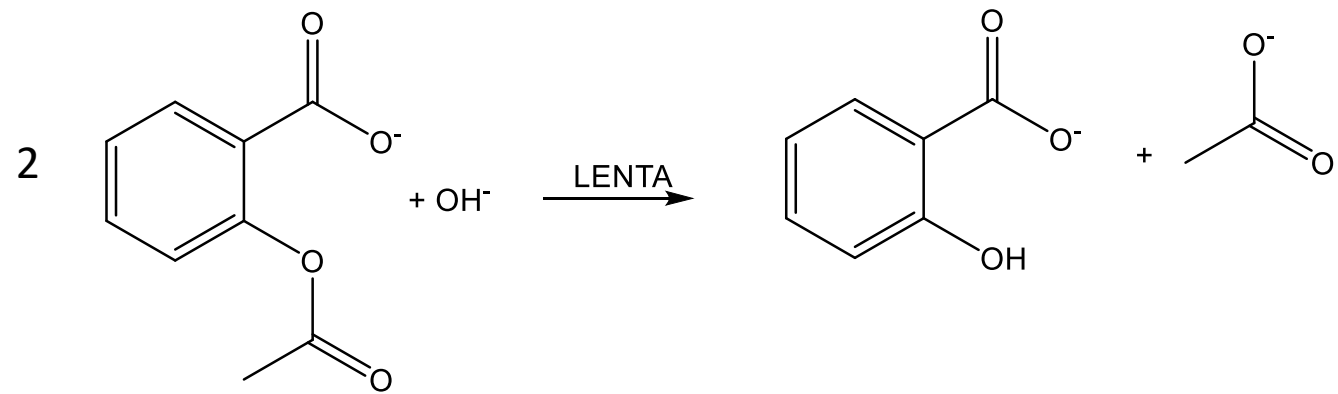
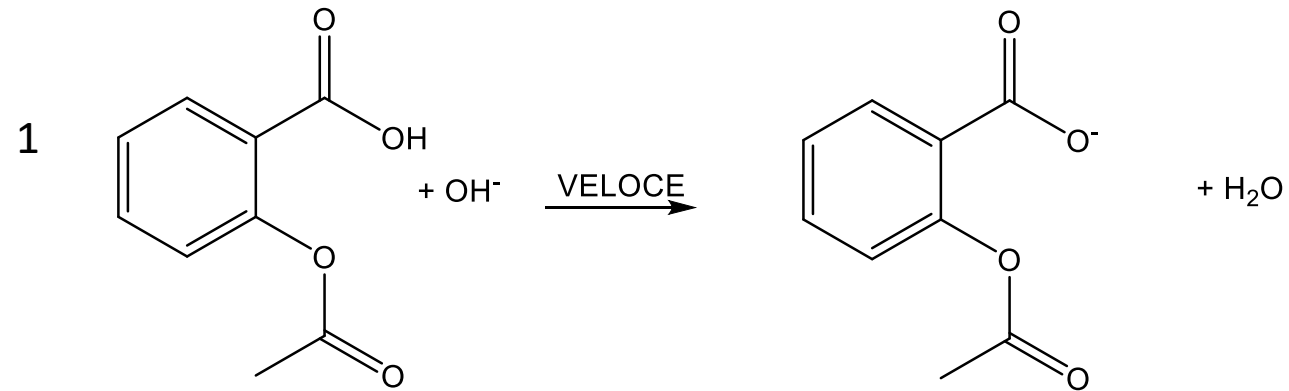
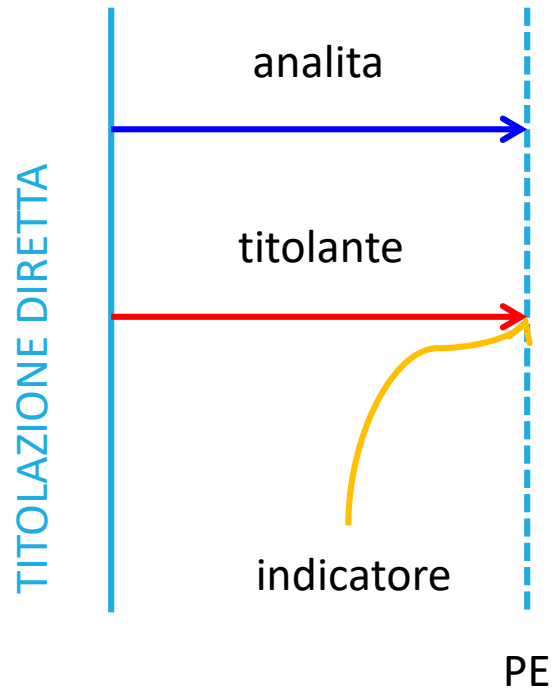


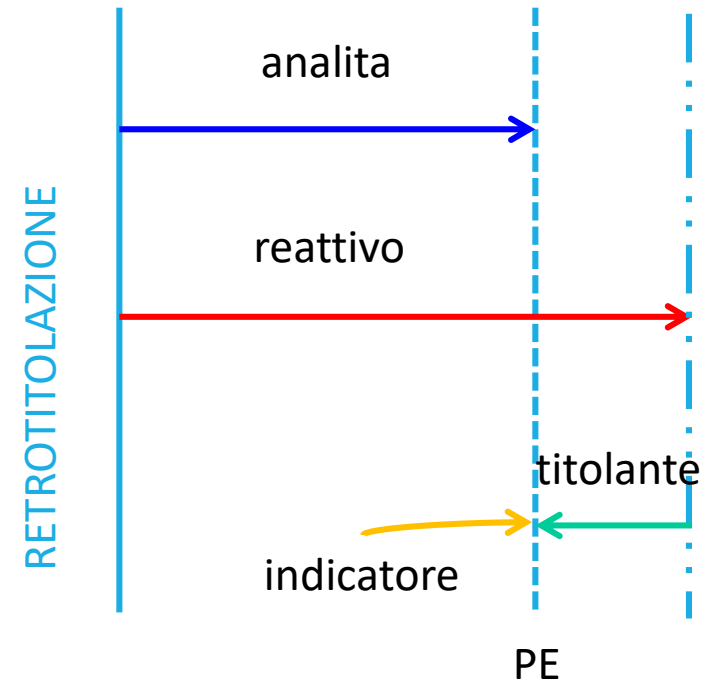
ASA: reazione con NaOH



ASA



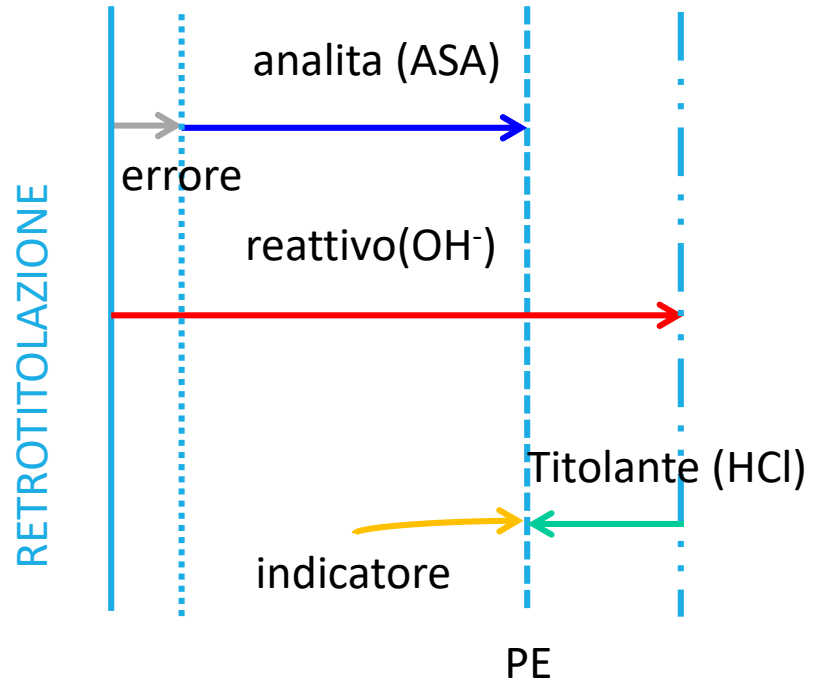
$$n \text{ eq } A = n \text{ eq } T$$



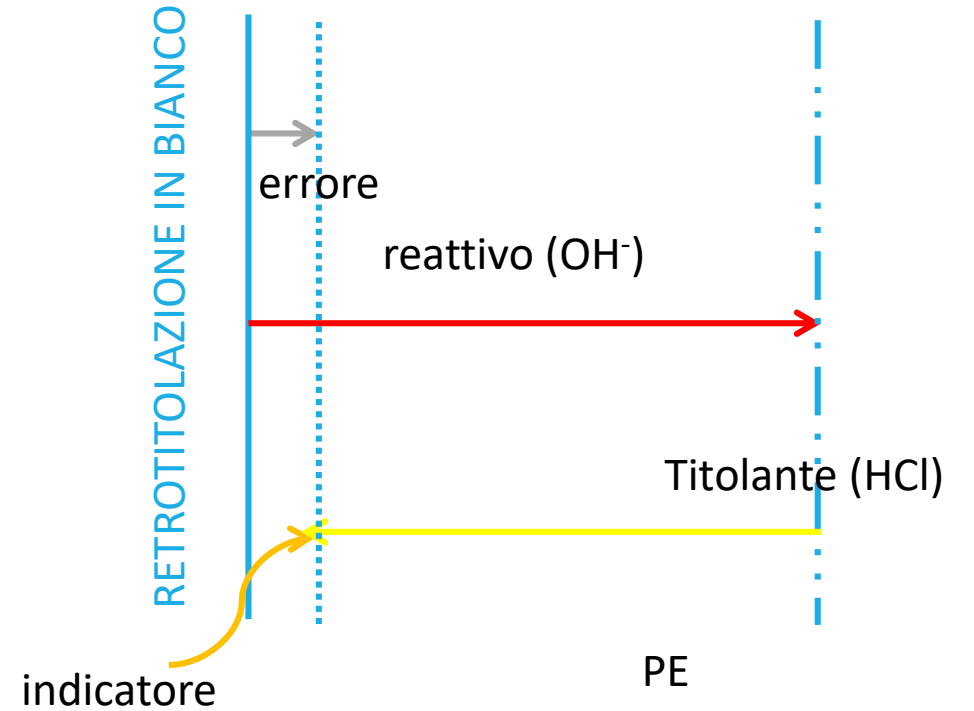
$$n \text{ eq } A = n \text{ eq } R - n \text{ eq } T$$

ASA

TITOLAZIONE



TITOLAZIONE IN BIANCO



$$n \text{ eq int (errore)} = n \text{ eq OH}^- - n \text{ eq HCl bianco}$$

errore dovuto a sostanze interferenti

$$n \text{ eq A} = n \text{ eq R} - n \text{ eq T} - n \text{ eq int (errore)}$$

ASA: calcoli

Analita ASA

Reattivo NaOH 0,5 M

Titolante HCl 0,5M

} stessa concentrazione!

$$n \text{ eq ASA} = n \text{ eq OH}^- - n \text{ eq HCl} - n \text{ eq int (errore)}$$

$$n \text{ eq int (errore)} = n \text{ eq OH}^- - n \text{ eq HCl bianco}$$

Dati:

Retrotitolazione $V(\text{OH}^-) - V(\text{H}^+) = V(\text{OH}^- + \text{errore})$

Bianco $V(\text{OH}^-) - V(\text{H}^+) = V(\text{OH}^- \text{ errore})$

} perché [NaOH] = [HCl]

$$V_{\text{eq}}(\text{OH}^-) = V(\text{OH}^- + \text{errore}) - V(\text{OH}^- \text{ errore})$$

$$g(\text{ASA})/\text{PE}(\text{ASA}) = [N(\text{OH}^-) * V_{\text{eq}}(\text{OH}^-)]/1000$$

$$g(\text{ASA}) = [N(\text{OH}^-) * V_{\text{eq}}(\text{OH}^-) * \text{PE}(\text{ASA})]/1000$$



Equivalente volumetrico in FU:
1 ml NaOH 0,5 M = 45,04 mg di ASA

$$\text{PE}(\text{ASA}) = \text{PM}(\text{ASA})/2$$