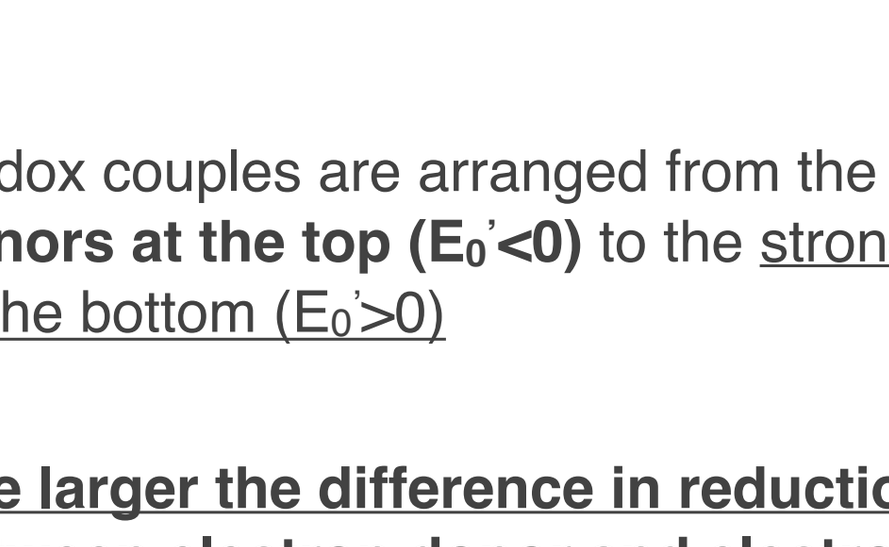
L03b

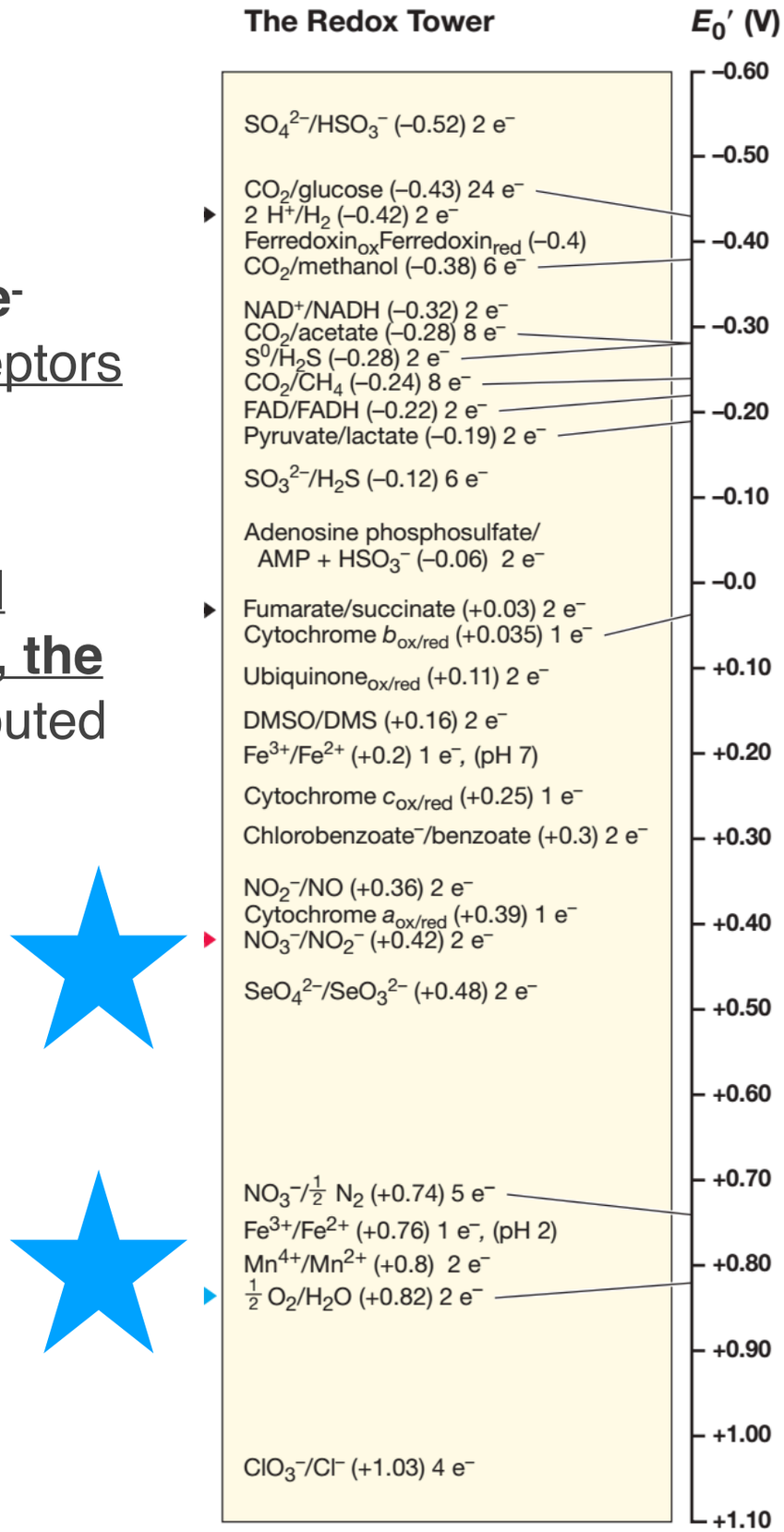
Recap L03

The Redox Tower

- Redox couples are arranged from the **strongest e- donors at the top (E_0'<0)** to the strongest e- acceptors at the bottom (E_0'>0)

- **The larger the difference in reduction potential between electron donor and electron acceptor, the more free energy is released** (ΔG_0' can be computed via Nernst equation from reduction potential)





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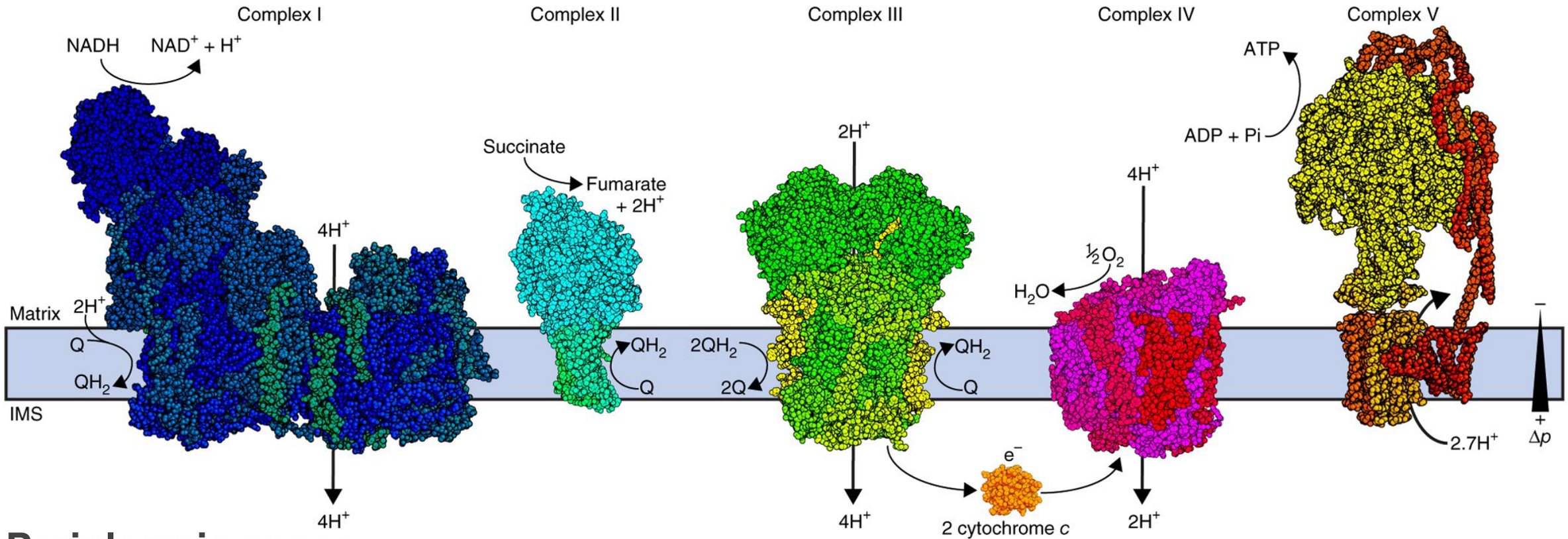
Electron transport chain (ETC), I

Cytoplasm

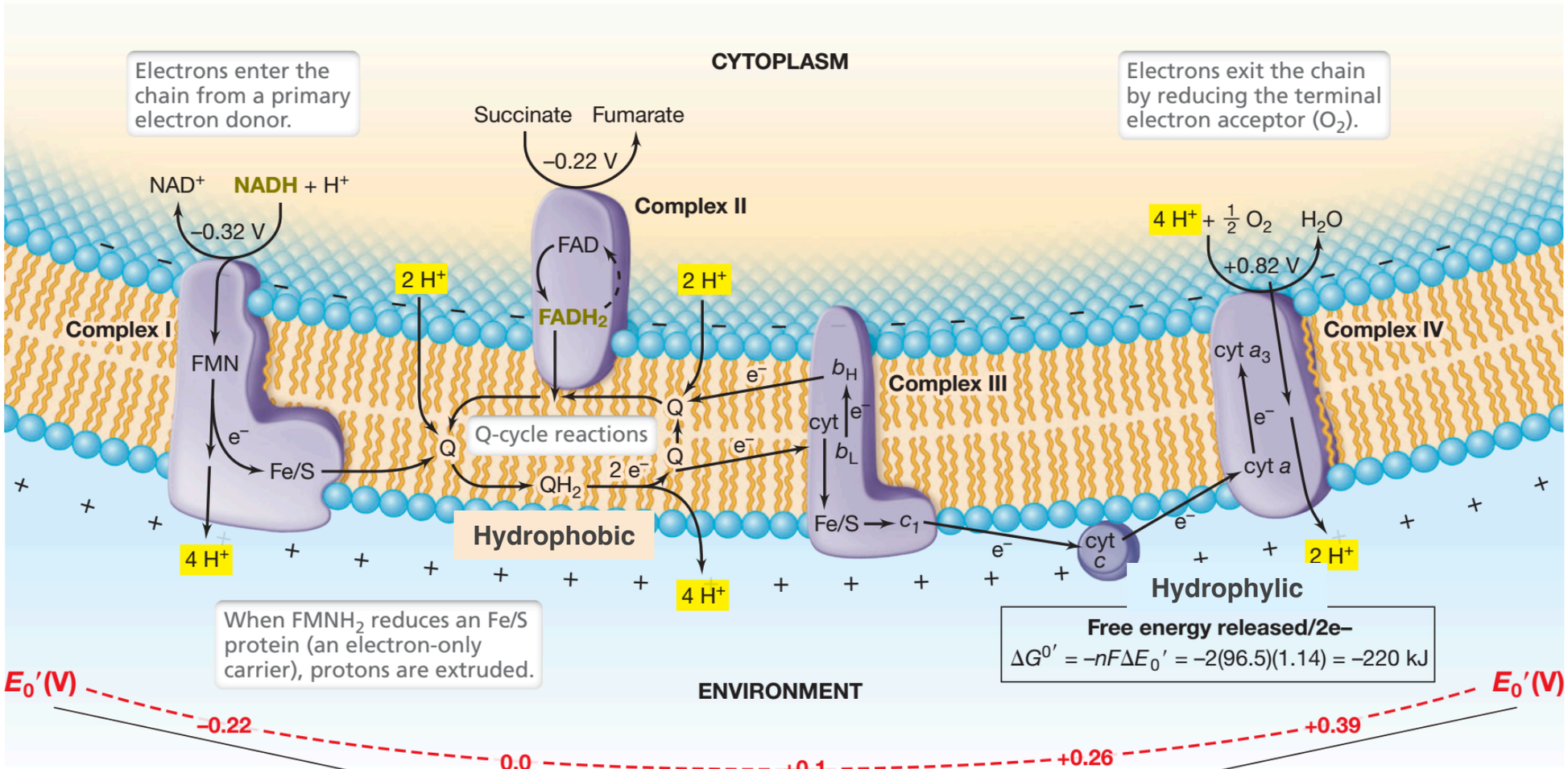


Periplasmic space

- In the membrane
- Intimate interaction between proteins (dehydrogenase, flavoproteins, iron-sulfur proteins) and diffusible molecules (quinons and cytochromes)
- Electrons are swapped
- Protons are pumped outside the cell (cytoplasm —> periplasmic space)

Electron transport chain, II

- A set of membrane-bound electron carriers (4) organized from **high to low redox** potentials —> **spontaneous** flow of electrons to the **terminal electron acceptor**
- The **membrane carriers are not structurally linked** so they can **diffuse** laterally in the membrane and collide with one another to promote the rapid exchange of electrons
- *Escherichia coli* uses lipophilic organic molecules called **quinones to electronically link a dehydrogenase enzyme complex to a specific terminal reductase**



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Electron transport chain, III

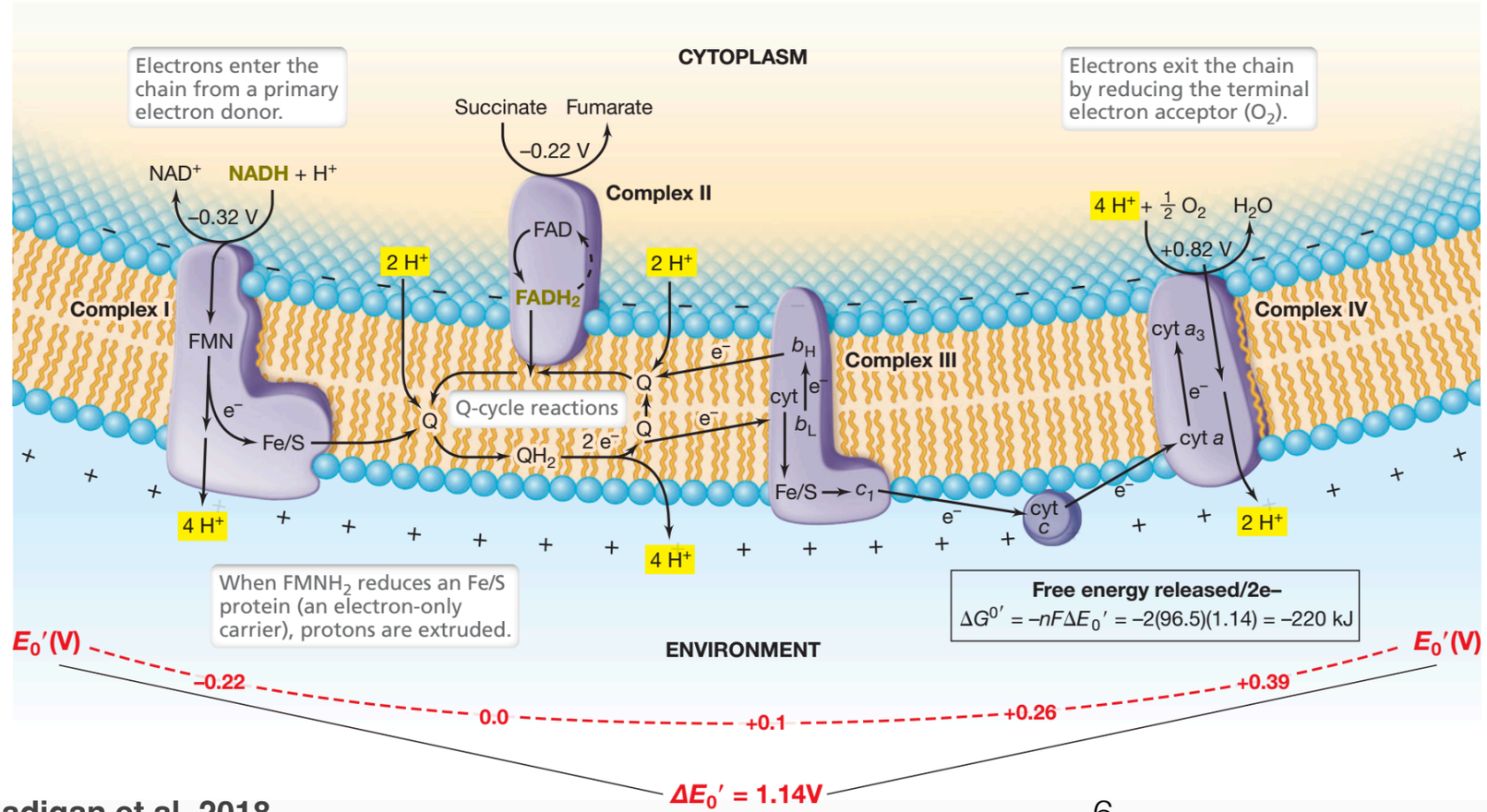
General features:

(1) **Carriers** are arranged in order of **increasingly more positive E_0'** (reduction potential)

(2) **Alternation of electron-only and electron-plus-proton carriers** in the chain

(3) Net result is **reduction of terminal electron acceptor** (such as O_2) + **generation proton motive force** (PMF, thanks to harnessing e^- flow)

(4) ATP production by PMF (ATP synthesis is driven by an ion gradient through the activity of ATP synthase)



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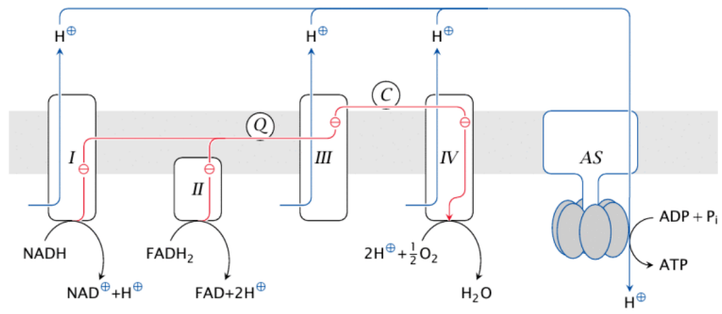
Environment

H+ flow



e- flow

Cytoplasm

Electron transport chain, III

General features:

(1) **Carriers** are arranged in order of **increasingly more positive E_0'** (reduction potential)

(2) **Alternation of electron-only and electron-plus-proton carriers** in the chain

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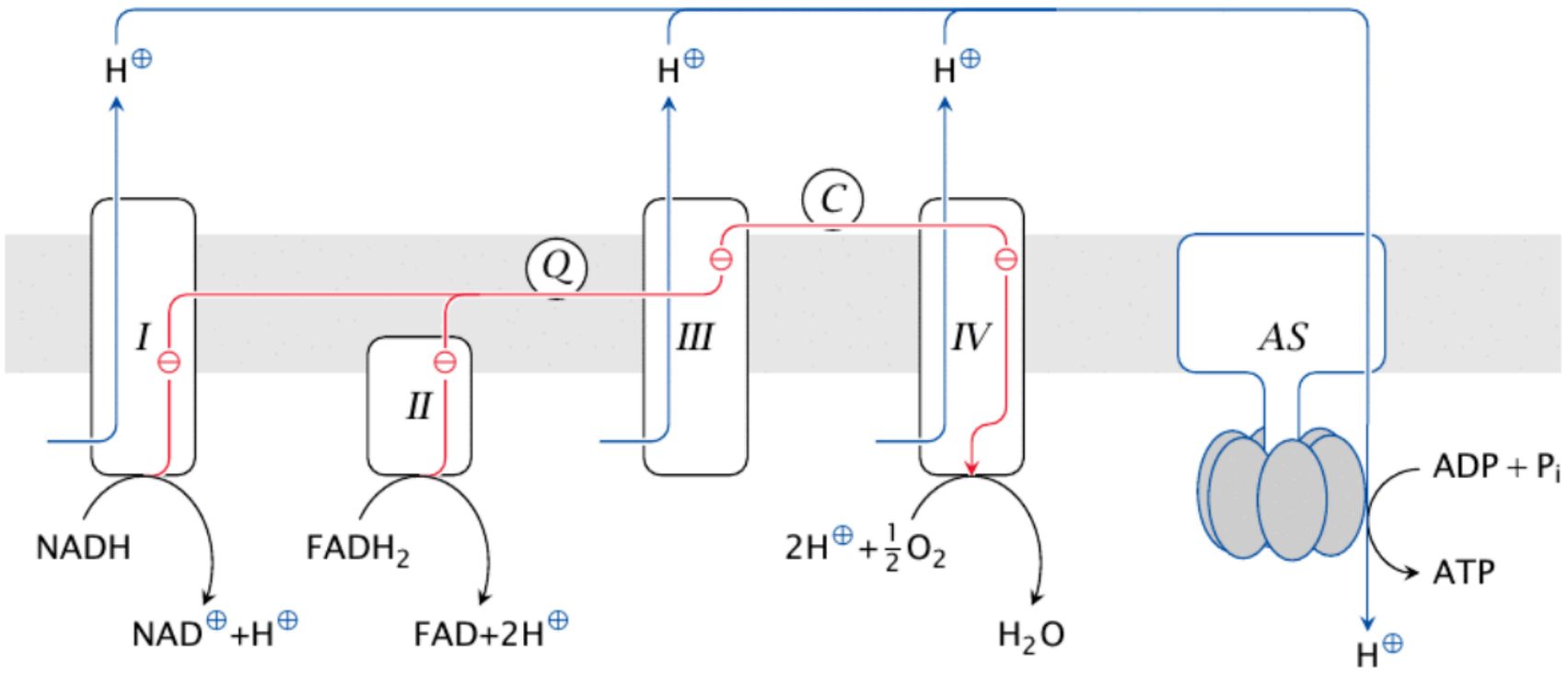

H+ flow

e- flow

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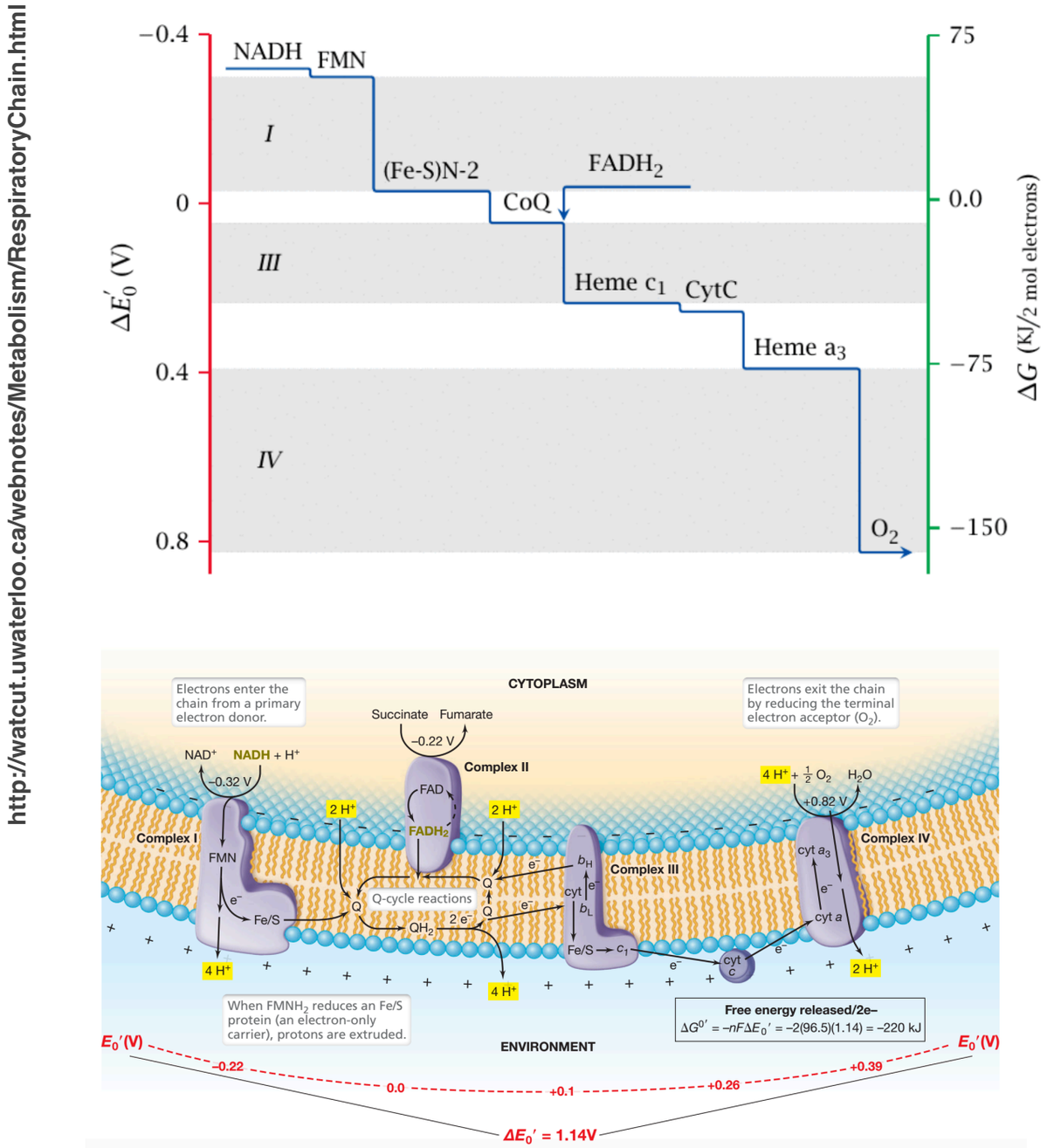
Structural orientation for ATP production

Redox potentials and free energies in the respiratory chain

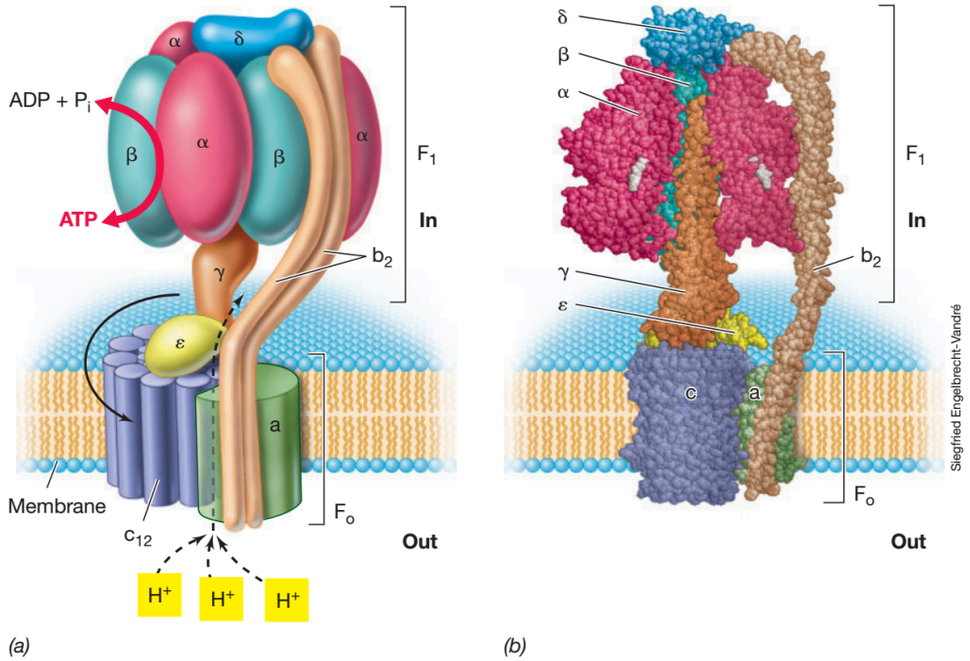


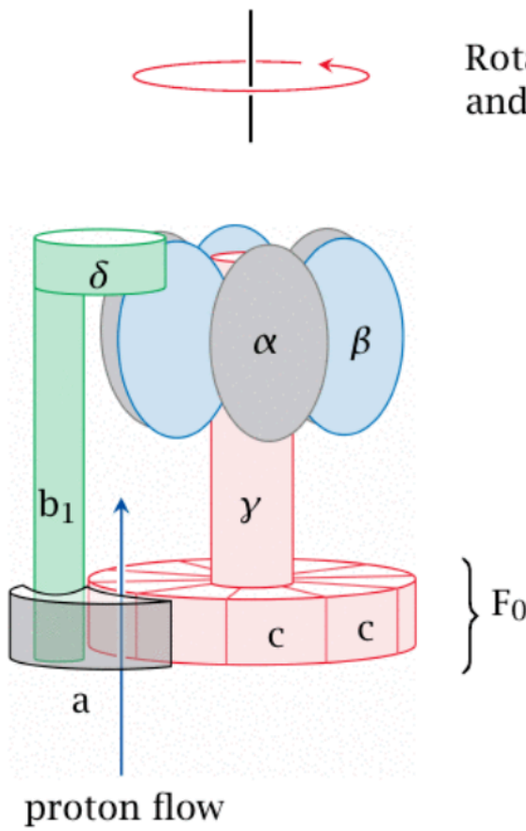
http://watcut.uwaterloo.ca/webnotes/Metabolism/RespiratoryChain.html

- Spontaneous flow of electrons (E_0')
- H^+ are separated from e^- across membrane (spatial localization ETC)
- Inner and outer surfaces of the membrane differ in charge, pH, and electrochemical potential
- Electrochemical potential is proton motive force (PMF) and energizes the membrane, much like a battery
- Only three of the four mentioned electron carriers are capable of transporting protons from the matrix to the intermembrane space: I, III, and IV

Madigan et al. 2018

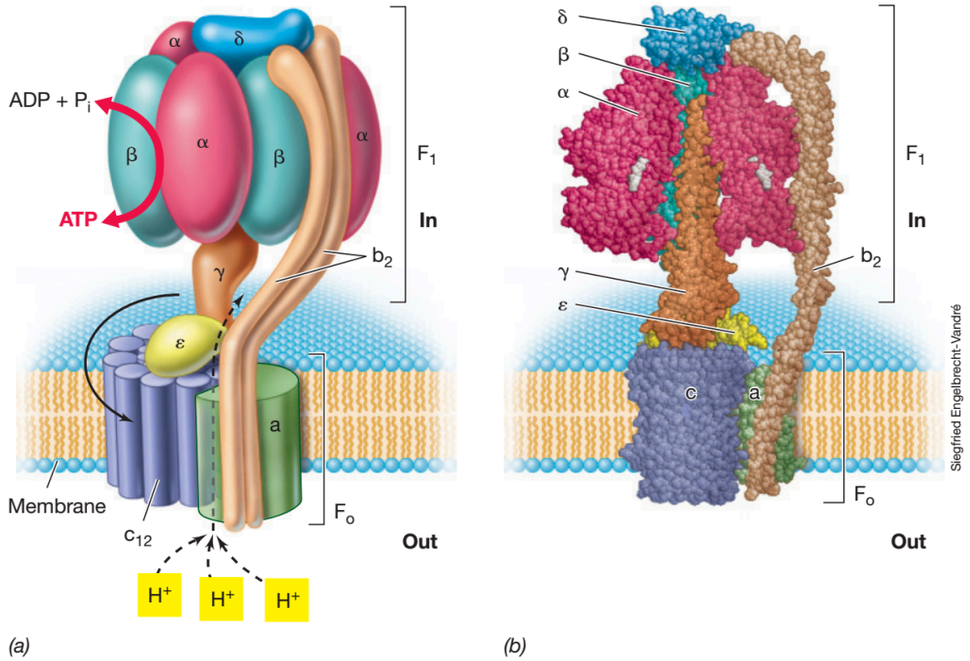
ATP production

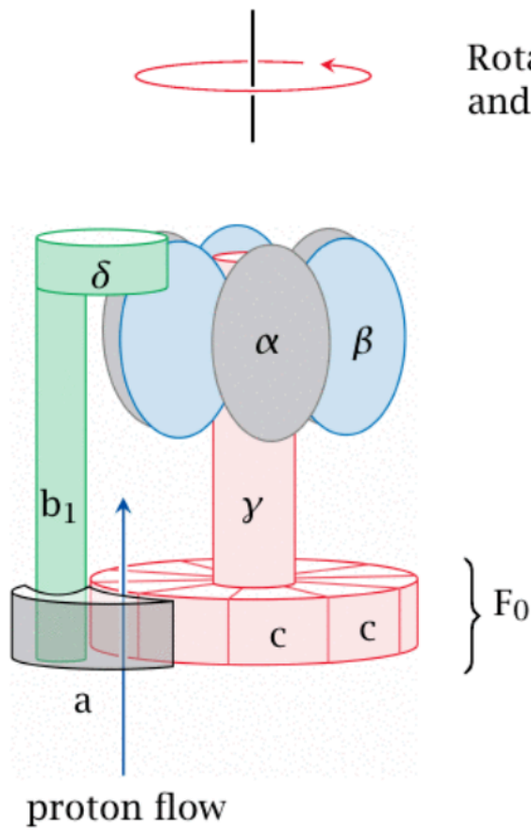




- **H^+ gradient that drives phosphorolation of ADP to ATP as well as several other important transport systems (nutrient transport, flagellar rotation, and other energy-requiring reactions)**
- **3 H^+ —> ATP (Noguchi et al., 2004): F1 is the catalytic complex responsible for the interconversion of ADP + Pi and ATP. Fo, the rotor, is integrated in the membrane**

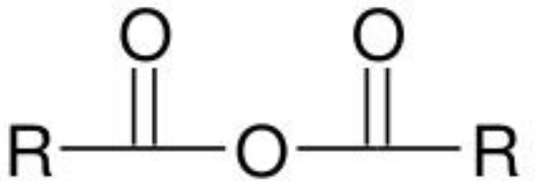
ATP production



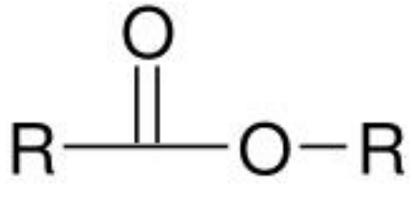


- **In analogy to how dissipation of the pmf applies torque that rotates the bacterial flagellum, the pmf also creates torque in the large membrane protein complex that synthesizes ATP**
- **This complex is called ATP synthase (ATPase)**
- **The activity of ATPase is driven by the pmf, and the formation of ATP from respiratory electron flow is called oxidative phosphorylation (contrast this with substrate-level phosphorylation in fermentation)**

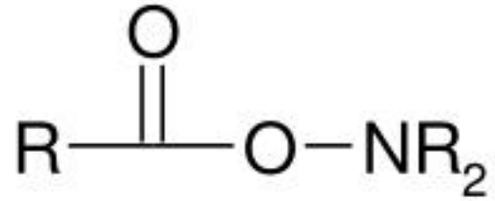
C and P: Anhydrides and Esters



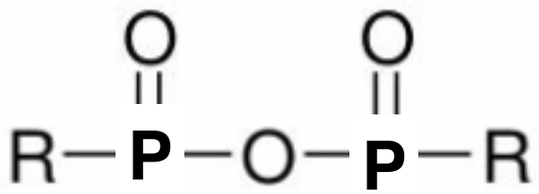
Anhydride



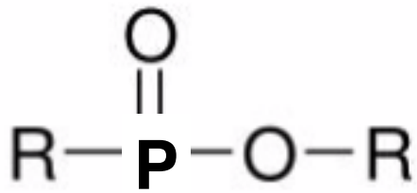
Ester



Amide



Anhydride



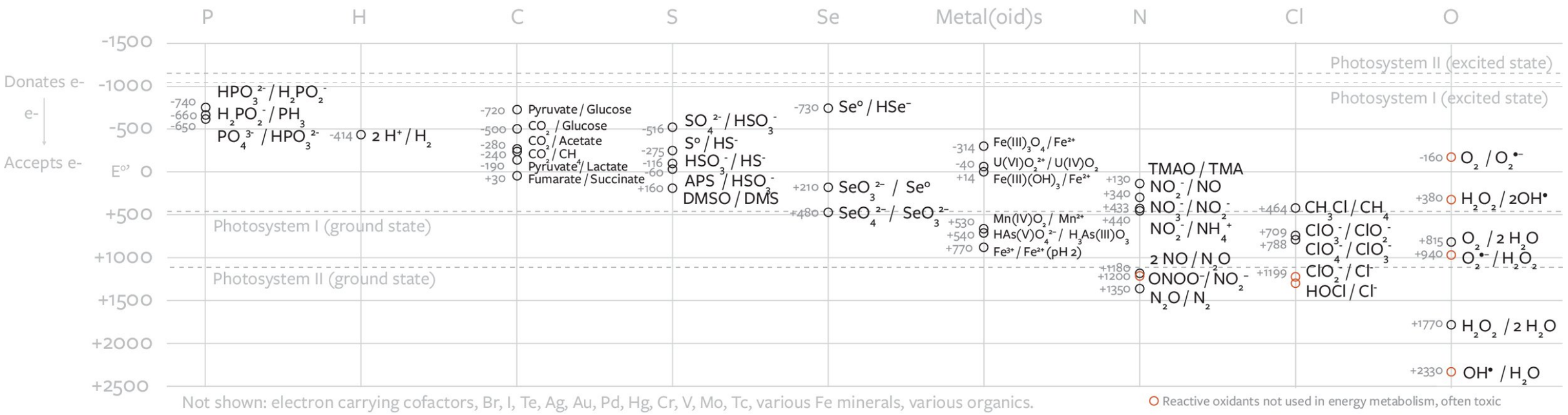
Ester

ATP

Microbial Redox couples

Redox couples and potentials (mV) for elements common in biology at pH 7 and temperature 25 C *

Redox potential indicates the propensity for a compound to transfer electrons to another compound. A more-negative redox potential means a compound is more likely to donate electrons (e-).
All of life gets its energy by capturing the change in potential energy from the transfer of electrons from the reducing compound to the oxidizing compound.



Not shown: electron carrying cofactors, Br, I, Te, Ag, Au, Pd, Hg, Cr, V, Mo, Tc, various Fe minerals, various organics.

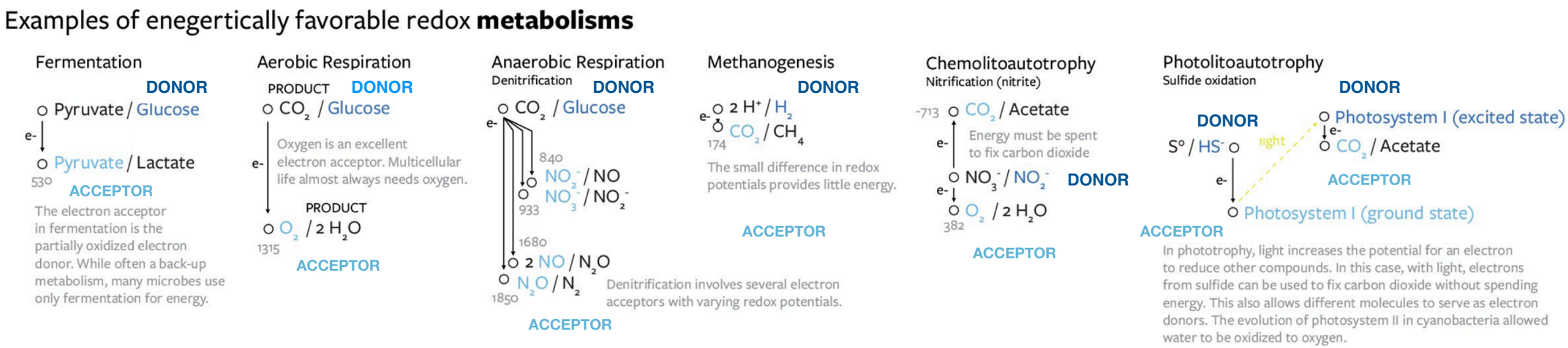
Reactive oxidants not used in energy metabolism, often toxic

- **Across periodic table**
- **P, H, C, S, Se, Fe, U, Mn, As, N, Cl, O**

* For teaching purposes only. Consult the scientific literature for exact values.

Image produced by Tyler Barnum @tylerbarnumphd

Microbial Redox couples structure the metabolism



* For teaching purposes only. Consult the scientific literature for exact values.

Image produced by Tyler Barnum @tylerbarnumphd

Microbial diversity and metabolic pathways to survive in the environment

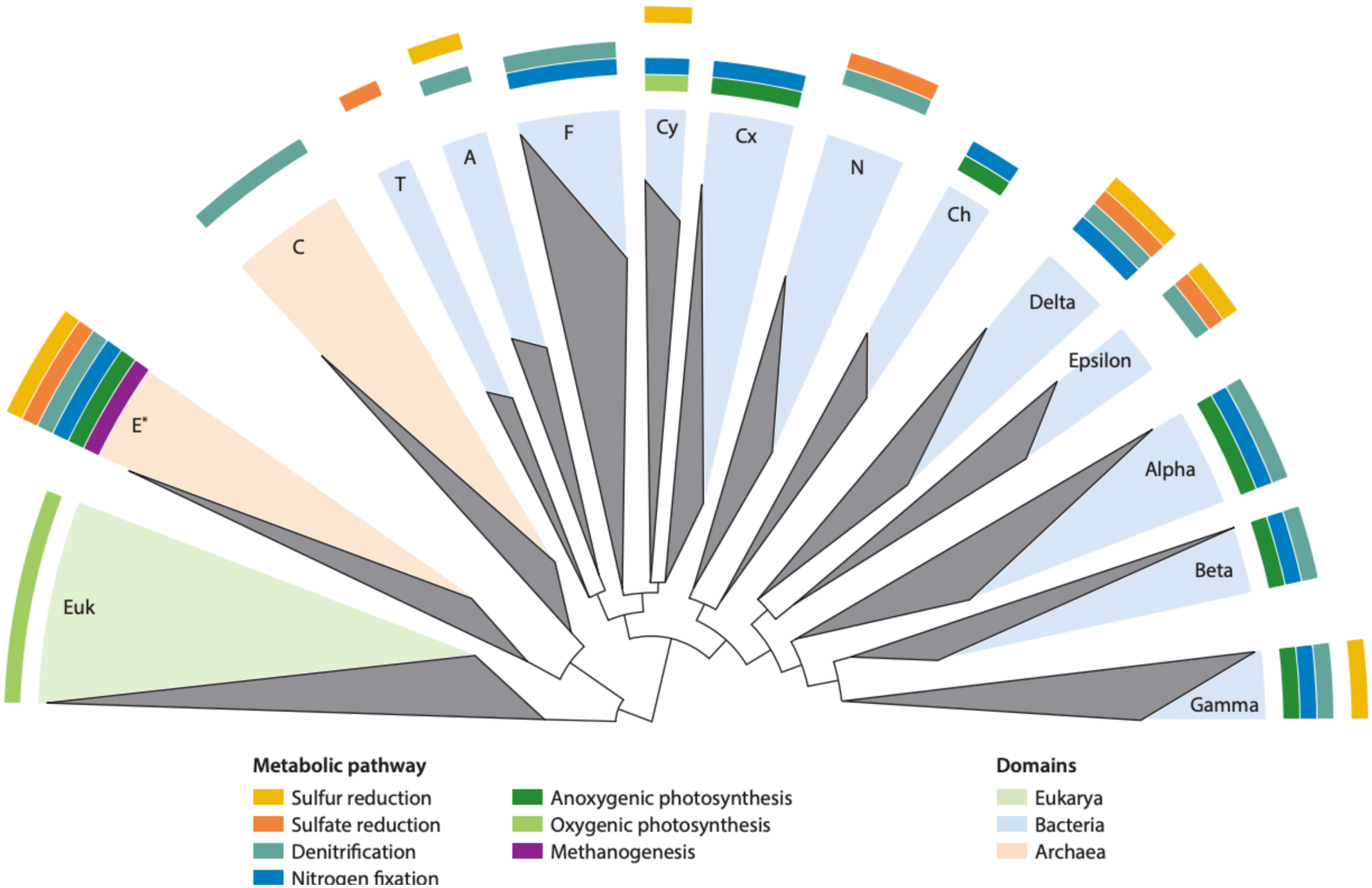
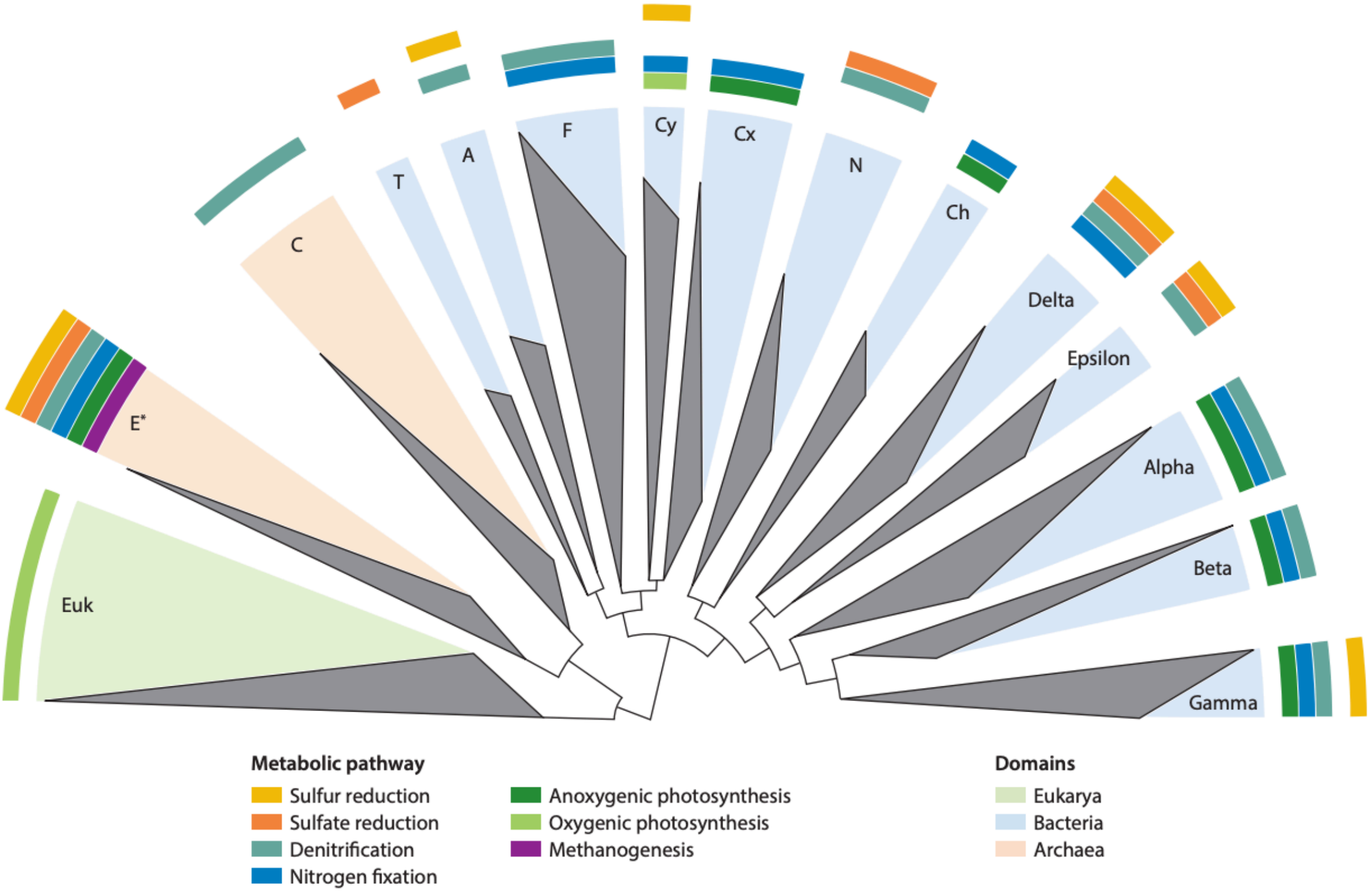


Figure 1

Distribution of selected metabolic pathways on the 16S rRNA tree of life. The tree (constructed with ARB; 104) was edited for clarity and shows selected bacterial and archaeal taxa. The area of each branch is proportional to the total number of 16S rRNA sequences present in the database. Metabolic pathways were assigned based on physiological data (**Supplemental Table 2**). Sulfate reduction includes sulfite and thiosulfate reduction pathways. **Euryarcheata* are capable of bacteriorhodpsin-based photosynthesis only. Abbreviations: A, *Aquificae*; Alpha, *Alphaproteobacteria*; Beta, *Betaproteobacteria*; C, *Crenarchaeota*; Ch, *Chlorobi*; Cx, *Chloroflexi*; Cy, *Cyanobacteria*; Delta, *Deltaproteobacteria*; E, *Euryarchaeota*; Epsilon, *Epsilonproteobacteria*; Euk, *Eukarya*; F, *Firmicutes*; Gamma, *Gammaproteobacteria*; N, *Nitrospirae*; T, *Thermodesulfobacteria*.

Jelen et al. 2016

Microbial diversity and metabolic pathways to survive in the environment

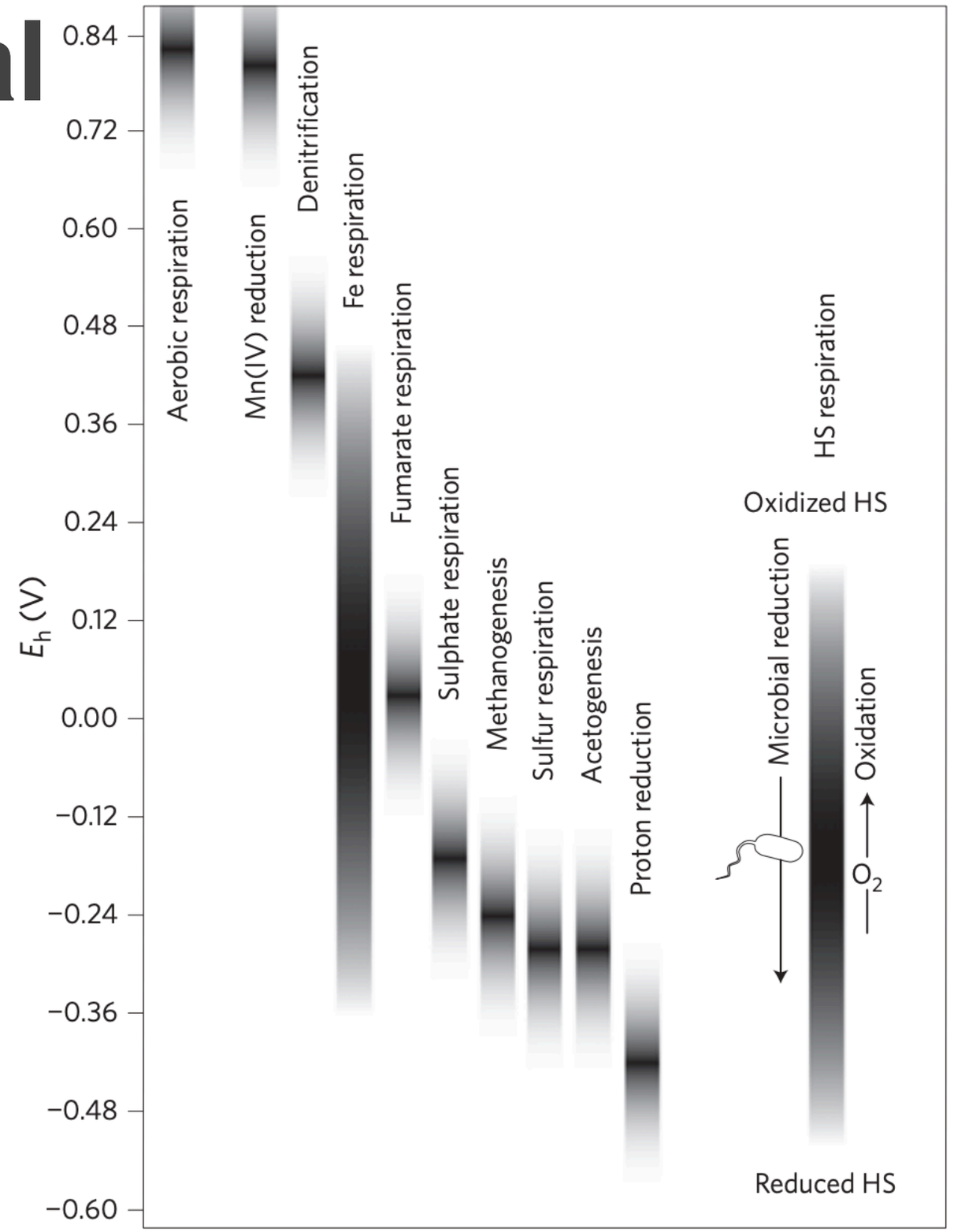


Jelen et al. 2016

Reduction potential ranges of microbial respiration

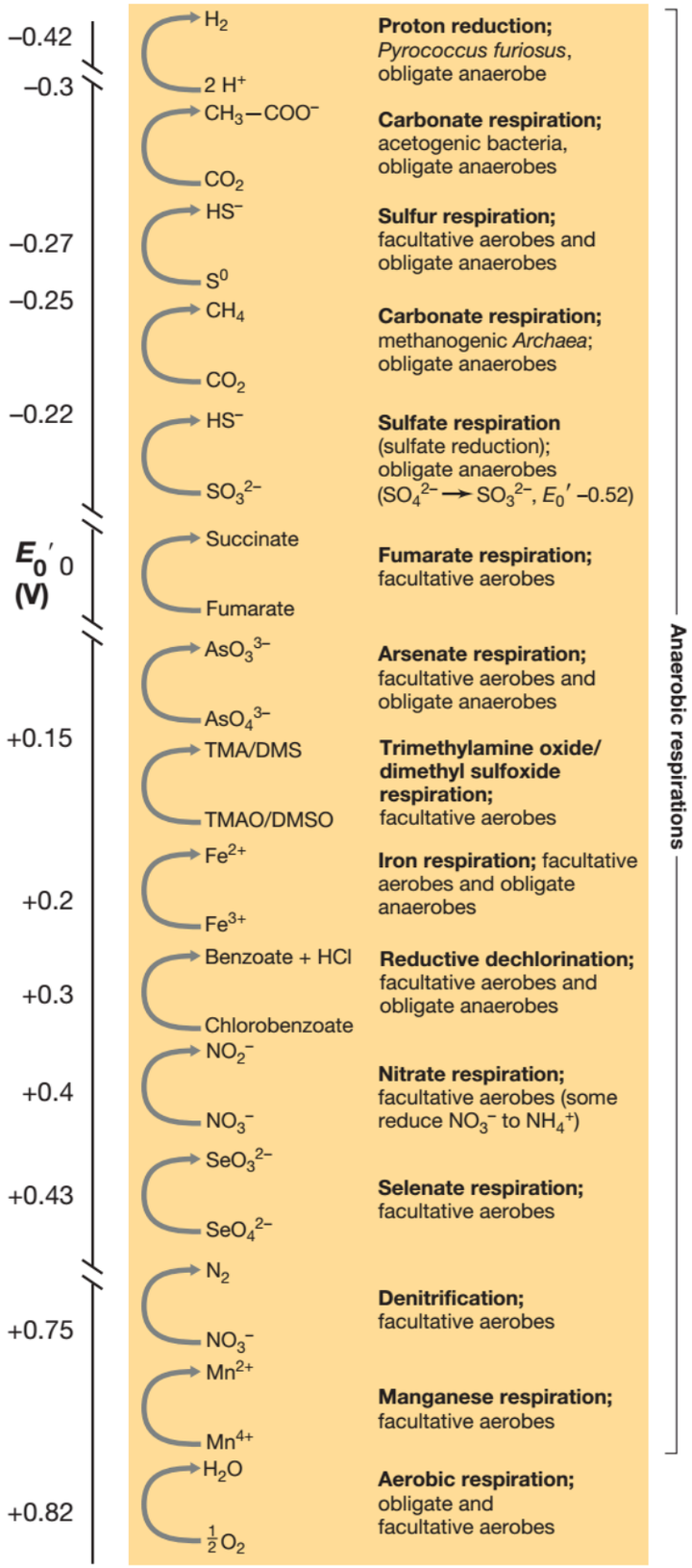
- The achievable energy yield of ETC depends on the difference in electrical potential between electron donor and acceptor

- Microbes able to respire in multiple ways will always choose available acceptors with the **biggest potential difference** to the donor (e.g., *E. coli* O2 > NO3-> fumarate)

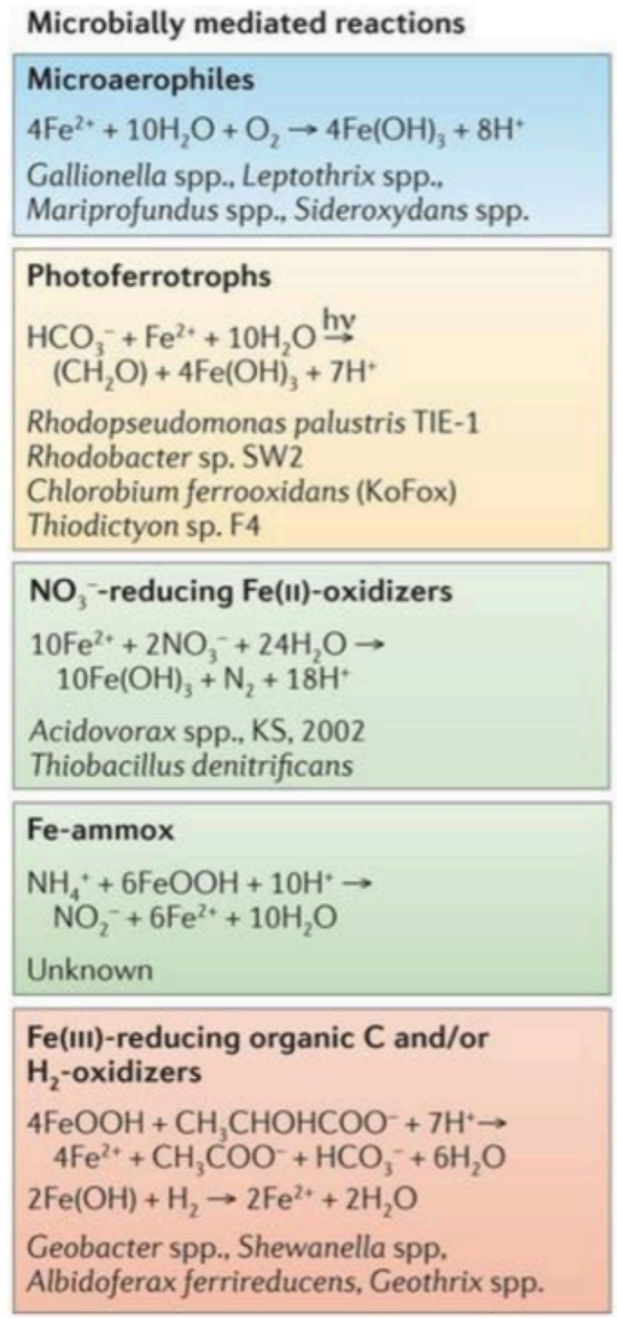


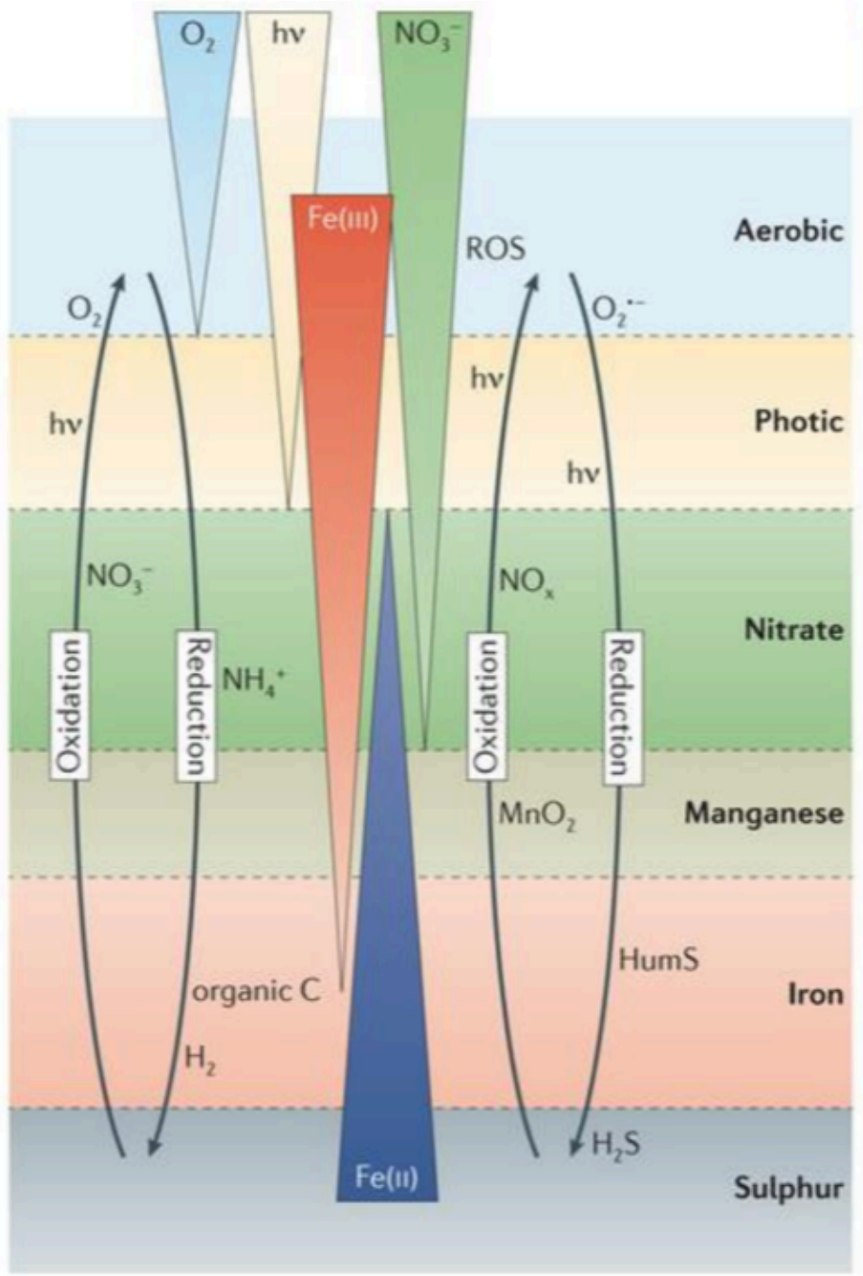
Klüpfel et al., 2014

Anaerobic respiration



Madigan et al. 2018





Melton et al., 2014

Fermentation/*Respiration*

- **Fermentation is a form of anaerobic catabolism in which organic compounds both donate electrons and accept electrons, and redox balance is achieved without the need for external electron acceptors**

- **ATP is made from these energy-rich compounds by substrate-level phosphorylation, a process whereby the energy-rich phosphate bond on the organic compound is transferred directly to ADP to form ATP**

- **Glucose fermentation into alcoholic or lactic acid: 2 ATP**

- *Respiration is a form of aerobic or anaerobic catabolism in which an organic or inorganic electron donor is oxidized with O_2 (in aerobic respiration) or some other compounds (in anaerobic respiration) functioning as electron acceptors*

- *ATP is made by PMF*

- *Glucose aerobic respiration into CO_2: 38 ATP*

Fermentation, II

TABLE 3.4 **Common fermentations and some of the organisms carrying them out**

Type	*Reaction (substrate → products)*	*Organisms*
Alcoholic	Hexose[a] → 2 ethanol + 2 CO_2	Yeast, *Zymomonas*
Homolactic	Hexose → 2 lactate$^-$ + 2 H^+	*Streptococcus*, some *Lactobacillus*
Heterolactic	Hexose → lactate$^-$ + ethanol + CO_2 + H^+	*Leuconostoc*, some *Lactobacillus*
Propionic acid	3 Lactate$^-$ → 2 propionate$^-$ + acetate$^-$ + CO_2 + H_2O	*Propionibacterium*, *Clostridium propionicum*
Mixed acid[b,c]	Hexose → ethanol + 2,3-butanediol + succinate^{2-} + lactate$^-$ + acetate$^-$ + formate$^-$ + H_2 + CO_2	Enteric bacteria including *Escherichia, Salmonella*, *Shigella*, *Klebsiella*, *Enterobacter*
Butyric acid[c]	Hexose → butyrate$^-$ + 2 H_2 + 2 CO_2 + H^+	*Clostridium butyricum*
Butanol[c]	2 Hexose → butanol + acetone + 5 CO_2 + 4 H_2	*Clostridium acetobutylicum*
Caproate/Butyrate	6 Ethanol + 3 acetate$^-$ → 3 butyrate$^-$ + caproate$^-$ + 2 H_2 + 4 H_2O + H^+	*Clostridium kluyveri*
Acetogenic	Fructose → 3 acetate$^-$ + 3 H^+	*Clostridium aceticum*

- Not all compounds are inherently fermentable, but sugars (e.g. glucose, other hexoses, most disaccharides, other relatively small sugars) —are fermentable

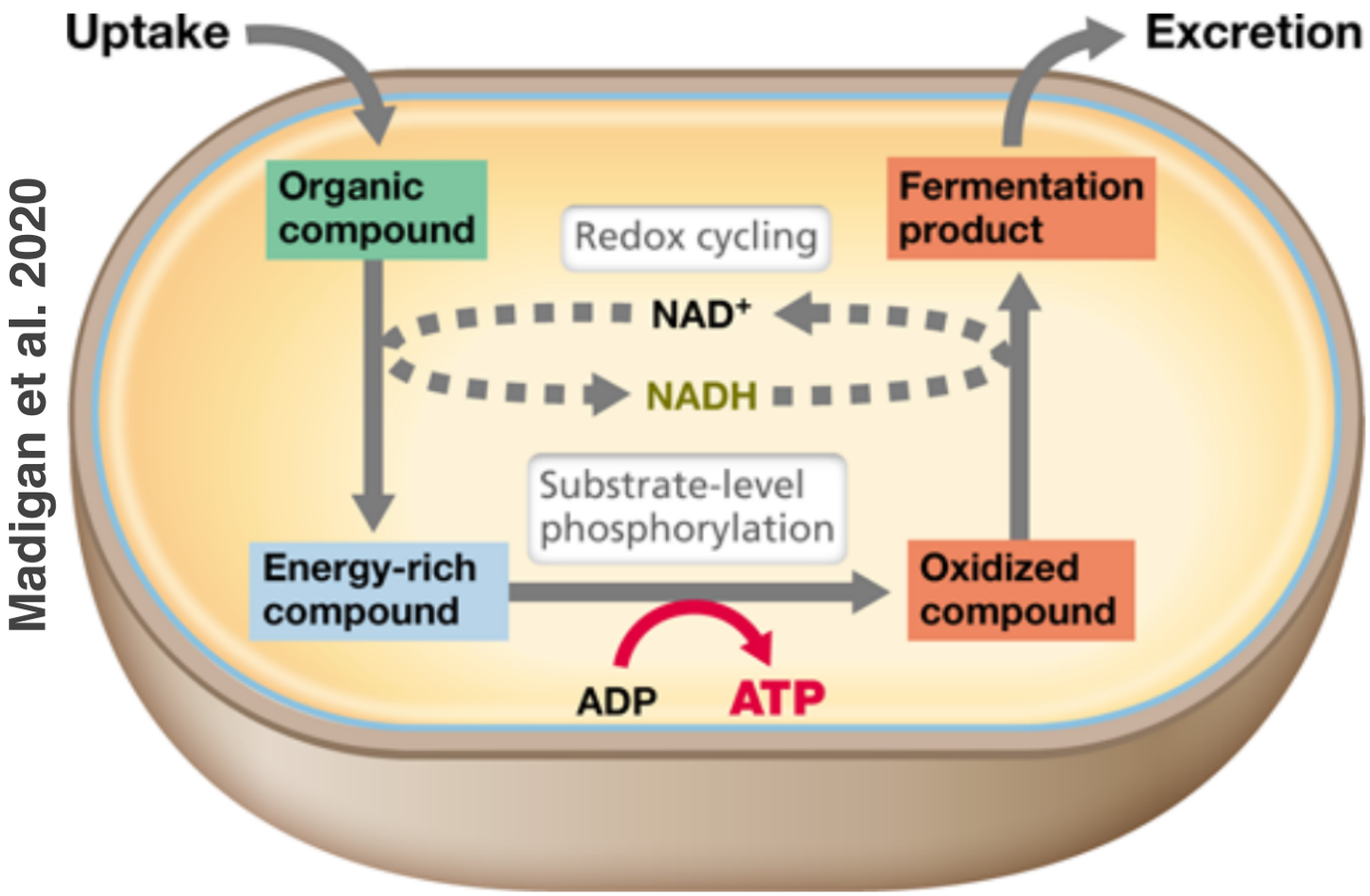
- Polysaccharides (e.g. cellulose, starch, chitin) are also fermentable by bacteria that produce enzymes that attack these large molecules and produce sugars from them if the latter are not glucose, they must first be converted to glucose before they enter glycolysis

- 2 net ATP molecules in glycolysis

- More ATP synthesis by substrate-level phosphorylation if fatty acid because the fatty acid is formed from its coenzyme-A precursor (energy-rich molecules)

Fermentation

Figure 3.14 The essentials of fermentation.



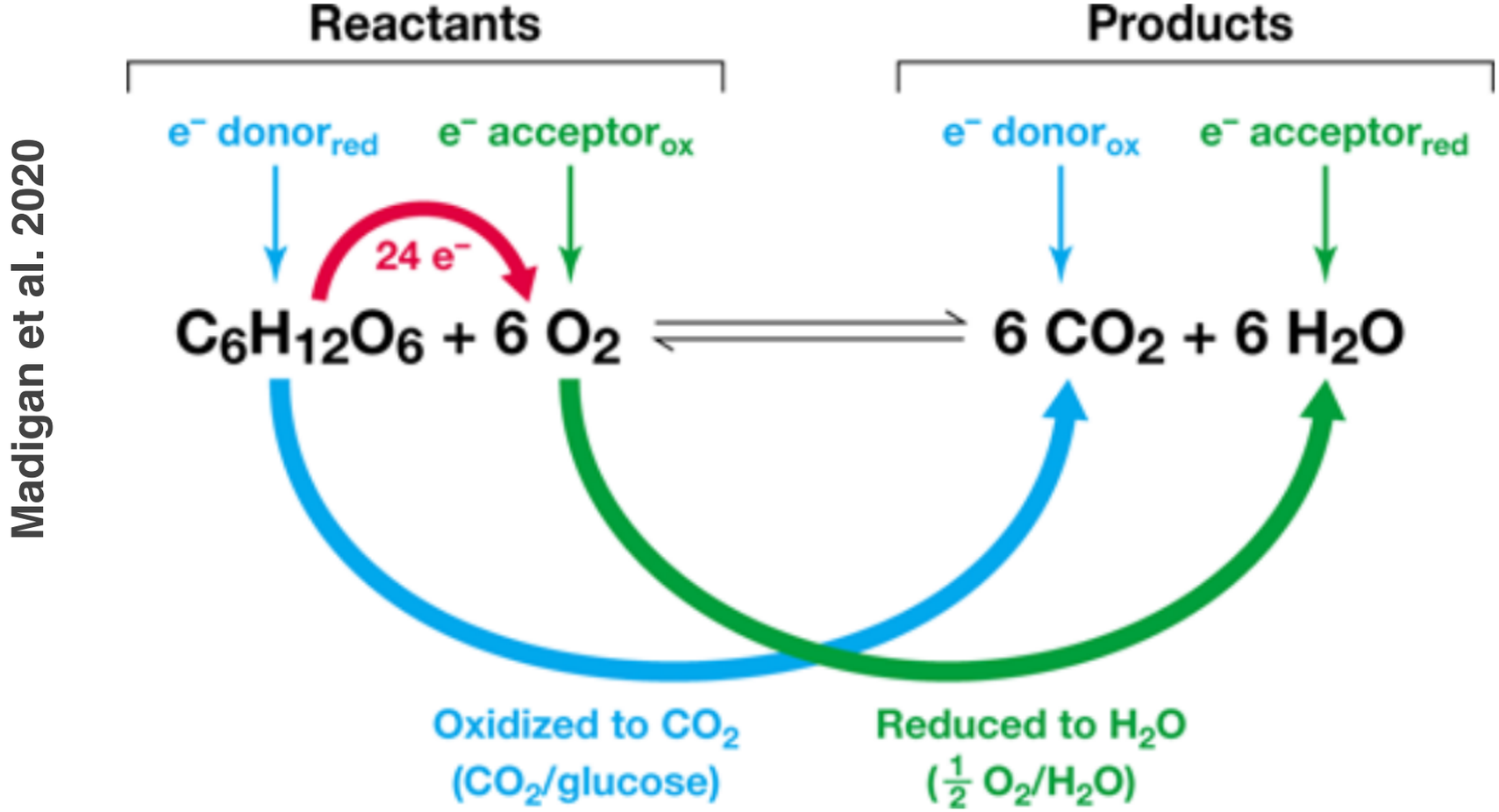
Madigan et al. 2020

- Both organic compounds accept and donate e-
- No need to external e- acceptor to achieve balance

- An organic compound is oxidized
- e- are recycled back to one of the oxidized organic products <u>because an external e- acceptor is lacking</u>
- Product is exceed from the cell and ATP is produced by substrate-level phosphorylation

Fundamentals in Metabolisms

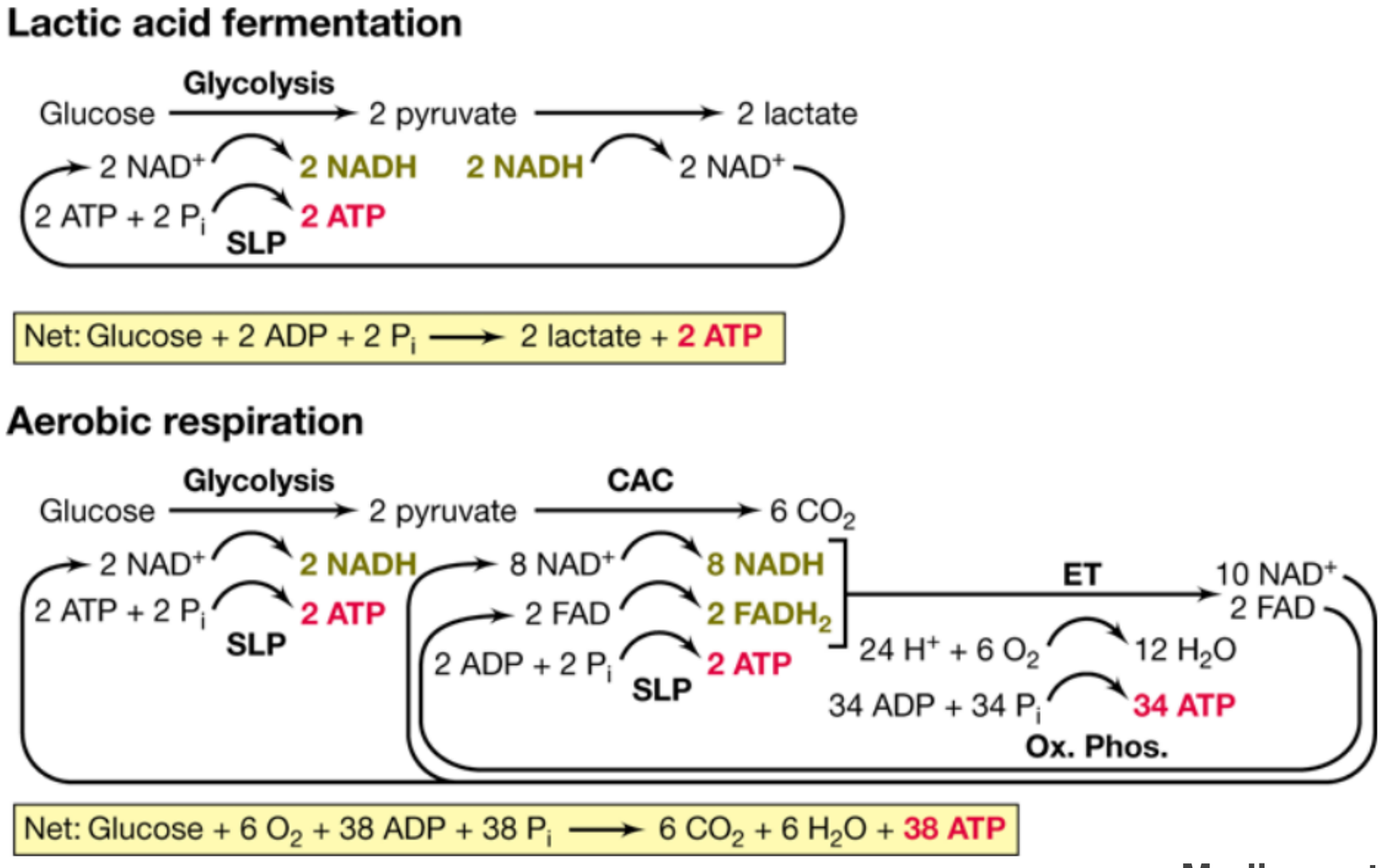
- Transfer e- and conserve energy
- Reactions are not performed in single-step —> consecutive reactions in different part of the cells
- Need of soluble e- carriers: NAD+/NADH, FAD+/FADH2



Substrate-Level-Phosphorylation

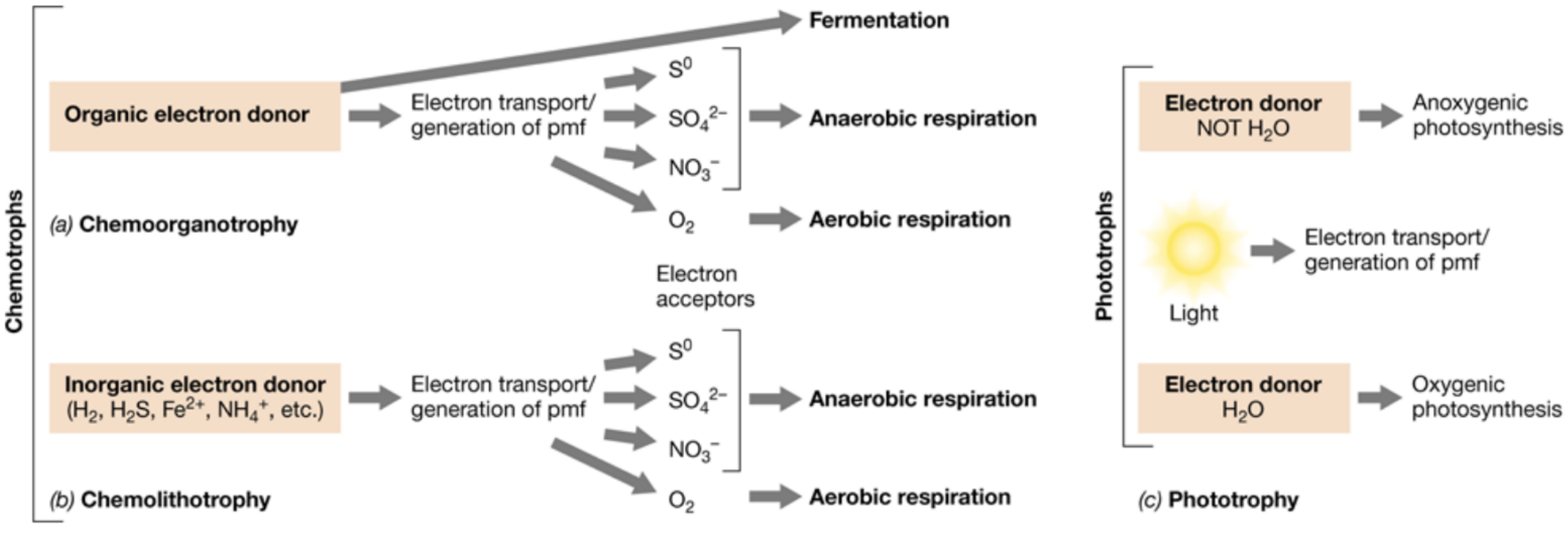
- Glycolysis can generate ATP in the absence of oxygen: anaerobic metabolism
- Glycolysis and citric acid cycle (CAC) result from substrate-level phosphorylation (SLP)
- SLP is distinct from oxidative phosphorylation that occurs in ETC
- Substrate-level phosphorylation refers to the formation of ATP from ADP and a phosphorylated intermediate, rather than from ADP and inorganic phosphate, Pi, as is done in oxidative phosphorylation (ET)

Figure 3.21 Energetics in fermentation and aerobic respiration.



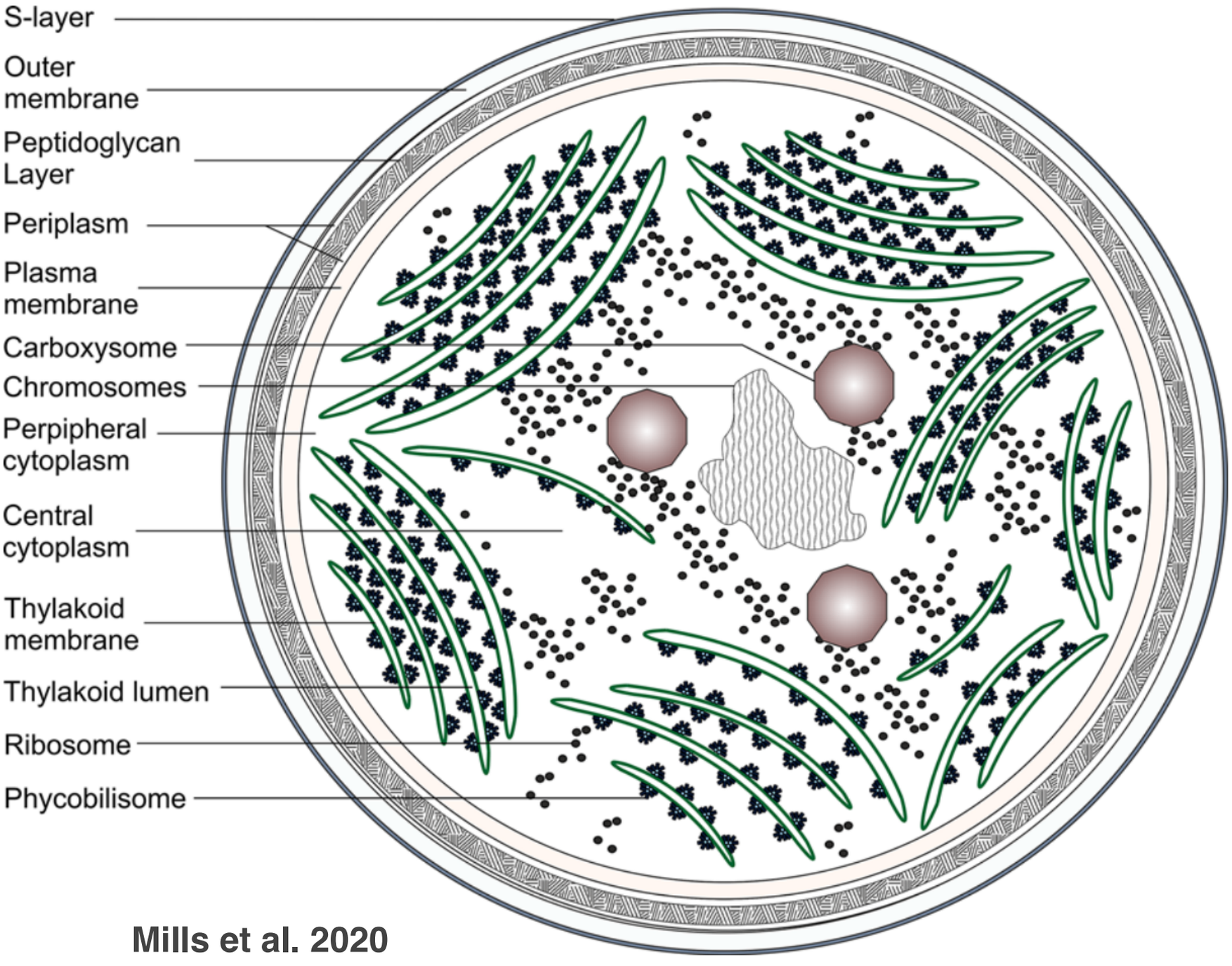
Madigan et al. 2020

Figure 3.22 Metabolic diversity and its relationship to oxygen.



Madigan et al. 2020

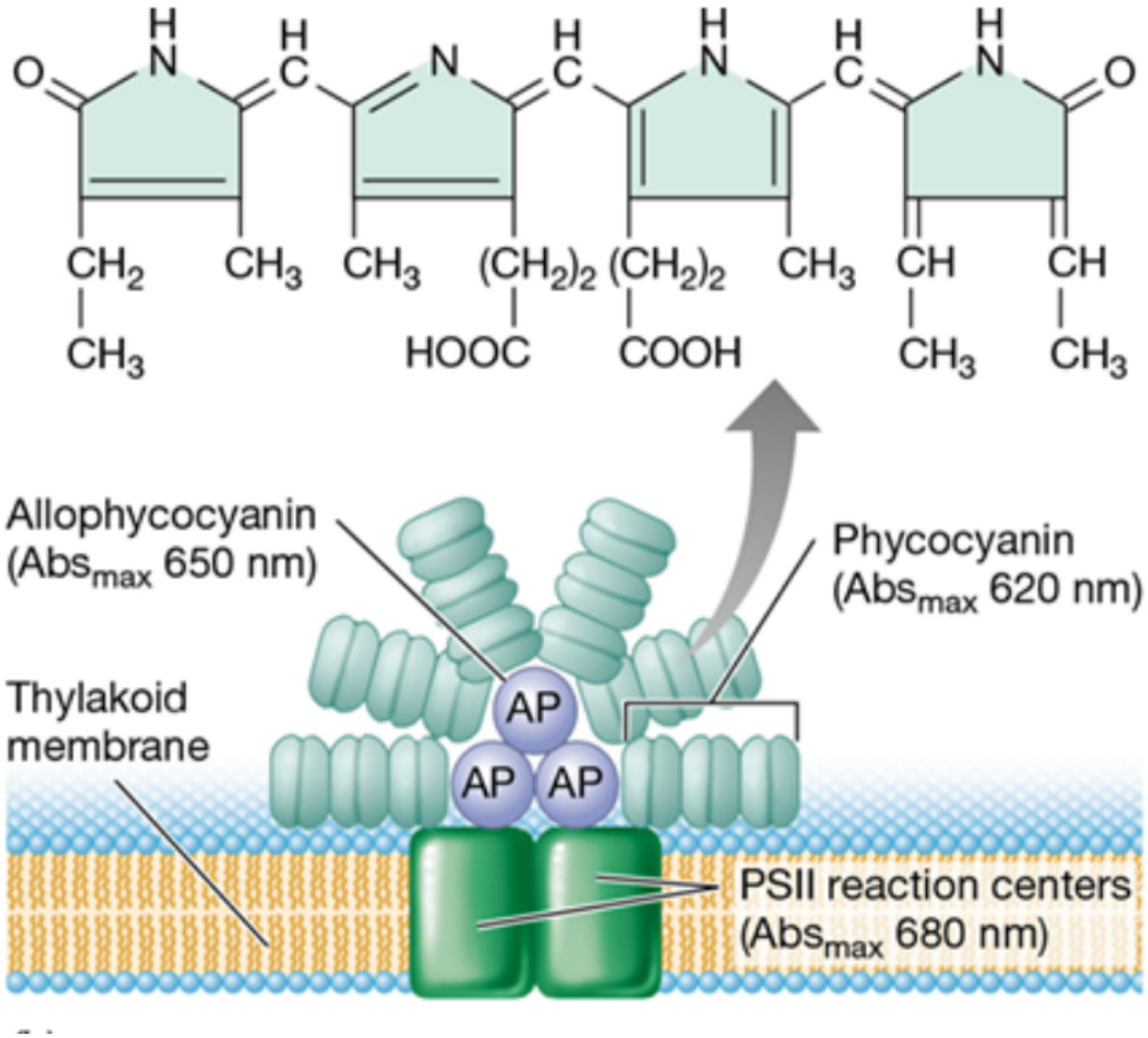
Photo Synthesis: Calvin–Benson–Bassham

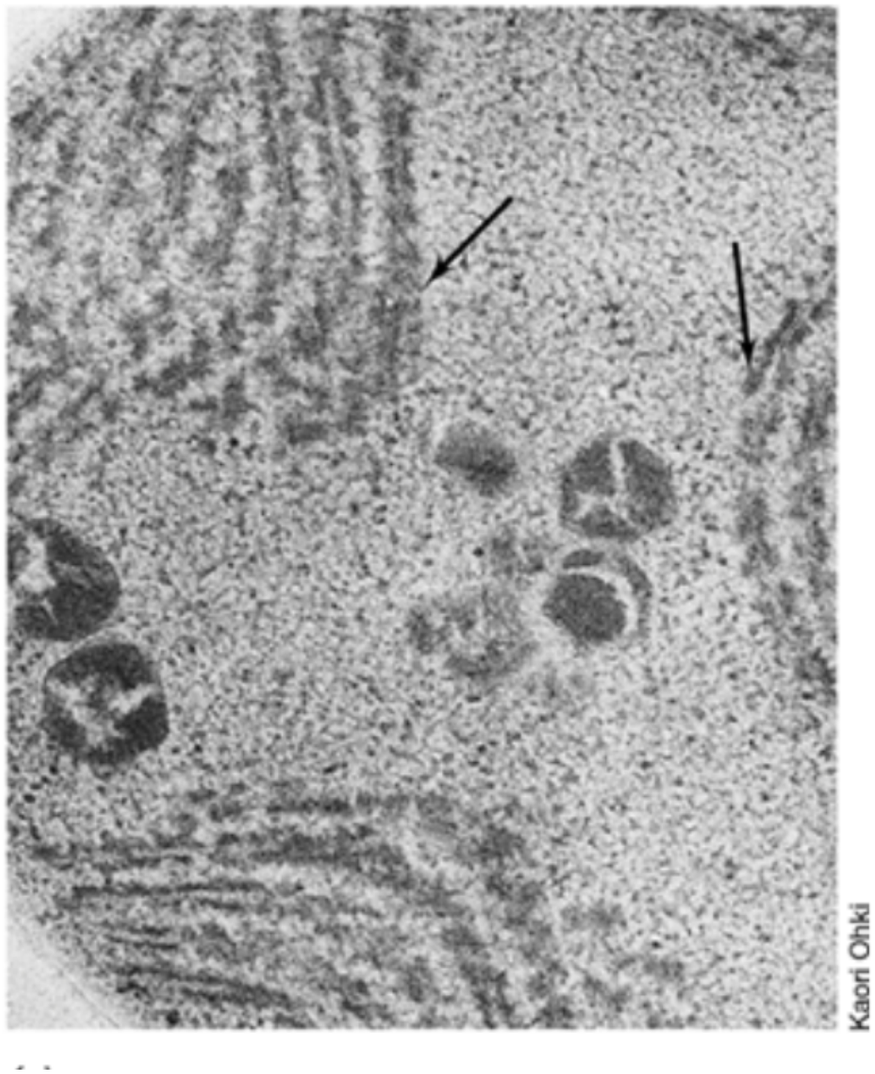


Mills et al. 2020

- Carboxysomes are made of polyhedral protein shells about 80 - 140 nm in diameter
- Concentrate carbon dioxide to overcome the inefficiency of RuBisCo (ribulose bisphosphate carboxylase/oxygenase)
- RuBisCO predominant enzyme in carbon fixation and the rate limiting enzyme in the Calvin-Benson-Bassham cycle

Oxygenic photosynthesis

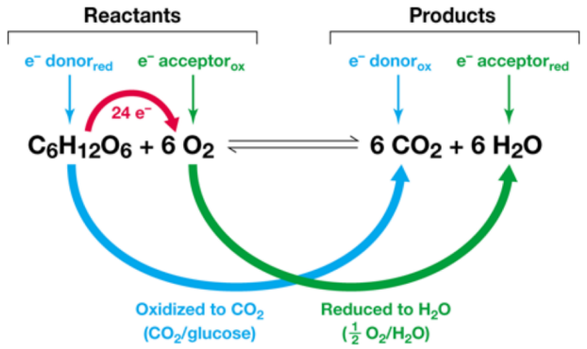




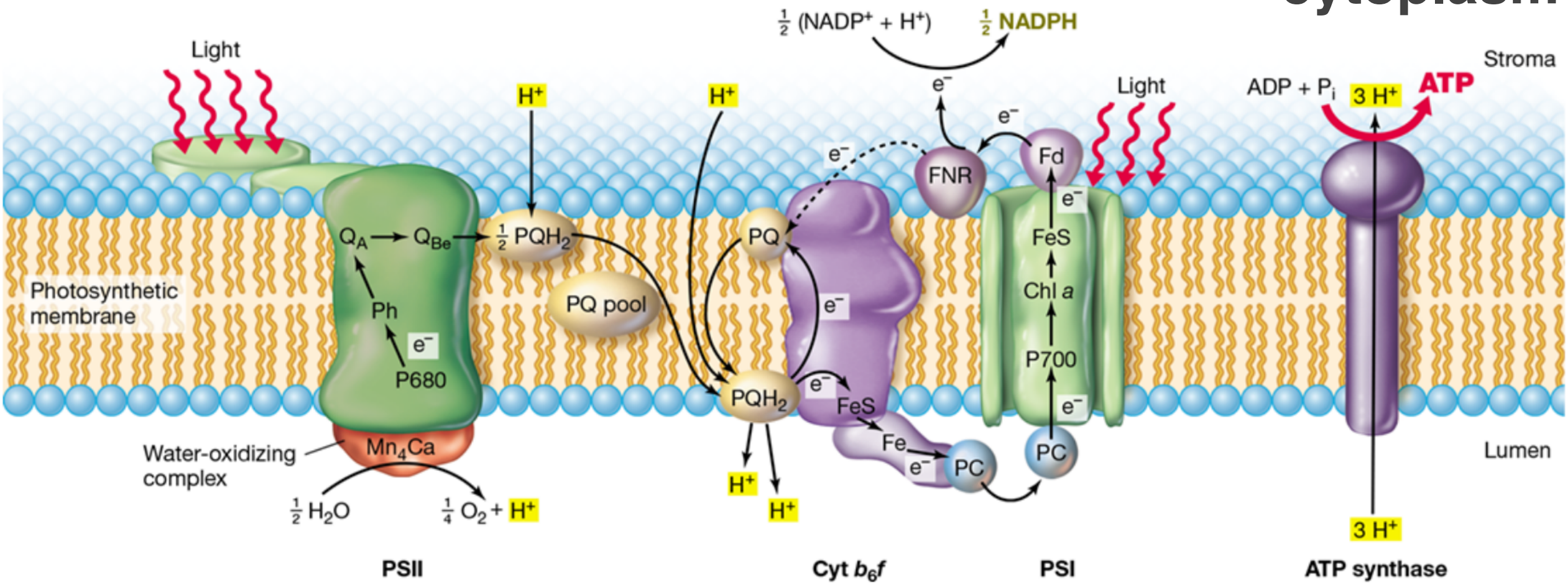
Madigan et al. 2020

- Physical location within the cell (Cyanobacteria)
- Bilayer w. proteins and complex that capture light, phycobilisome

Photo



cytoplasm

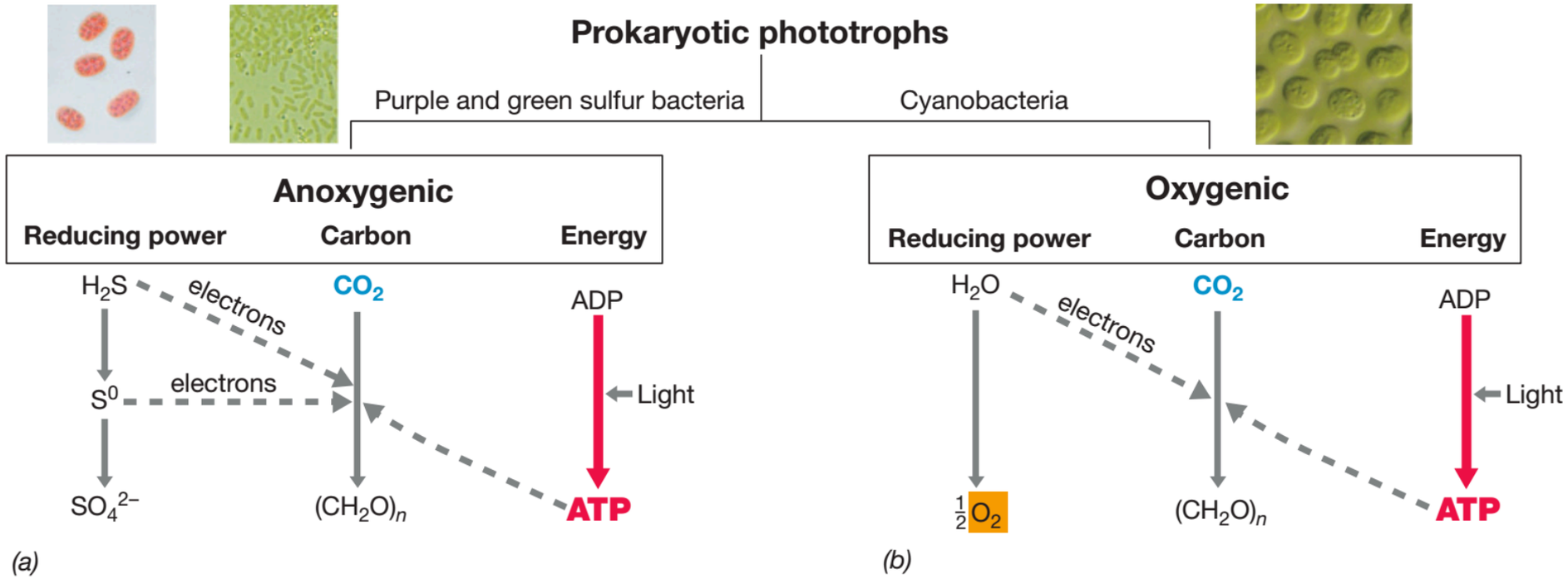


- Splitting of H2O

Madigan et al. 2020

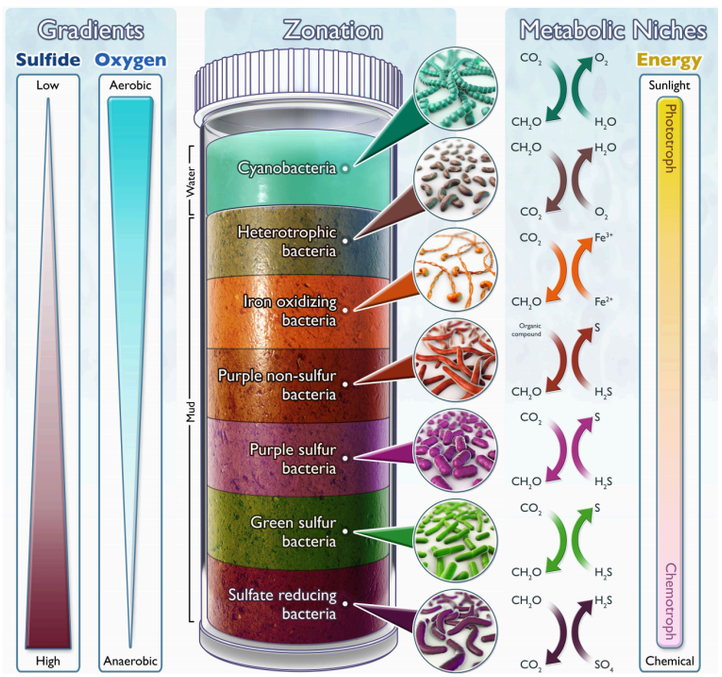
- Generation H+ motive force
- Generation of NADPH —> **C fixation (from CO2)** via Calvin–Benson–Bassham cycle
- ATP production

Light driven processes



Madigan et al. 2018

Winogradsky columns



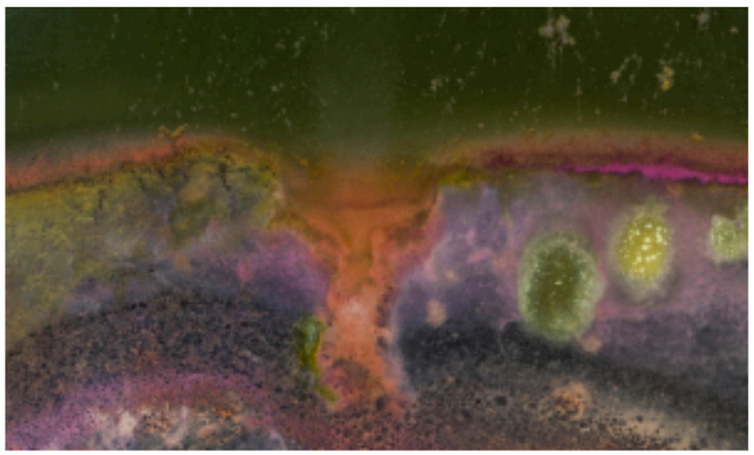


Figure 3. The upper sediment interface on day 15. Aerobic cyanobacteria and algae (upper aqueous phase), yellow-orange microaerophilic iron-oxidizing bacteria, and anaerobic green and purple photosynthetic bacteria develop into layered communities.

From HHMI

Energy generating metabolic pathways

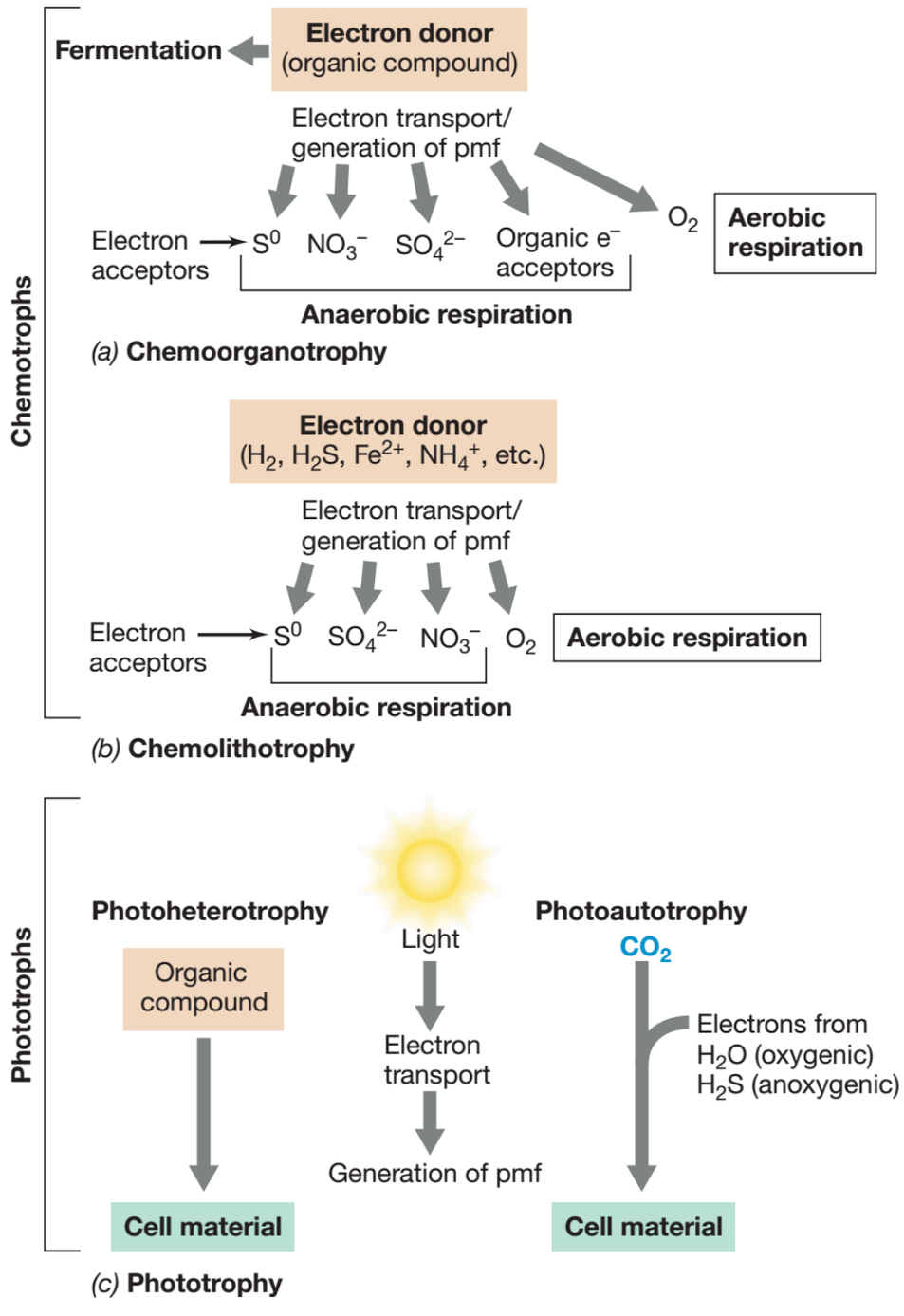
•Oxygenic Photosynthesis
ATP and NADPH are made in large amounts
Produces oxygen as a bi-product during splitting of water for reducing power

•Anoxygenic Photosynthesis
ATP made in large amounts
Reduction of NADP does not involve water; hence no oxygen produced

•Aerobic Respiration
ATP and NADH are made in abundance
Requires oxygen

•Anaerobic Respiration
Lower ATP yield than aerobic respiration; NAD easily reduced
Requires electron acceptor other than oxygen

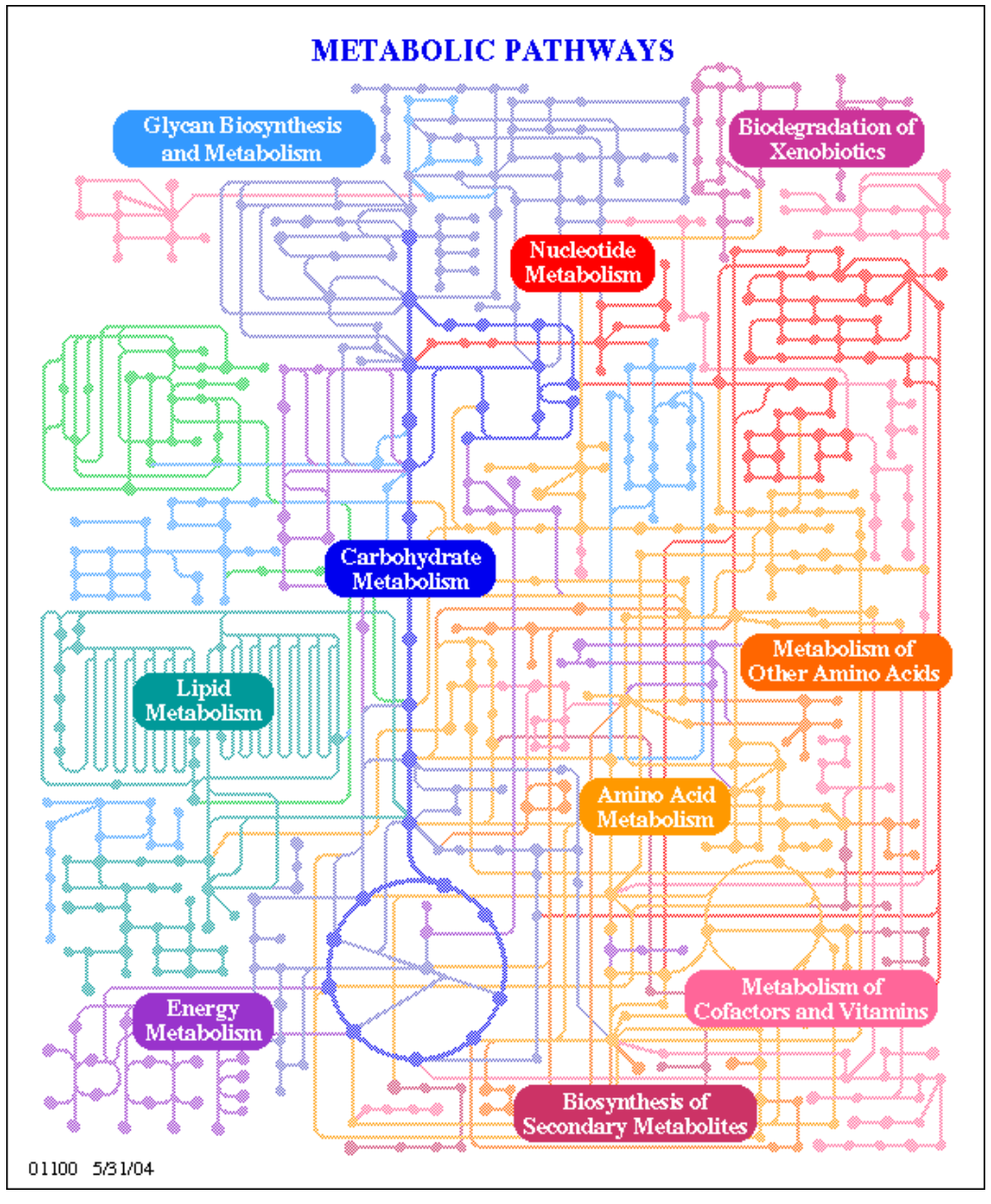
Fermentation
Little ATP, no net NAD reduction, MOST SIMPLE SYSTEM



Madigan et al. 2018

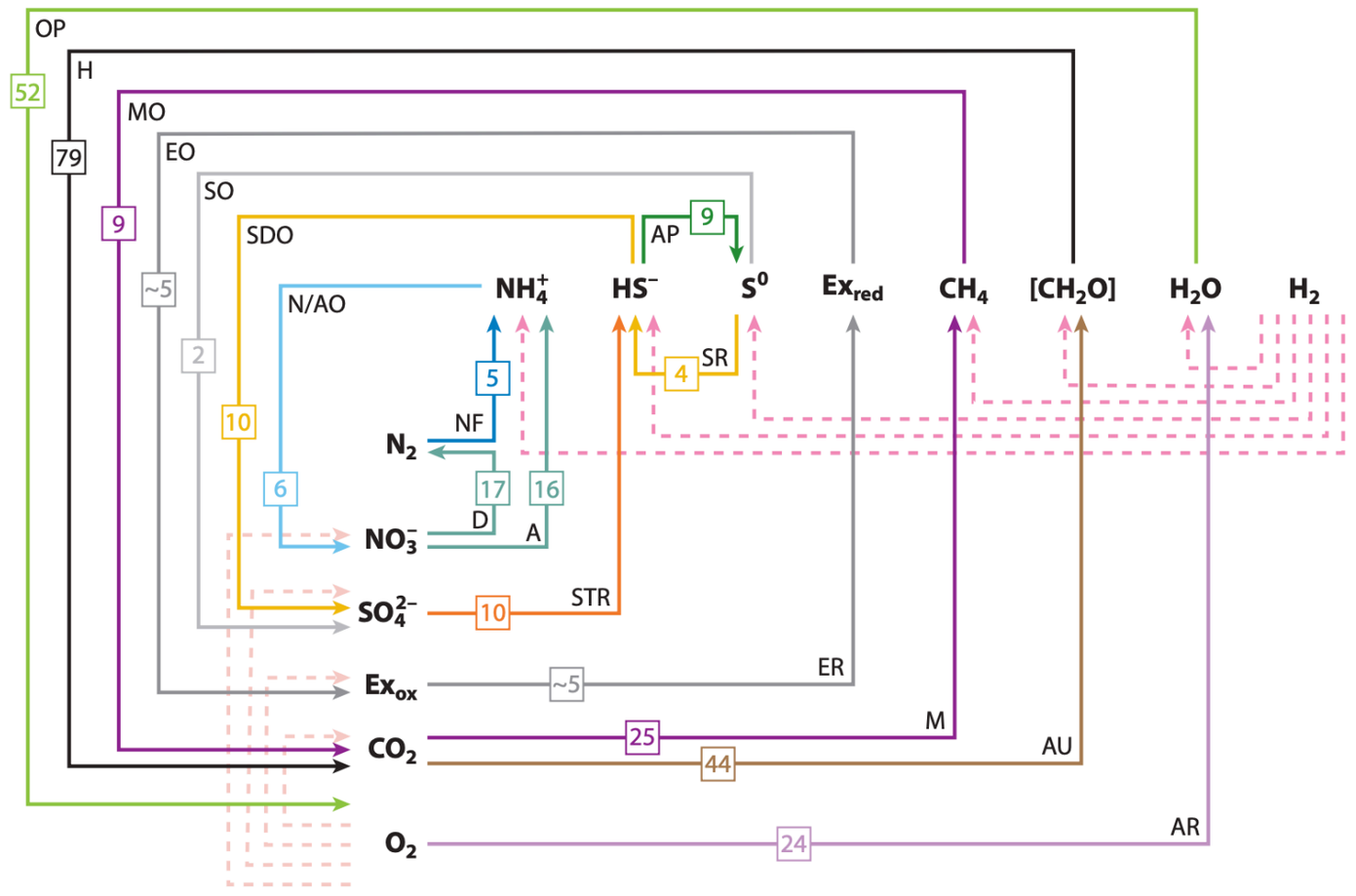
Integrative approach

Metabolic pathways evolved to utilize available substrates produced as end products of other types of microbial metabolism, either by modification of existing metabolic pathways or by using established ones in reverse



https://www.genome.jp/

Oxidative reaction



Reductive reaction

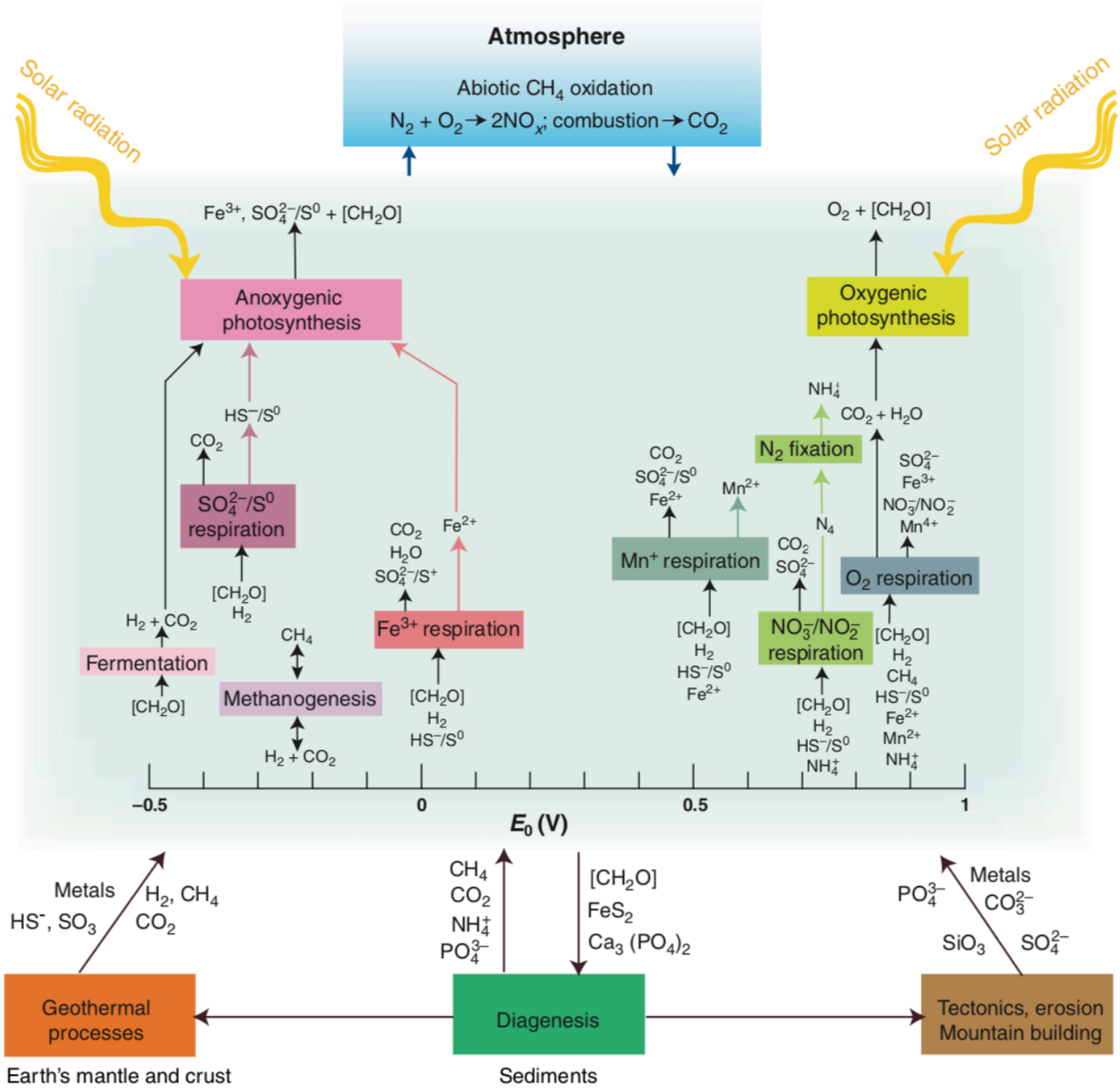
Jelen et al. 2016

A, ammonification; AP, anoxygenic photosynthesis; AR, aerobic respiration; AU, autotrophy; D, denitrification; Exox, other elements oxidation; Exred, other elements reduction; H, heterotrophy; M, methanogenesis; MO, methane oxidation/methanotrophy; N/AO, nitrification/ammonia oxidation; NF, nitrogen fixation; OP, oxygenic photosynthesis; SDO, sulfide oxidation; SO, sulfur oxidation; SR, sulfur reduction; STR, sulfate reduction

Energy conservation

- The achievable energy gain (Gibbs free energy, ΔG) of ETC depends on the redox potential difference (ΔE) of all reactions between electron donor and acceptor
- Microbes able to respire in multiple ways will always choose available acceptors with the biggest potential difference to the donor (e.g., *E. coli* O2 > NO3-> fumarate)
- **Cellular metabolism coordinate the production, management and re-distribution of carbon building blocks and energy (ATP and NADPH) between various electron and carbon sinks**
- ATP and NAD(P)H are **essential energy carriers** for numerous biochemical reactions occurring
- With the exception of fermentation, in which substrate-level phosphorylation occurs all other mechanisms of microbial energy conservation are linked to the proton motive force (or gradient of sodium ions, Na+, instead of protons)
- Whether electrons come from the oxidation of organic or inorganic chemicals or are mediated by light-driven processes, in both respiration and photosynthesis, **energy conservation is the result of electron transport reactions and the formation of a PMF —> ATP**
- **The oxidation of NADH and FADH, to NAD+ and FAD**, respectively, is linked to **energy conservation** via ETC

Biosphere model of energy fluxes and elemental cycles



Falkowski, Fenchel and Delong, 2008

Microbial microscale actions structure planet-scale functioning