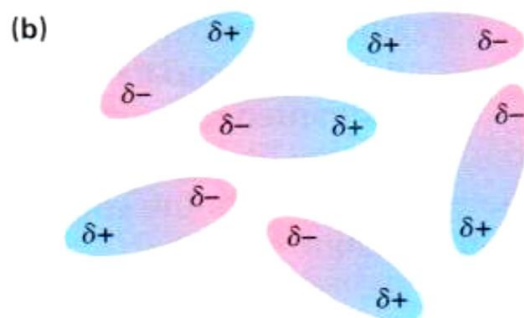
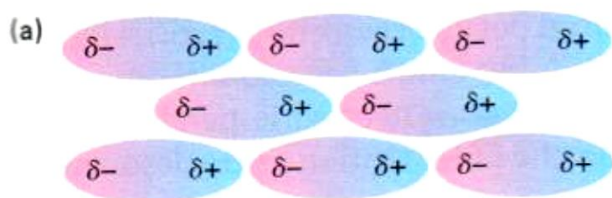
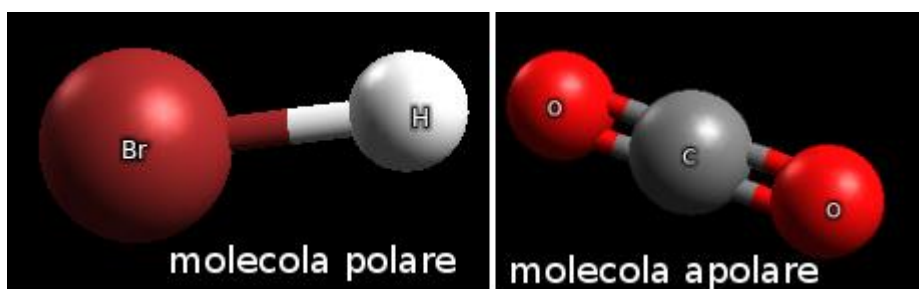
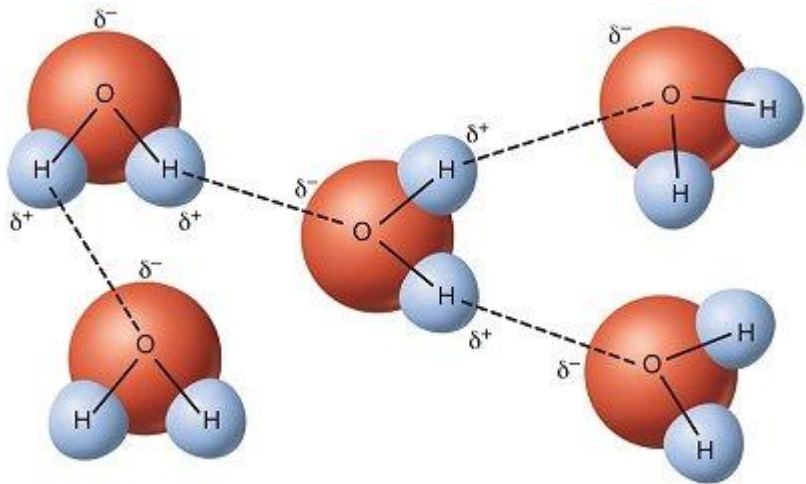
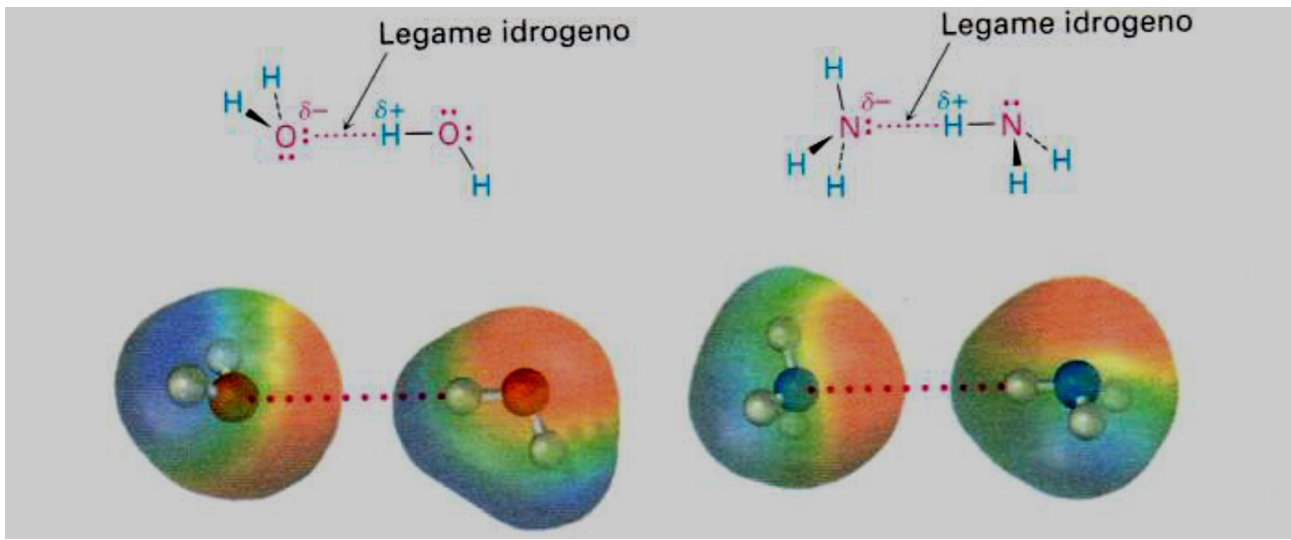


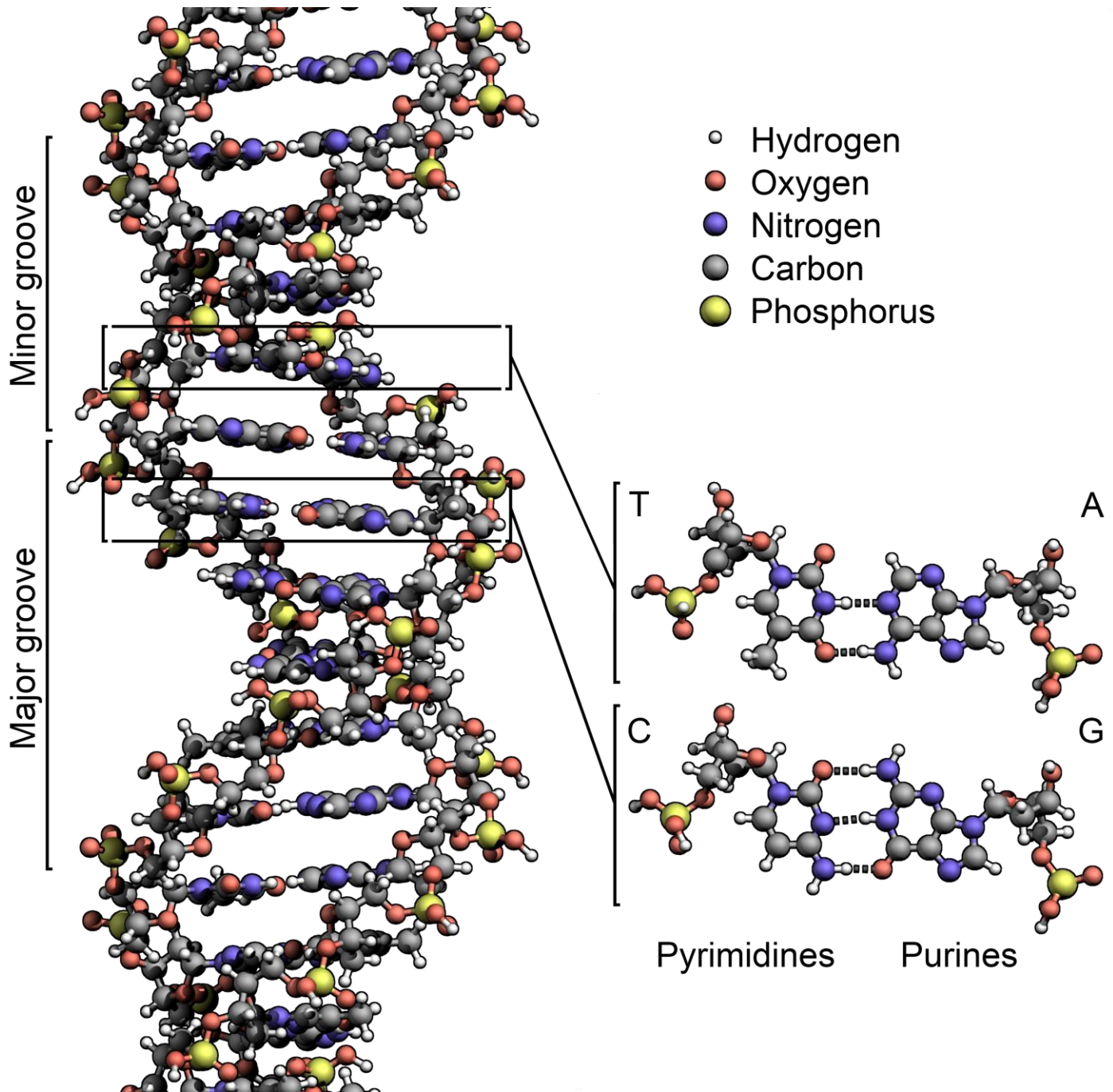
Legami di van der Waals



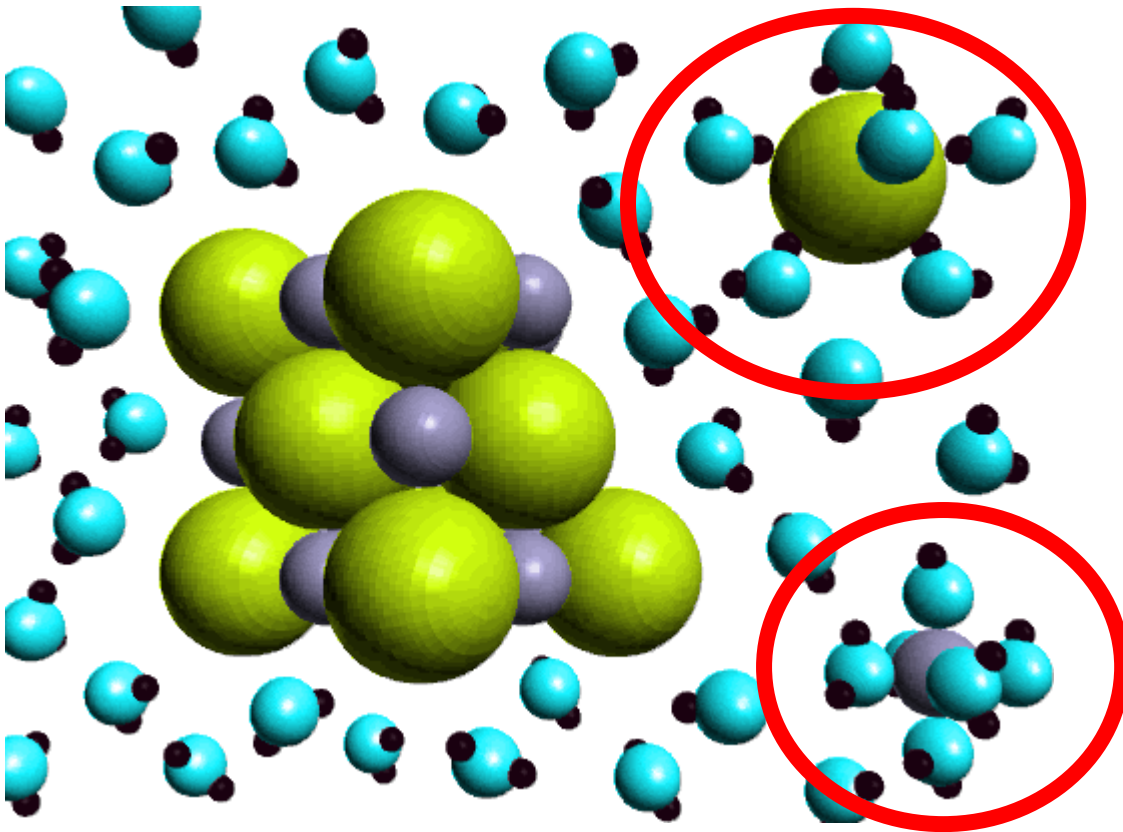
Interazione dipolo-dipolo



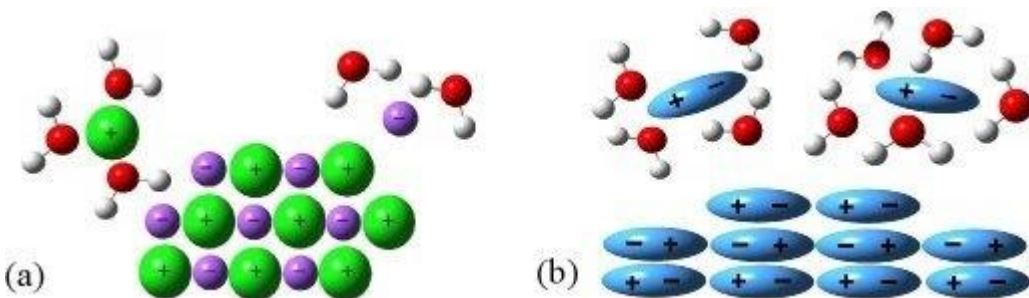
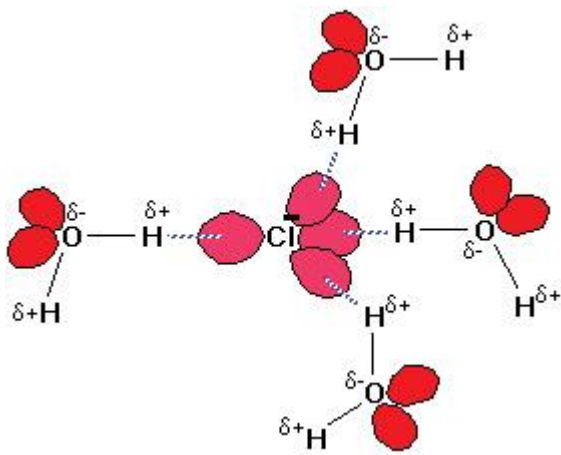




Solubilizzazione

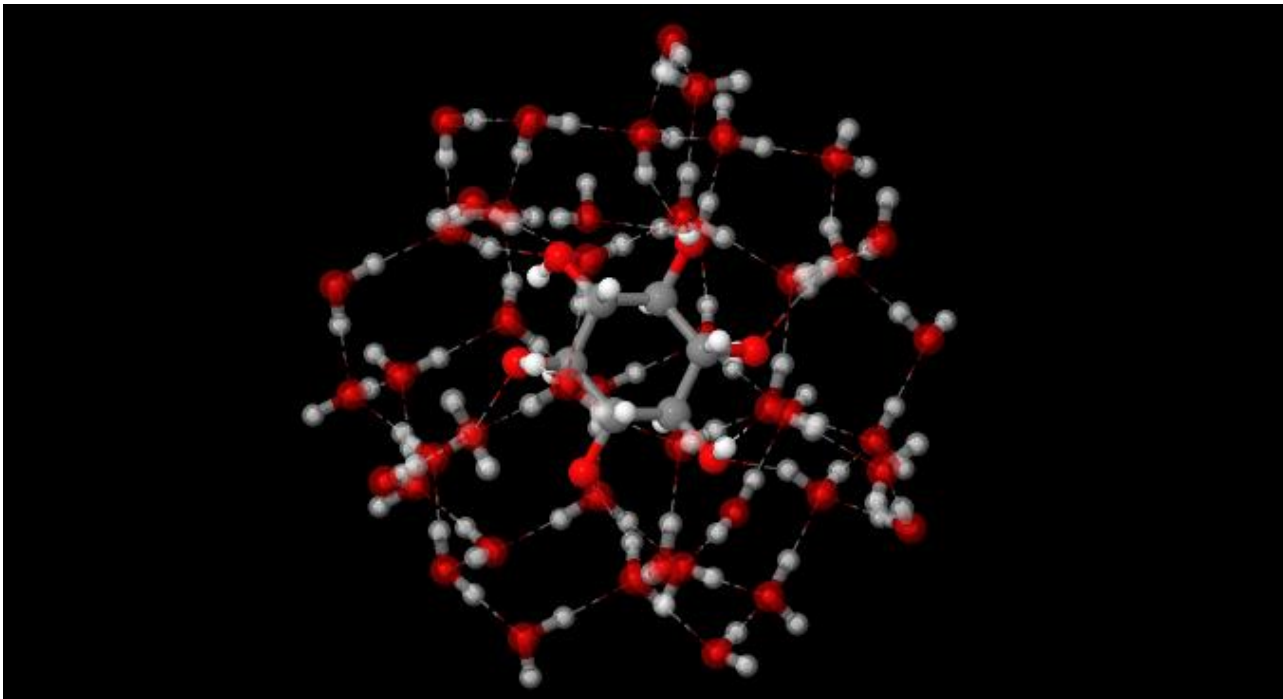
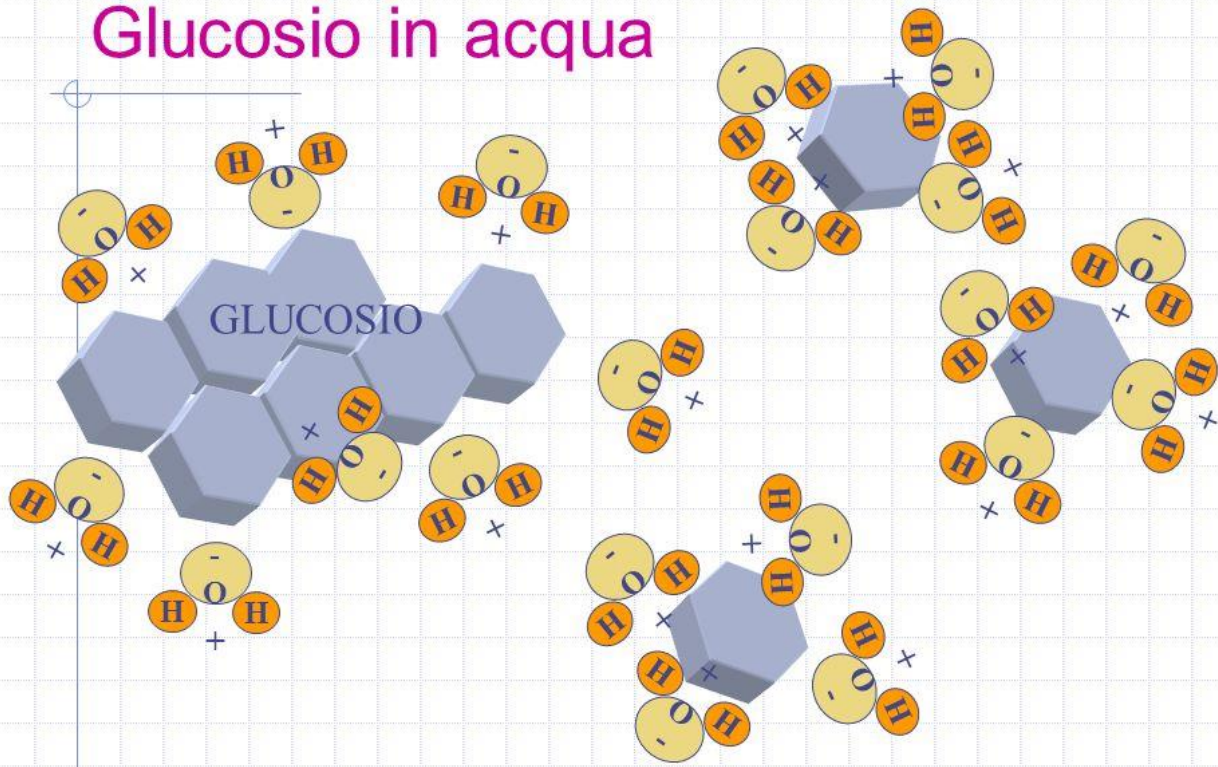


L'immagine mostra un cristallo di Cloruro di sodio che viene dissolto da molecole d'acqua. Il polo negativo delle molecole d'acqua circonda gli ioni Na^+ , mentre il polo positivo gli ioni Cl^- .



a) Solvatazione di ioni. B) solvatazione di molecolare polari.

Glucosio in acqua



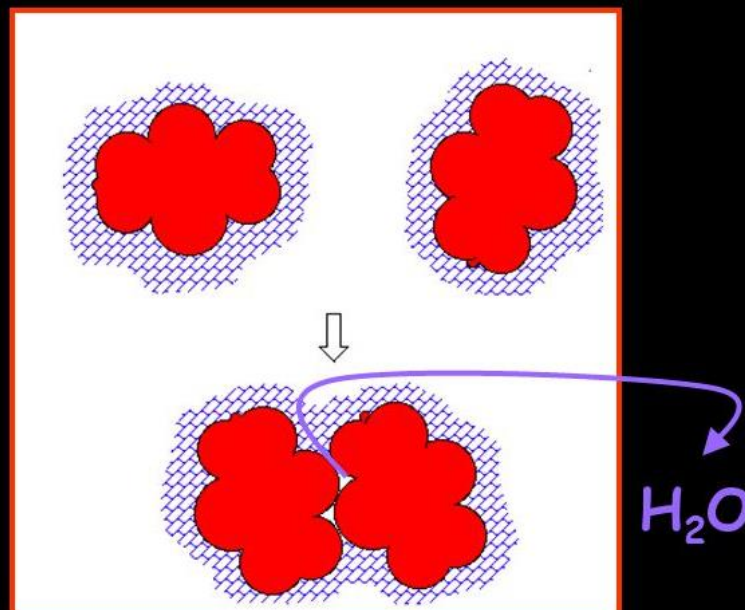
interazione idrofobica

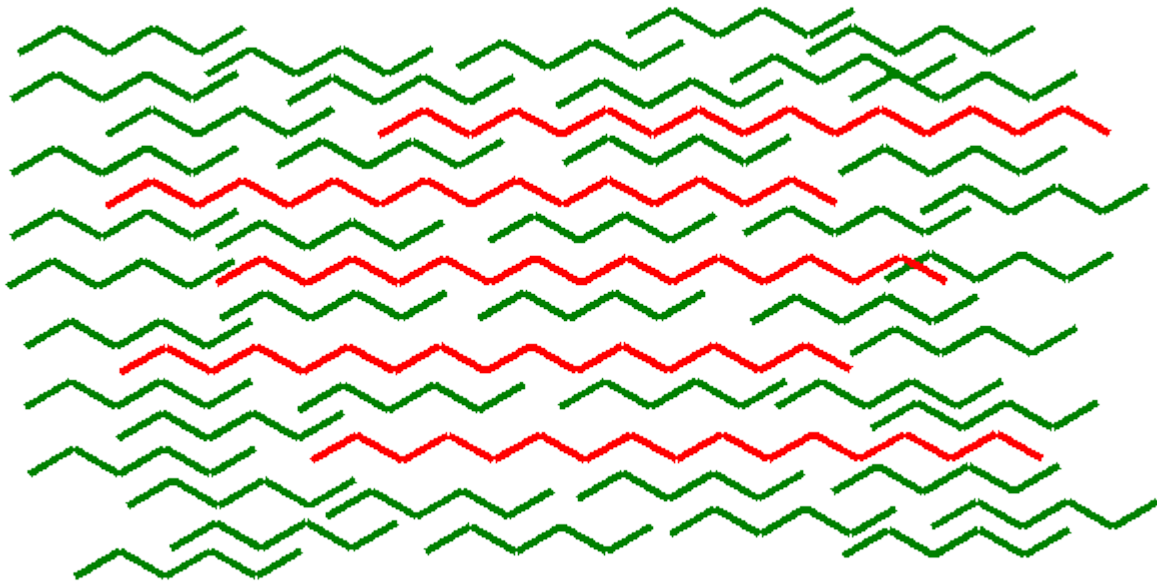
→ non si tratta realmente di un tipo di legame, ma di una interazione fra l'acqua e una molecola apolare, modulata da variazioni di entropia



Le molecole d'acqua che circondano molecole apolari formano uno stretto reticolo di molecole connesse da ponti H, in cui moti e orientazioni sono fortemente limitati

Ciò si traduce nella tendenza che le molecole apolari mostrano di unirsi tra di loro escludendo l'acqua...





.20).
rsità del-
dere che

MACRO

Olio

Acqua

MICRO

A diagram illustrating the concept of "like dissolves like". On the left, a flask contains two distinct liquid layers: a yellow layer on top labeled "Olio" (oil) and a clear layer on the bottom labeled "Acqua" (water). A green arrow points from the oil layer to a microscopic view on the right. The microscopic view shows a clear interface between two phases. The top phase contains green, spherical particles, and the bottom phase contains brown, spherical particles. A red arrow points from the water layer in the flask to the bottom phase in the microscopic view. The text "OGNI SIMILE SCIoglie IL SUO SIMILE!" is written at the bottom of the diagram.

OGNI SIMILE SCIoglie IL SUO SIMILE!