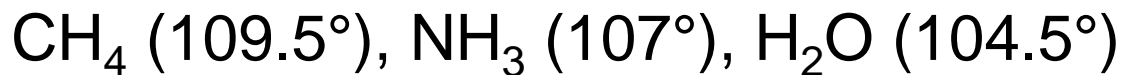


Table 2.6 The basic arrangement of regions of electron density according to the VSEPR model

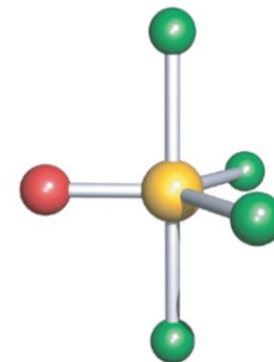
Number of electron regions	Arrangement
2	Linear
3	Trigonal planar
4	Tetrahedral
5	Trigonal bipyramidal
6	Octahedral

Repulsione

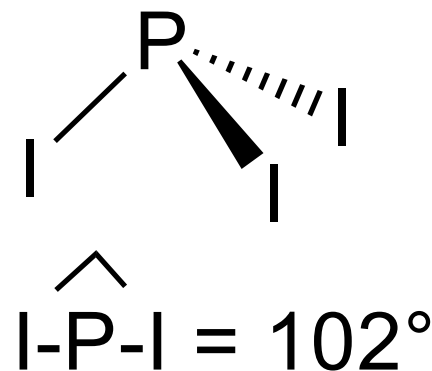
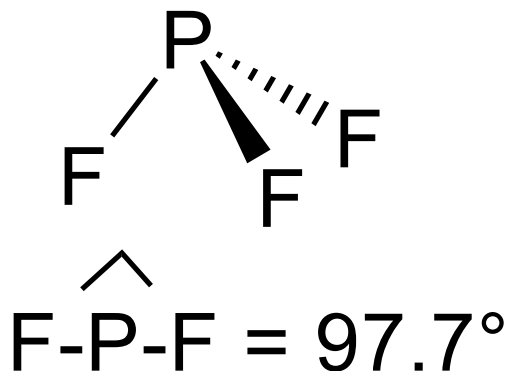
coppia solitaria – coppia solitaria > coppia solitaria – coppia di legame > coppia di legame – coppia di legame



legame triplo – legame singolo > legame doppio – legame singolo > legame singolo – legame singolo



coppie di legame verso atomi (o gruppi) elettronegativi occupano meno spazio di quelle verso atomi più elettropositivi



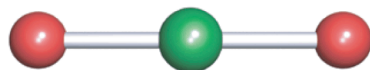
numero sterico = numero di coppie di legame e di non legame intorno all'atomo centrale

n.s. = elettroni di valenza (\pm carica)/8 + resto/2

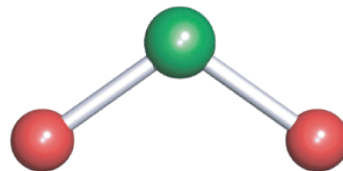


*Il numero sterico determina la **geometria delle coppie**, che non necessariamente coincide con la **geometria molecolare***

2-Coordinate

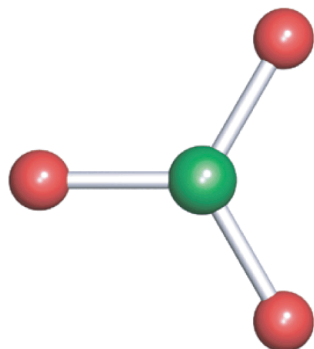


Linear

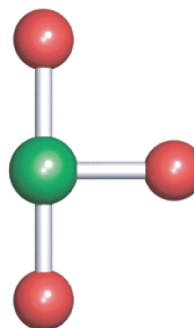


Bent

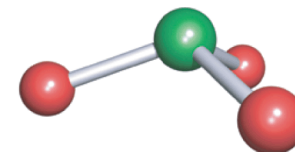
3-Coordinate



Trigonal planar



T-shaped

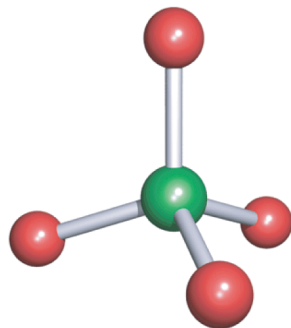


Trigonal pyramidal

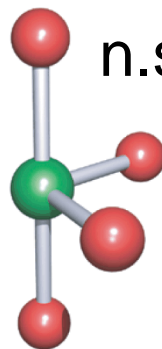


n.s. = 5

4-Coordinate

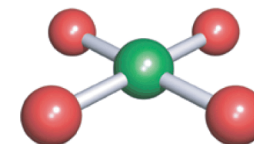
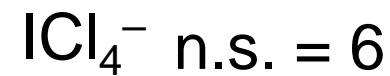


Tetrahedral



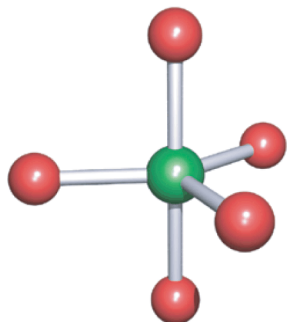
n.s. = 5

Disphenoidal

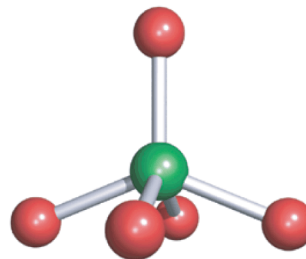


Square planar

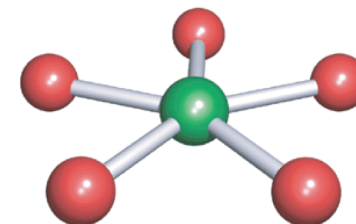
5-Coordinate



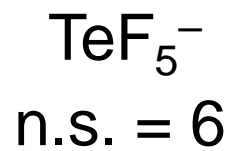
Trigonal bipyramidal



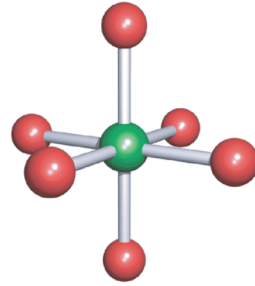
Square-based pyramidal



Pentagonal planar

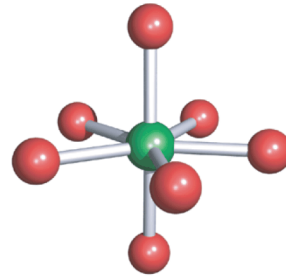


6-Coordinate



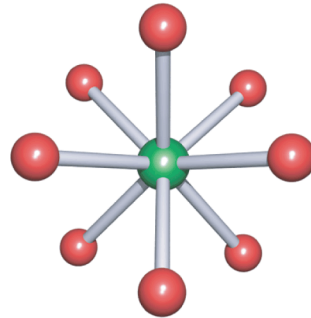
Octahedral

7-Coordinate



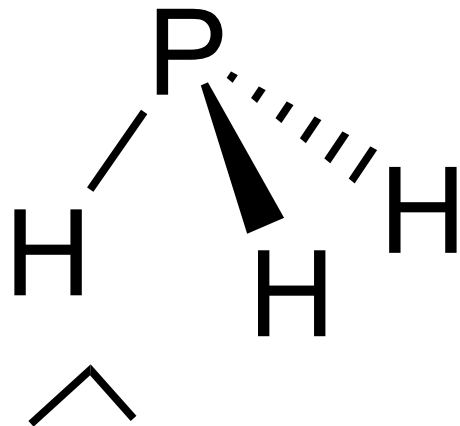
Pentagonal bipyramidal

8-Coordinate



Square antiprismatic

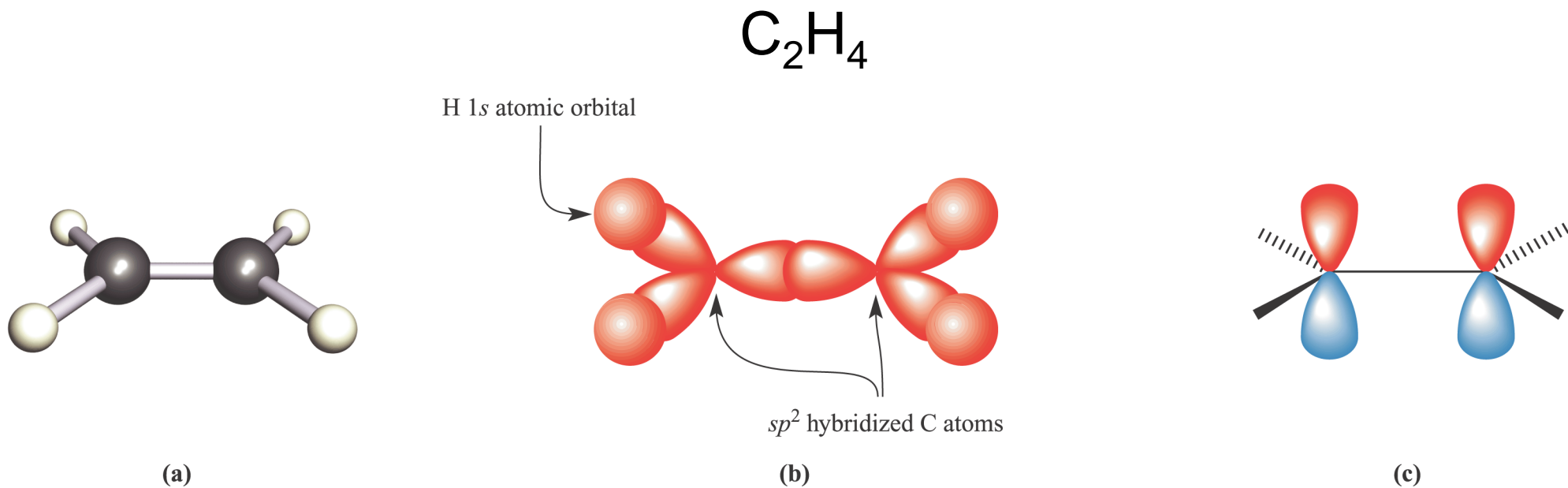
Coppie solitarie stereochimicamente inattive



H-P-H ca. 90°

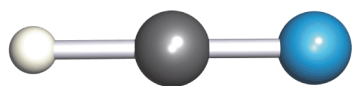


Molecole poliatomiche secondo la teoria VB

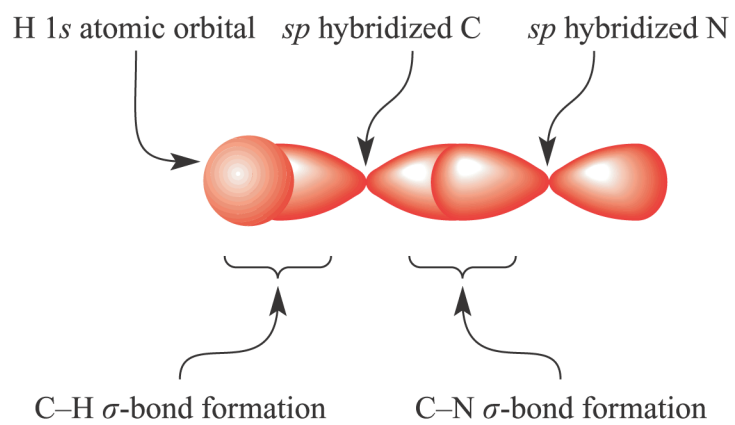


598 kJ mol^{-1} vs $2 \times 346 \text{ kJ mol}^{-1}$

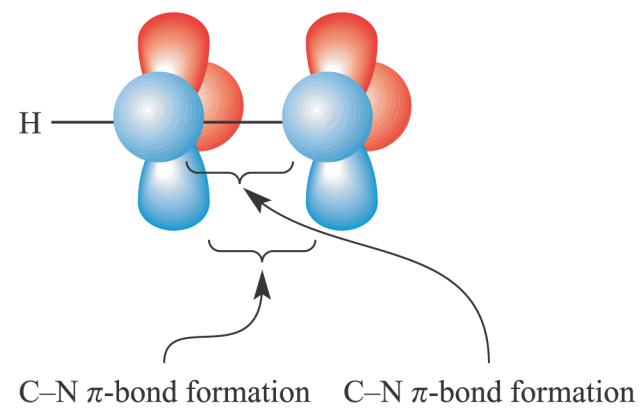
HCN



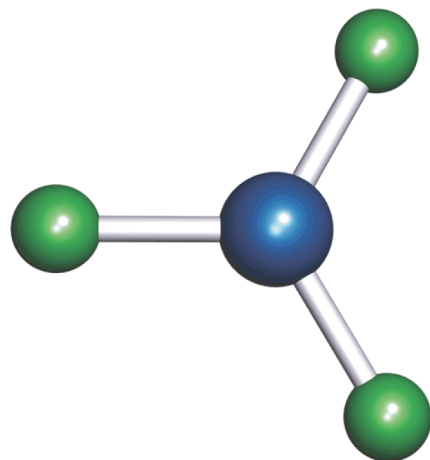
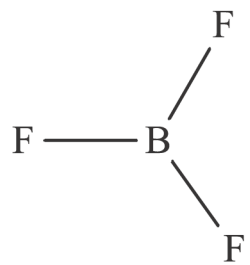
(a)



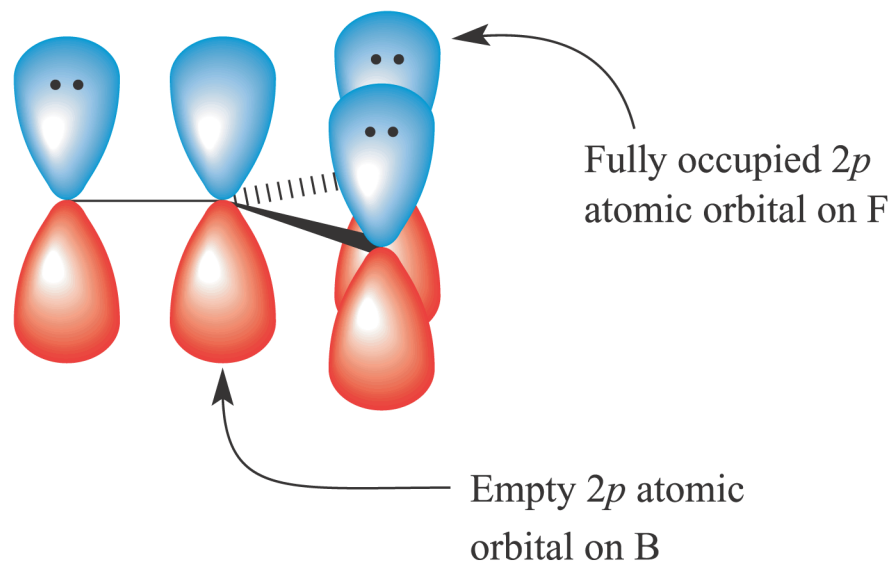
(b)



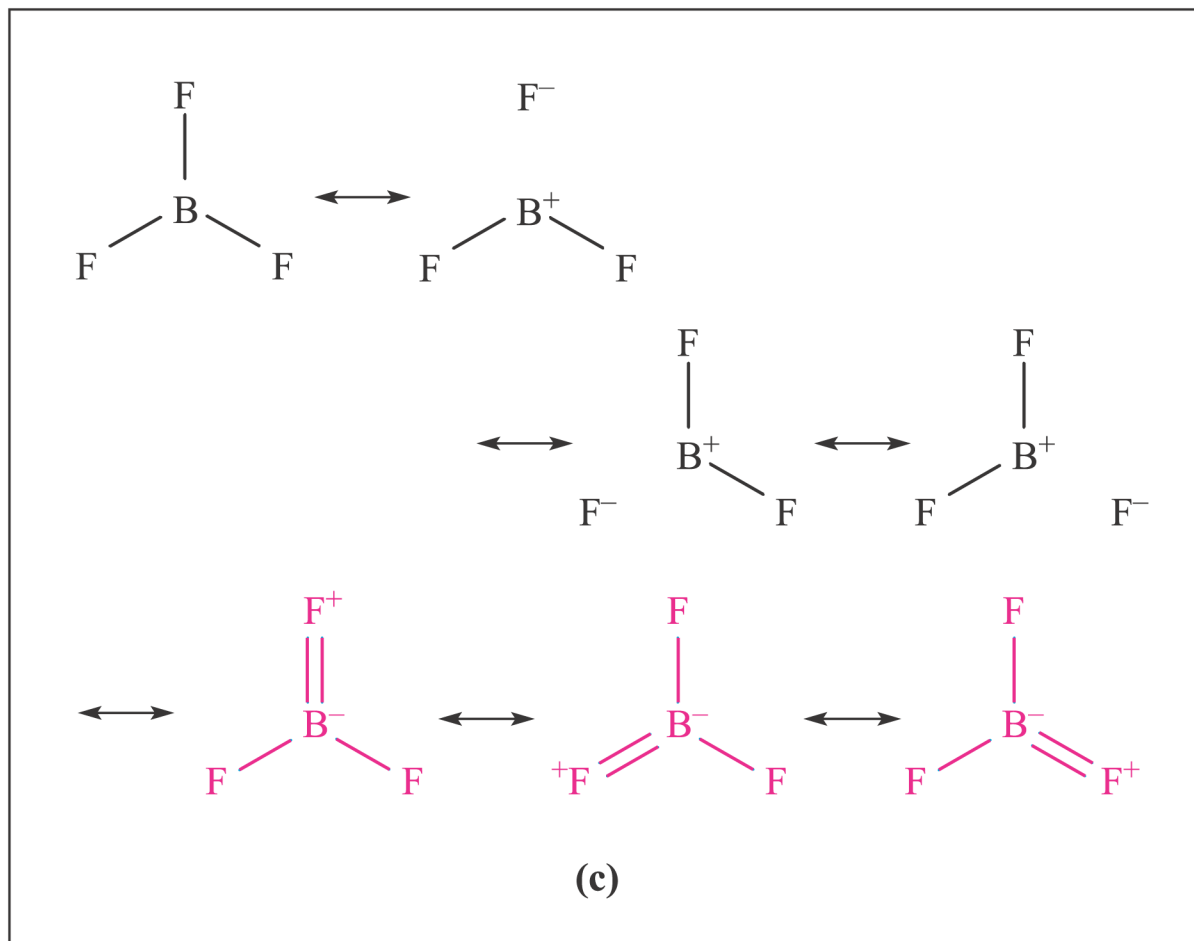
(c)



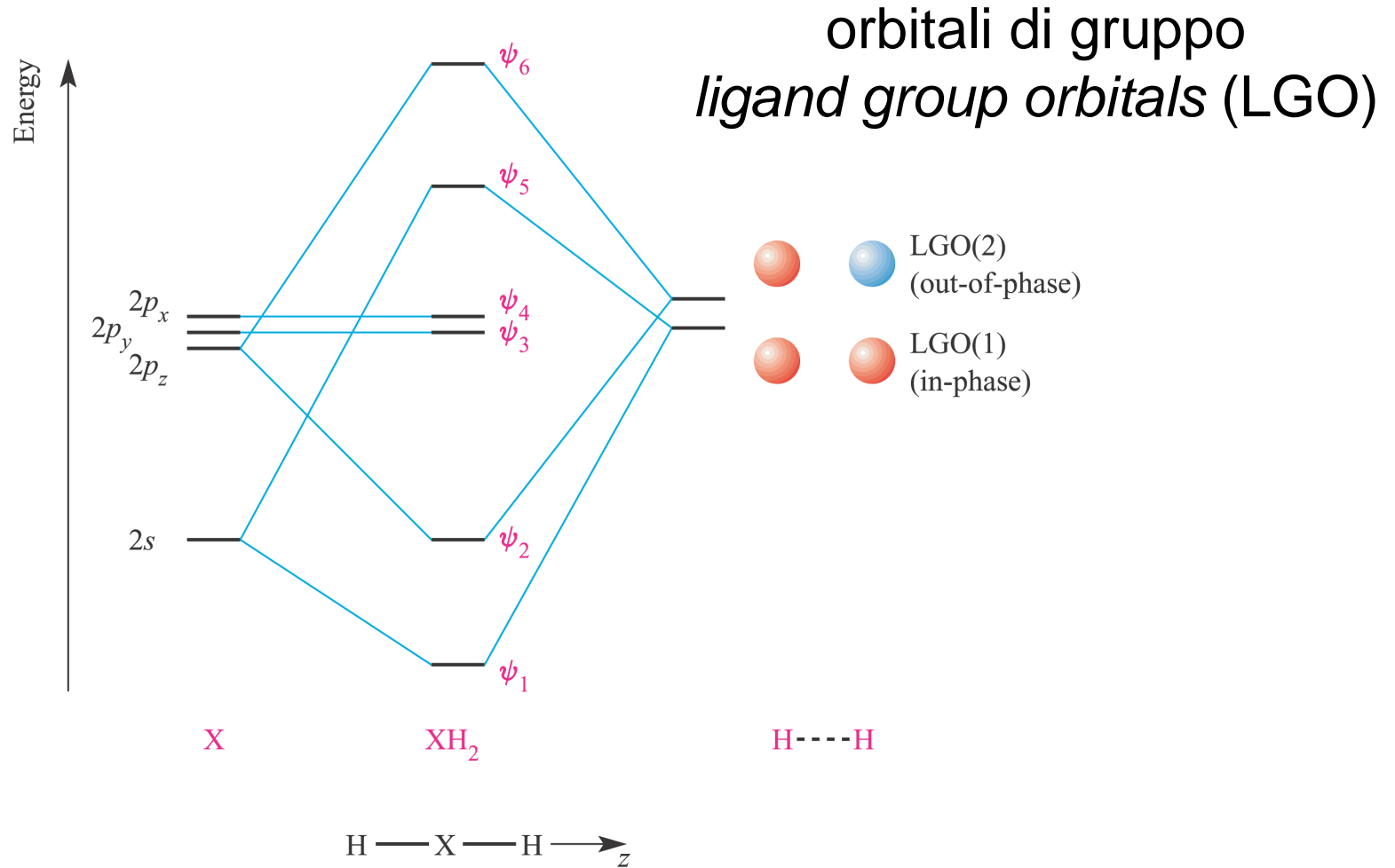
(a)



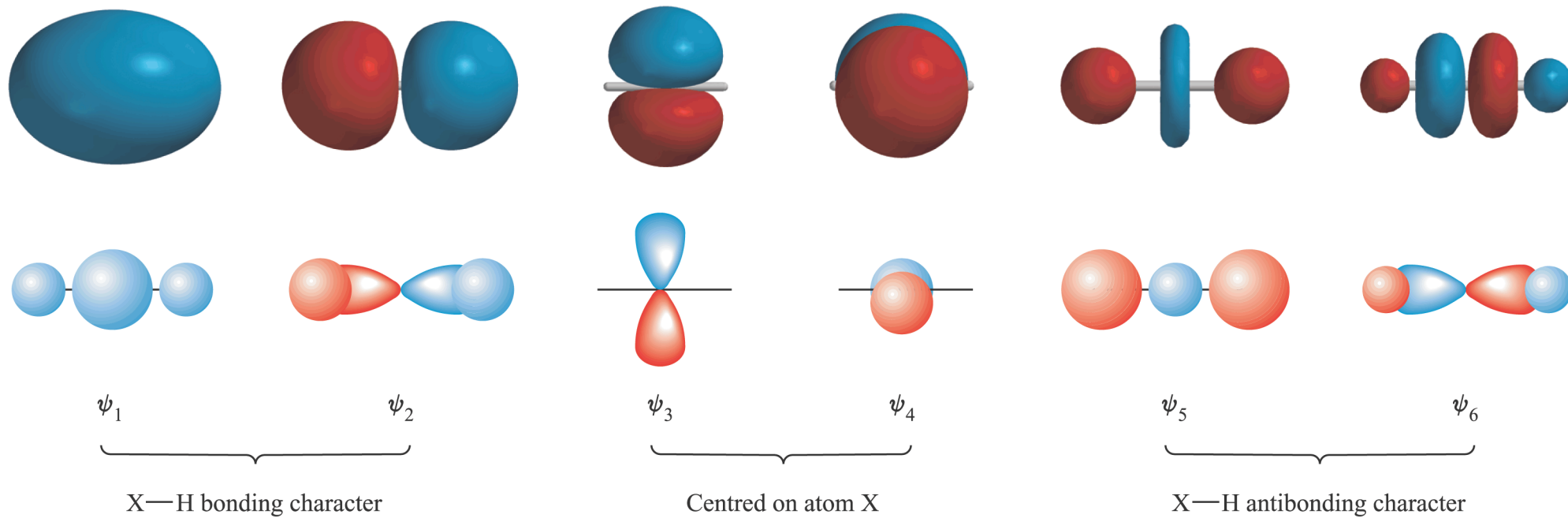
(b)

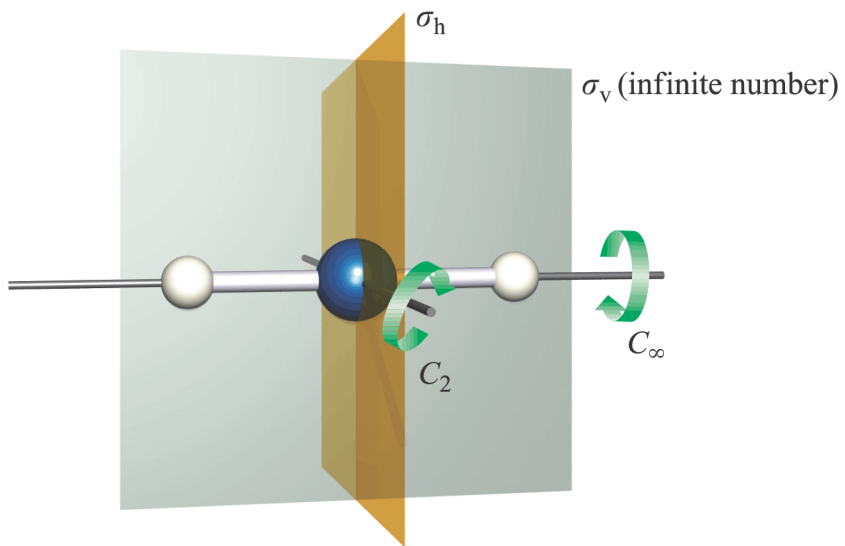


Molecole poliatomiche secondo la teoria MO



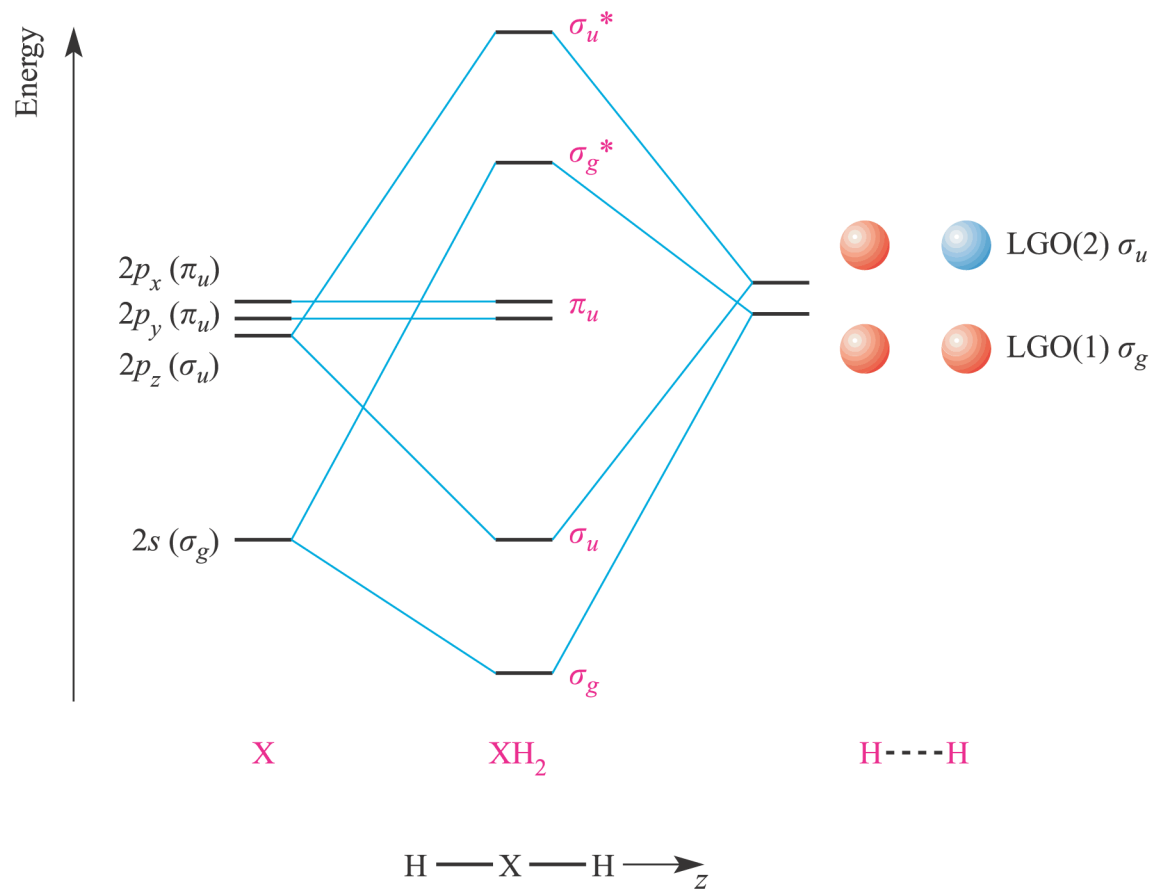
BeH₂



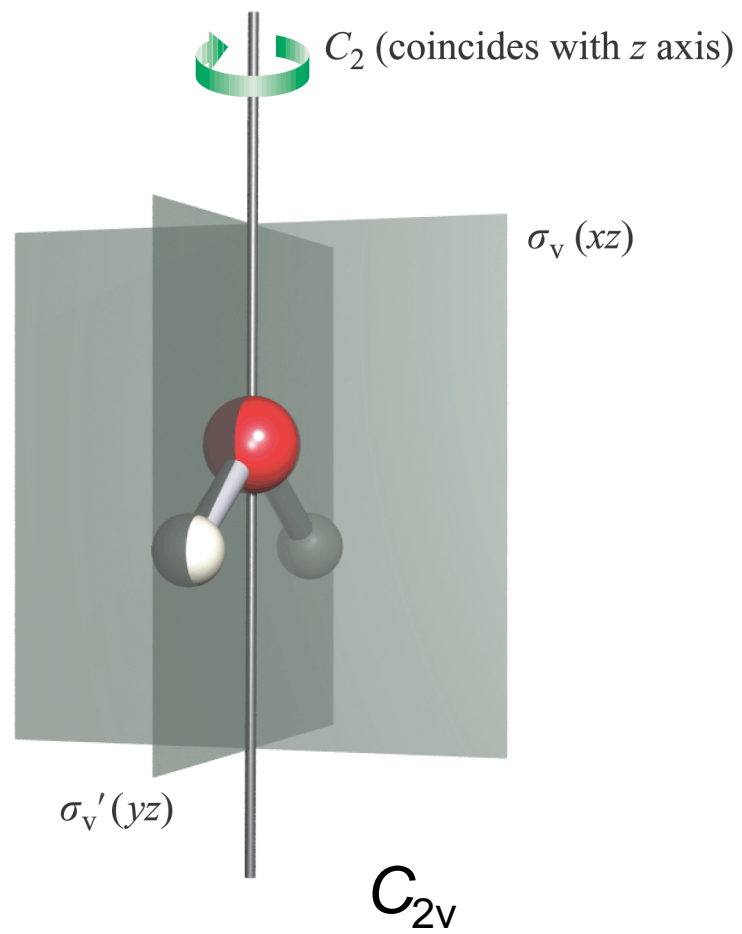


$D_{\infty h}$

(a)



(b)



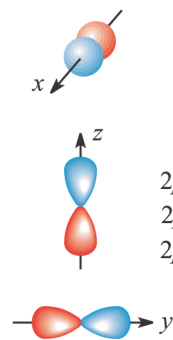
C_{2v}	E	C_2	$\sigma_v(xz)$	$\sigma_v'(yz)$		
A_1	1	1	1	1	z	x^2, y^2, z^2
A_2	1	1	-1	-1	R_z	xy
B_1	1	-1	1	-1	x, R_y	xz
B_2	1	-1	-1	1	y, R_x	yz

2H

2	0	0	2
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$A_1 + B_2$

Energy ↑

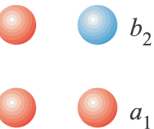


$2p_x (b_1)$
 $2p_z (a_1)$
 $2p_y (b_2)$

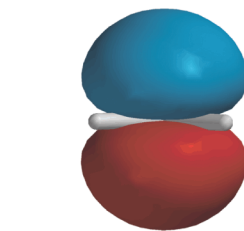
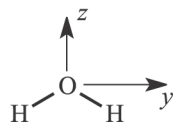
$2s (a_1)$

O

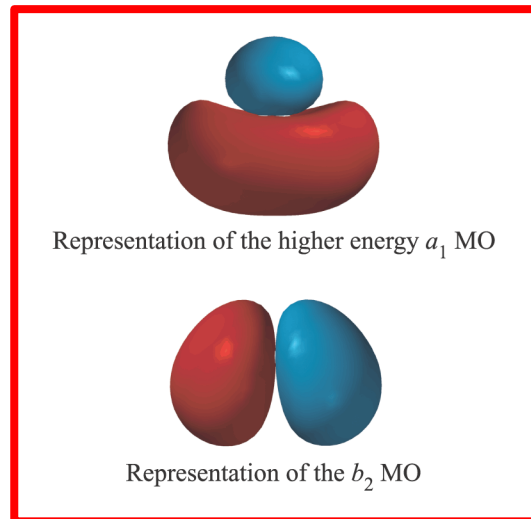
H₂O



H---H

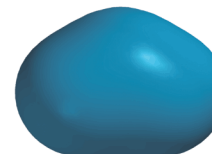


Representation of the b_1 MO



Representation of the higher energy a_1 MO

Representation of the b_2 MO

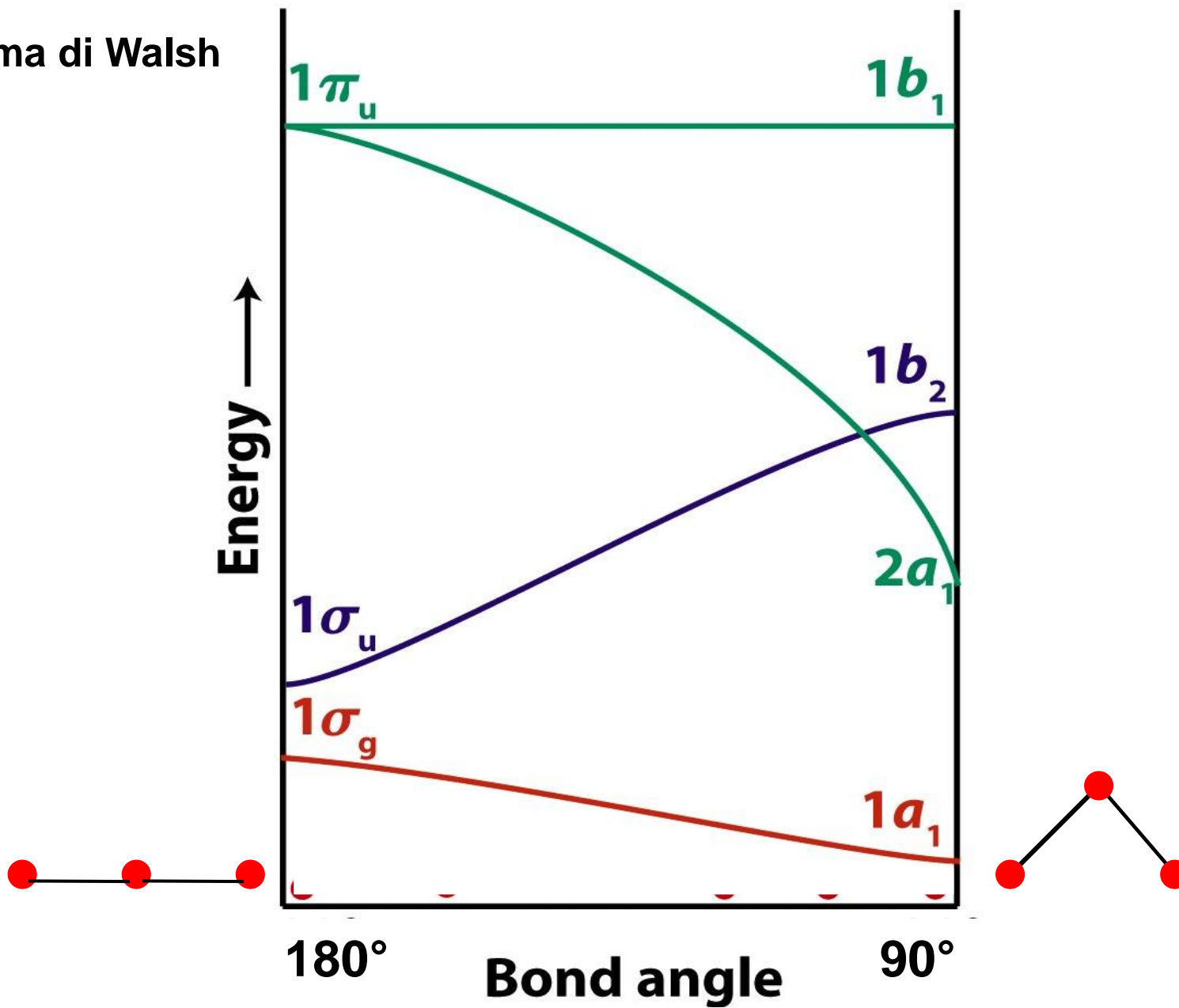


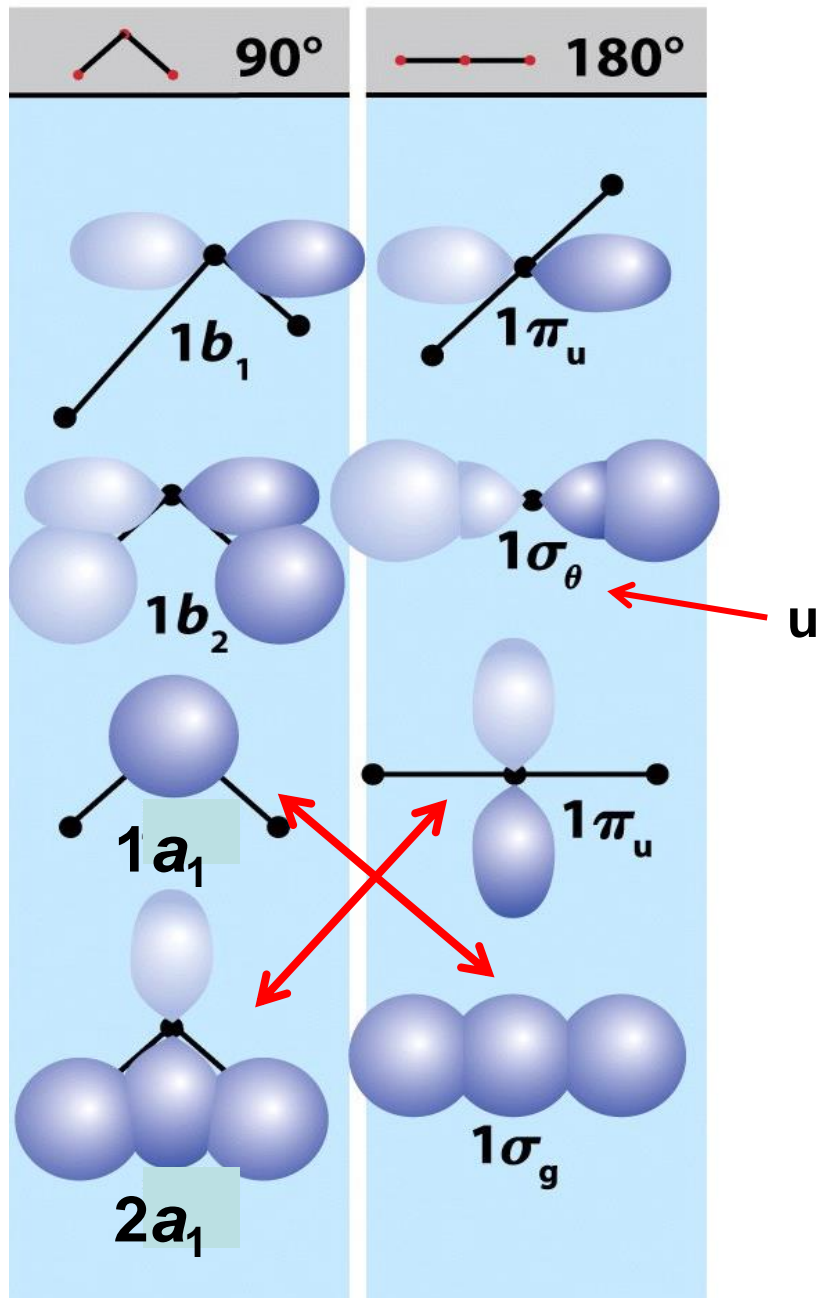
Representation of the lower energy a_1 MO

n.l.

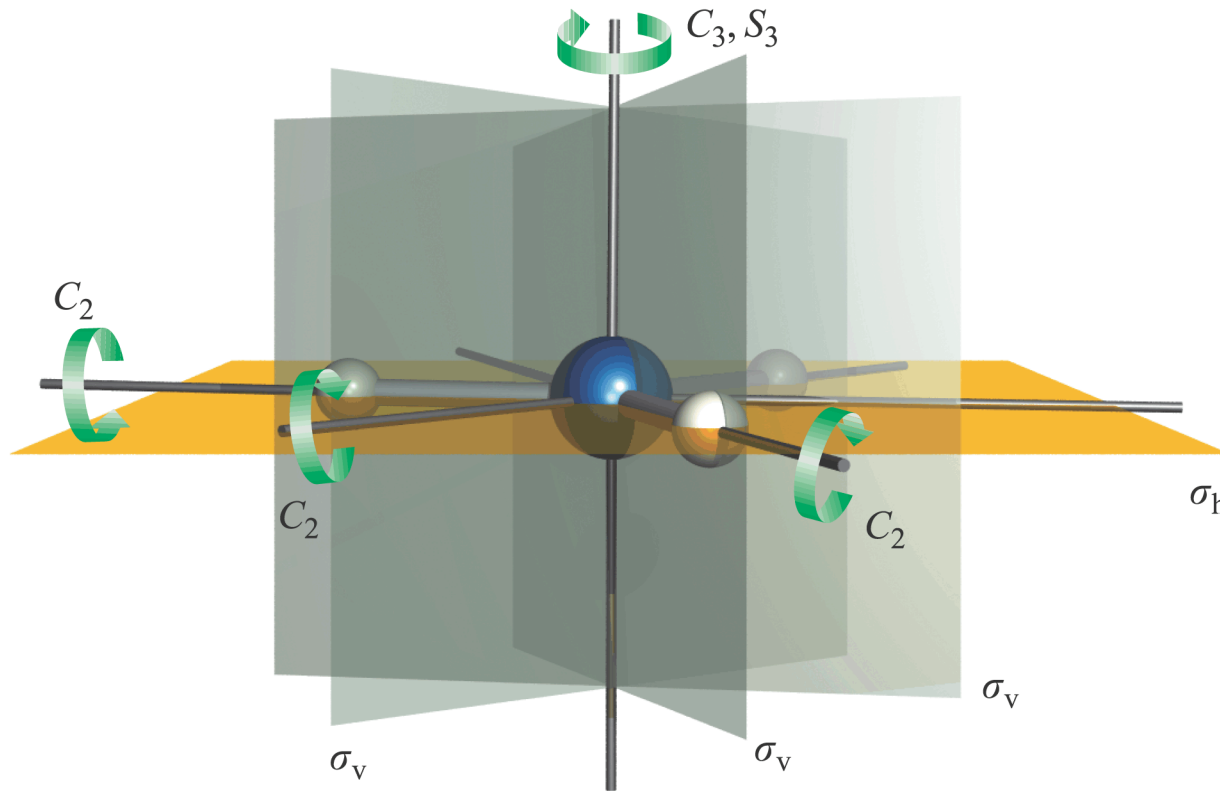
n.l.

Diagramma di Walsh





BH₃, gruppo puntuale D_{3h}

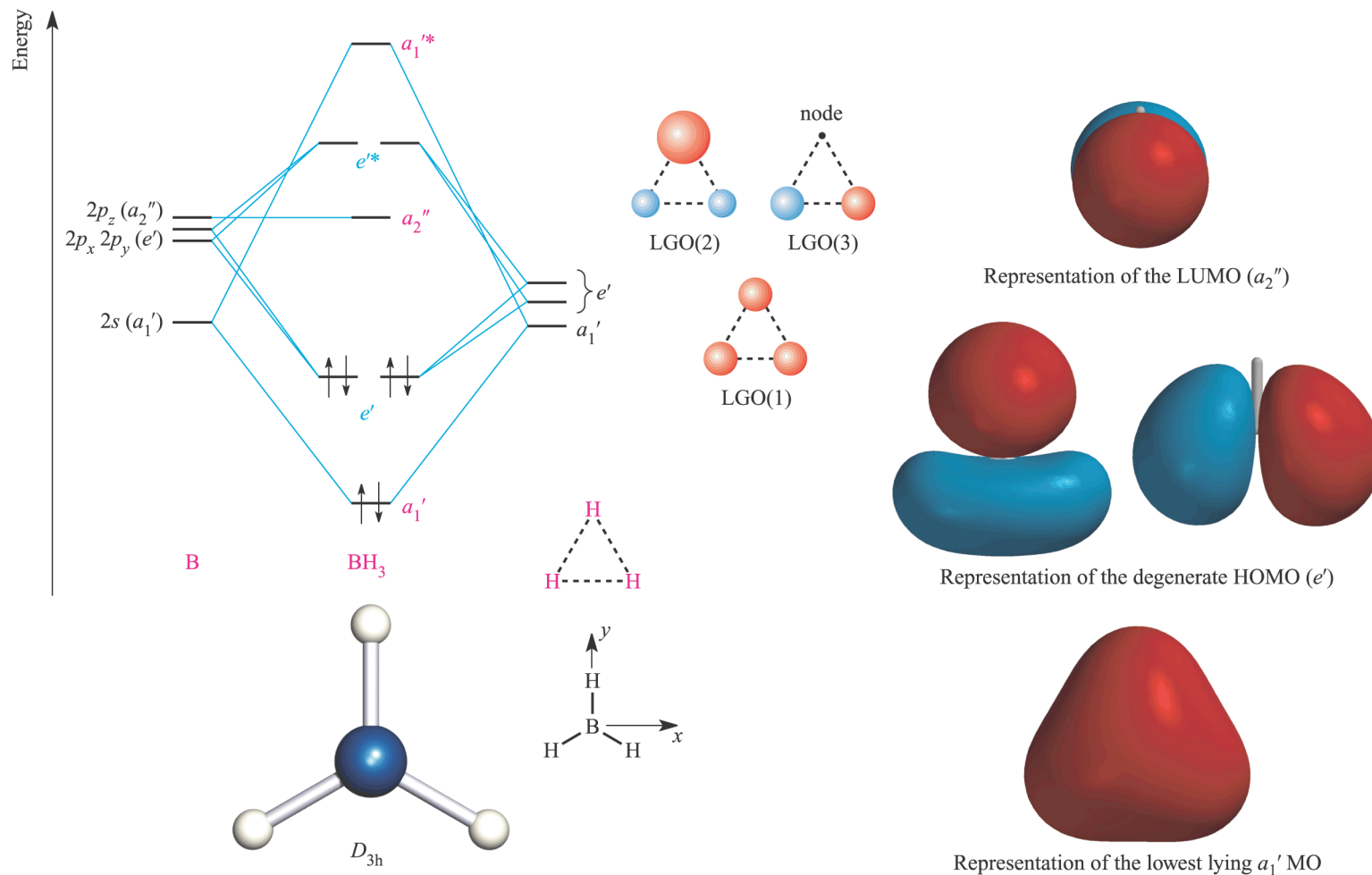


D_{3h}	E	$2C_3$	$3C_2$	σ_h	$2S_3$	$3\sigma_v$		
A_1'	1	1	1	1	1	1		$x^2 + y^2, z^2$
A_2'	1	1	-1	1	1	-1	R_z	
E'	2	-1	0	2	-1	0	(x, y)	$(x^2 - y^2, xy)$
A_1''	1	1	1	-1	-1	-1		
A_2''	1	1	-1	-1	-1	1	z	
E''	2	-1	0	-2	1	0	(R_x, R_y)	(xz, yz)

3H

3	0	1	3	0	1
---	---	---	---	---	---

$A_1' + E'$



$$\Psi(a_1') = 1/\sqrt{3} \times (\Psi_1 + \Psi_2 + \Psi_3)$$

(LGO1)

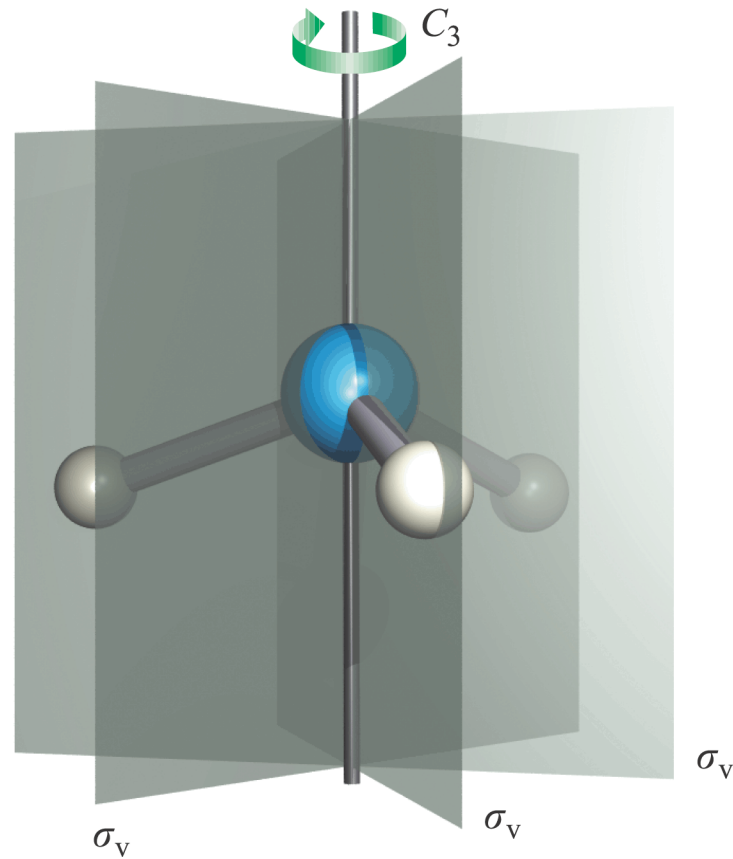
$$\Psi(e')_1 = 1/\sqrt{6} \times (2\Psi_1 - \Psi_2 - \Psi_3)$$

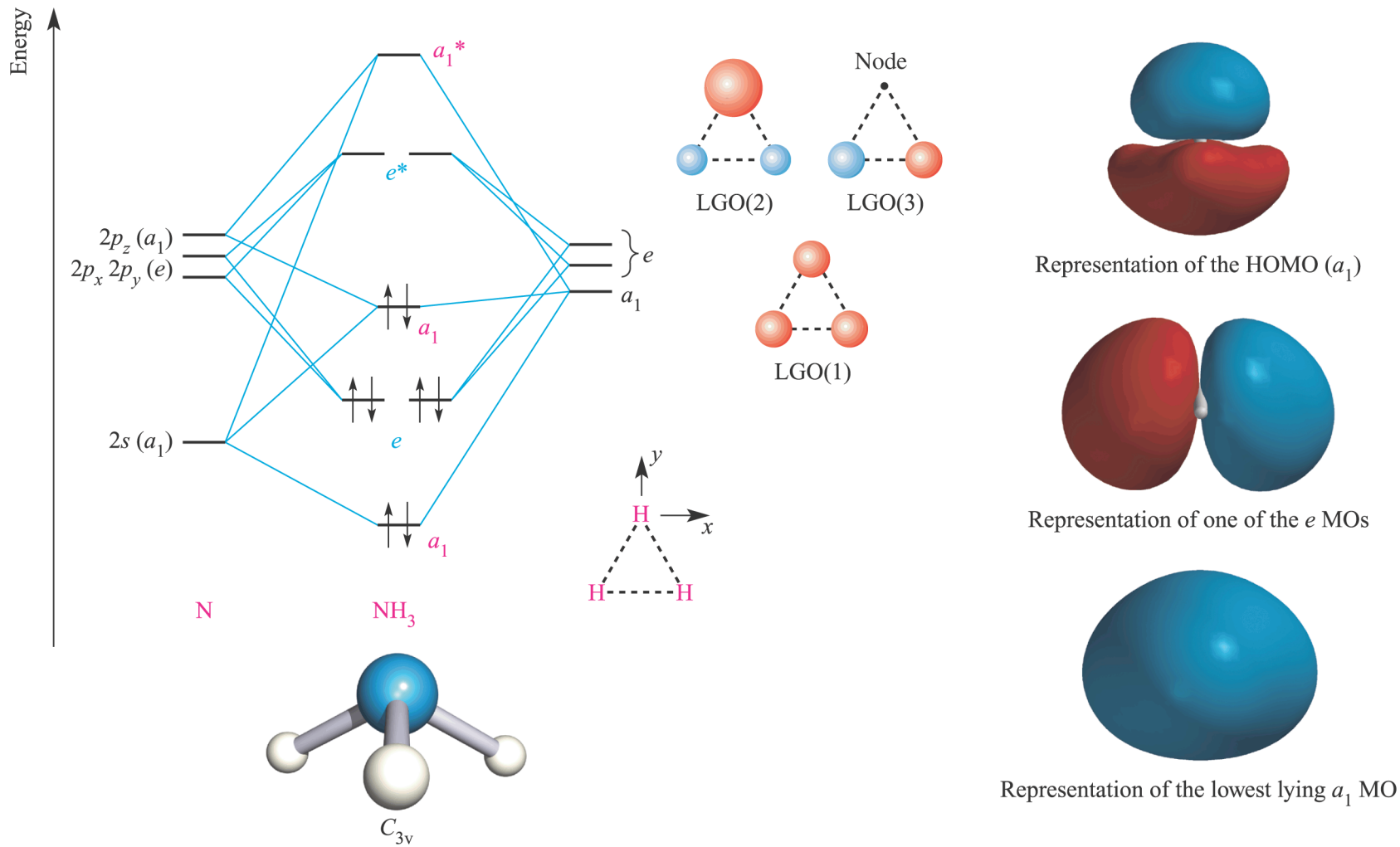
(LGO2)

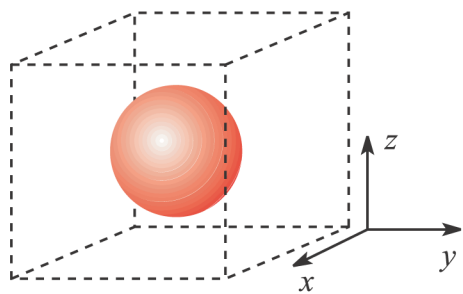
$$\Psi(e')_2 = 1/\sqrt{2} \times (\Psi_2 - \Psi_3)$$

(LGO3)

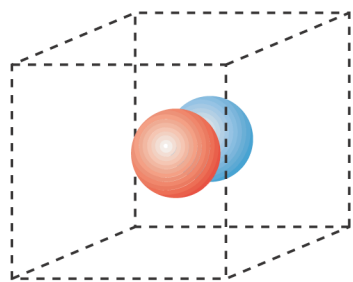
NH₃, gruppo puntuale C_{3v}



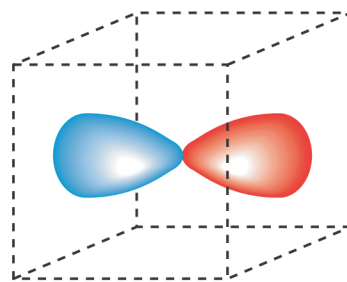




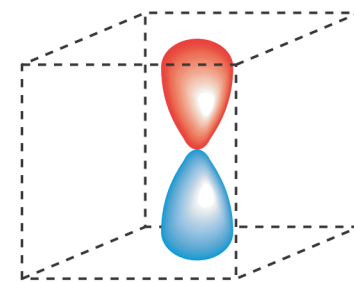
$2s(a_1)$



$2p_x(t_2)$

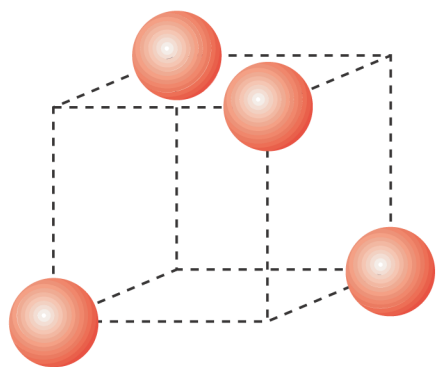


$2p_y(t_2)$

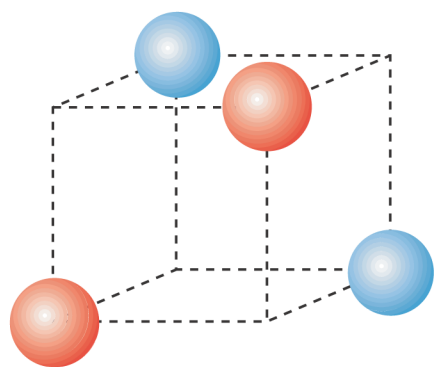


$2p_z(t_2)$

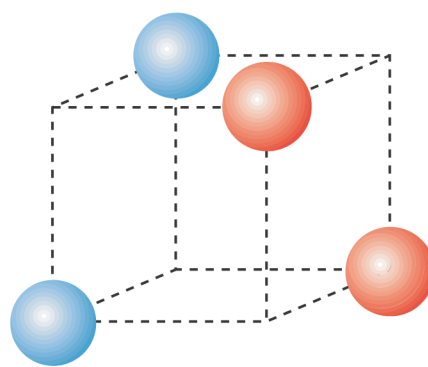
(a)



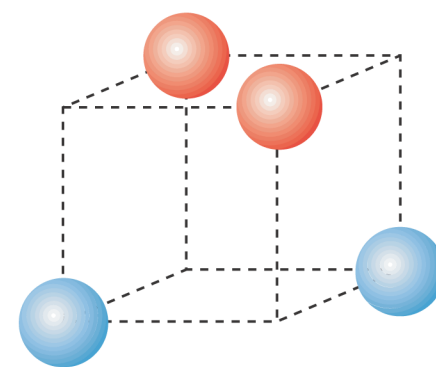
$LGO(1)(a_1)$



$LGO(2)(t_2)$

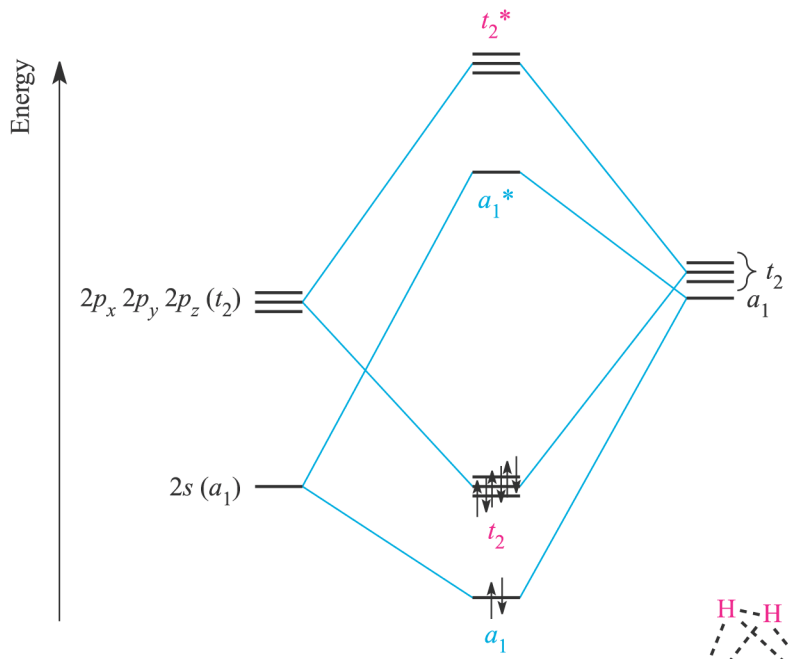


$LGO(3)(t_2)$



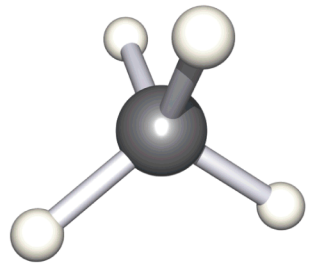
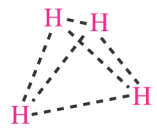
$LGO(4)(t_2)$

(b)

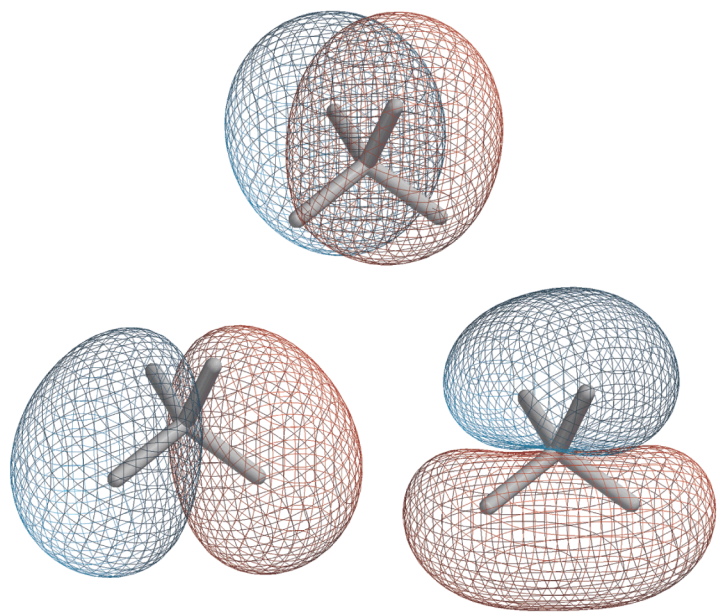


C

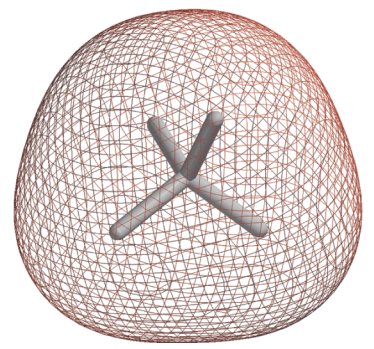
CH₄



T_d



Representations of the triply degenerate (*t*₂) bonding MOs

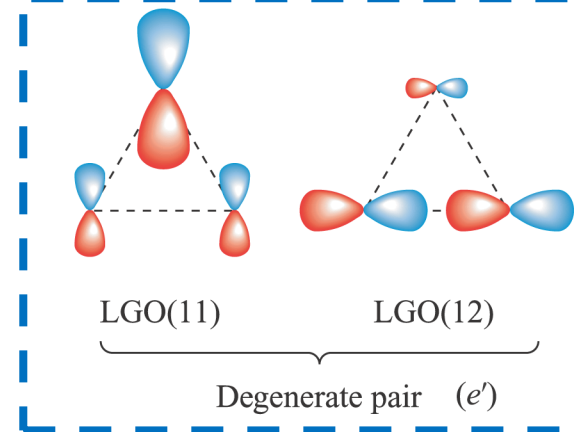
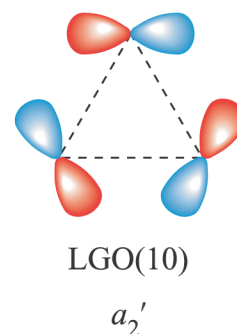
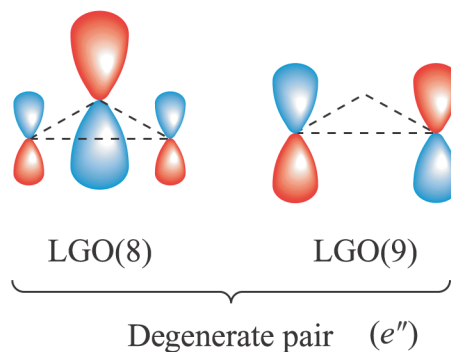
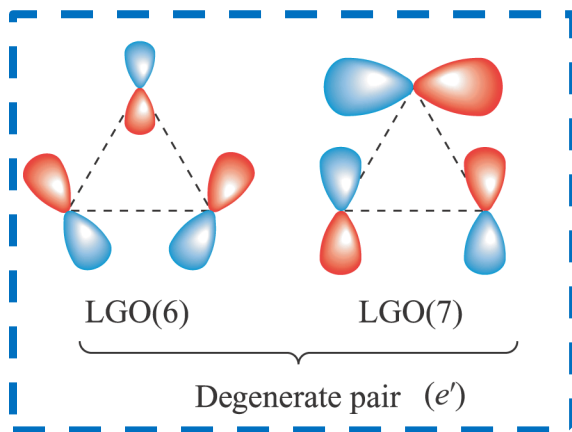
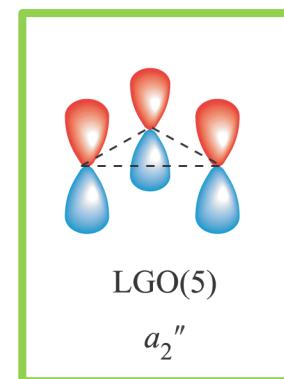
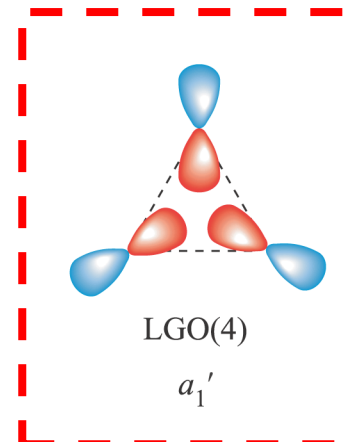
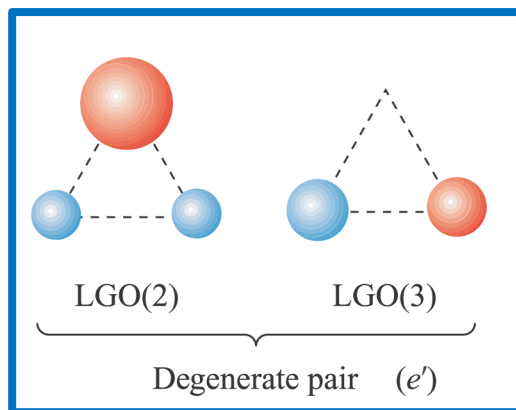
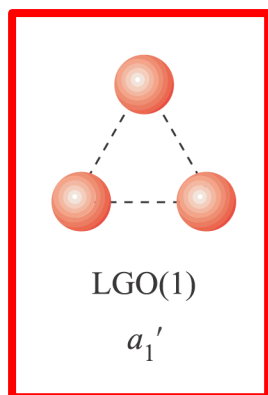
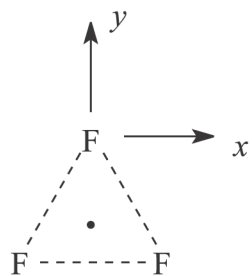


Representation of the *a*₁ bonding MO



<i>D</i> _{3h}	<i>E</i>	<i>2C</i> ₃	<i>3C</i> ₂	<i>σ</i> _h	<i>2S</i> ₃	<i>3σ</i> _v		
<i>A</i> ₁ '	1	1	1	1	1	1		<i>x</i> ² + <i>y</i> ² , <i>z</i> ²
<i>A</i> ₂ '	1	1	-1	1	1	-1	<i>R</i> _z	
<i>E</i> '	2	-1	0	2	-1	0	(<i>x</i> , <i>y</i>)	(<i>x</i> ² - <i>y</i> ² , <i>xy</i>)
<i>A</i> ₁ ''	1	1	1	-1	-1	-1		
<i>A</i> ₂ ''	1	1	-1	-1	-1	1	<i>z</i>	
<i>E</i> ''	2	-1	0	-2	1	0	(<i>R</i> _x , <i>R</i> _y)	(<i>xz</i> , <i>yz</i>)

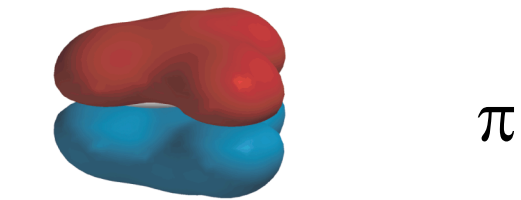
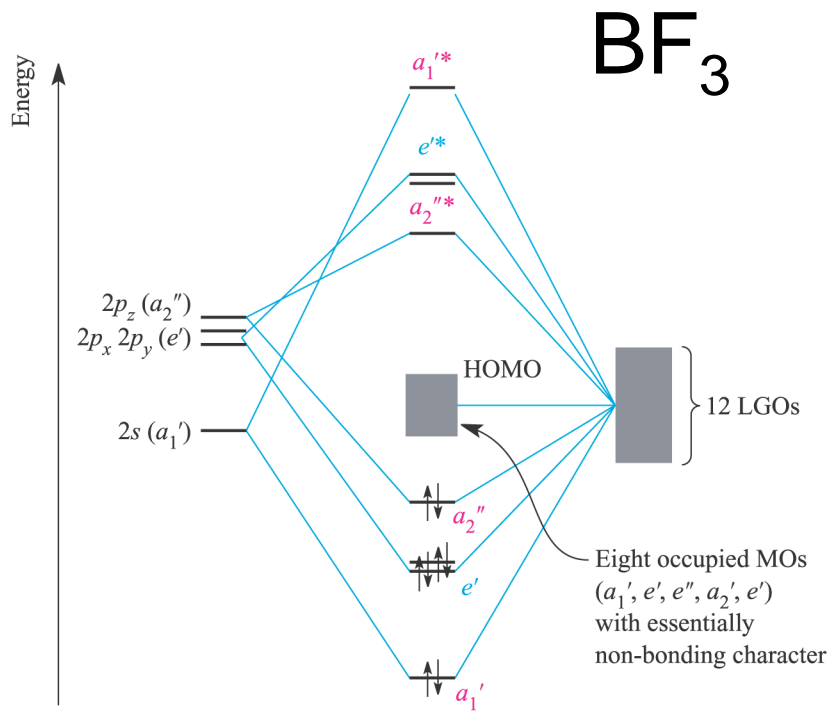
BF₃



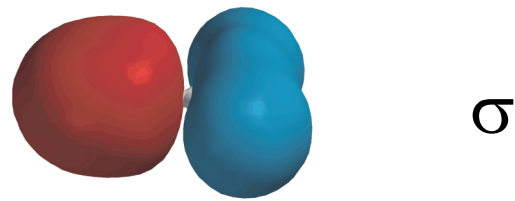
tre $2p_z$

E	C_3	C_2	σ_h	S_3	σ_v
3	0	-1	-3	0	1

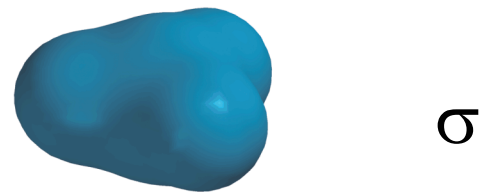
$A_2'' + E''$



Representation of the a_2'' MO

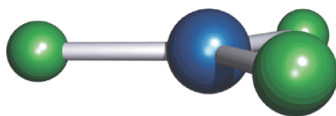


Representation of one of the e' MOs



Representation of the a_1' MO

B BF₃ F₃ fragment



D_{3h}

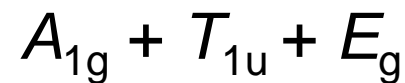
Ordine di legame 4/3

Molecole ipervalenti: SF₆ (gruppo O_h)

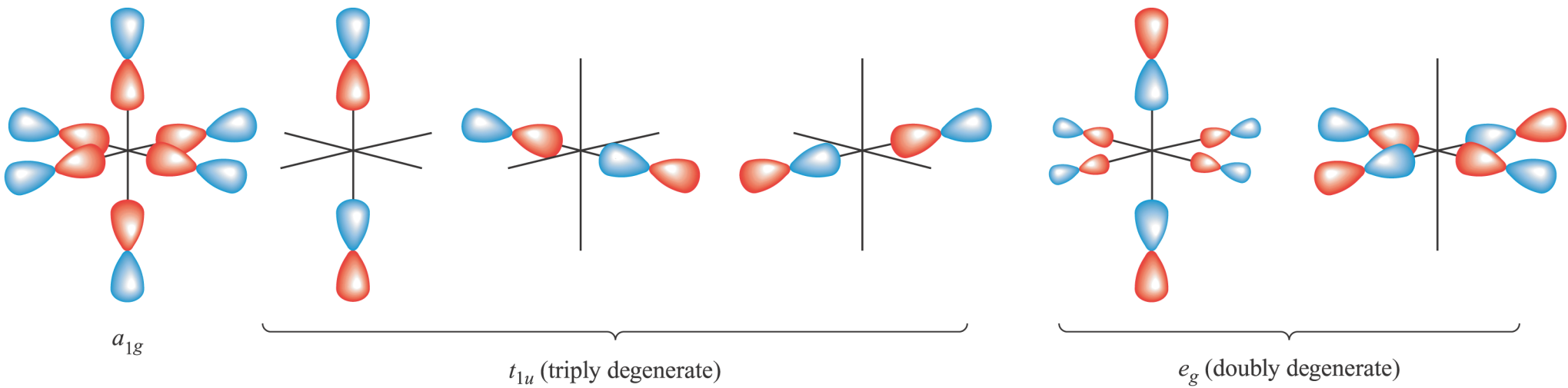
O _h	E	8C ₃	6C ₂	6C ₄	3C ₂ (= C ₄ ²)	i	6S ₄	8S ₆	3σ _h	6σ _d
A _{1g}	1	1	1	1	1	1	1	1	1	1
A _{2g}	1	1	-1	-1	1	1	-1	1	1	-1
E _g	2	-1	0	0	2	2	0	-1	2	0
T _{1g}	3	0	-1	1	-1	3	1	0	-1	-1
T _{2g}	3	0	1	-1	-1	3	-1	0	-1	1
A _{1u}	1	1	1	1	1	-1	-1	-1	-1	-1
A _{2u}	1	1	-1	-1	1	-1	1	-1	-1	1
E _u	2	-1	0	0	2	-2	0	1	-2	0
T _{1u}	3	0	-1	1	-1	-3	-1	0	1	1
T _{2u}	3	0	1	-1	-1	-3	1	0	1	-1

sei 2p_z radiali

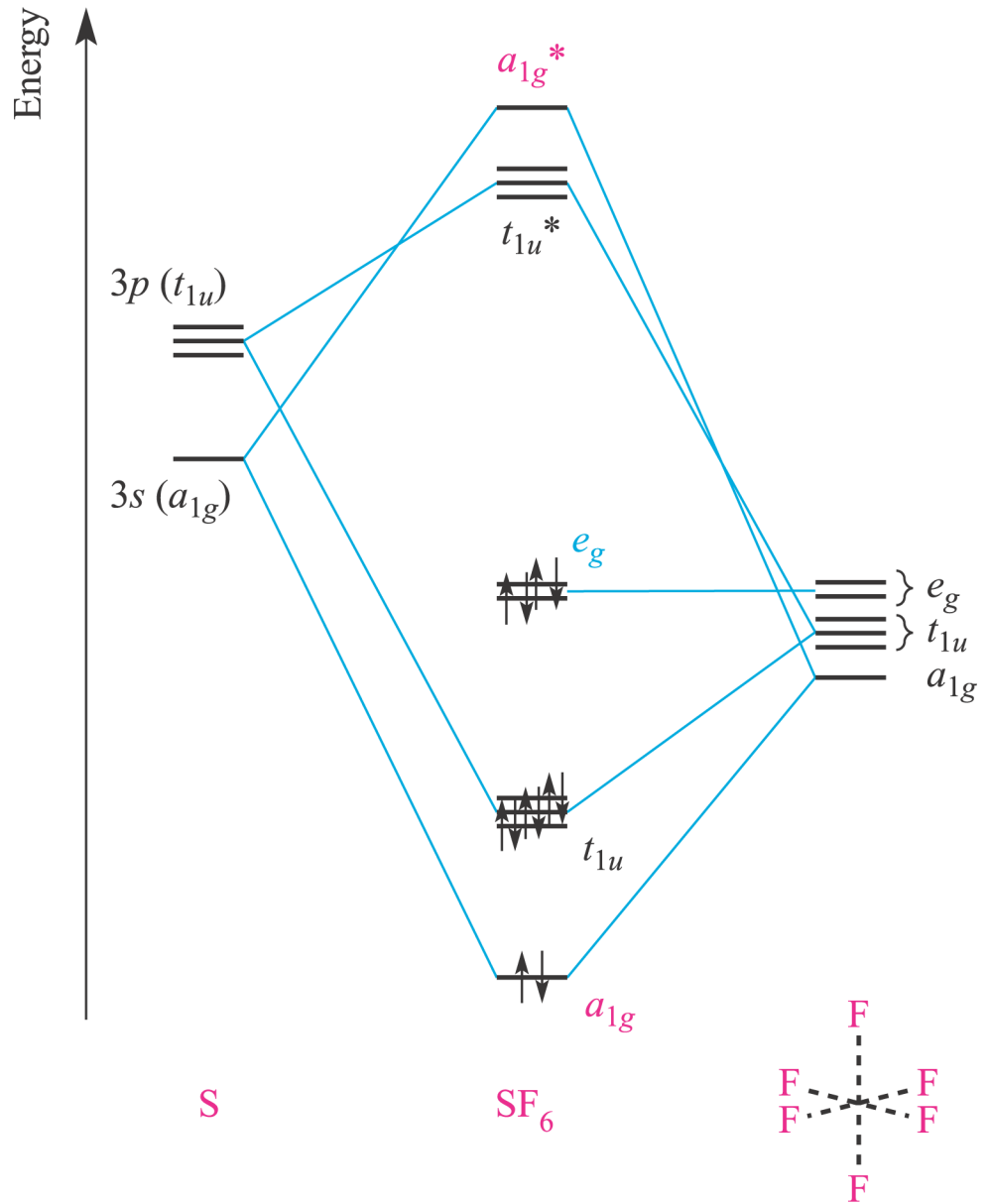
E	8C ₃	6C ₂	6C ₄	3C ₂	i	6S ₄	8S ₆	3σ _h	6σ _d
6	0	0	2	2	0	0	0	4	2



LGO del frammento F_6 in SF_6



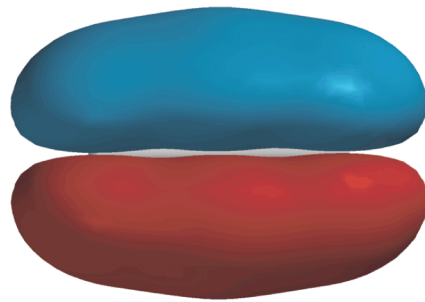
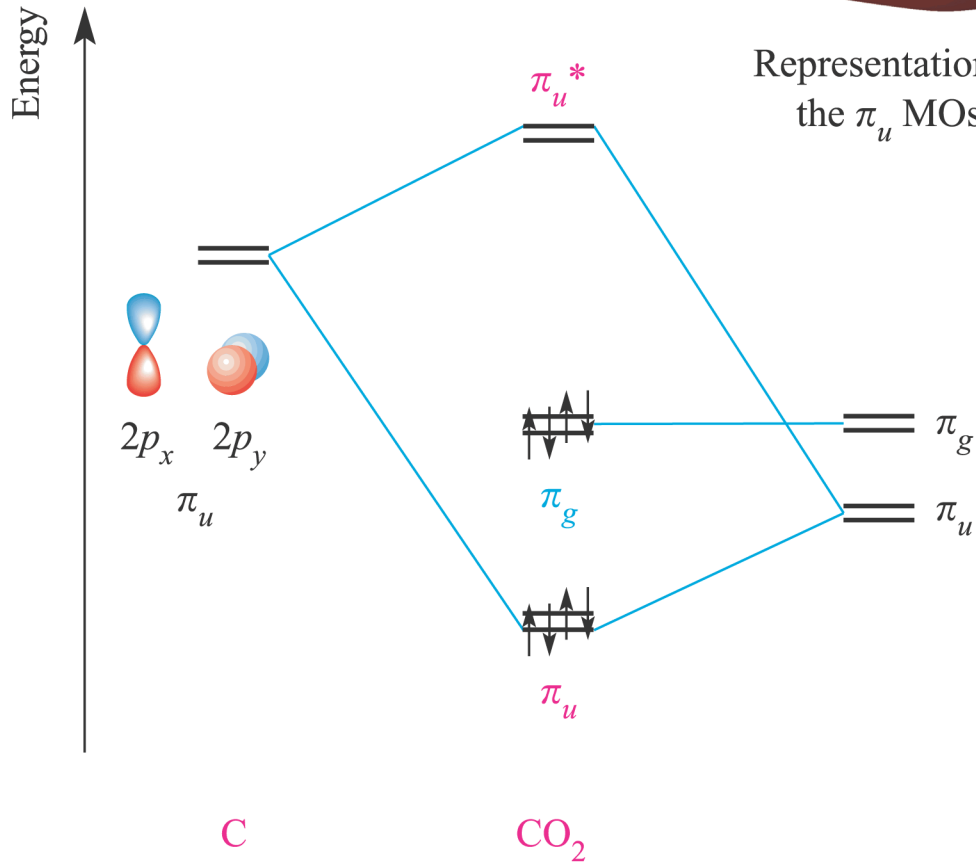
nonlegame



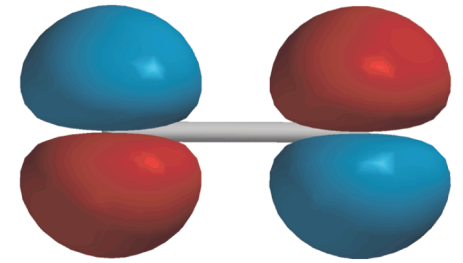
Ordine di legame 2/3,
nessun orbitale d

CO₂

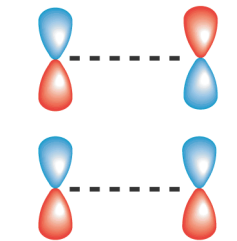
Diagrammi MO parziali



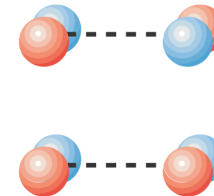
Representation of one of the π_u MOs in CO₂



Representation of one of the π_g MOs in CO₂



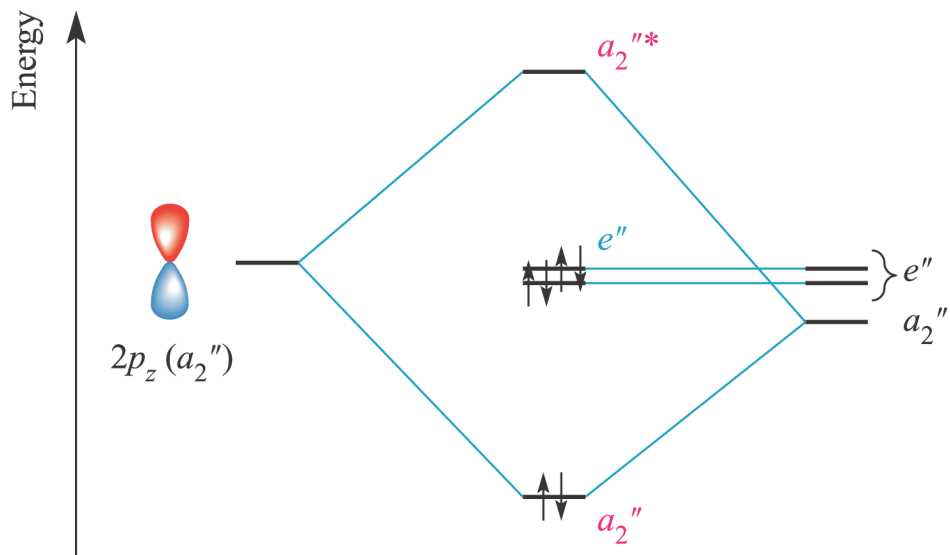
Combination of $2p_x$ orbitals



Combination of $2p_y$ orbitals

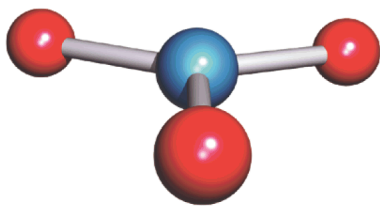


NO₃⁻

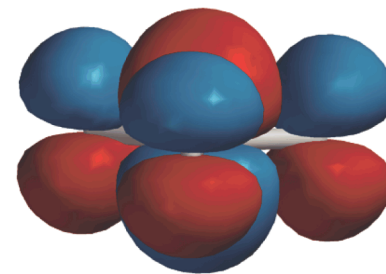
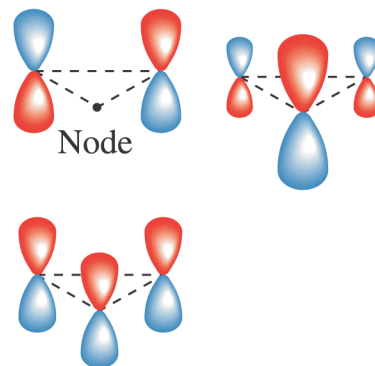


N

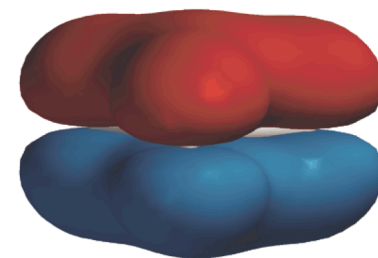
[NO₃]⁻



D_{3h}

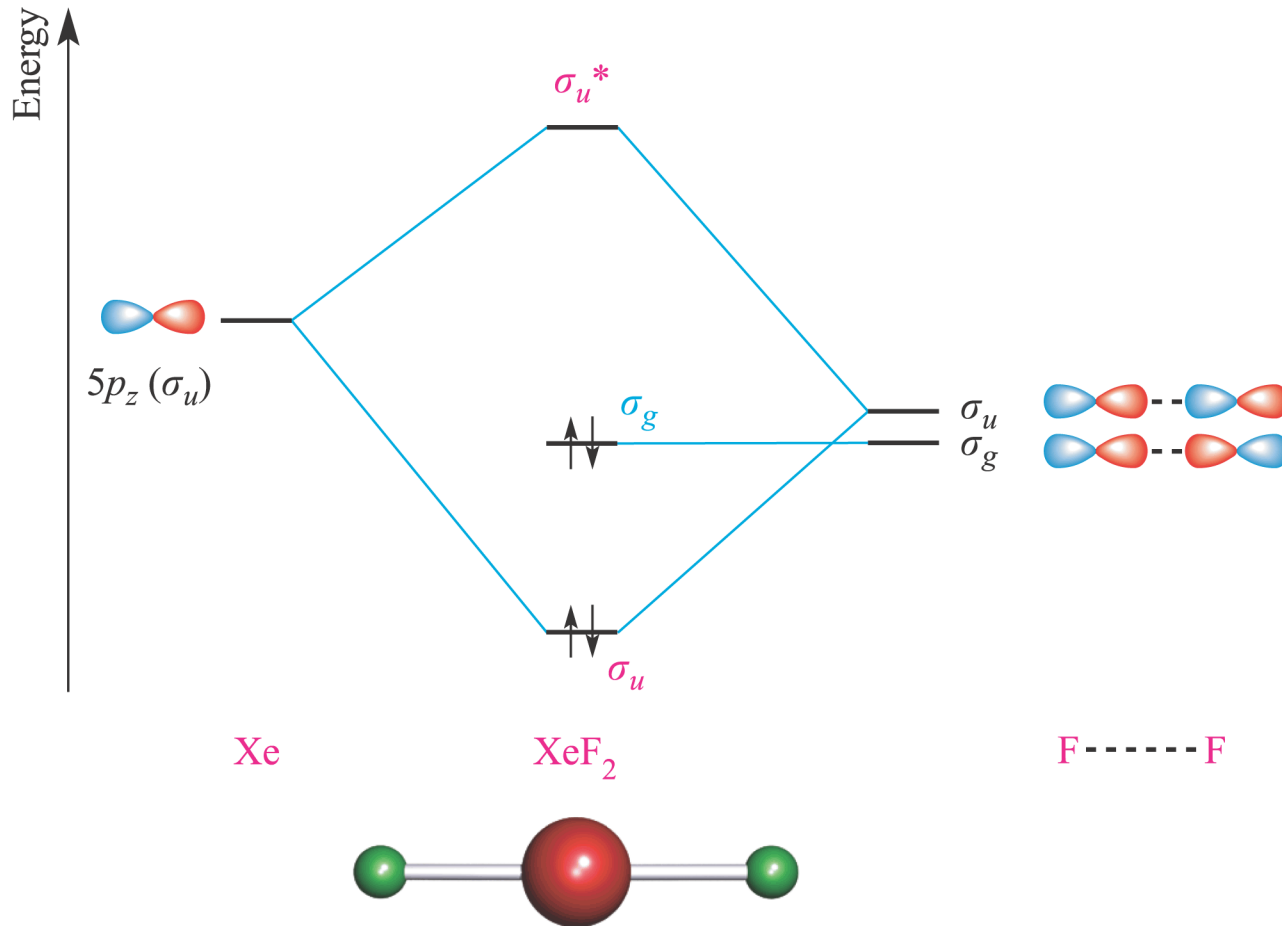


A representation of the $a_2''^*$ MO

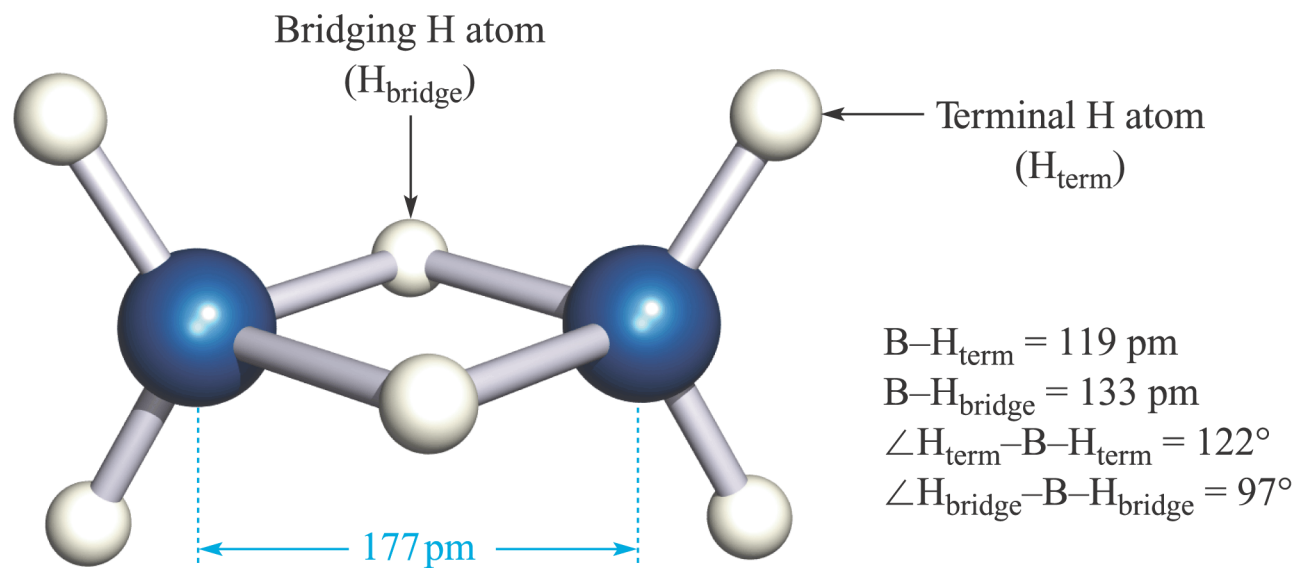


A representation of the a_2'' MO showing the delocalization of π -character over the N and O centres

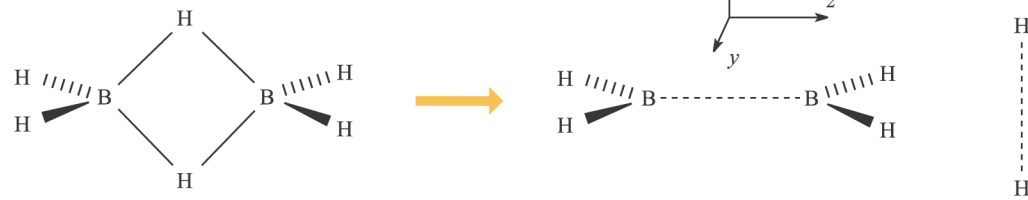
Interazione 3c – 2e



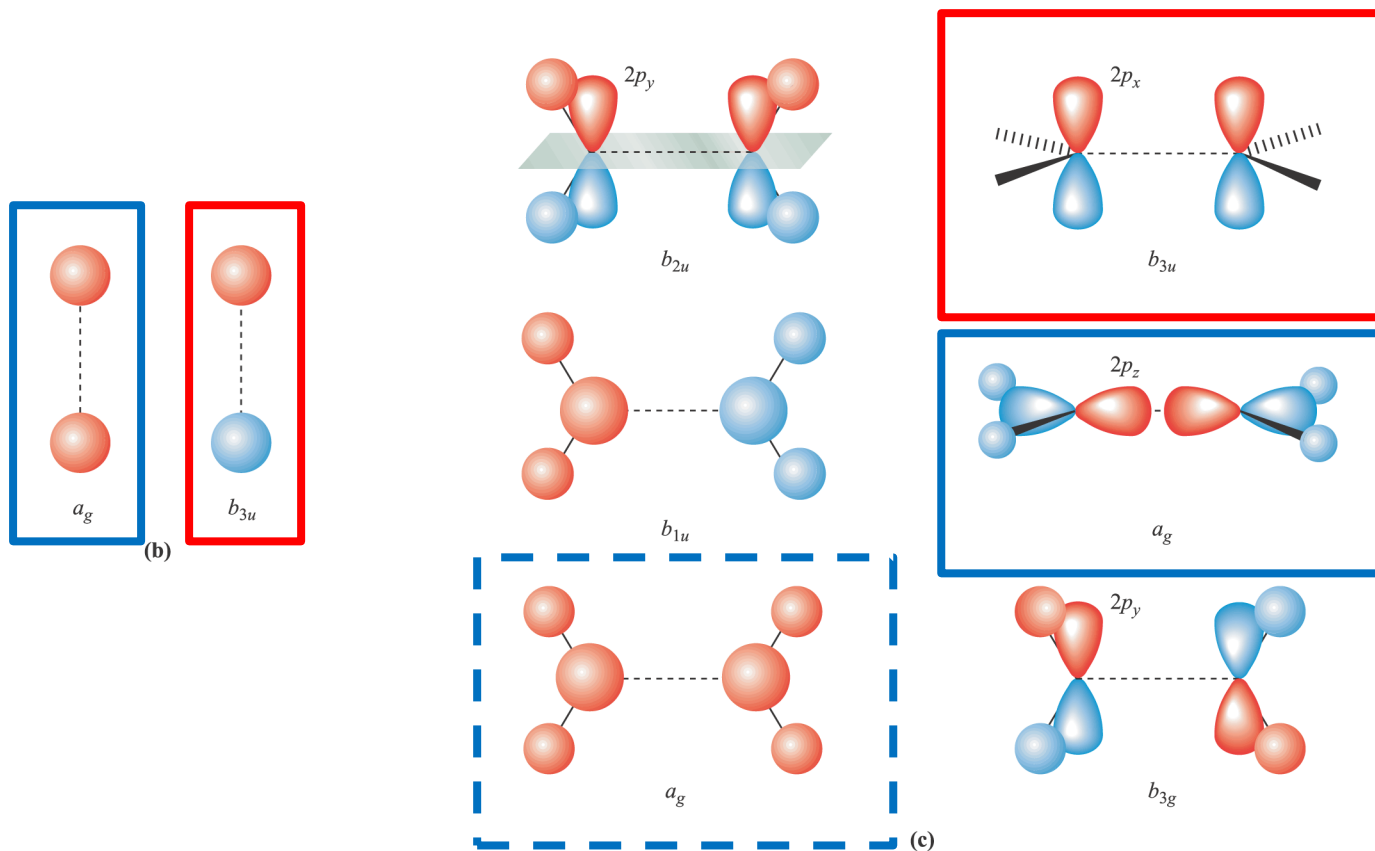
Diborano, B_2H_6



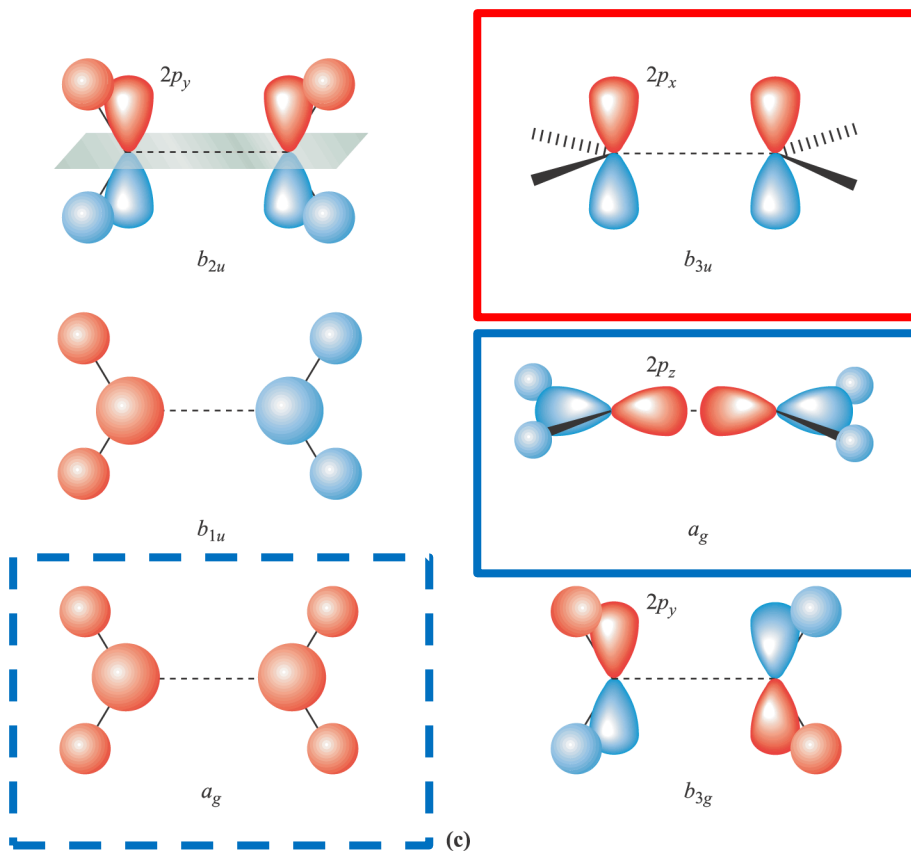
D_{2h}



(a)

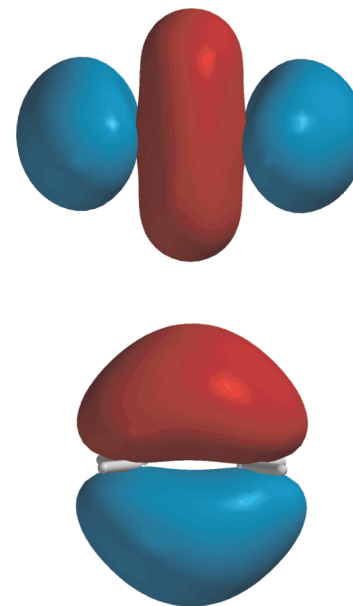
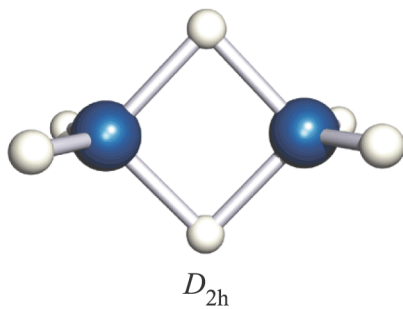
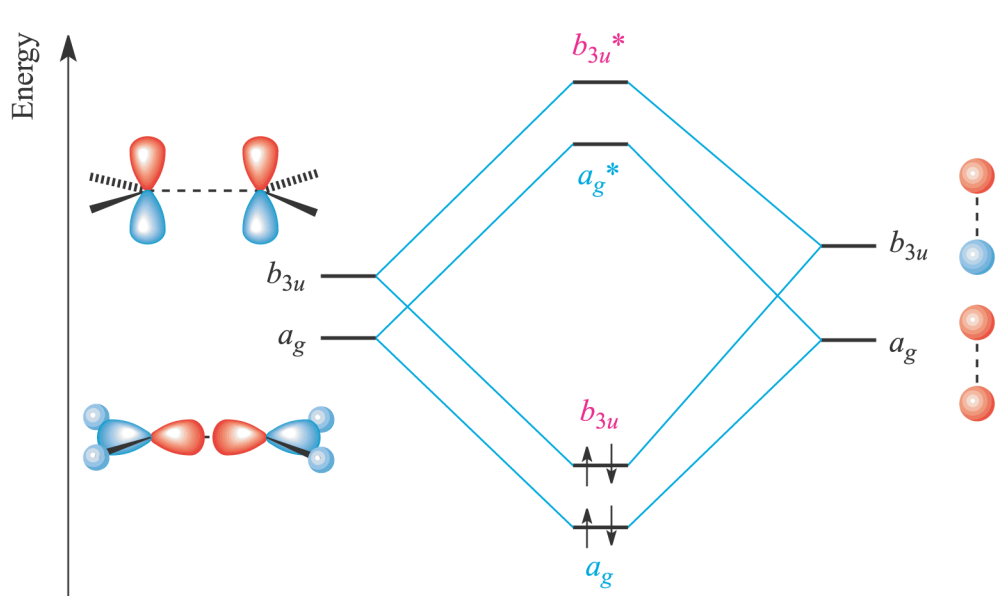


(b)



(c)

6 LGO (su 12) a energia più bassa del frammento B_2H_4



Representation of the a_g (top) and b_{3u} MOs which contain B–H–B bonding character