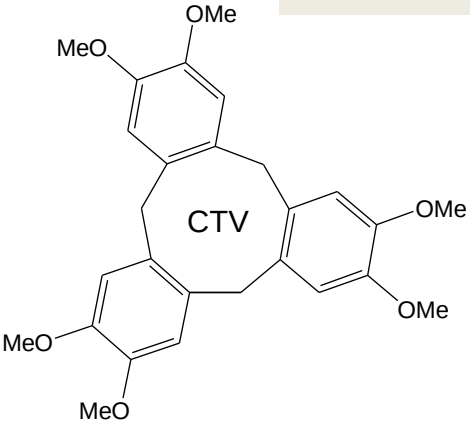
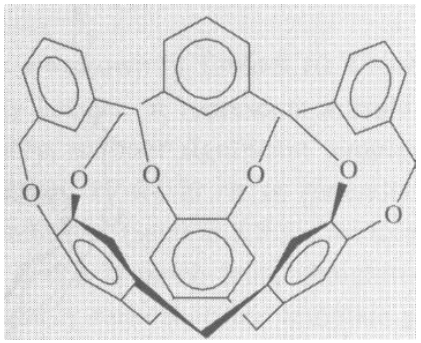
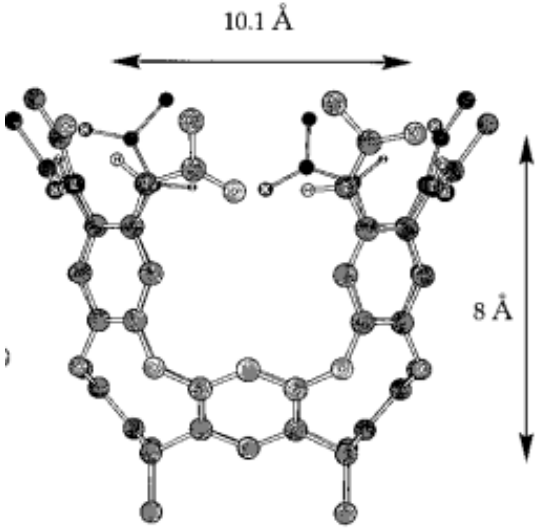
Cavitandi

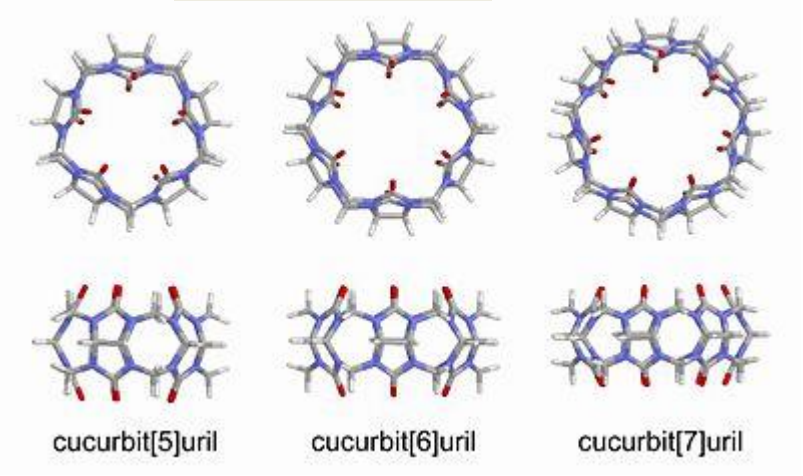
ciclotriveratrilene



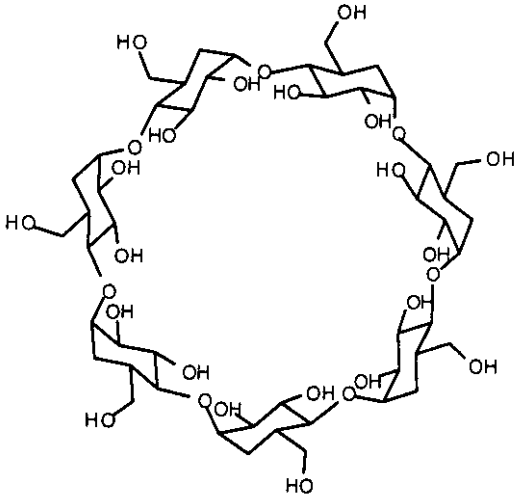


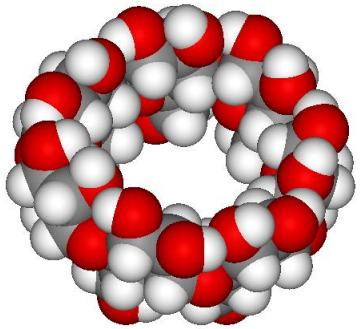


Cucurbiturili



Ciclodestrine





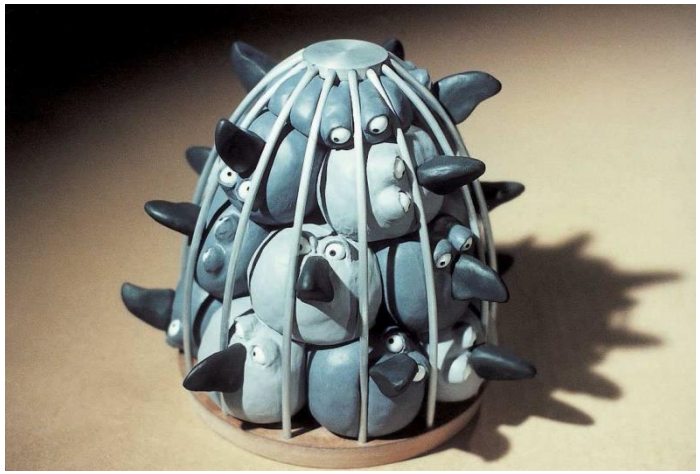
Capsule Molecolari

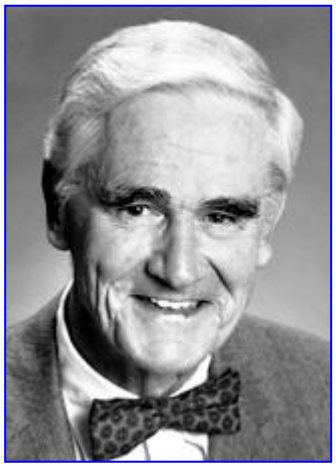
Unione di due cavitandi

Connessione covalente
Legame idrogeno
Legame di coordinazione

Pre-organizzazione

Protezione dal solvente esterno
Rallentamento delle cinetiche di scambio
Stabilizzazione di specie reattive
Reazioni catalitiche
Drug delivery





Carcerando:

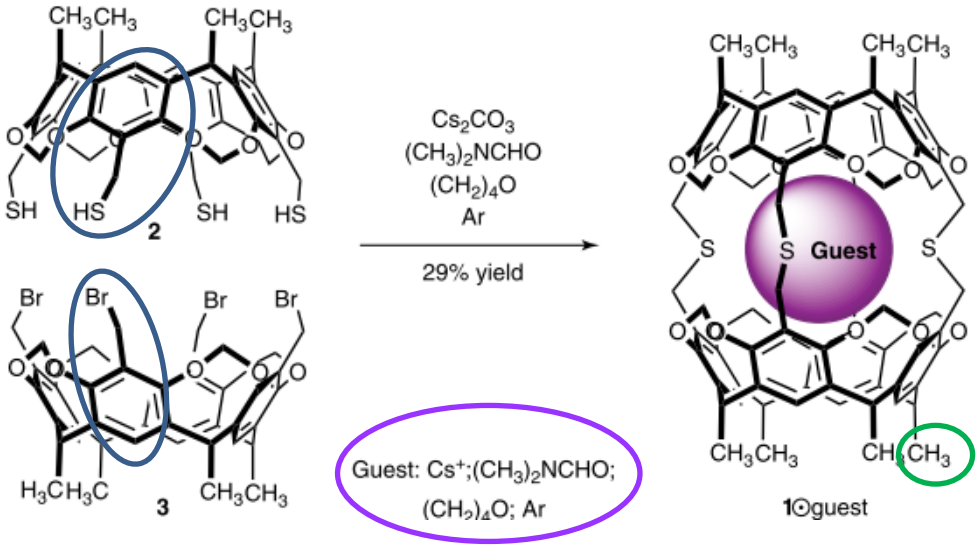
Contenitore molecolare chiuso (capsula) che definisce cavità sferica, i guest sono intrappolati (all'atto della sintesi) entrata e uscita solo per rottura di legame covalente, i.e. velocità di scambio virtualmente nulla

Carcerandi

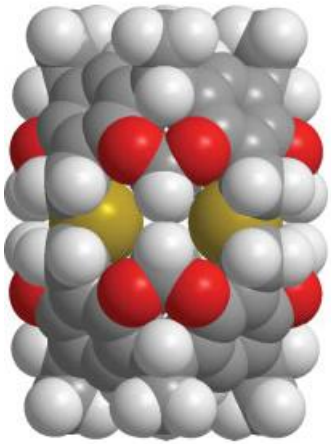
Benzil-tiolo

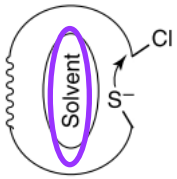
benzil cloruro
(o bromuro)

Alta dliuz



Insolubilità = caratterizz via IR, FAB-MS, analisi elementare, test chimici
FAB-MS dei carciplessi

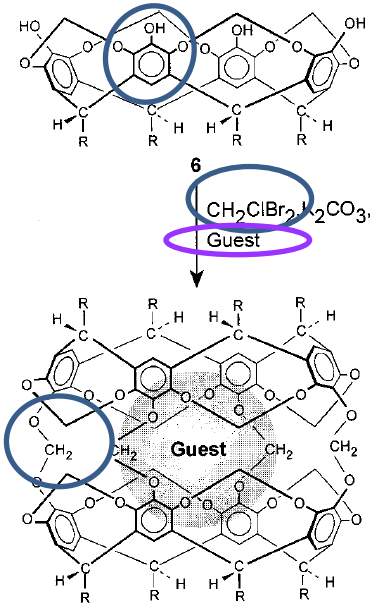


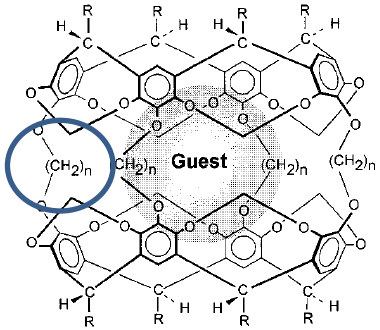


Carcerandi

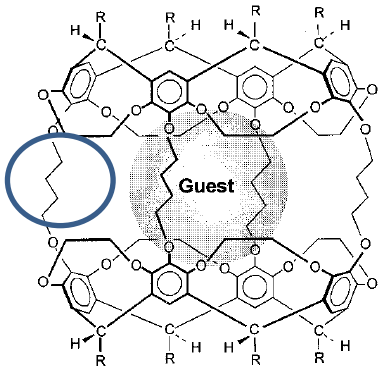
fenolo

bromo-clorometano

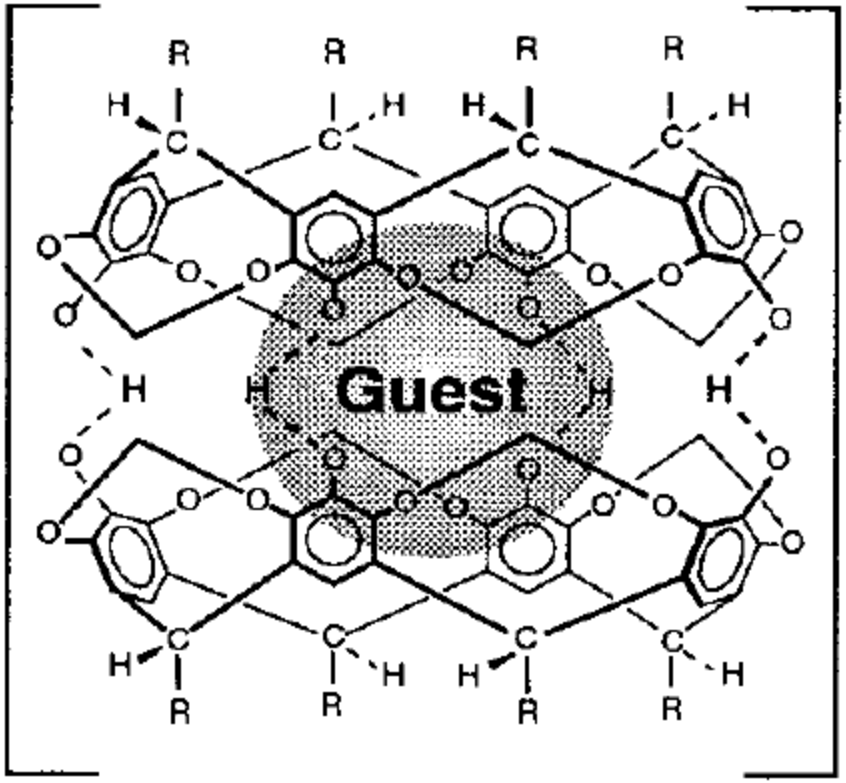




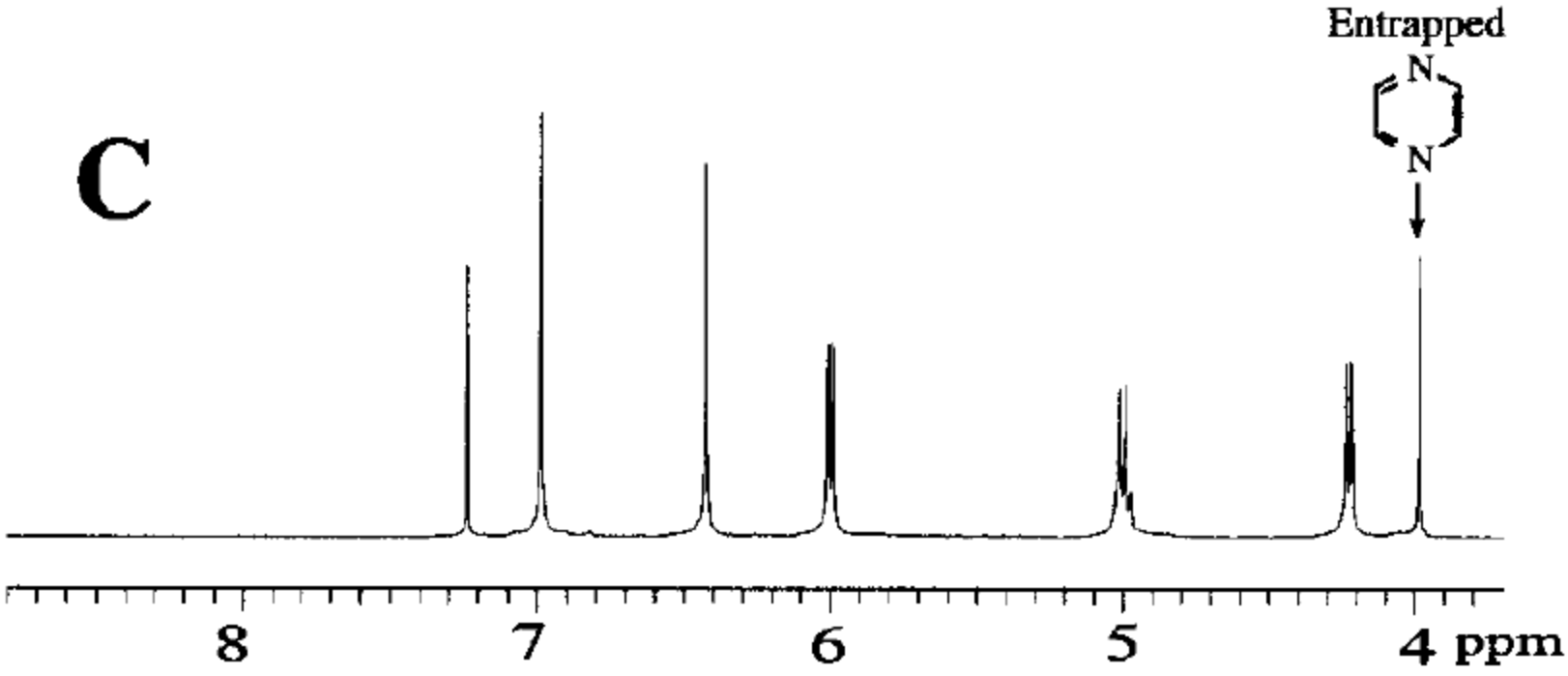
11•Guest n = 2 R = CH_2CH_2Ph
12•Guest n = 3 or $C_{11}H_{23}$

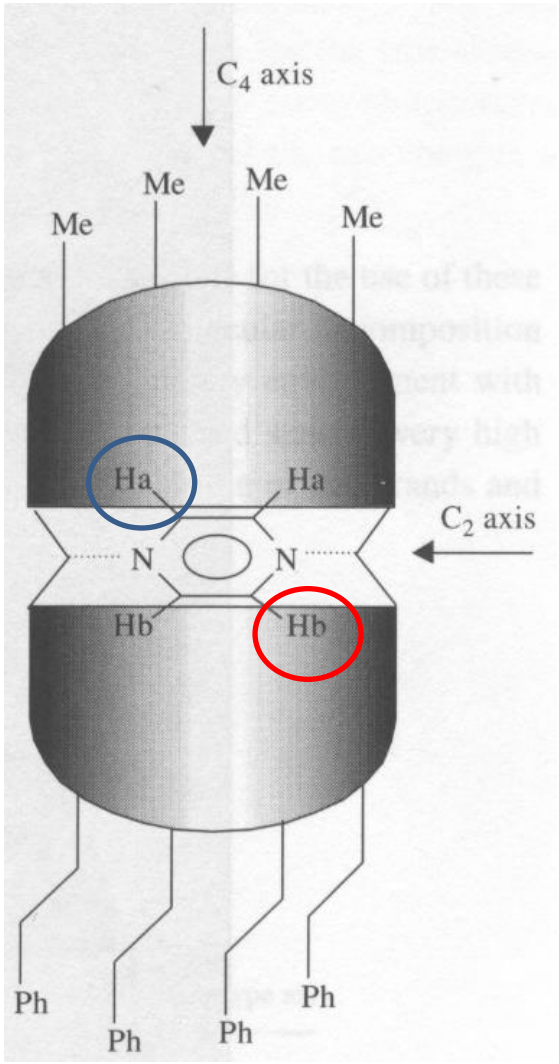


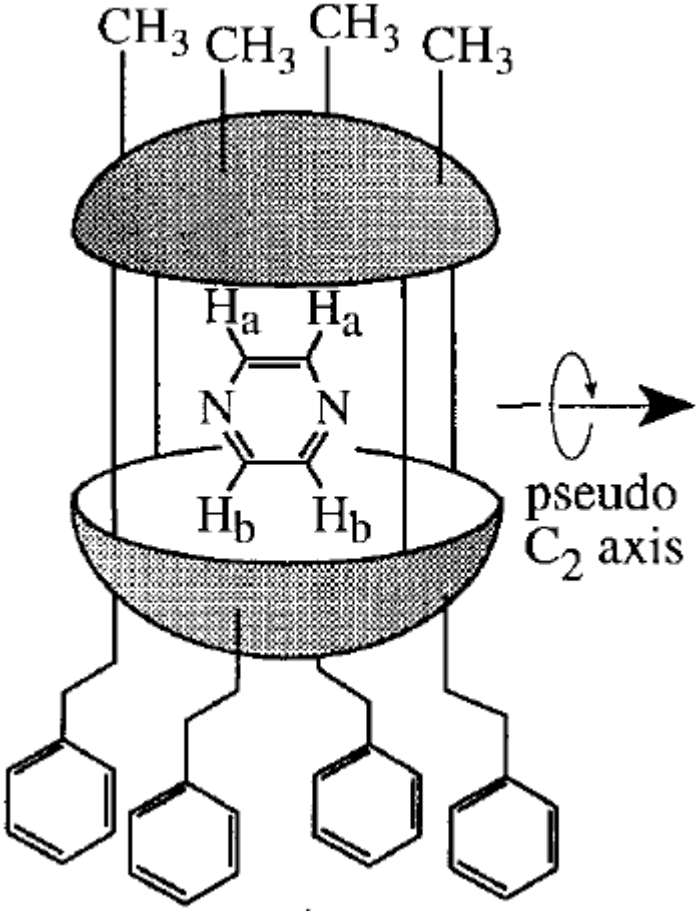
13•Guest R = $(CH_2)_4CH_3$



Complex **3**•Guest

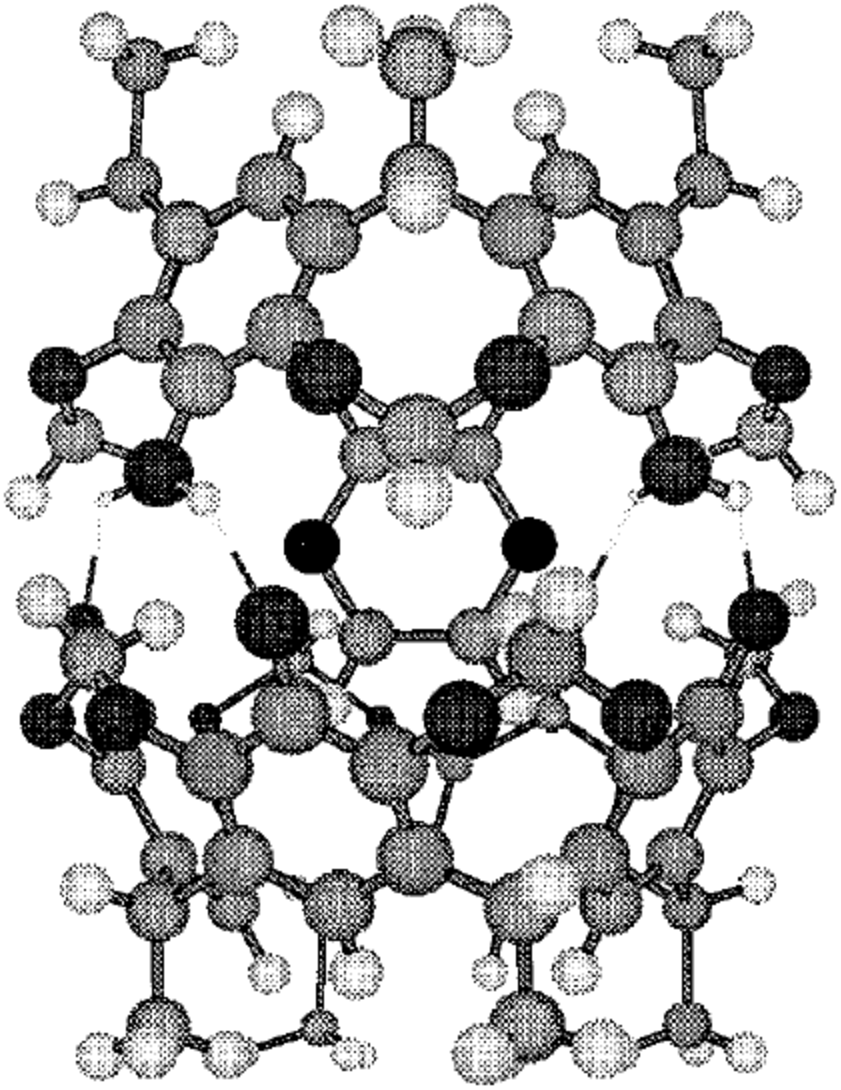

C
Entrapped
N
N
8
7
6
5
4 ppm


C4 axis
Me
Me
Me
Me
Ha
Ha
C2 axis
N
N
Hb
Hb
Ph
Ph
Ph
Ph


CH3 CH3 CH3 CH3
Ha Ha
N N
Hb Hb
pseudo
C2 axis

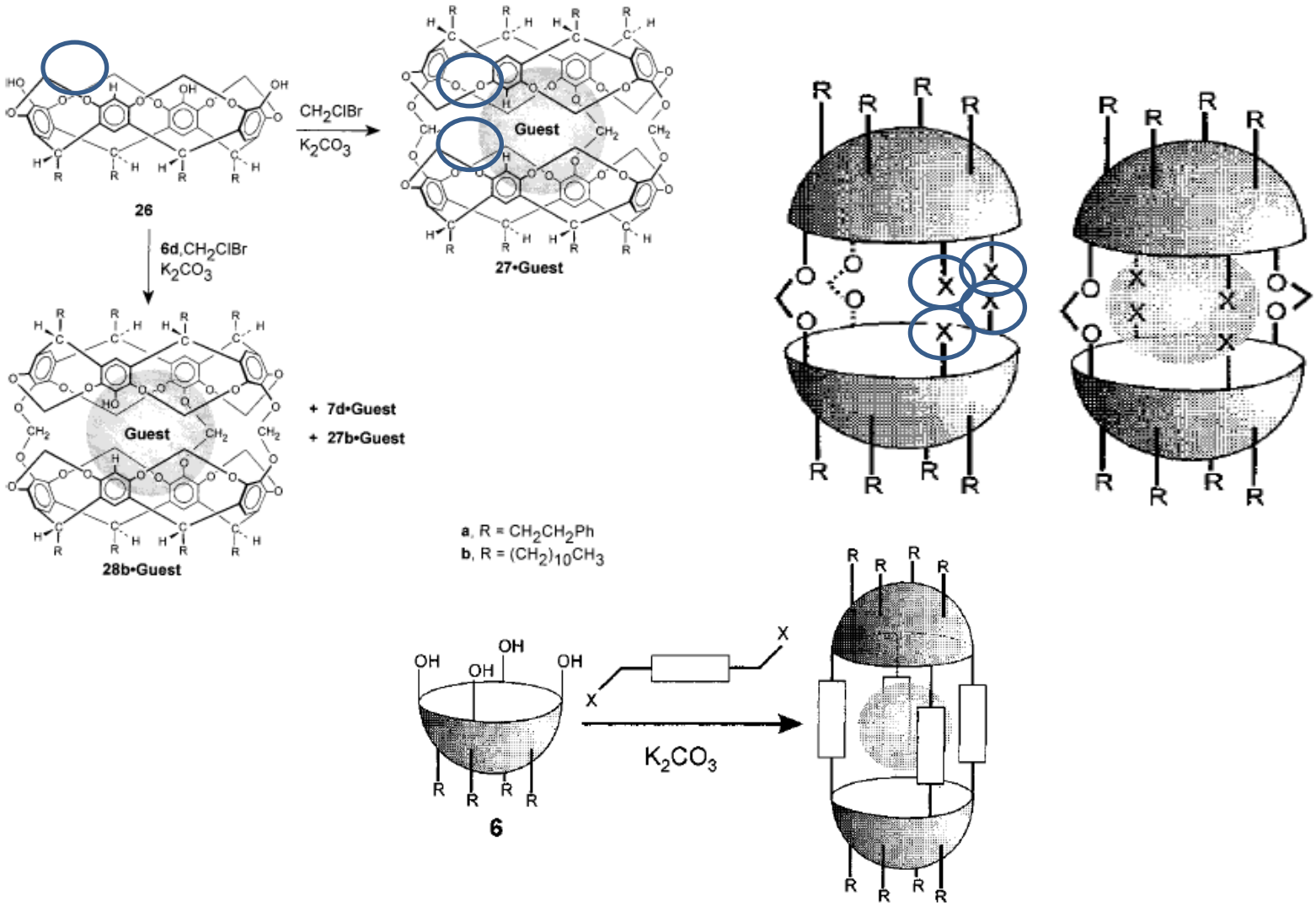
tively. The signals for pyrazine in asymmetric complex **3c**•pyrazine consisted of two *meta*-split doublets at 4.31 ($J = 1$ Hz) and 4.35 ppm ($J = 1$ Hz).[16] This confirms that pyrazine is oriented in complex **3b**•pyrazine with its nitrogens at the equator and its hydrogens extending into the bowls (structure **A**, Figure 3). The activation energy for rotation of pyrazine about the pseudo-C_2 axes in asymmetric complex **3c**•pyrazine was measured by variable-temperature ^{1}H NMR spectroscopy to be 18 kcal/ mol,[17] which agrees well with the 19 kcal/mol activation

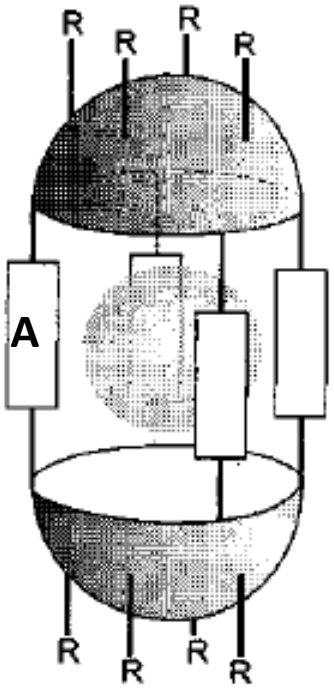
(17) The activation barrier for rotation of pyrazine about the pseudo-C_2 axes of asymmetric complex **3c**•pyrazine was calculated to be 18.3 kcal/ mol based on a coalescence temperature (T_c) of 353 K and separation of the signals ($\Delta\delta_{Hz}$) of 14.3 Hz using the following equation: $\Delta G_c^{\ddagger} = RT_c$-$[22.96 + \ln(T_c/\Delta\delta_{Hz})]$ where $\Delta G_c^{\ddagger}$ is the activation barrier in kcal/mol; T_c is the temperature of coalescence, and $\Delta\delta_{Hz}$ is the separation of the signals in Hz. See: Abraham, R. J.; Fisher, J.; Loftus, P. *Introduction to NMR Spectroscopy*; Wiley: New York, 1990; pp 194−197.



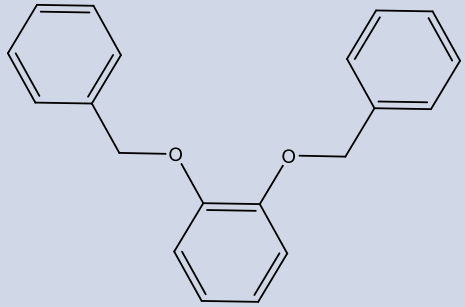
Emicarcerando:

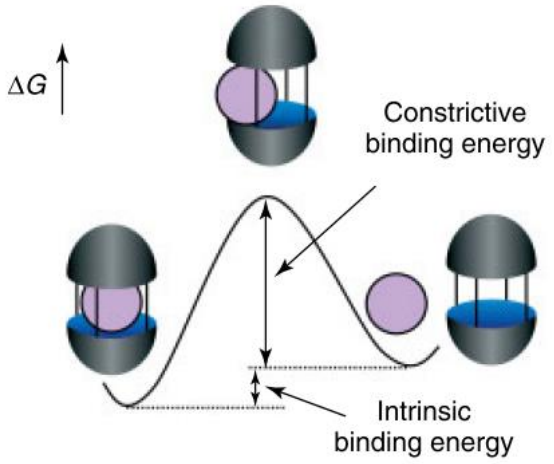
Contenitore molecolare chiuso (capsula) che definisce cavità sferica, i guest sono intrappolati (all'atto della sintesi) - entrata e uscita senza rottura di legame covalente, i.e. velocità di scambio misurabile





A
$O–CH_2–O$
$S–CH_2–S$
$O–(CH_2)_4–O$
$O–(CH_2)_5–O$
$O–(CH_2)_6–O$
$O–(CH_2)_2O(CH_2)_2–O$



guests. In host−guest systems, there are two energetic quantities associated with guest binding—constrictive and intrinsic binding. The constrictive binding[11,12] is the activation energy required for a guest to enter the host, while intrinsic binding is the change in energy upon formation of the host−guest complex from free host and guest. Intrinsic binding energy determines the equilibrium constant for binding; the intrinsic binding energy plus the constrictive binding energy determines the kinetic barrier to decomplexation.[12]

Intrinisic binding, the free energy of complexation, depends on the magnitude of the noncovalent interactions between the guest and the host's inner surface.

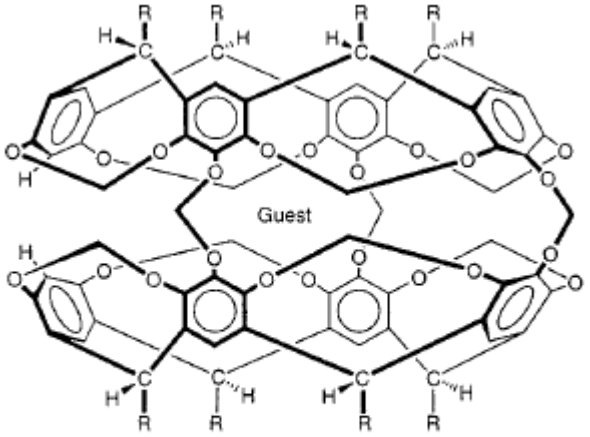
Constrictive binding, activation energy required for a guest to enter OR exit the inner cavity of a hemicarcerand through a size restricting portal in the host's skin.

Constrictive binding: aumenta con le dimensioni del guest, diminuisce con le dimensioni dei portali, e con l'aumento della flessibilità dei linker (T).

Gating: Gating in host-guest chemistry refers to conformational changes that occur in response to some stimulus, resulting in opening or closing of a physical barrier that controls access to the host.

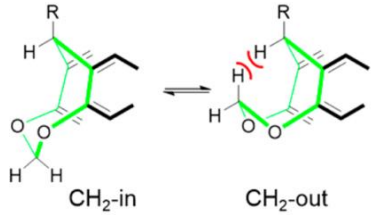
- **Thermal Gating**
- **Stimulated Gating**

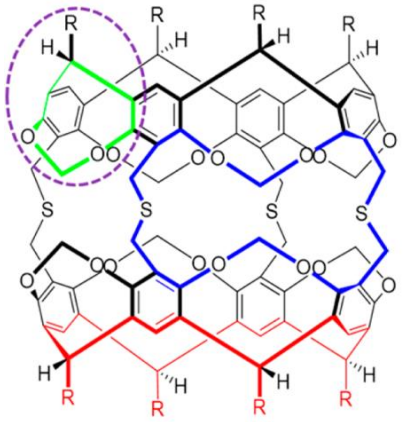
Gate mechanisms (molecular mechanics calculations) – **French door**
chair-to-boat transition of the methylene bridges, calculated barrier 22 kcal/mol.



12⊙ Guest

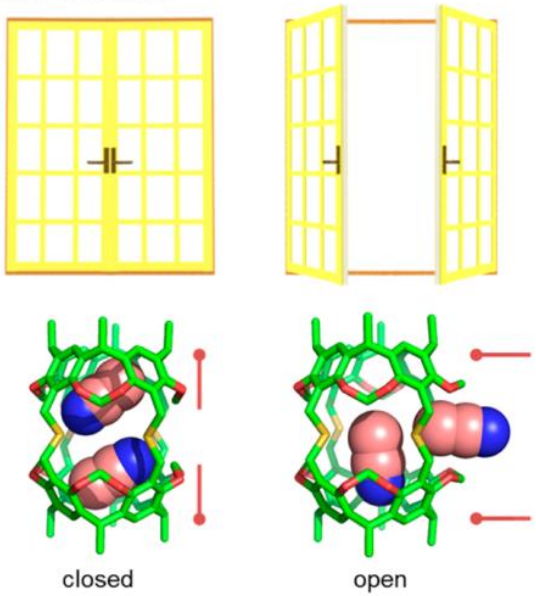
R = CH_2CH_2Ph



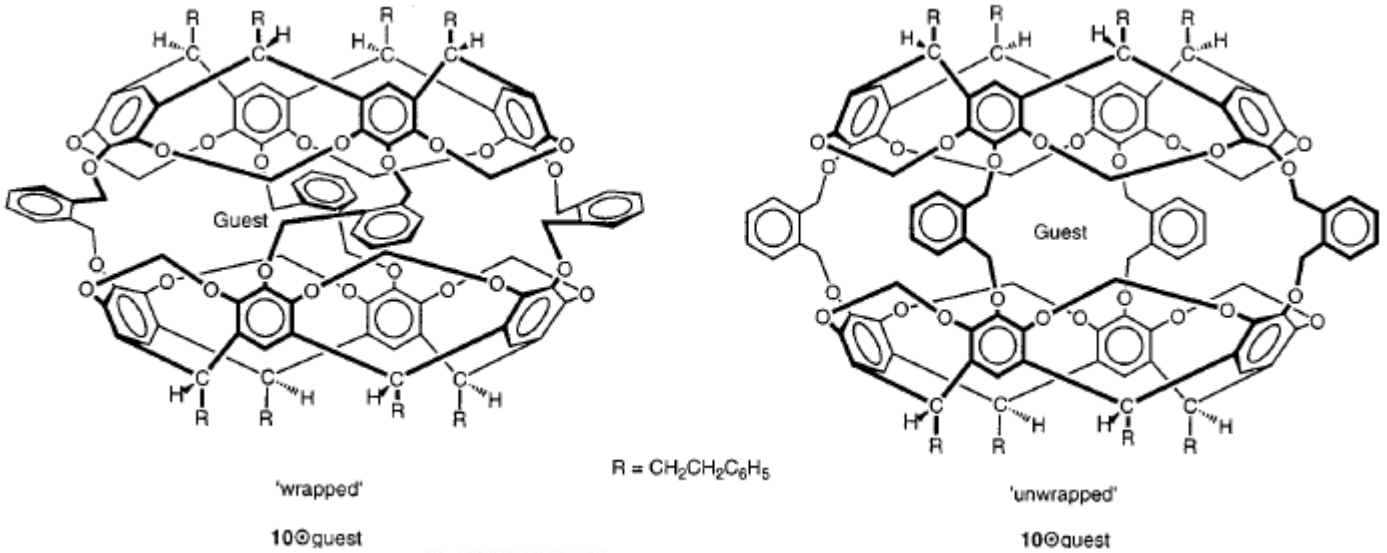


1a. R = $(CH_2)_2Ph$
1b. R = Me

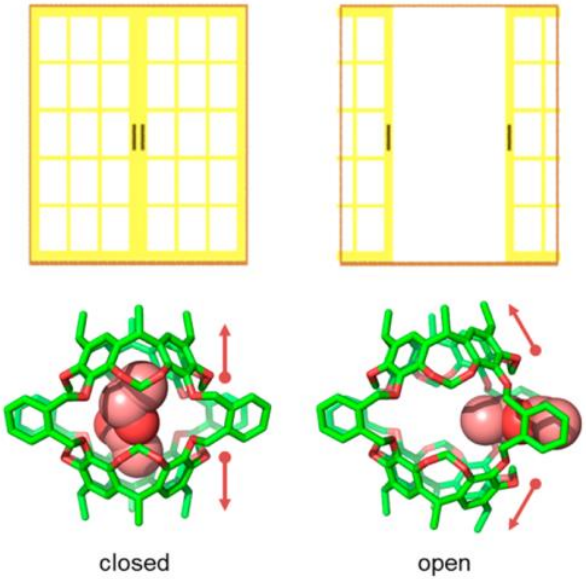
a. French door



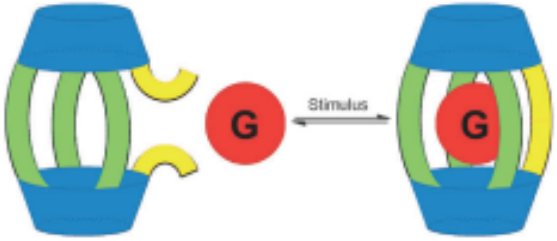
Gate mechanisms (molecular mechanics calculations) – **Sliding door**
twisting and untwisting of the two host cavitands – measured barrier (VT NMR) 12.6 kcal/mol



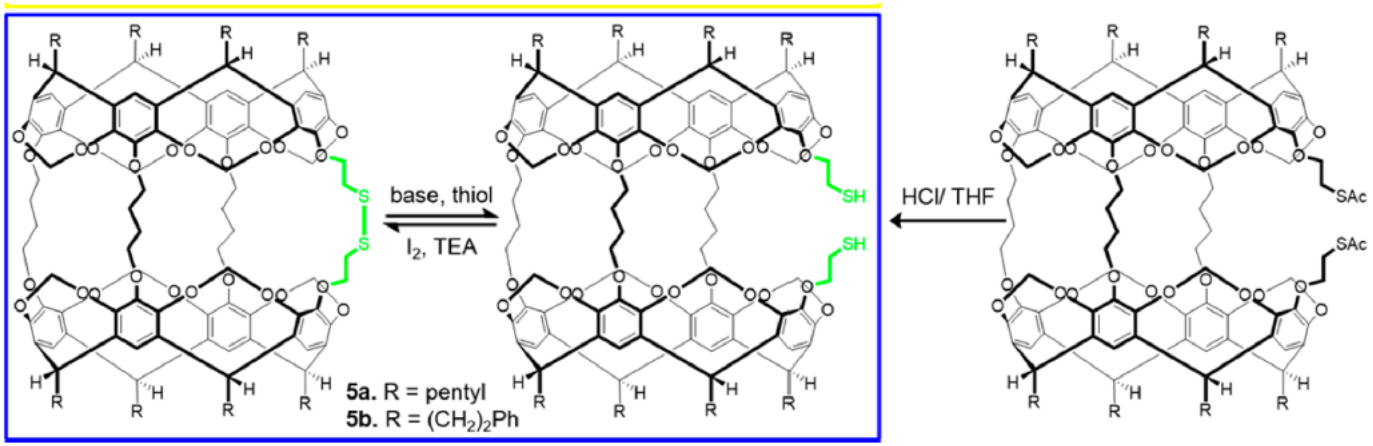
b. Sliding door



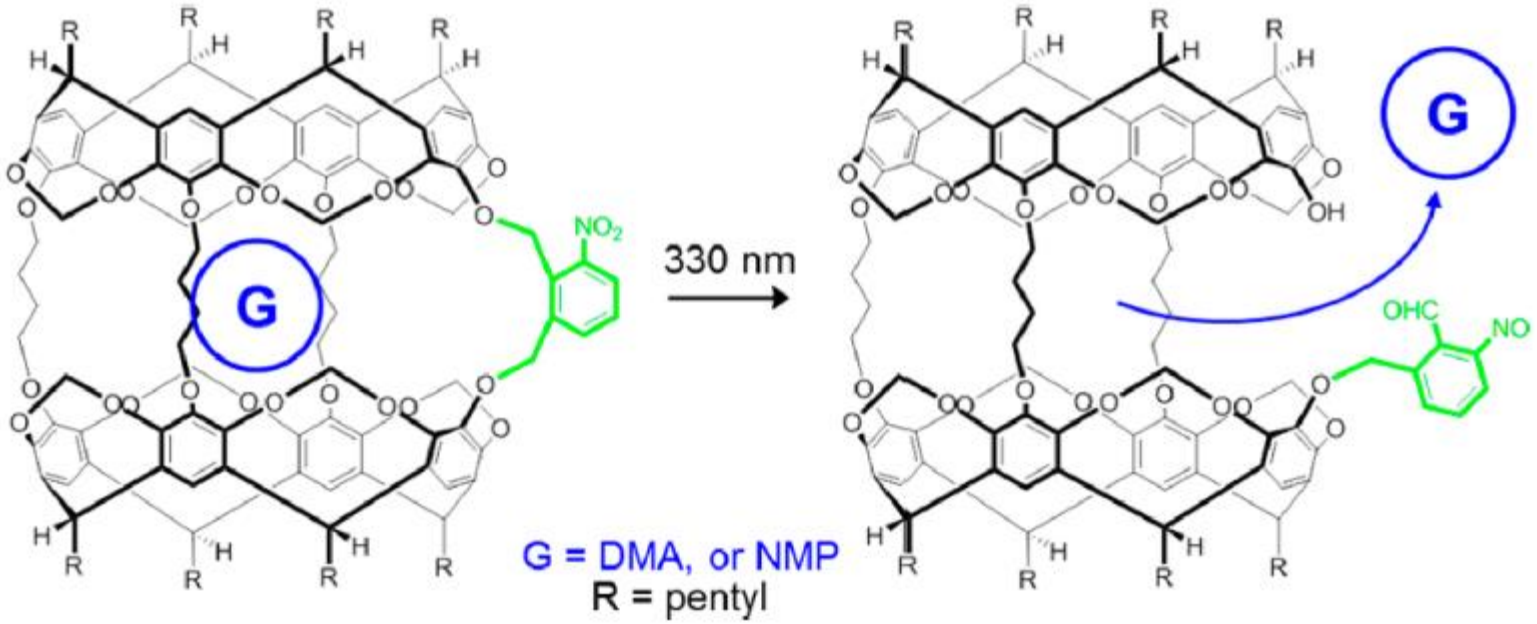
Stimulated Gating



Redox Gating



Photochemical Gating



Host–Guest Chemistry

DOI: 10.1002/anie.201205376

Reversible Photochemically Gated Transformation of a Hemicarcerand to a Carcerand**

*Hao Wang, Fang Liu, Roger C. Helgeson, and Kendall N. Houk**

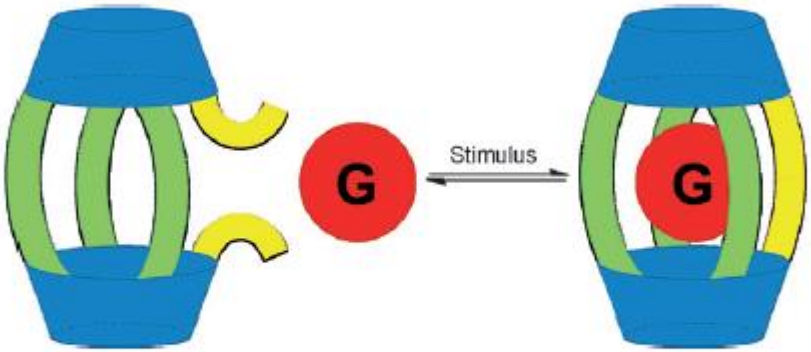
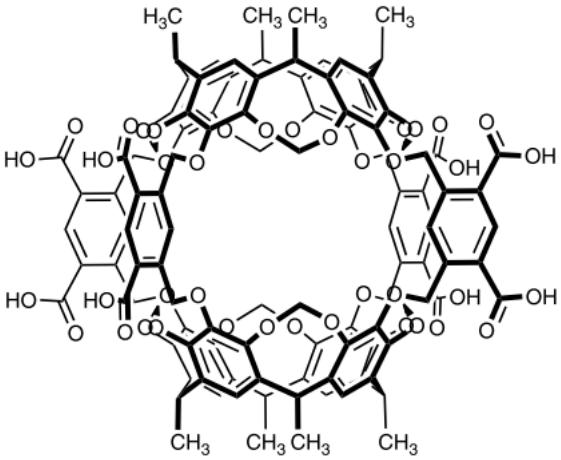
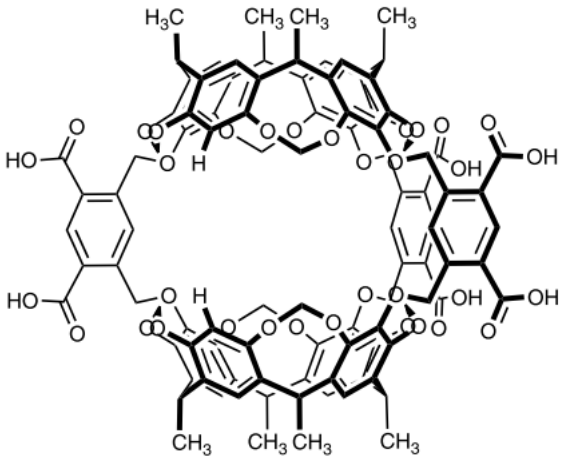


Figure 1. Gating in container molecules converts a hemicarcerand (left) into a carcerand (right).

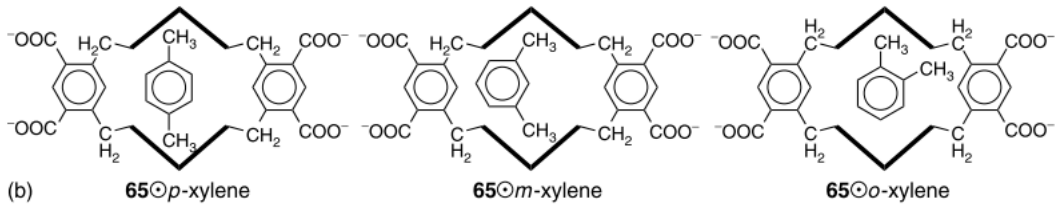


Water soluble octa-acid hemicarcerand:

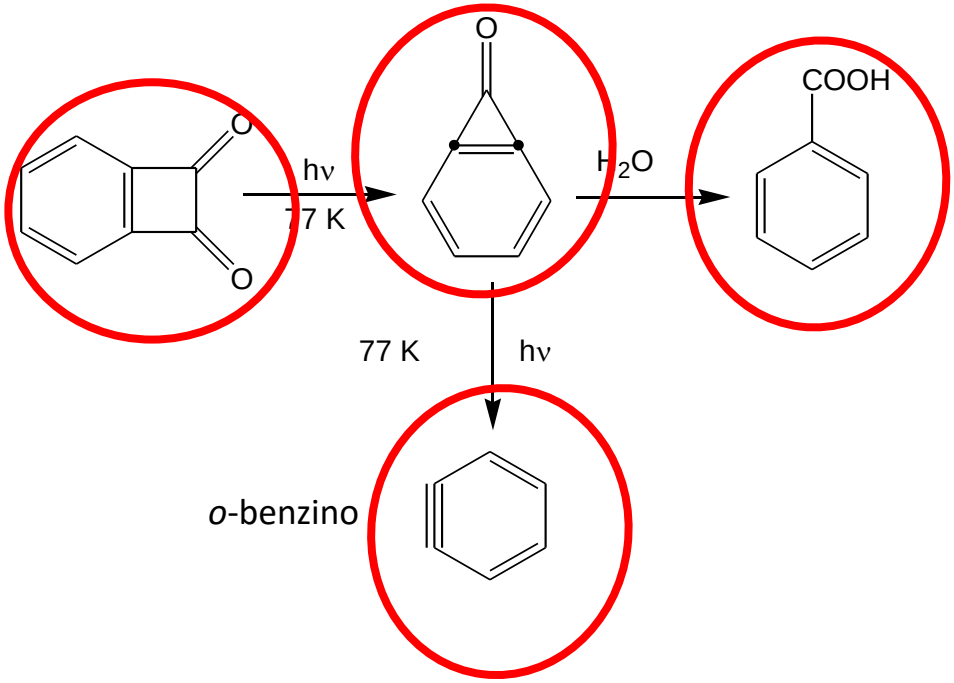
Hydrophobic effect (higher than cyclodextrines)!



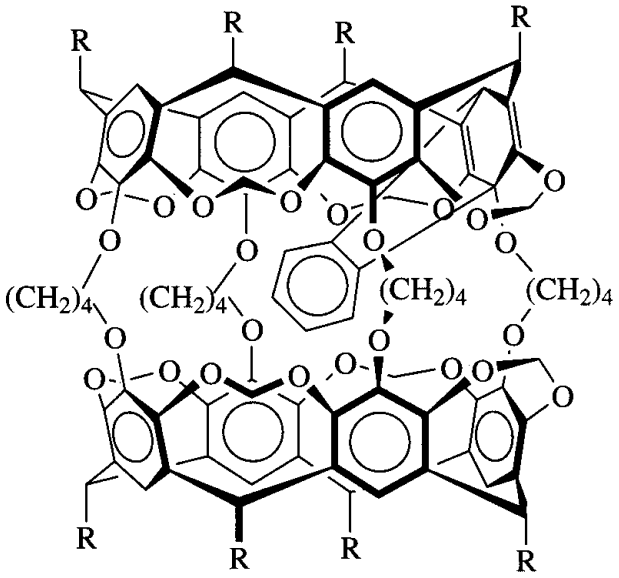
CH-π interactions for isomeric xylenes or dimetoxybenzenes direct the order of affinity:
meta > para >> ortho

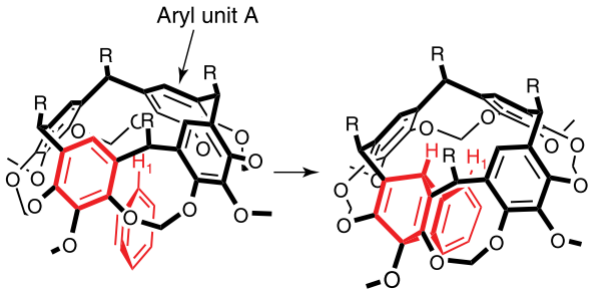


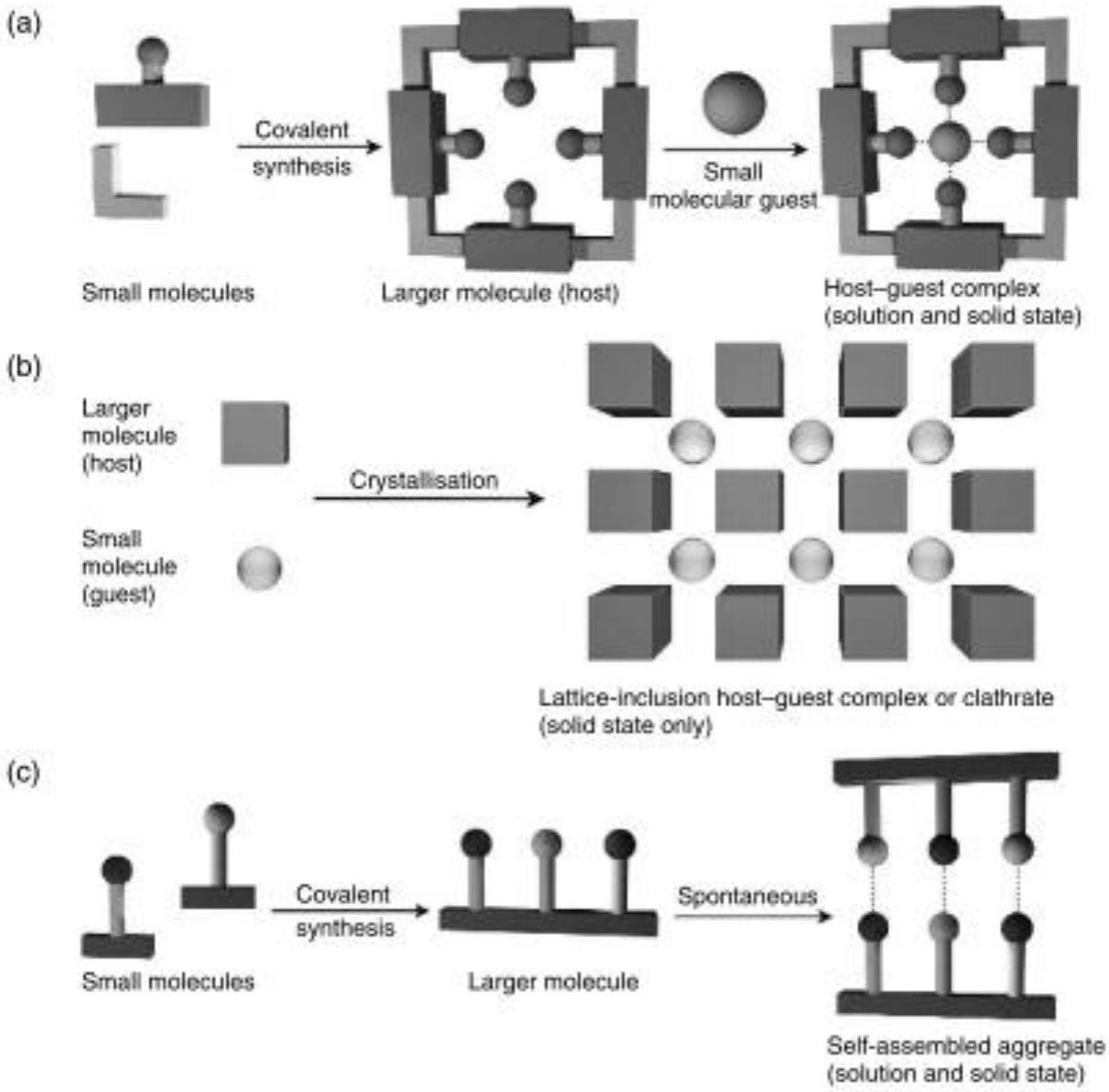
Benzociclobutendione Benzociclopropenone



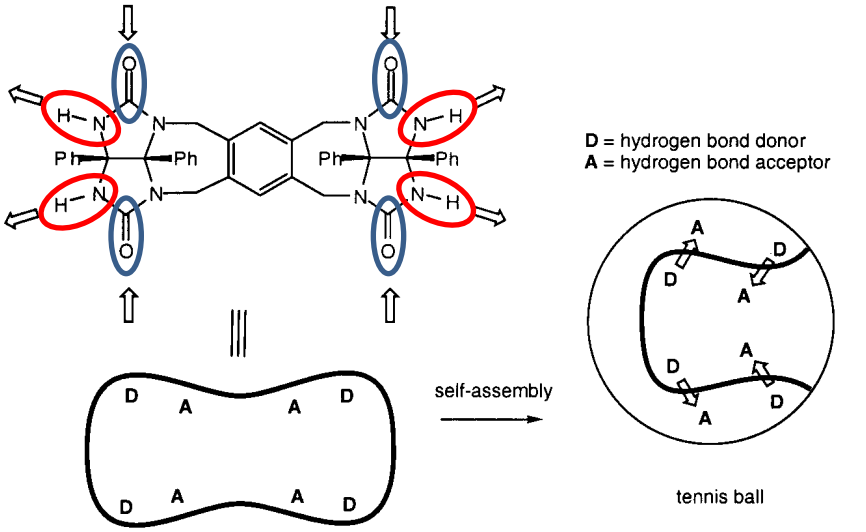
^{1}H e ^{13}C NMR a bassa T



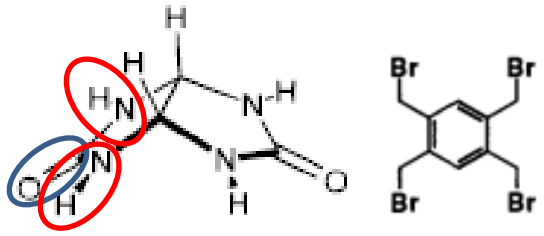



(a)
Covalent
synthesis
Small
molecular guest
Small molecules
Larger molecule (host)
Host–guest complex
(solution and solid state)
(b)
Larger
molecule
(host)
Small
molecule
(guest)
Crystallisation
Lattice-inclusion host–guest complex or clathrate
(solid state only)
(c)
Covalent
synthesis
Spontaneous
Small molecules
Larger molecule
Self-assembled aggregate
(solution and solid state)

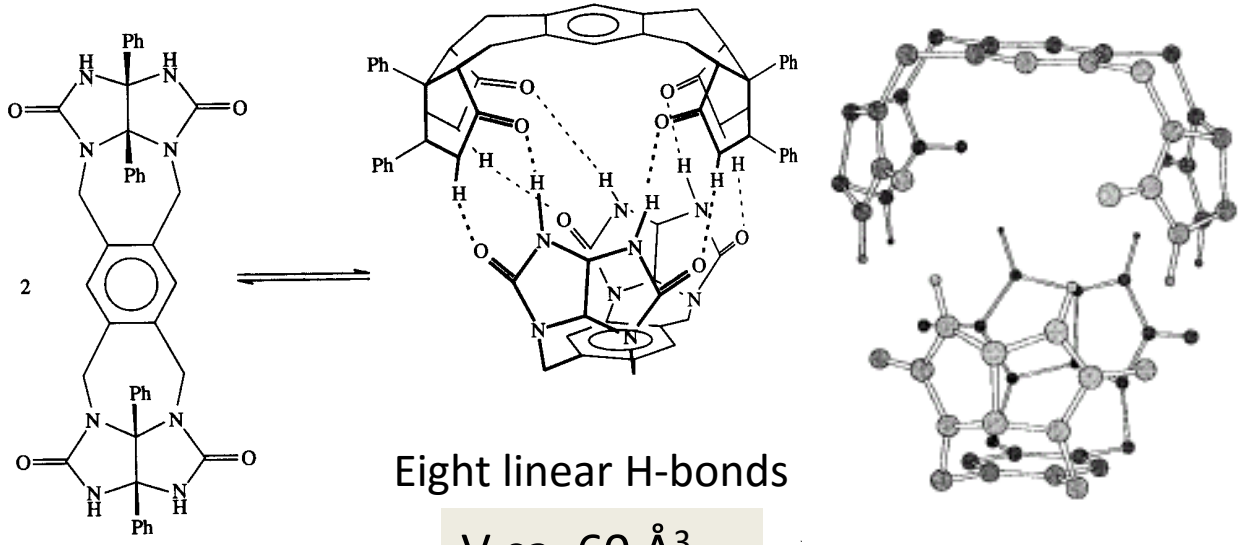
Tennis-ball



The formation of a molecular tennis ball

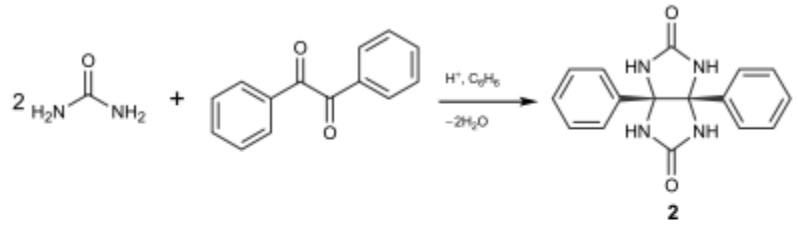


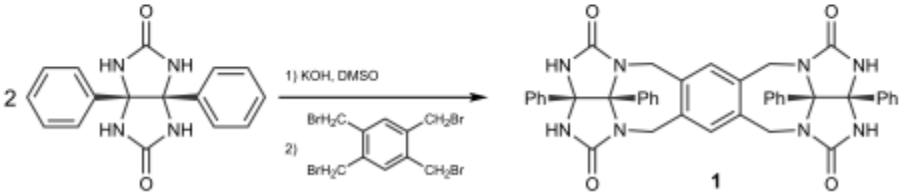
unità glicolurile

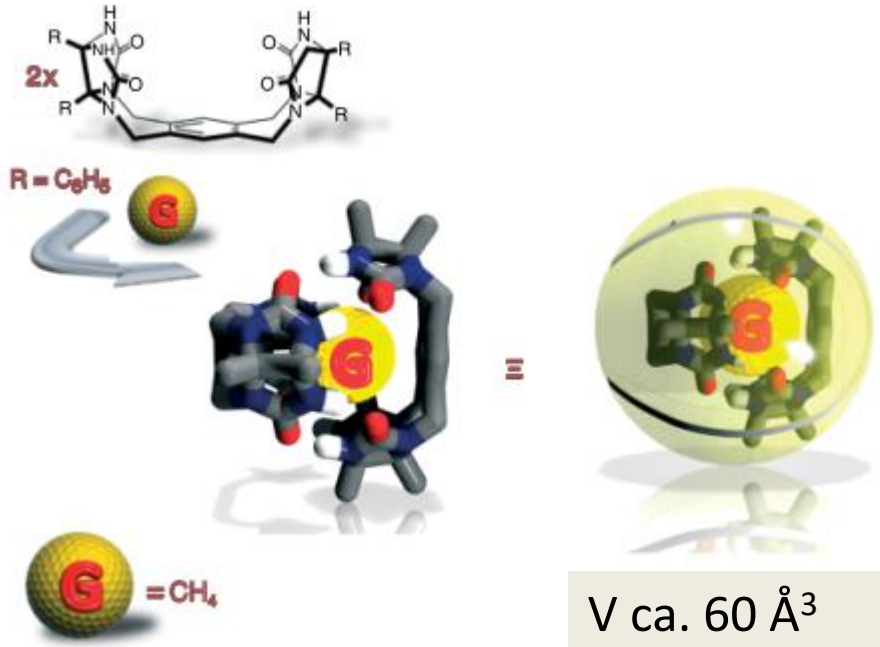


Eight linear H-bonds

V ca. 60 Å^3

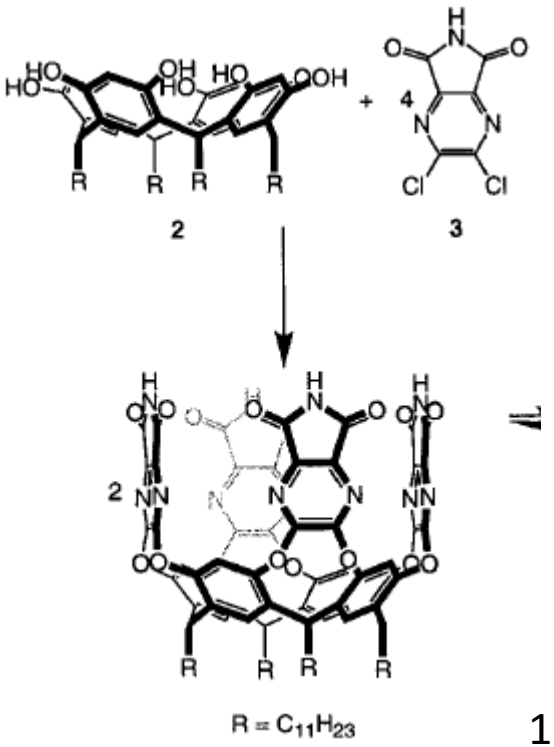

2
H2N
NH2
+
H+, C6H6
−2H2O
HN
NH
O
2


2
1) KOH, DMSO
2)
BrH2C
CH2Br
Ph
1

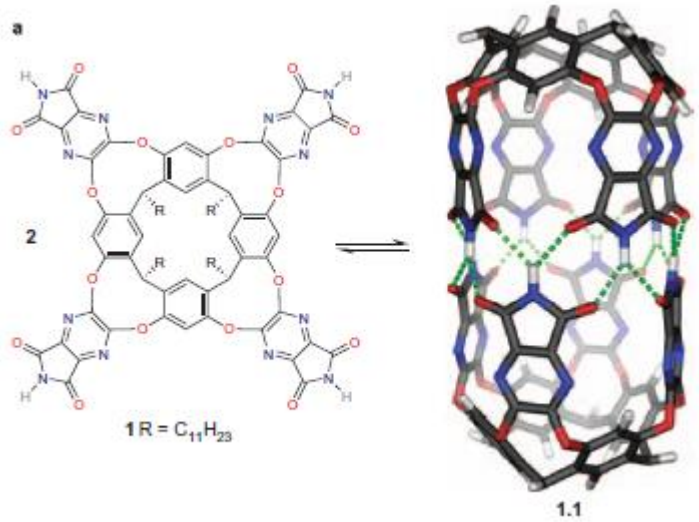

2x
R = C6H5
G
= CH4

V ca. 60 Å^3

Molecular Cylinder

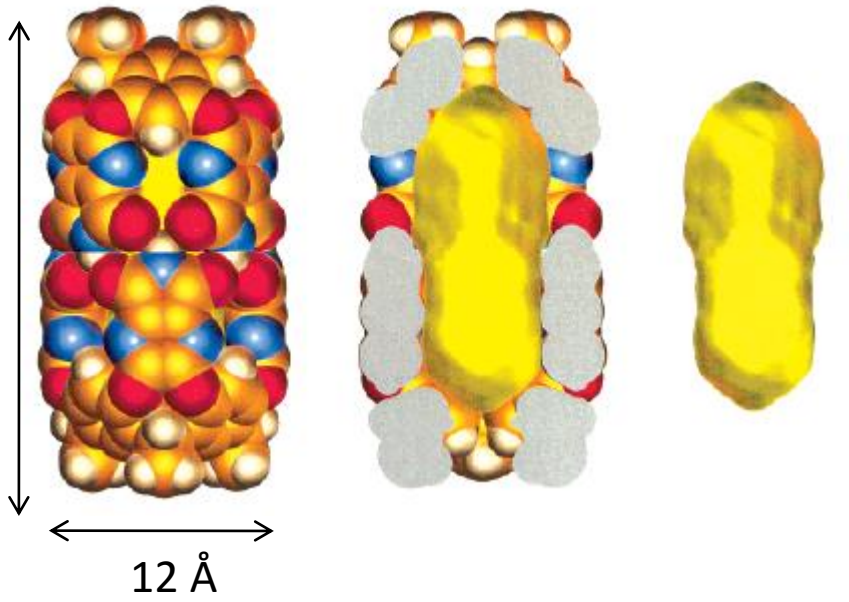


16 Å



Eight bifurcated H-bonds

V *ca.* 420 Å^3



From molecular mechanics calculations:

the encapsulated guest(s) occupy approximately 55% of the available space (same occupancy inside most weakly interacting organic solvents).

Stability deacreases at higher or lower space occupancy.

The ins and outs of molecular encapsulation

Liam C. Palmer and Julius Rebek, Jr.*

Org. Biomol. Chem., 2004, **2**, 3051–3059

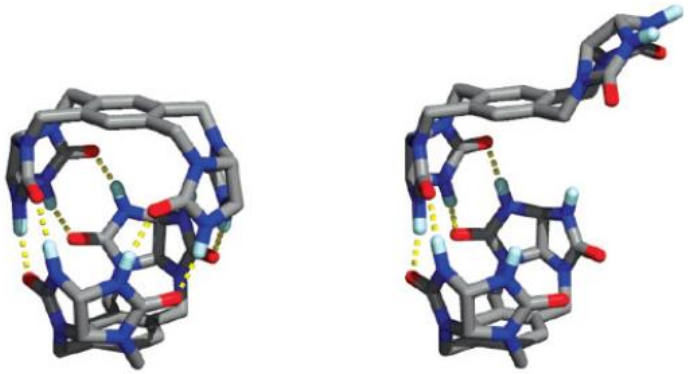
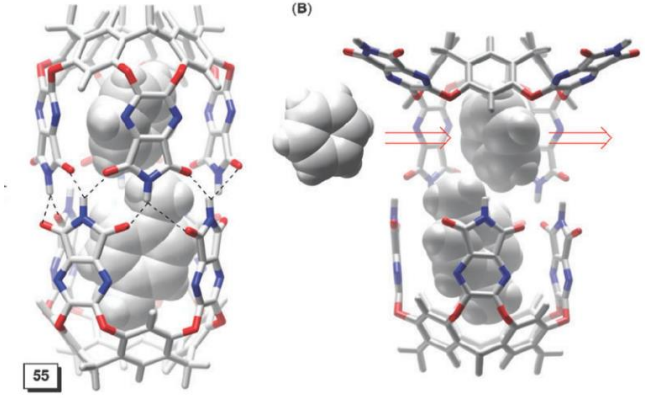


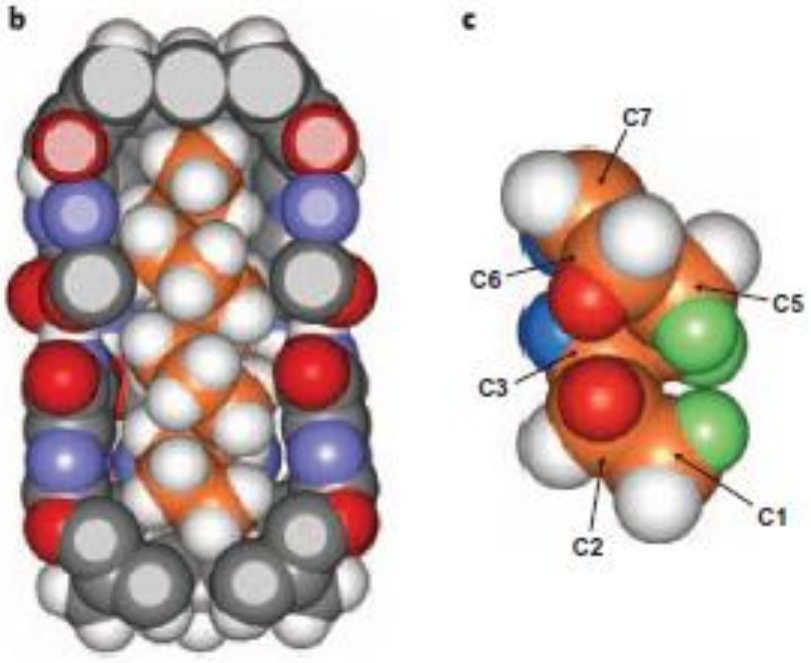
Fig. 5 The intact tennis ball, left, and the conformation with an open flap, right. The structure on the right represents intermediate of a possible "gating" mechanism.

A priori, two plausible exchange mechanisms can be imagined: (1) a dissociative mechanism in which the two capsule subunits completely separate, or (2) a gating mechanism[22] whereby one "flap" of the tennis ball opens by ring inversion of the seven-membered ring (as shown in Fig. 5).

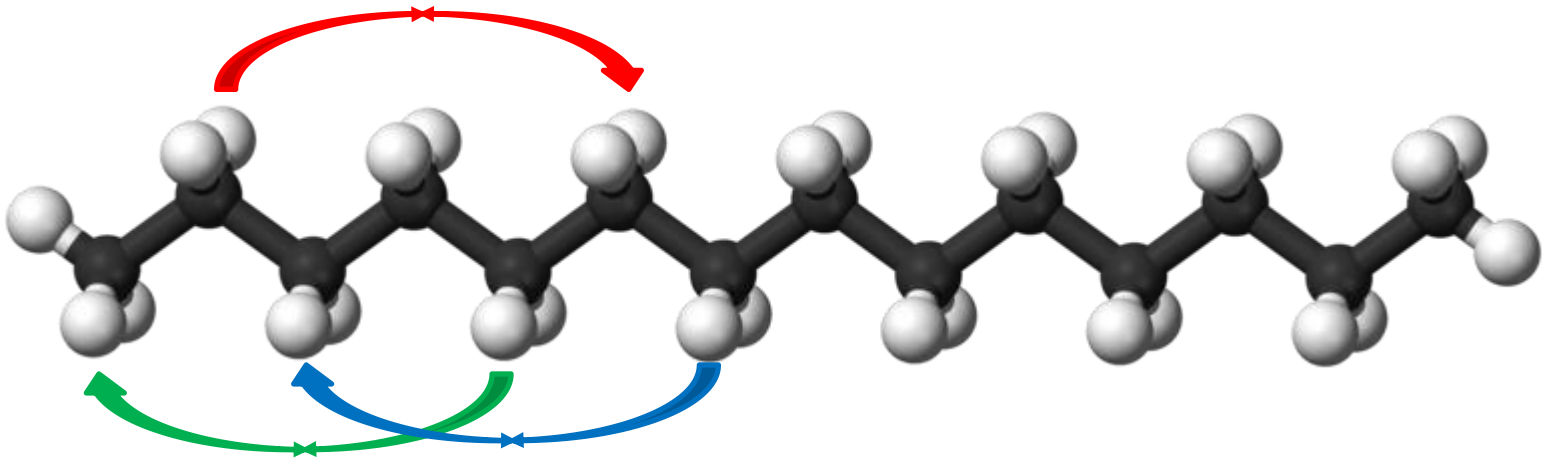


Small guests like toluene can escape by opening of a single flap, left. Larger guests like dimethylbiphenyl require opening of two or more flaps, right.

Contrasting the mechanism in apolar media, the presence of methanol seems to cause large guest exchange by complete dissociation of the capsule subunits. Complexes of small guests are not stable under these conditions.



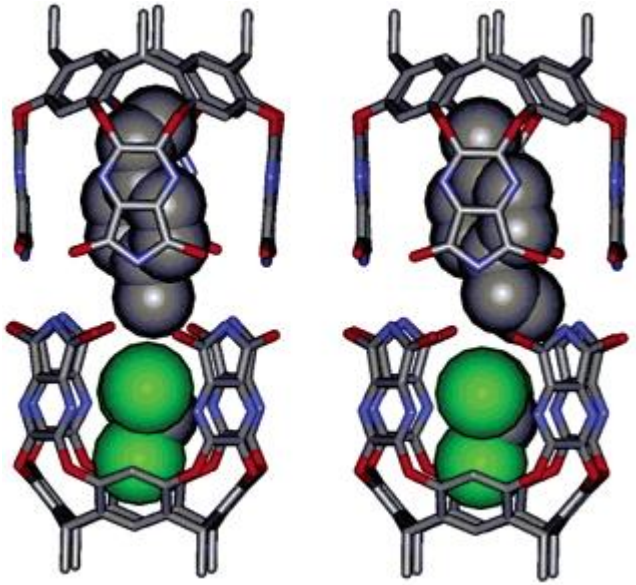
Model structure: incapsulation of coiled alkanes - tetradecane



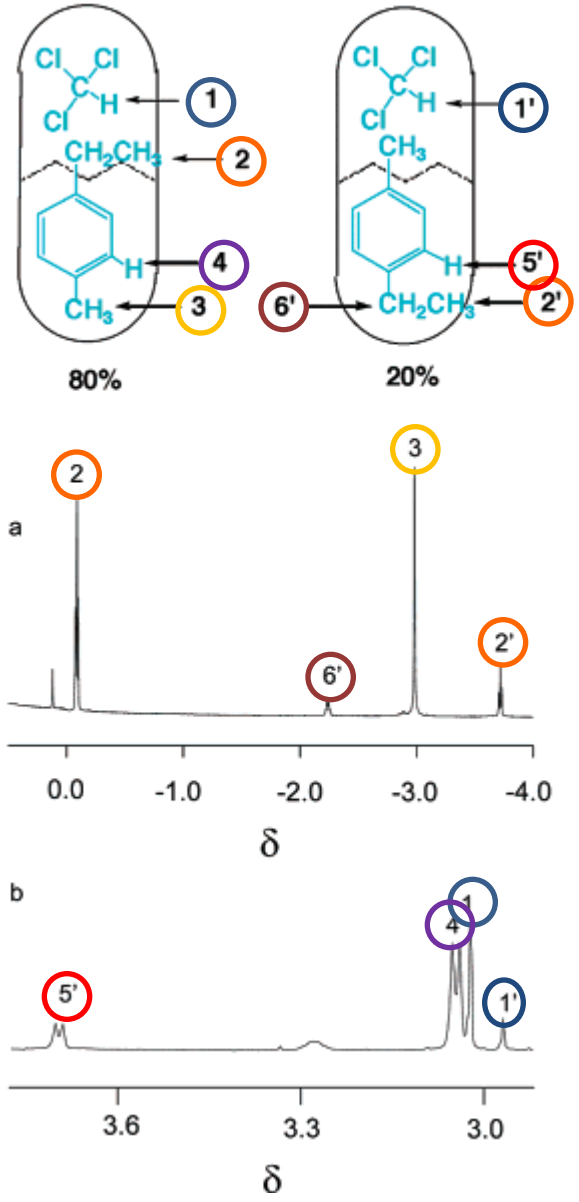
Social Isomers:

A new form of isomerism arises when two different guests are confined to a cylindrical, self-assembled host capsule. The shape and dimensions of the capsule prevent the guests from exchanging positions or tumbling on the NMR time scale. The phenomenon depends on matching guest size and shape with that of the host and on the interaction of the two guest molecules. The orientation of one guest depends on the presence and nature of the other.

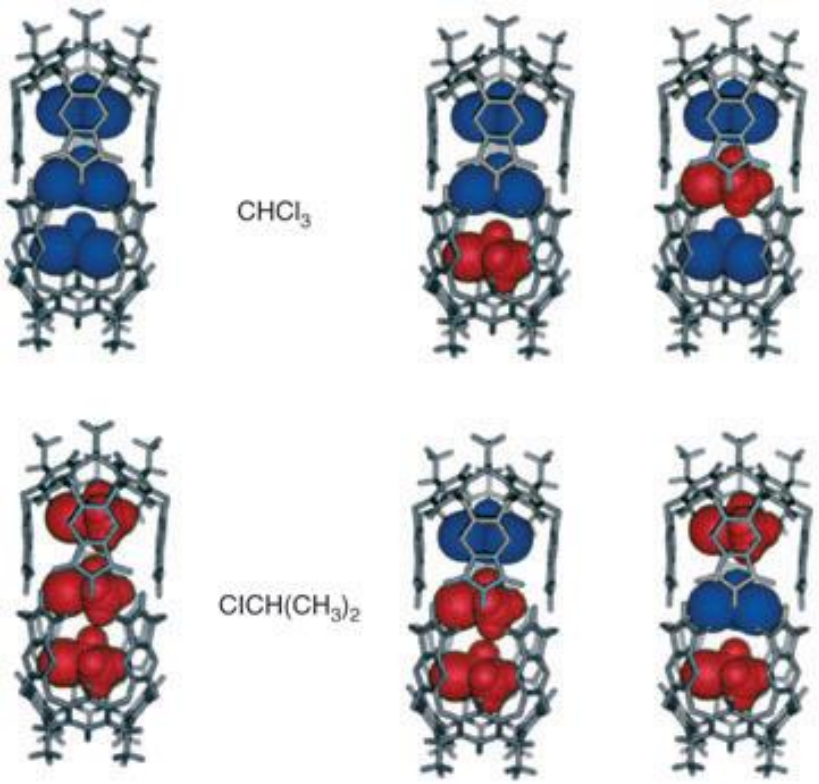
Social Isomers



MM optimized structures:
cloroformio e *para*-etiltoluene



Constellation Isomers

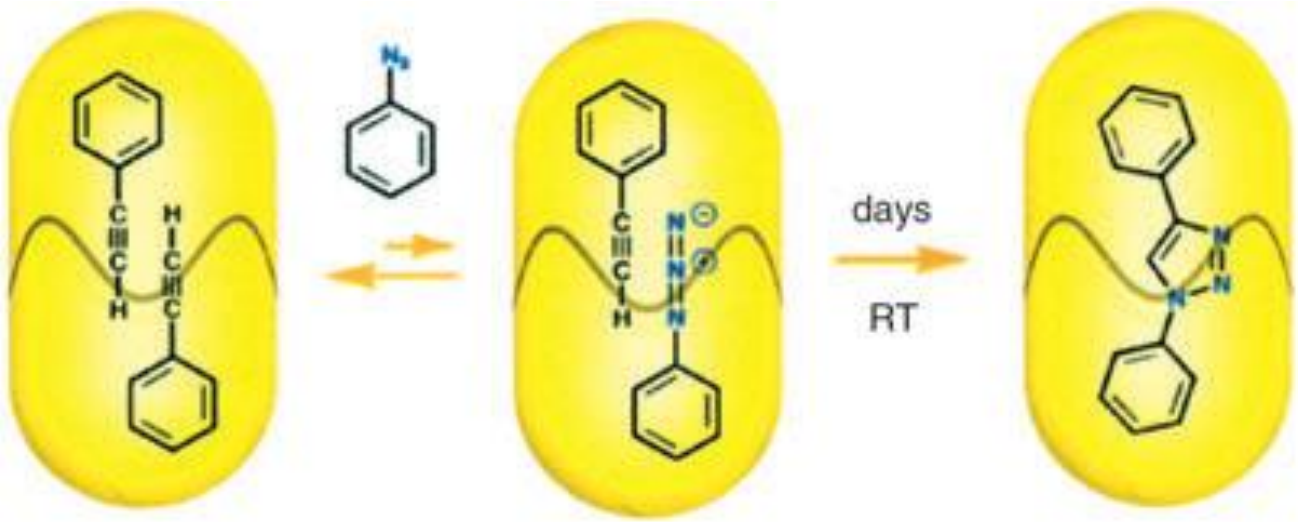


MM optimized structures:
cloroformio e *iso*-propilcloruro

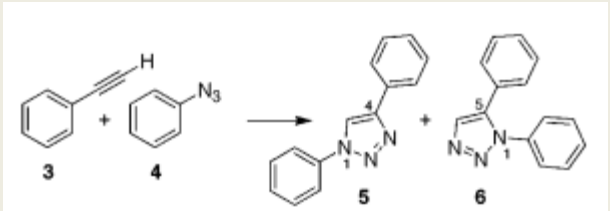
In addition to being able to preserve highly labile species, they may serve as catalysts and accelerate reactions inside their inner cavity by either concentrating the reactants leading to higher effective concentrations or TS stabilization or by preorganising them inside the capsule.

They may create a micro-environment in which two encapsulated reactants are held together in a orientation that differs from their most reactive arrangment in solution (or gas phase) leading to products that are disfavoured in equivalent solution phase reactions.

Reattività nelle capsule molecolari

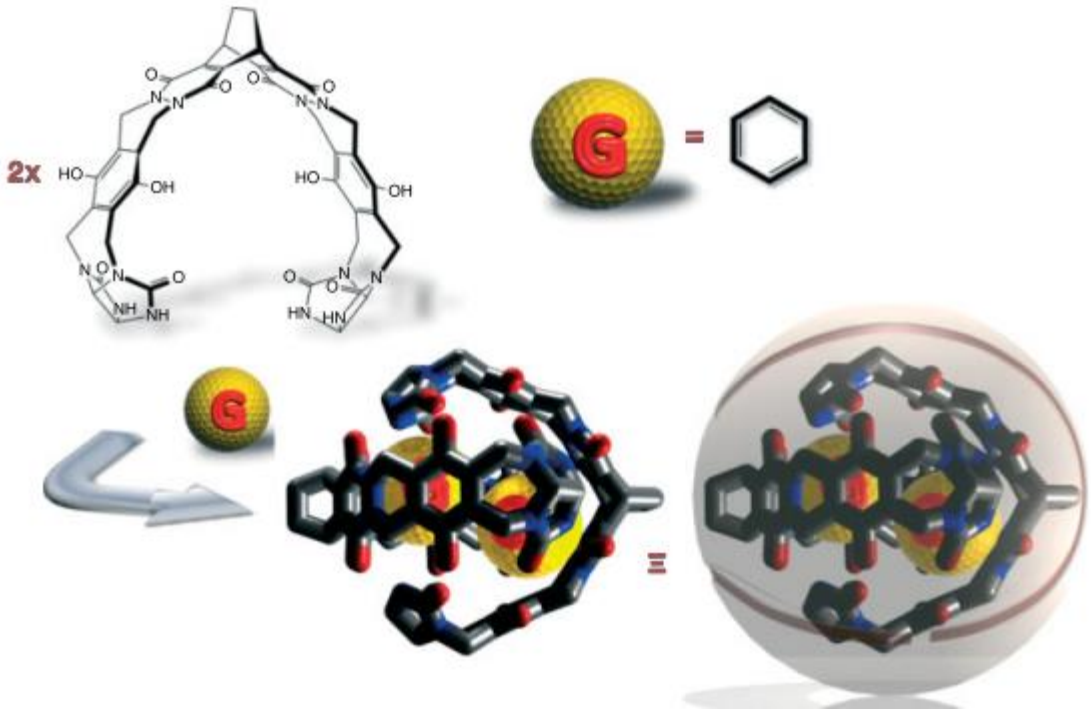


Cicloaddizione 1,3 regioselettiva di fenilacetilene e fenilazide:



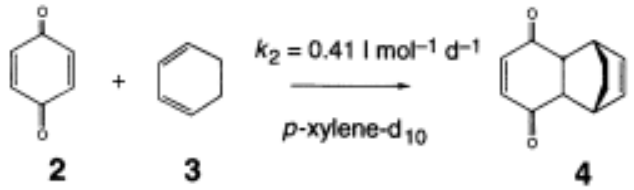
Volume definito = [ ] 4M vs mM
Tempo di contatto = 1 s vs 1 ns
Solvatazione fissa

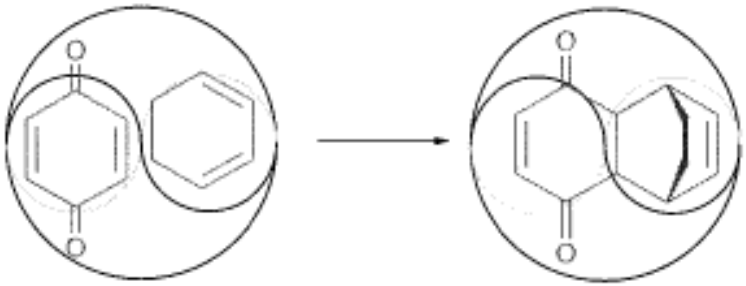
Soft Ball

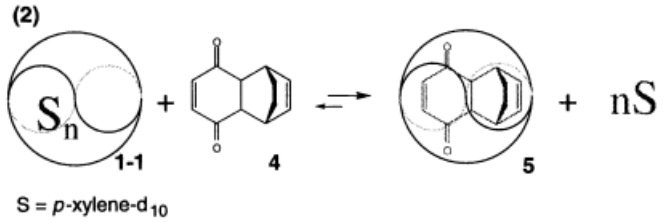


V ca. 400 Å^3

Reattività nelle capsule molecolari

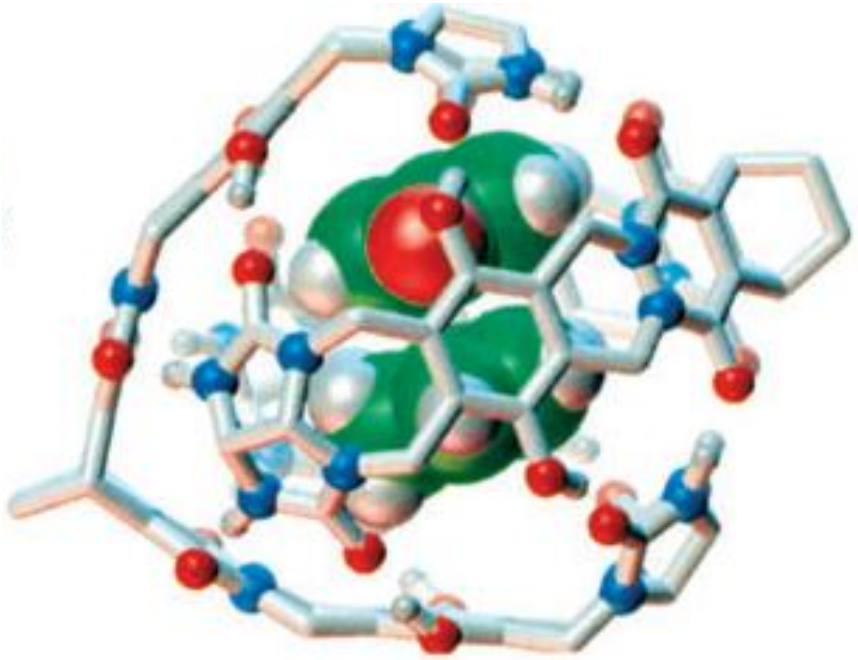






Reattività nelle capsule molecolari

Cicloaddizione Diels-Alder acceleraz di ca. 200 volte
[ ] = 5M
Solvatazione
Tempo di contatto

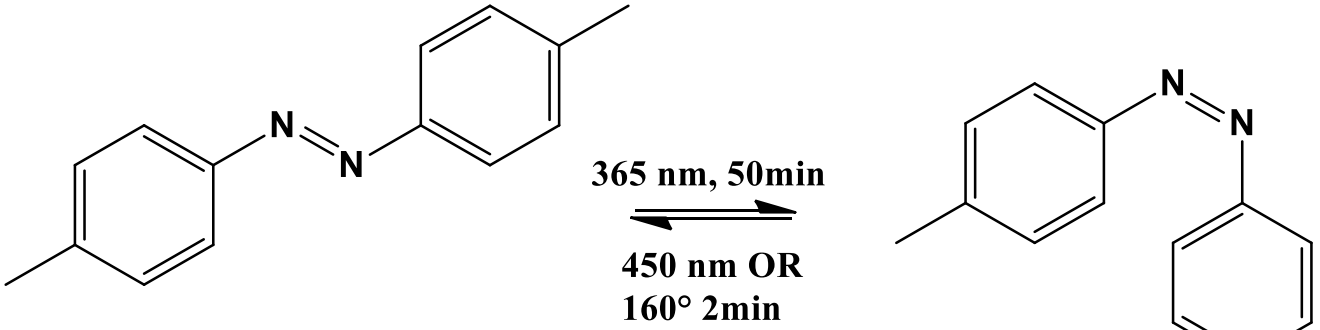


Host–Guest Chemistry

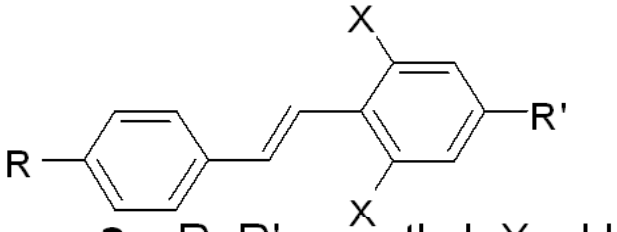
Angew. Chem. Int. Ed. 2010, 49, 3192 –3195

Photochemical Control of Reversible Encapsulation

*Henry Dube, Dariush Ajami, and Julius Rebek, Jr.**



trans-4,4'-dimethylazobenzene (*trans*-**1**) (*cis*-**1**)



R, R' = methyl, X = H