
Me_6TREN



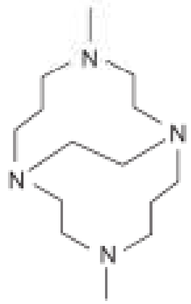
PMDETA



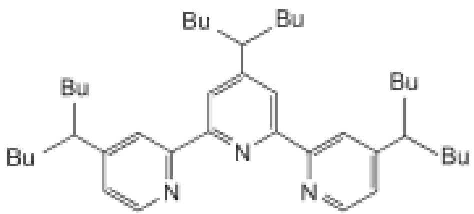
DOIP



Me_4Cyclam



Cyclam-B



TERPY

TABLE 8.1. Commercially Available Polymers Synthesized with Complex Coordination Catalysts

Polymer	Principal Stereochemistry	Typical Uses
Plastics		
Polyethylene, high density (HDPE)	—	Bottles, drums, pipe, conduit, sheet, film, wire and cable insulation
Polyethylene, ultrahigh molecular weight (UHMWPE)	—	Surgical prostheses, machine parts, heavy-duty liners
Polypropylene	Isotactic	Automobile and appliance parts, rope, cordage, webbing, carpeting, film
Poly(1-butene)	Isotactic	Film, pipe
Poly(4-methyl-1-pentene)[a]	Isotactic	Packaging, medical supplies, lighting
Polystyrene	Syndiotactic	Specialty plastics
1,4-Polybutadiene	*trans*	Metal can coatings, potting compounds for transformers
1,4-Polyisoprene	*trans*	Golf ball covers, orthopedic devices
Ethylene-1-alkene[b] copolymer (linear low-density polyethylene, LLDPE)	—	Blending with LDPE, packaging film, bottles
Ethylene-propylene block copolymers (polyallomers)	Isotactic	Food packaging, automotive trim, toys, bottles, film, heat-sterilizable containers
Polydicyclopentadiene[c]	—	Reaction injection molding (RIM) structural plastics
Elastomers		
1,4-Polybutadiene	*cis*	Tires, conveyer belts, wire and cable insulation, footware
1,4-Polyisoprene	*cis*	Tires, footware, adhesives, coated fabrics
Poly(1-octenylene) (polyoctenamer)[c]	*trans*	Blending with other elastomers
Poly(1,3-cyclopentenylenevinylene) (norbornene polymer)[c]	*trans*	Molding compounds, engine mounts, car bumper guards
Polypropylene (amorphous)	—	Asphalt blends, sealants, adhesives, cable coatings
Ethylene–propylene copolymer (EPM, EPR)	—	Impact modifier for polypropylene, car bumper guards
Ethylene–propylene–diene copolymer (EPDM)	—	Wire and cable insulation, weather stripping, tire side walls, hose, seals

[a] Usually copolymerized with small amounts of 1-pentene.

[b] 1-Butene, 1-hexene, and 1-octene.

[c] Synthesized by ring-opening metathesis polymerization of the corresponding cycloalkene.

Table 4.2 Initiation modes of various monomers

Monomer	Initiator			
	Free radical	Anionic	Cationic	Co-ordination
Ethylene	√			√
Propylene (and other α-olefins, R)				√
Isobutylene			√	
Styrene	√	√	√	√
Butadiene and isoprene (R)	√	√		√
Acrylates and methacrylates (COOR)	√	√		
Acrylonitrile (CN)	√	√		
Vinyl ethers (OR)			√	
Vinyl halides (Hal)	√			
Fluorocarbons (e.g. TFE, $CF_2{=}CF_2$)	√			
Vinyl esters (e.g. acetate, $OCOCH_3$)	√			
Formaldehyde ($CH_2{=}O$)			√	
Formaldehyde trimer (trioxan)		√		√
Ethylene oxide (CH_2–CH_2, O)		√		√
Cyclic ethers (e.g. THF)			√	√
Cyclic lactams and lactones (CONH, $(CH_2)_n$; COO, $(CH_2)_n$)		√		√
Cyclic siloxanes ($R_2SiO_{3\ or\ 4}$)		√		
Cycloalkenes and cycloalkynes				√
Alkynes (≡–R)				√

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