

Luminescenza

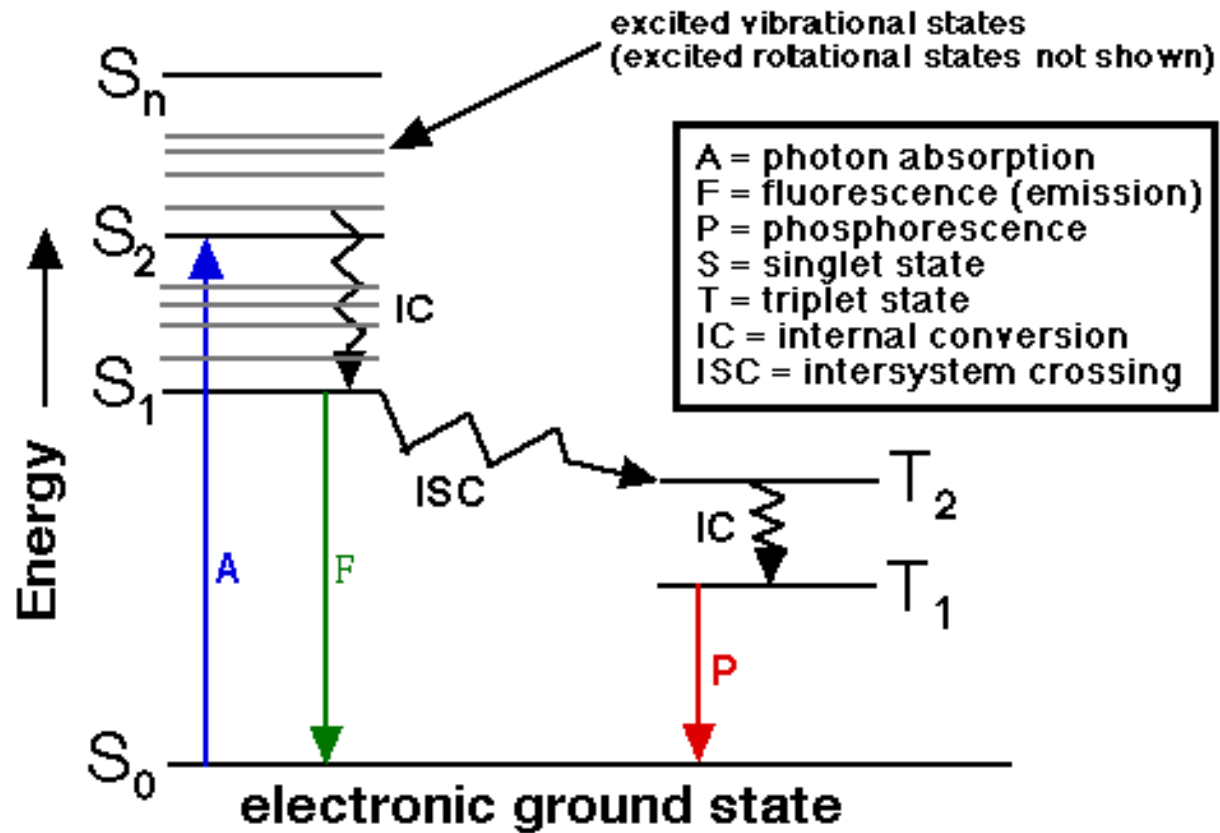
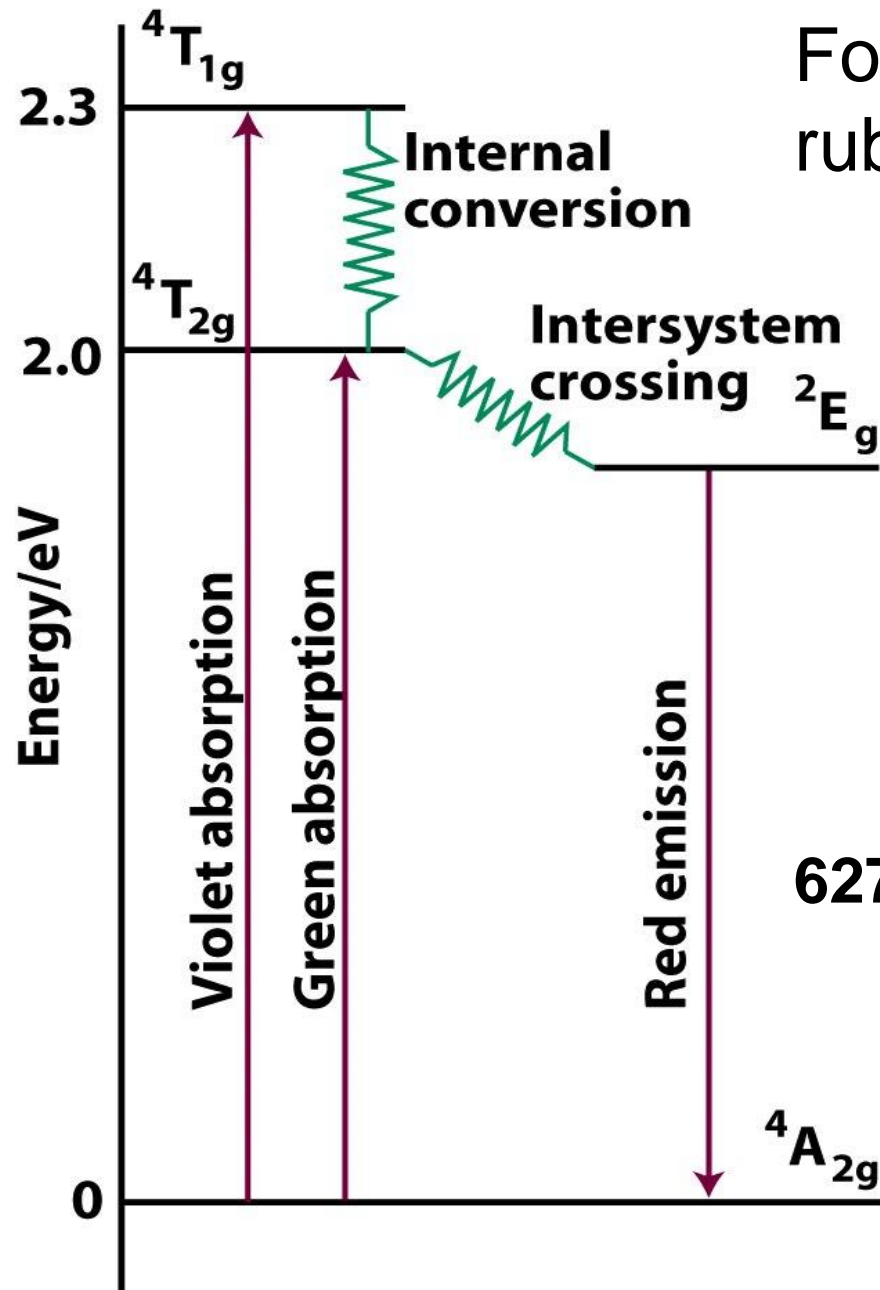
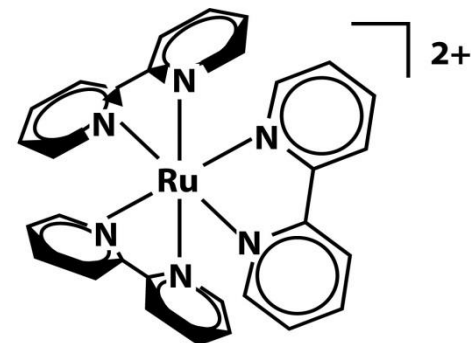
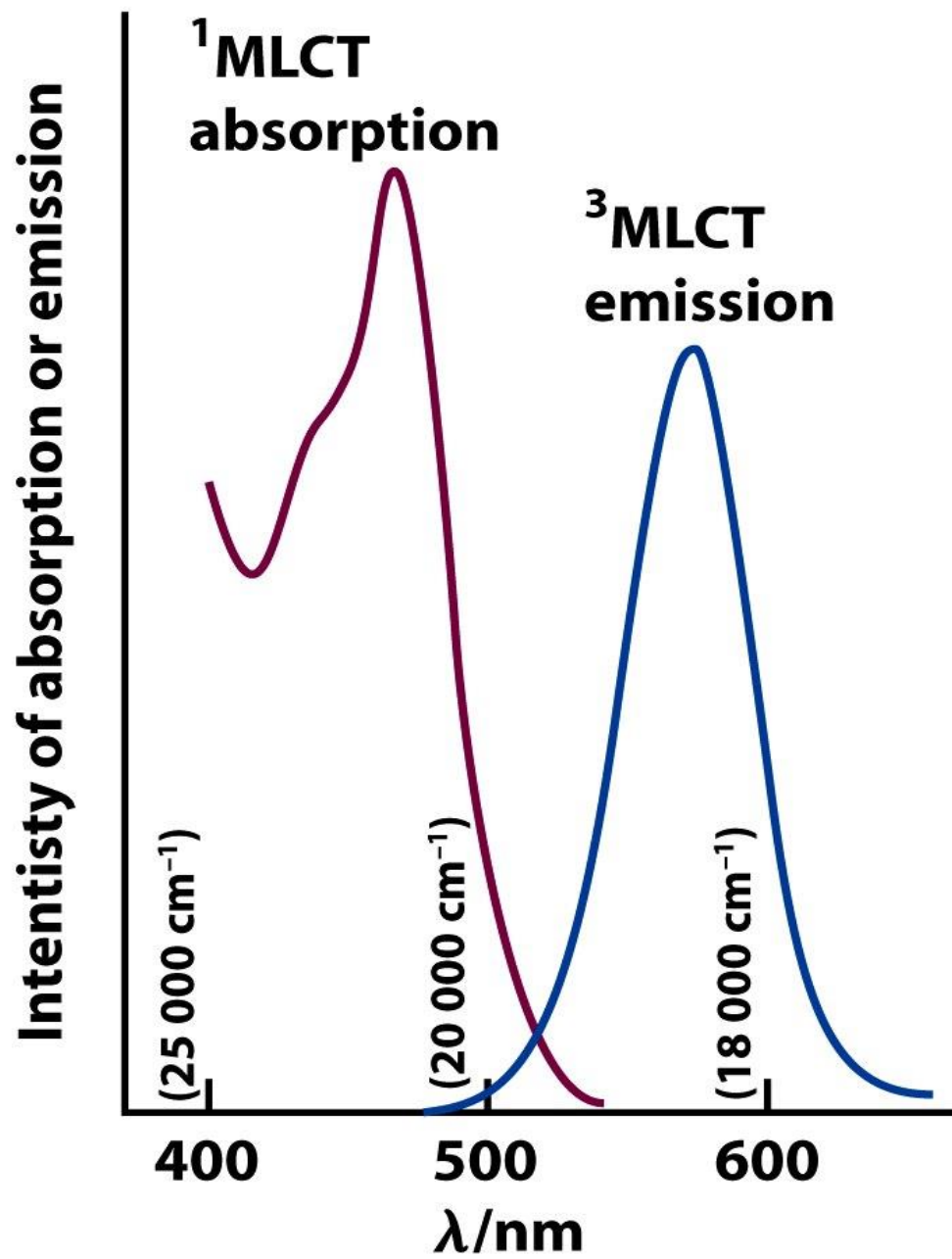


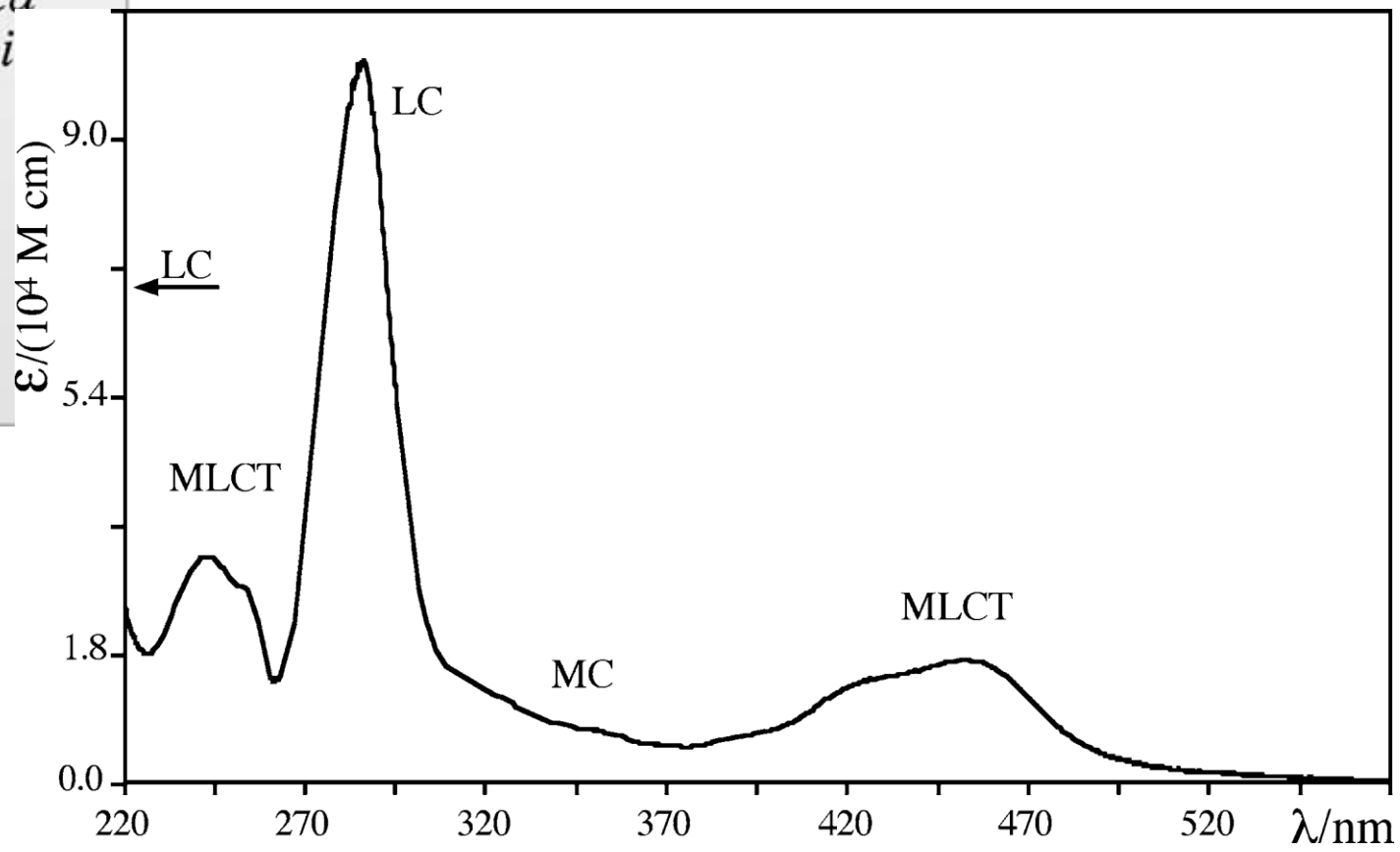
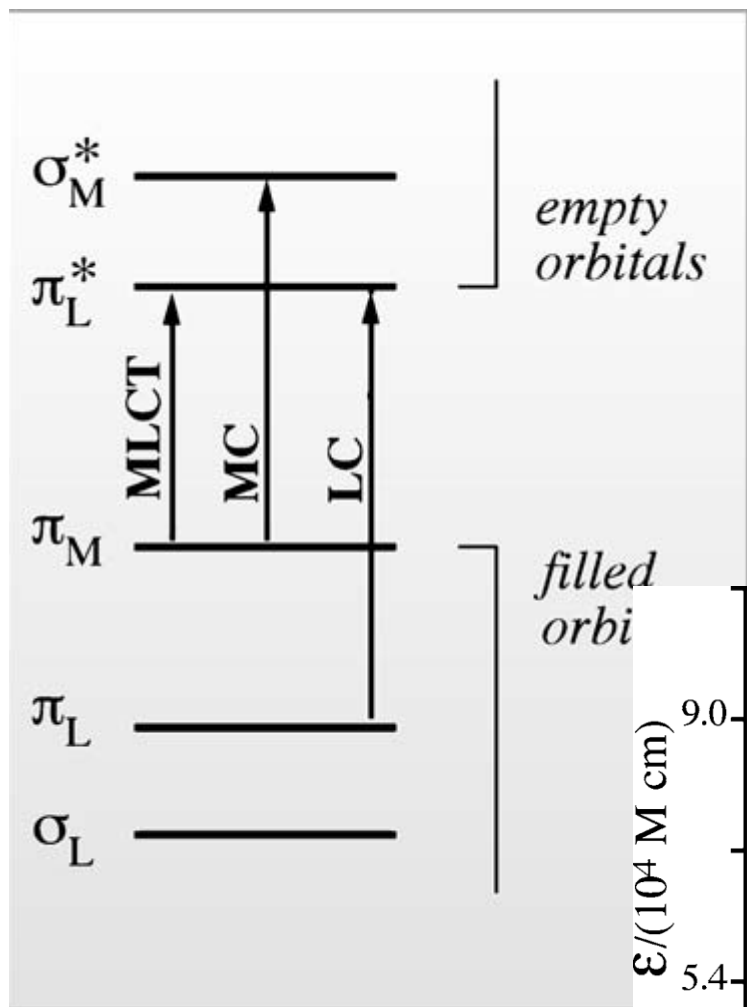
Diagramma di Jablonski

Fosforescenza del rubino ($\text{Cr}^{3+}(\text{d}^3)$)

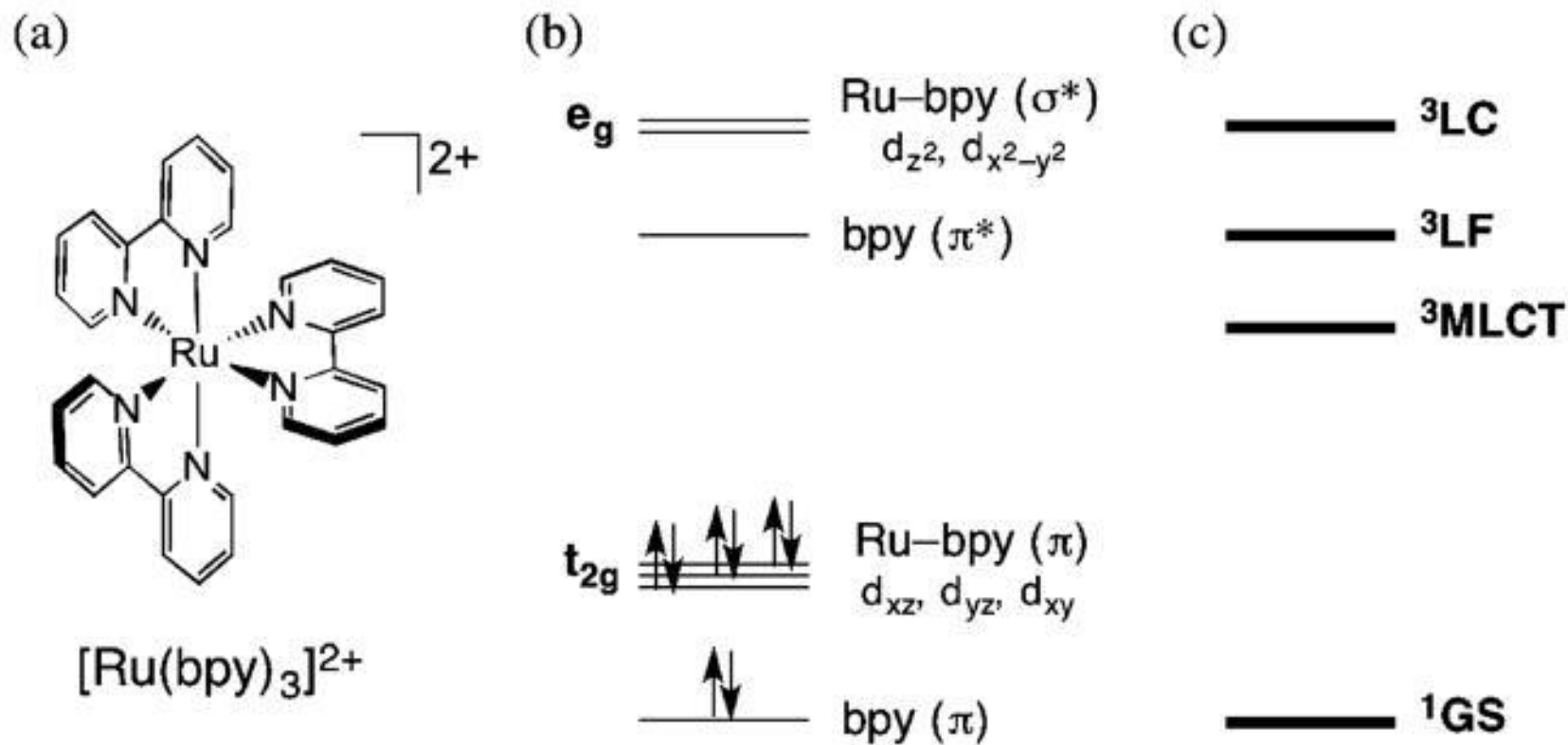


627 nm

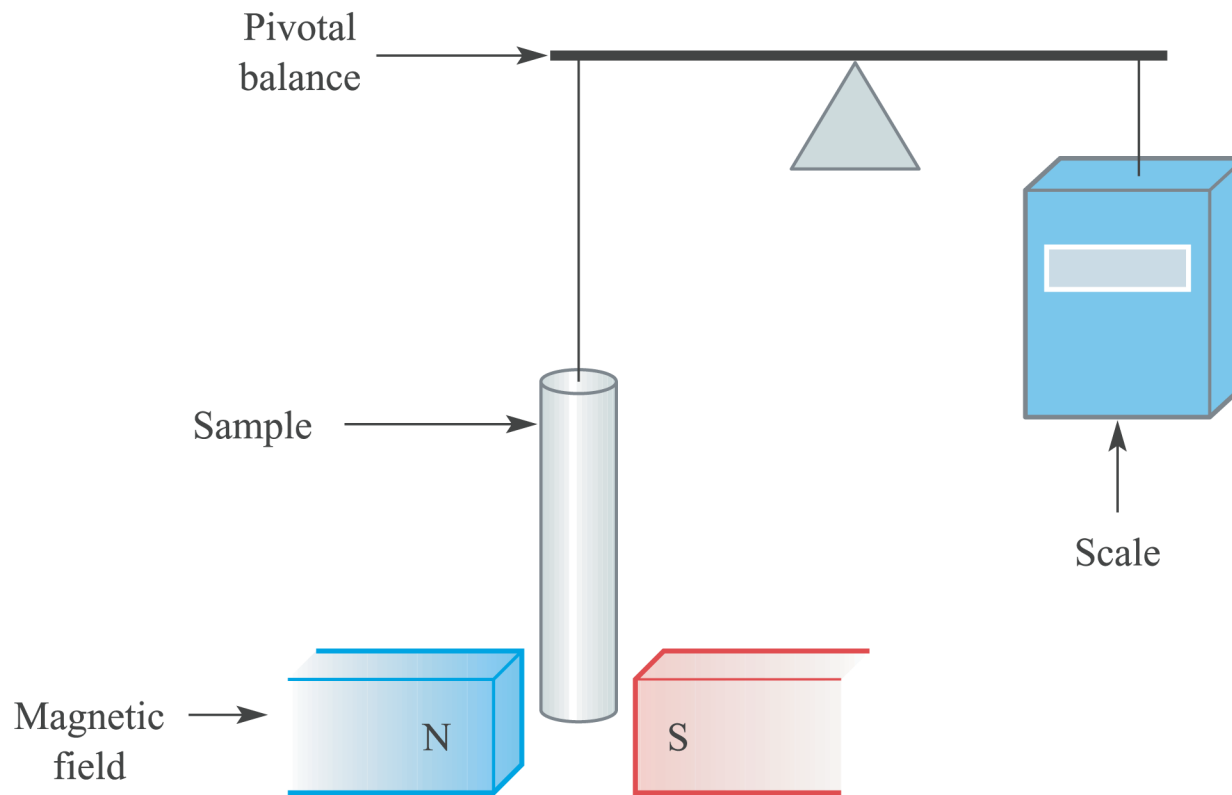




Diagrammi semplificati degli MO di frontiera e degli stati di $[\text{Ru}(\text{bpy})_3]^{2+}$



Schema di una bilancia di Gouy



Approssimazione *spin-only*

$$\mu = 2\sqrt{S(S+1)} \quad S = n/2 \quad \mu = \sqrt{n(n+2)}$$

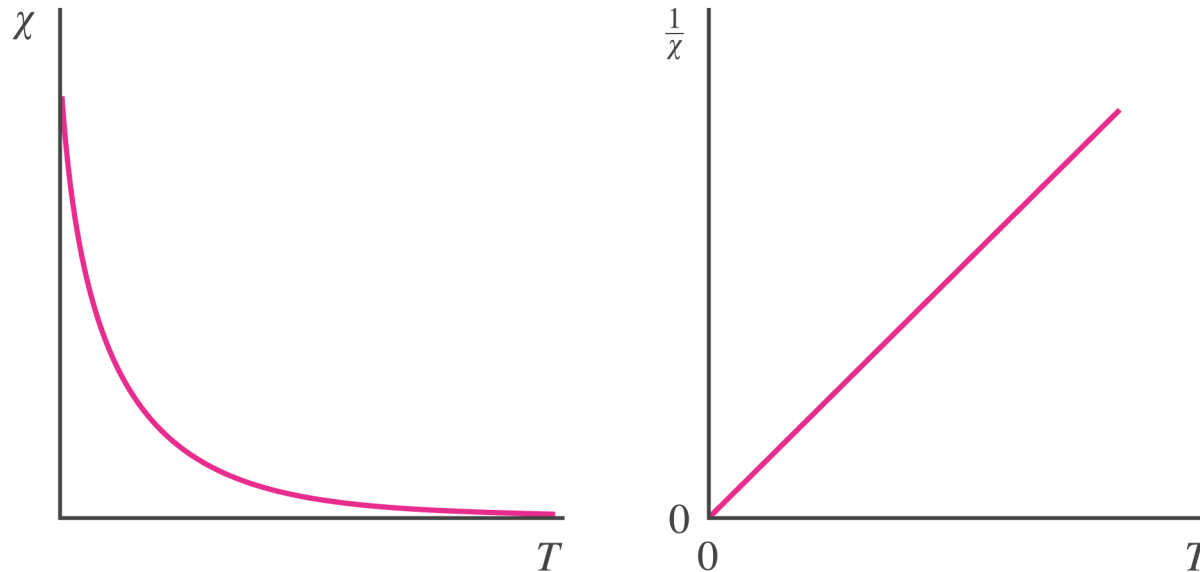
$$1 \text{ BM (o } \mu_B) = 9.274 \times 10^{-24} \text{ J T}^{-1}$$

Metal ion	d^n configuration	S	$\mu_{\text{eff}}(\text{spin-only}) / \mu_B$	Observed values of μ_{eff} / μ_B
Sc ³⁺ , Ti ⁴⁺	d^0	0	0	0
Ti ³⁺	d^1	$\frac{1}{2}$	1.73	1.7–1.8
V ³⁺	d^2	1	2.83	2.8–3.1
V ²⁺ , Cr ³⁺	d^3	$\frac{3}{2}$	3.87	3.7–3.9
Cr ²⁺ , Mn ³⁺	d^4	2	4.90	4.8–4.9
Mn ²⁺ , Fe ³⁺	d^5	$\frac{5}{2}$	5.92	5.7–6.0
Fe ²⁺ , Co ³⁺	d^6	2	4.90	5.0–5.6
Co ²⁺	d^7	$\frac{3}{2}$	3.87	4.3–5.2
Ni ²⁺	d^8	1	2.83	2.9–3.9
Cu ²⁺	d^9	$\frac{1}{2}$	1.73	1.9–2.1
Zn ²⁺	d^{10}	0	0	0

$$\mu_{\text{eff}} = 2.83\sqrt{\chi_m T}$$

χ_m = suscettività magnetica molare

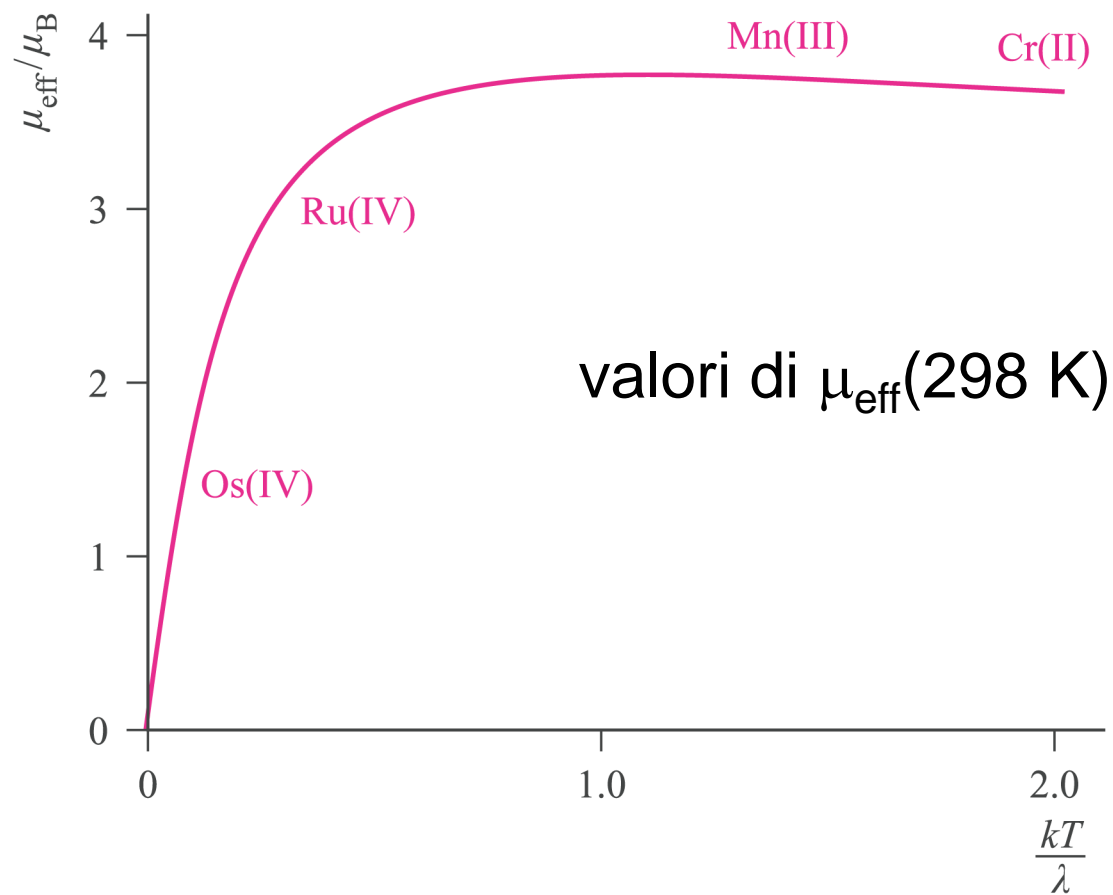
Materiale paramagnetico



$$\chi = C/T$$

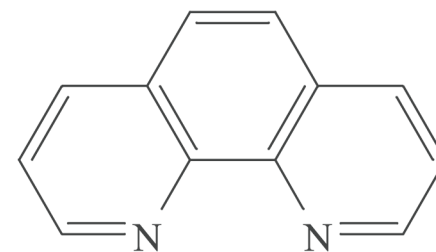
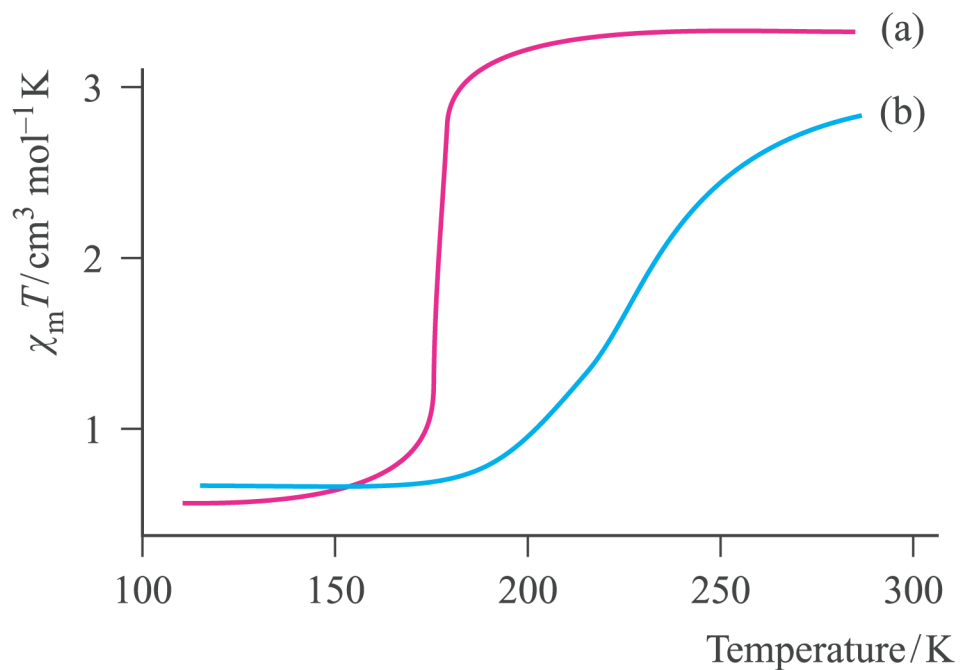
Legge di Curie

Diagramma di Kotani per la configurazione t_{2g}^4

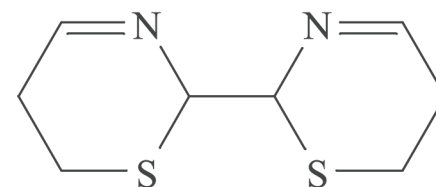


λ = costante di accoppiamento spin-orbita

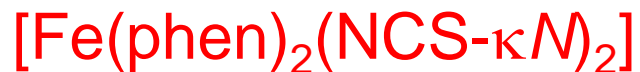
Spin crossover in complexes of Fe(II)



phen = 1,10-phenanthroline



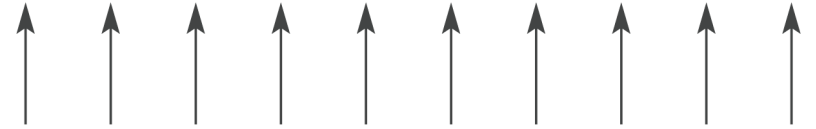
btz = 2,2'-bi-4,5-dihydrothiazine



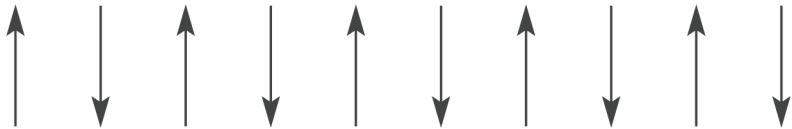
paramagnetismo



ferromagnetismo



antiferromagnetismo

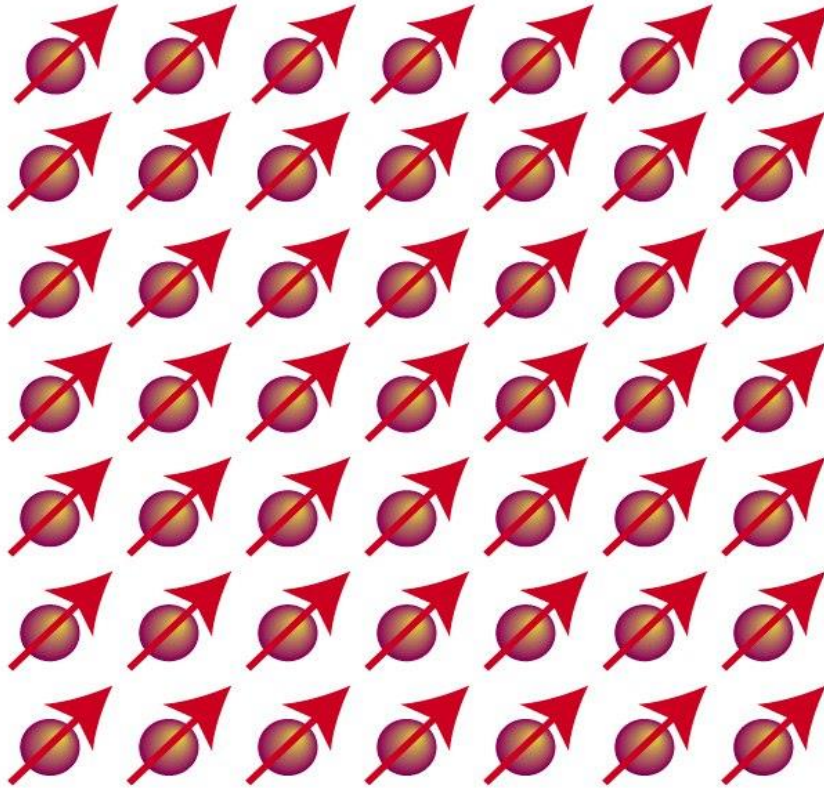


Domini

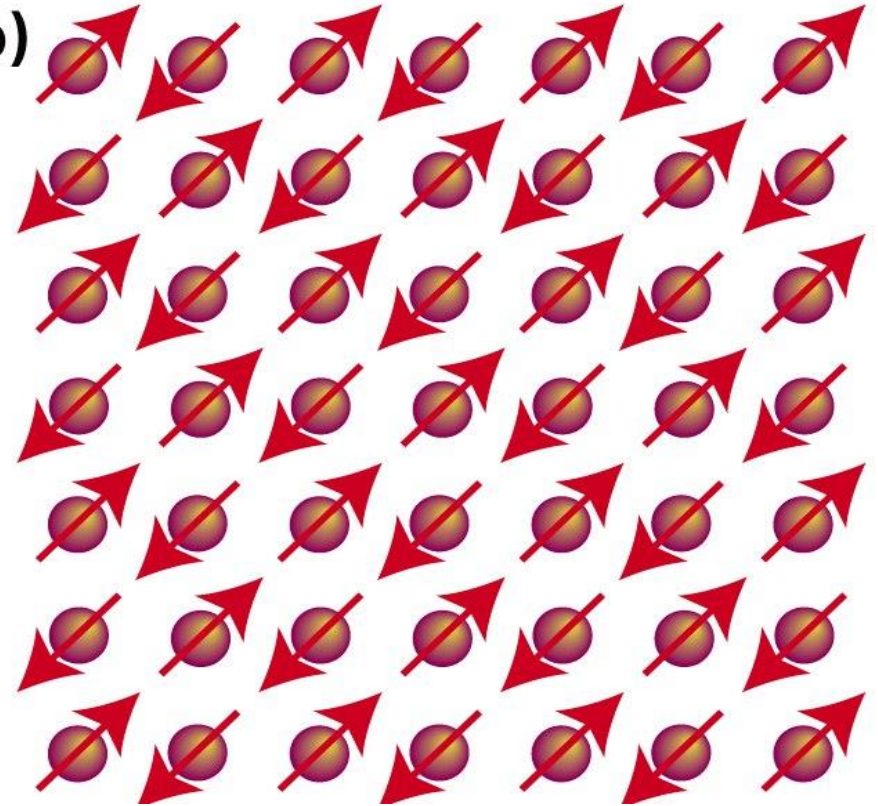
ferromagnetismo

antiferromagnetismo

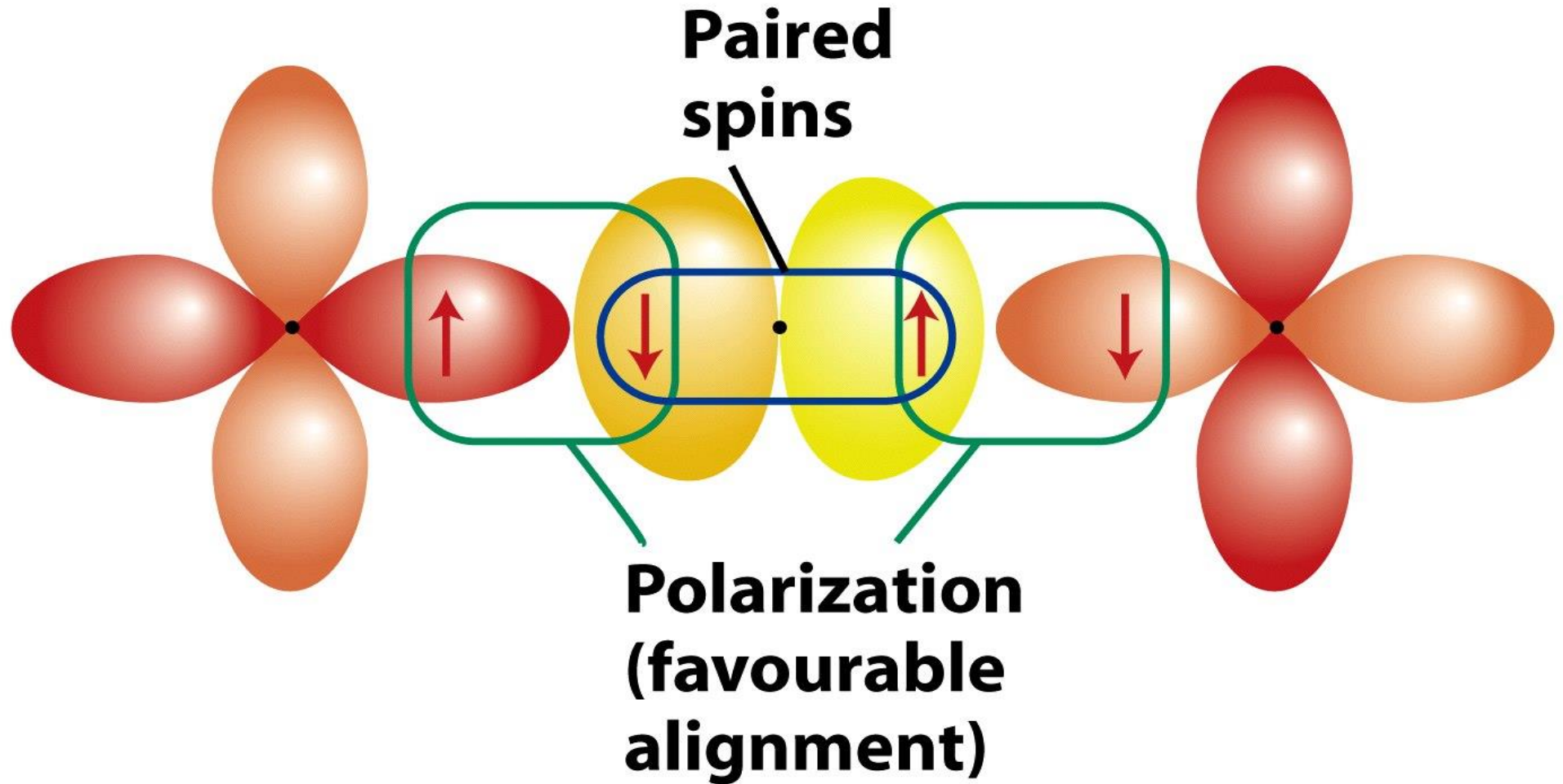
(a)



(b)

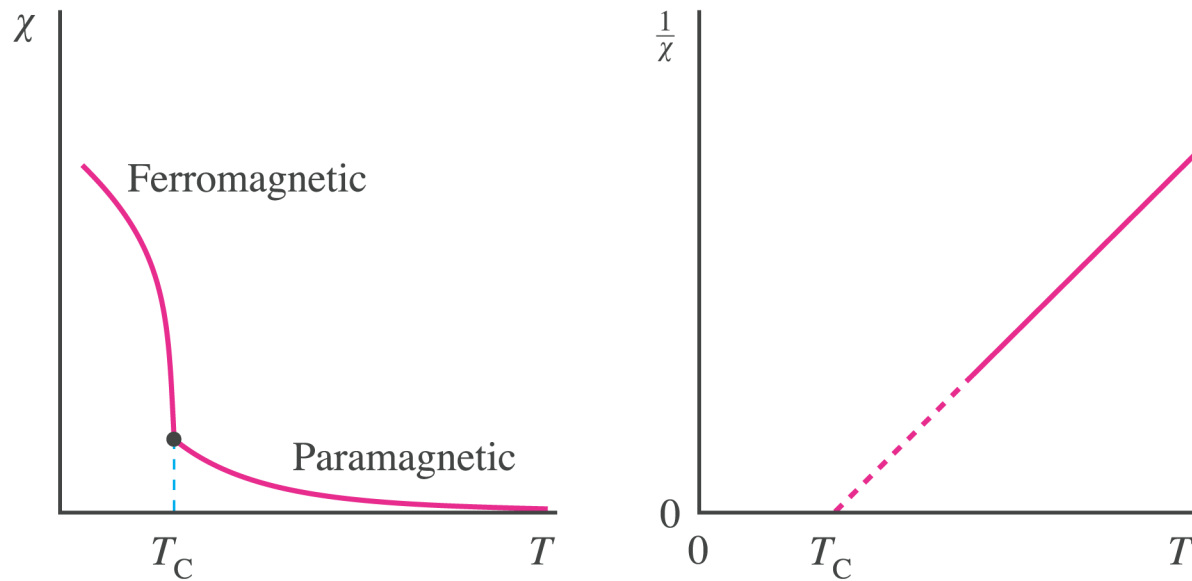


Accoppiamento antiferromagnetico fra due centri metallici mediato dalla polarizzazione di spin di un legante a ponte



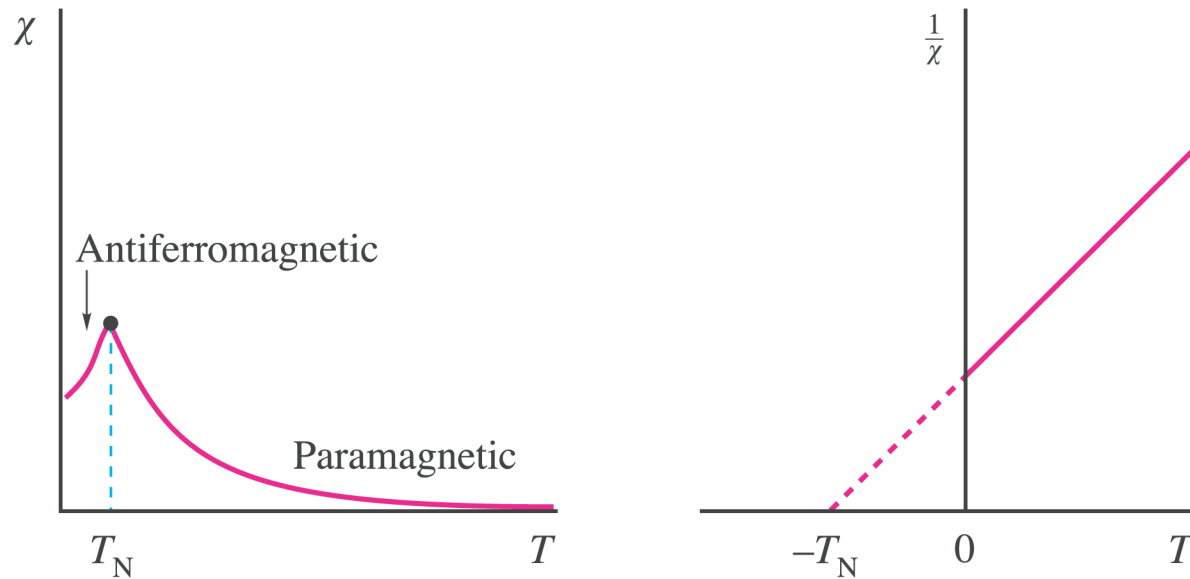
Meccanismo di superscambio

Materiali ferromagnetici

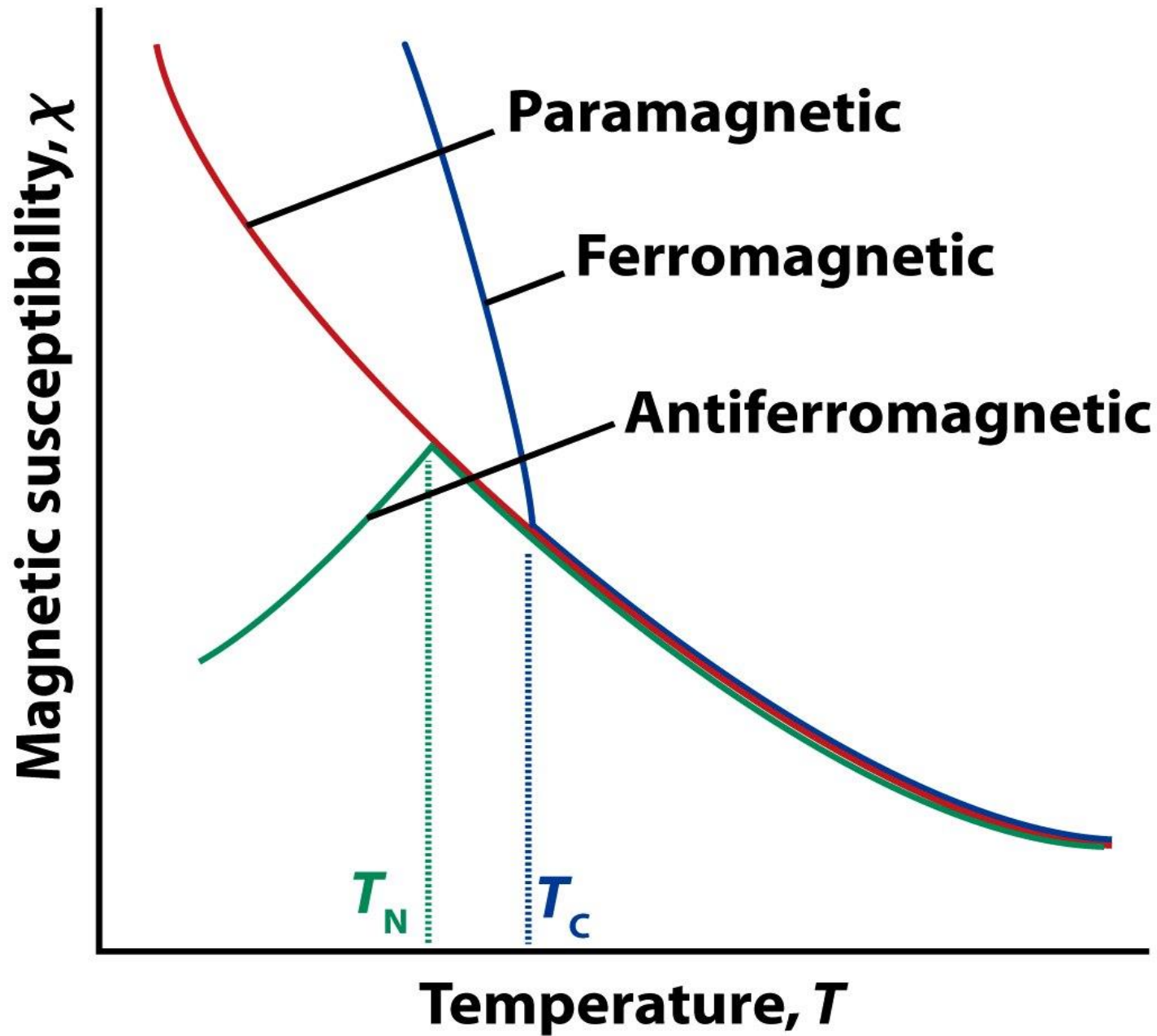


Temperatura di Curie, T_C

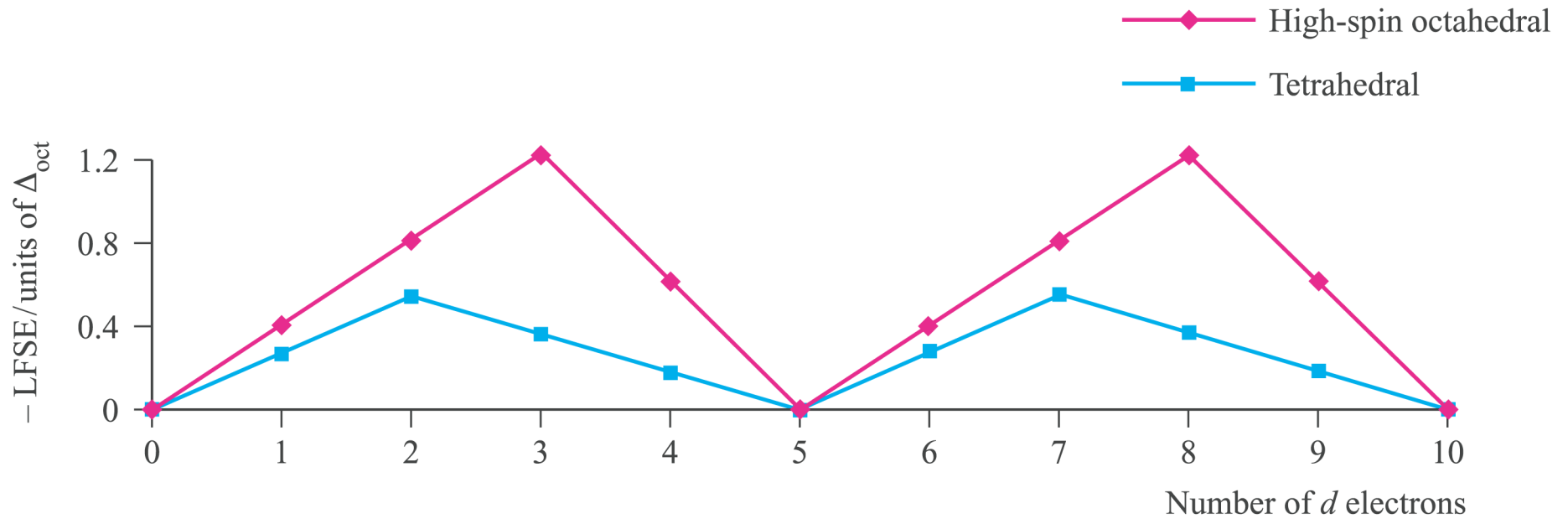
Materiali anti-ferromagnetici



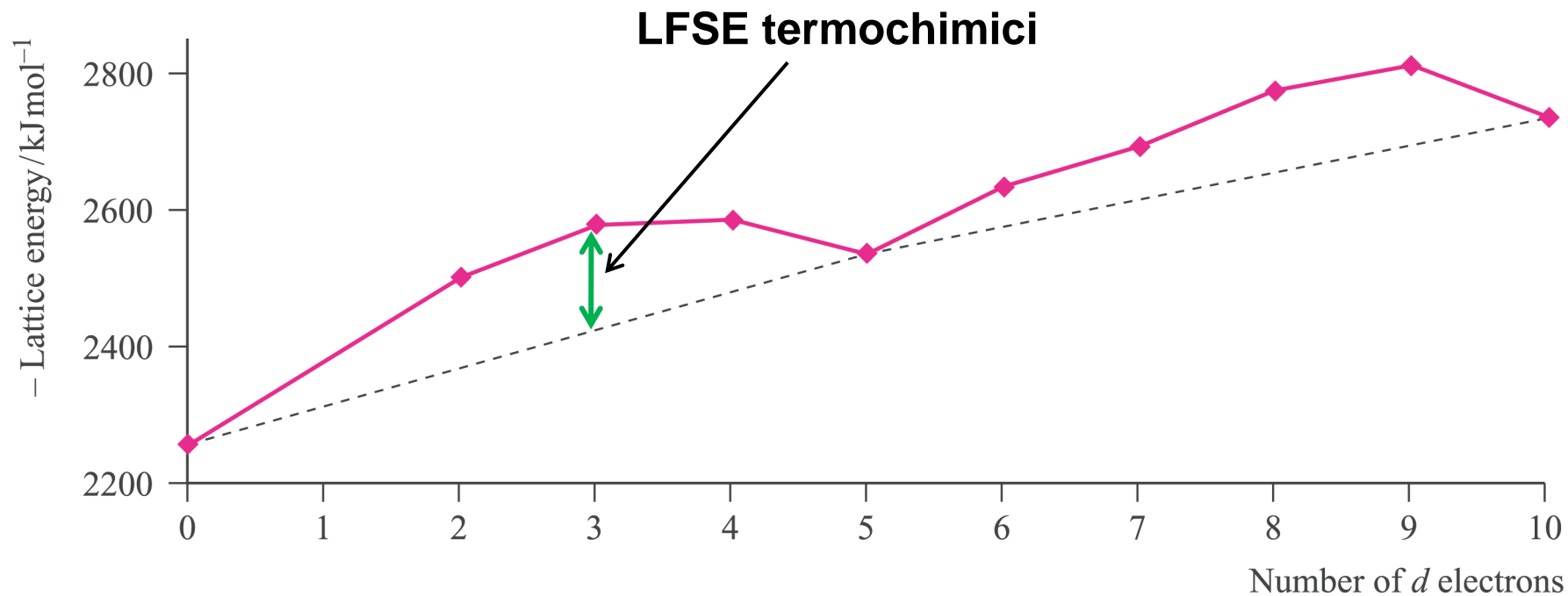
Temperatura di Néel, T_N



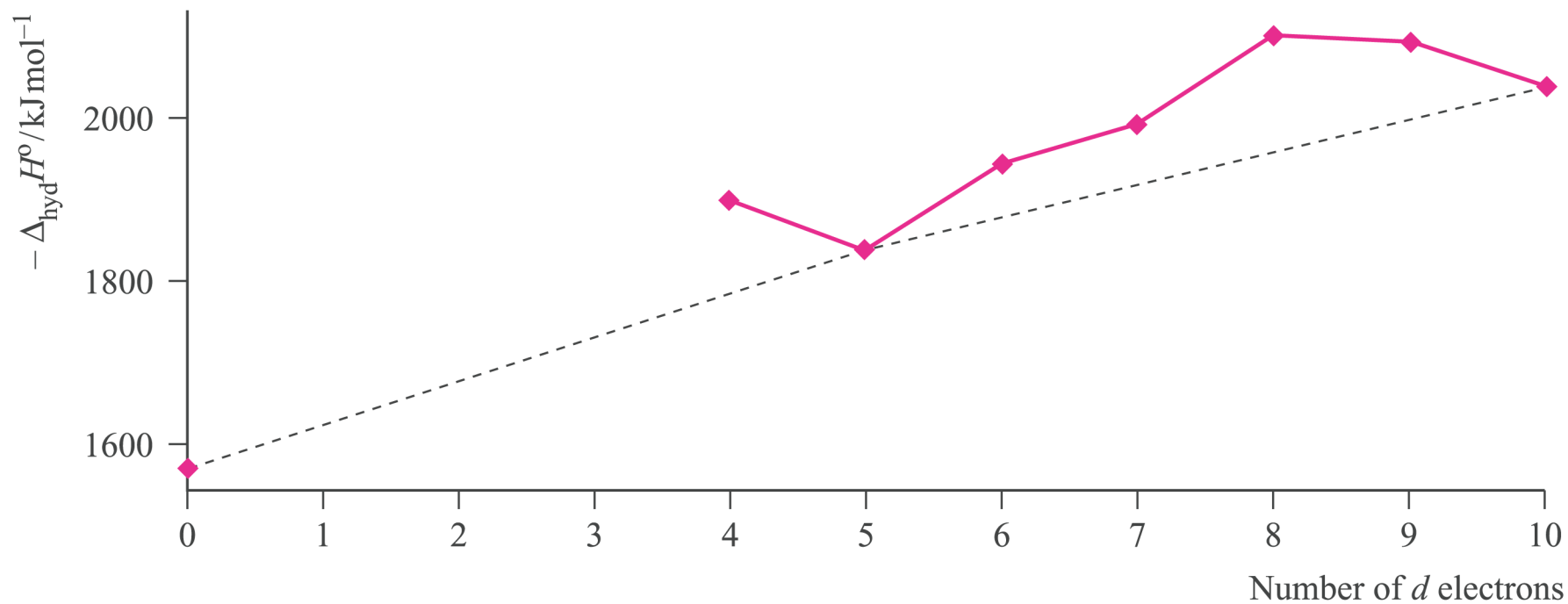
LFSE in funzione di Δ_O



Energie reticolari per MCl_2



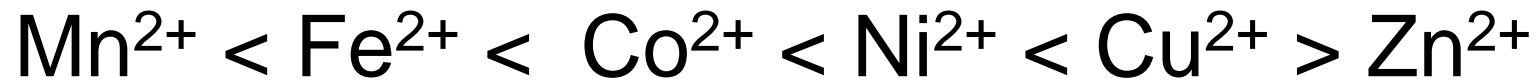
Entalpie di idratazione per ioni M^{2+}



Costanti di stabilità per complessi ottaedrici alto spin

	d ⁵	d ⁶	d ⁷	d ⁸	d ⁹	d ¹⁰
Metal ion	Mn ²⁺	Fe ²⁺	Co ²⁺	Ni ²⁺	Cu ²⁺	Zn ²⁺
log β ₃ for [M(en) ₃] ²⁺	5.7	9.5	13.8	18.6	18.7	12.1
log β for [M(EDTA)] ²⁻	13.8	14.3	16.3	18.6	18.7	16.1

Serie di Irving – Williams



Costanti di stabilità progressive per la sostituzione di H_2O con NH_3

