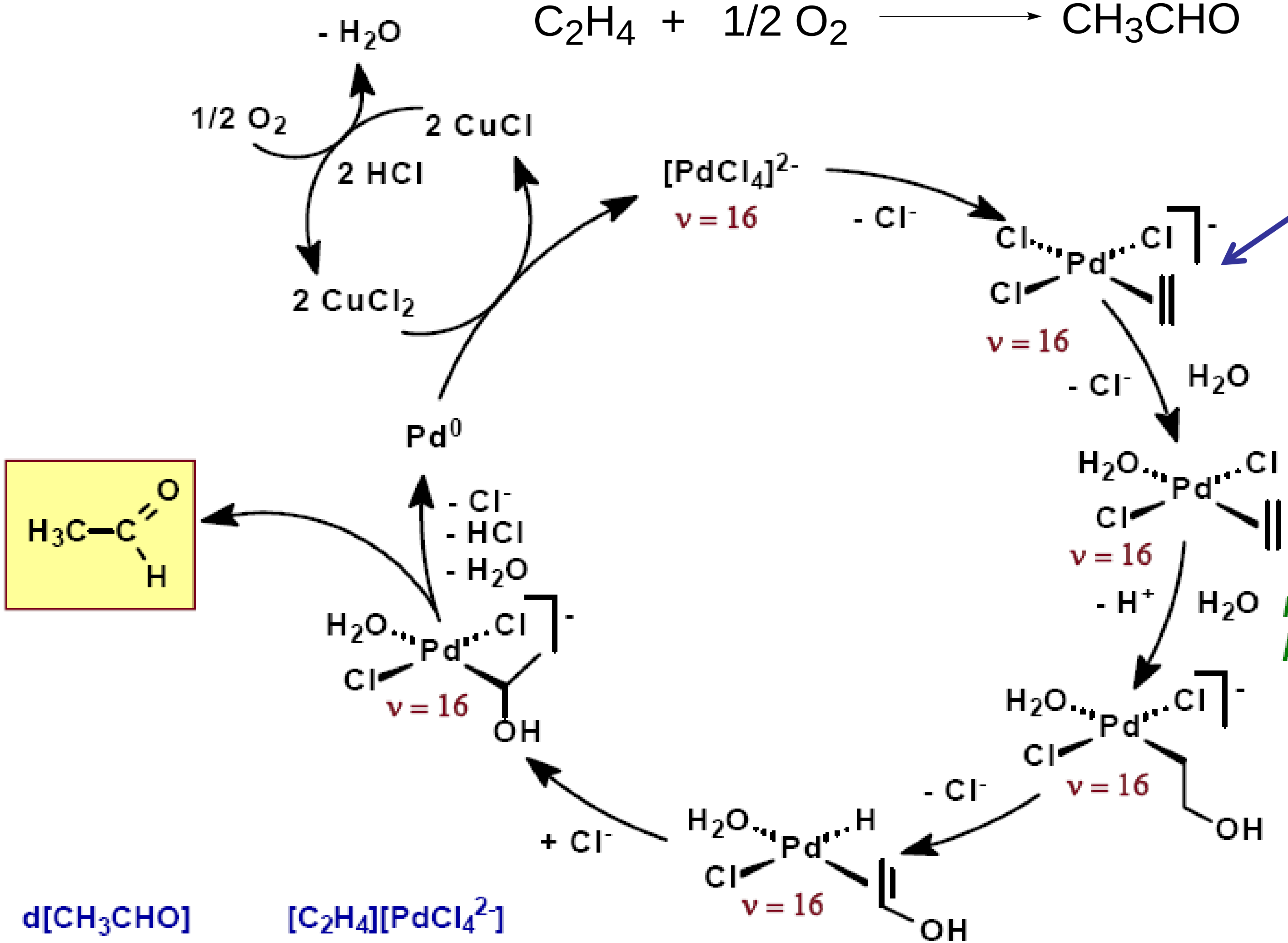
# Synthesis of acetic acid

## The Wacker process

$$C_2H_4 + 1/2\ O_2 \xrightarrow{cat} CH_3CHO \qquad \Delta G = -197\ kJ\ mol^{-1}$$



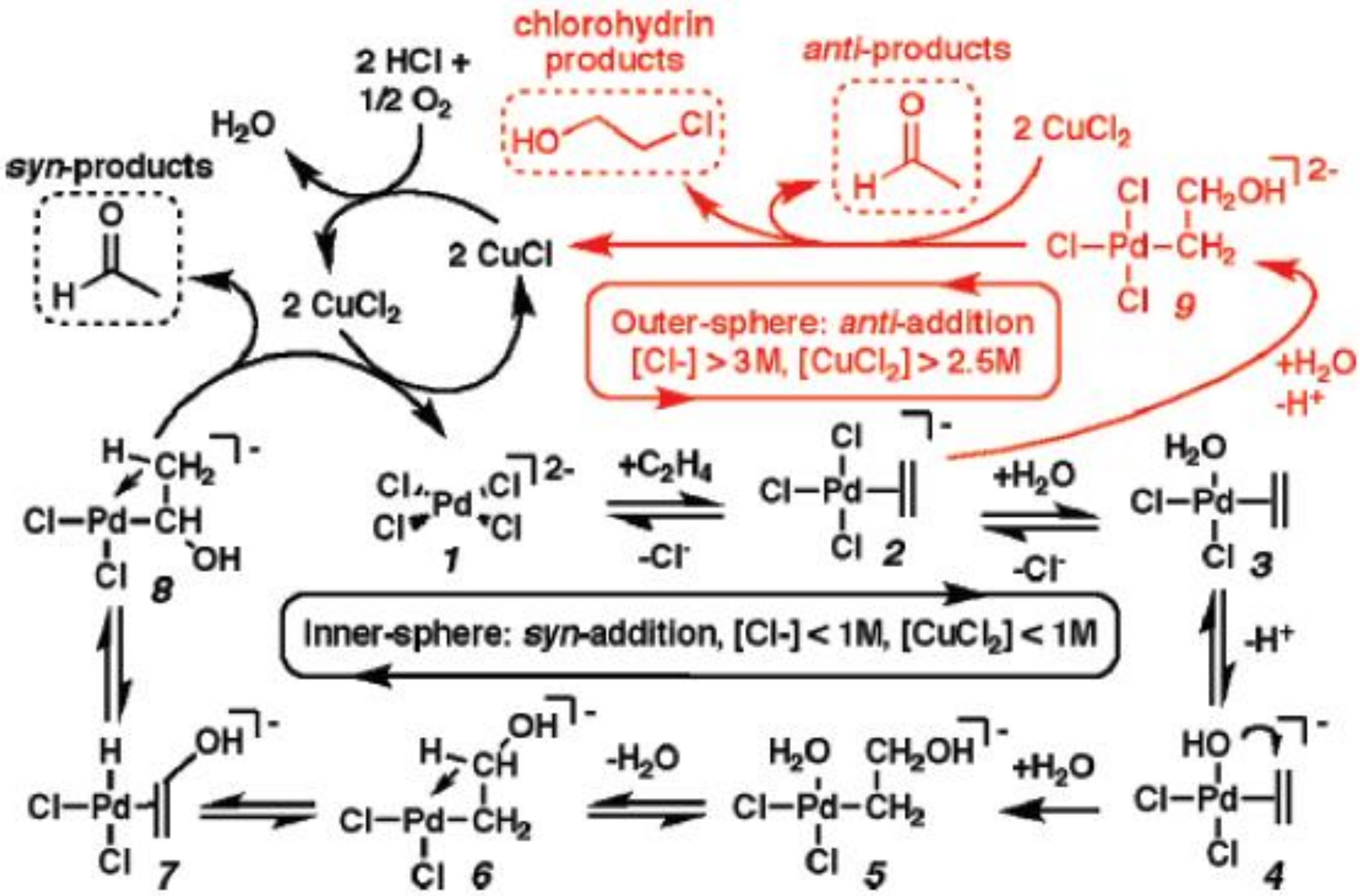
analogo del sale di Zeise

*nucleophilic attack by $H_2O$*

◀

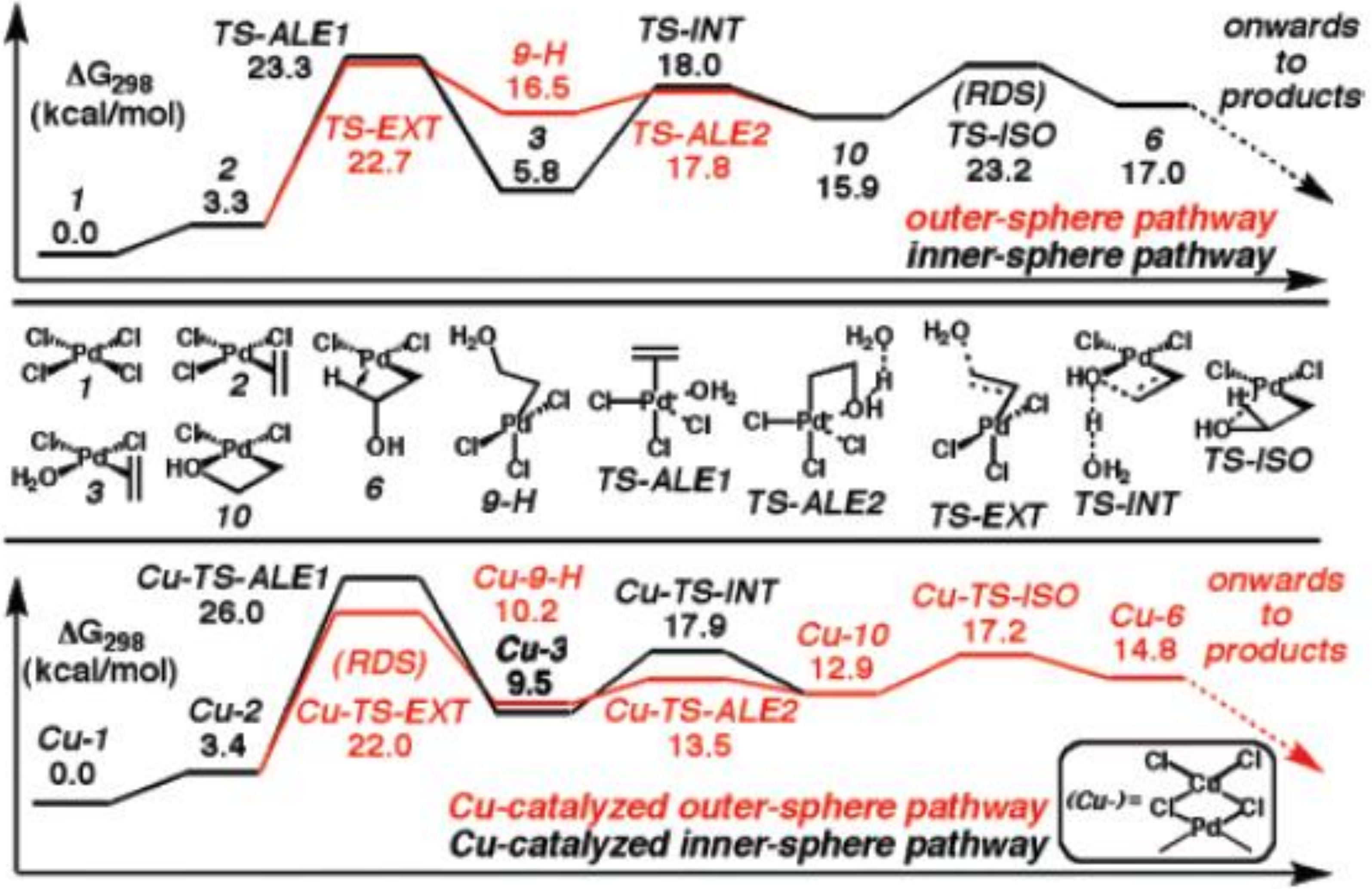
$$\frac{d[CH_3CHO]}{dt} = k\,\frac{[C_2H_4][PdCl_4^{2-}]}{[H^+][Cl^-]^2}$$

# *The recently reported mechanism*

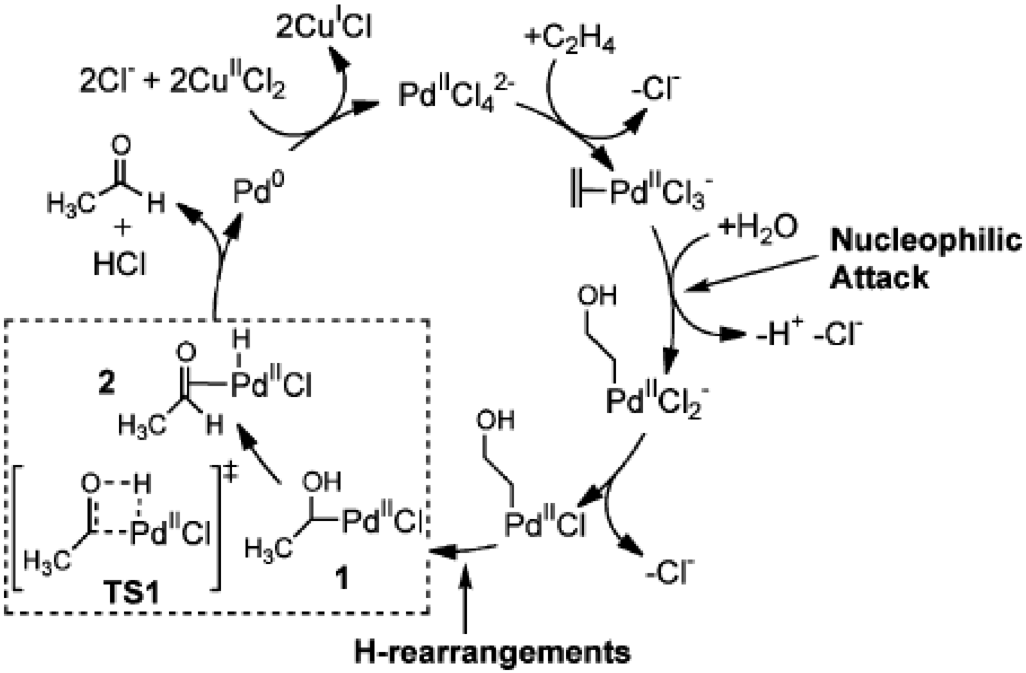


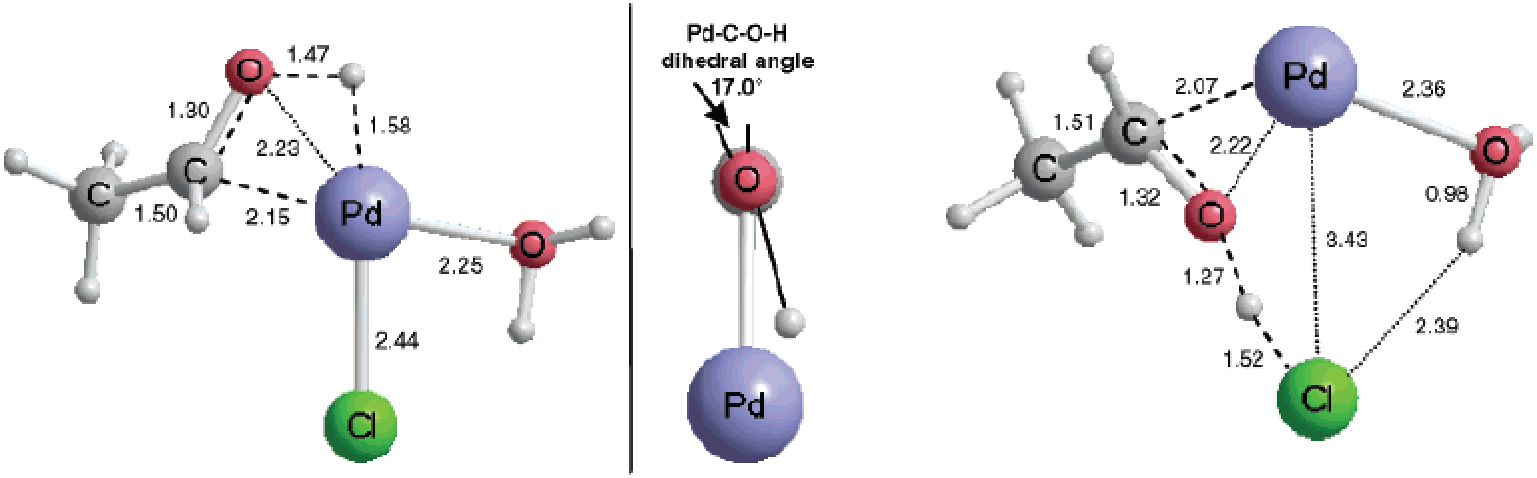
J. Oxgaard, W. A. Goddard, III et al *J. Am. Chem. Soc.* 2007, *1291*, 12342.

# The recently reported mechanism



# *The recently reported mechanism*





# Oxidation of Pd(0) to Pd(II) by the Cu(II) salt

$$Pd(0) + 2\,[CuCl_4]^{2-} \longrightarrow Pd^{2+} + 2\,[CuCl_2]^{-} + 4\,Cl^{-}$$

$$2\,[CuCl_2]^{-} + 1/2\,O_2 + 2\,H^{+} + 4\,Cl^{-} \longrightarrow 2\,[CuCl_4]^{2-} + H_2O$$

## One stage process

**Conversion limited to 40 % → Gas recycle necessary**

**High purity gases are needed: pure $O_2$; 99.9 % ethene**

**Ti-equipment needed (highly corrosive HCl in oxidative medium)**
**Low pressure**

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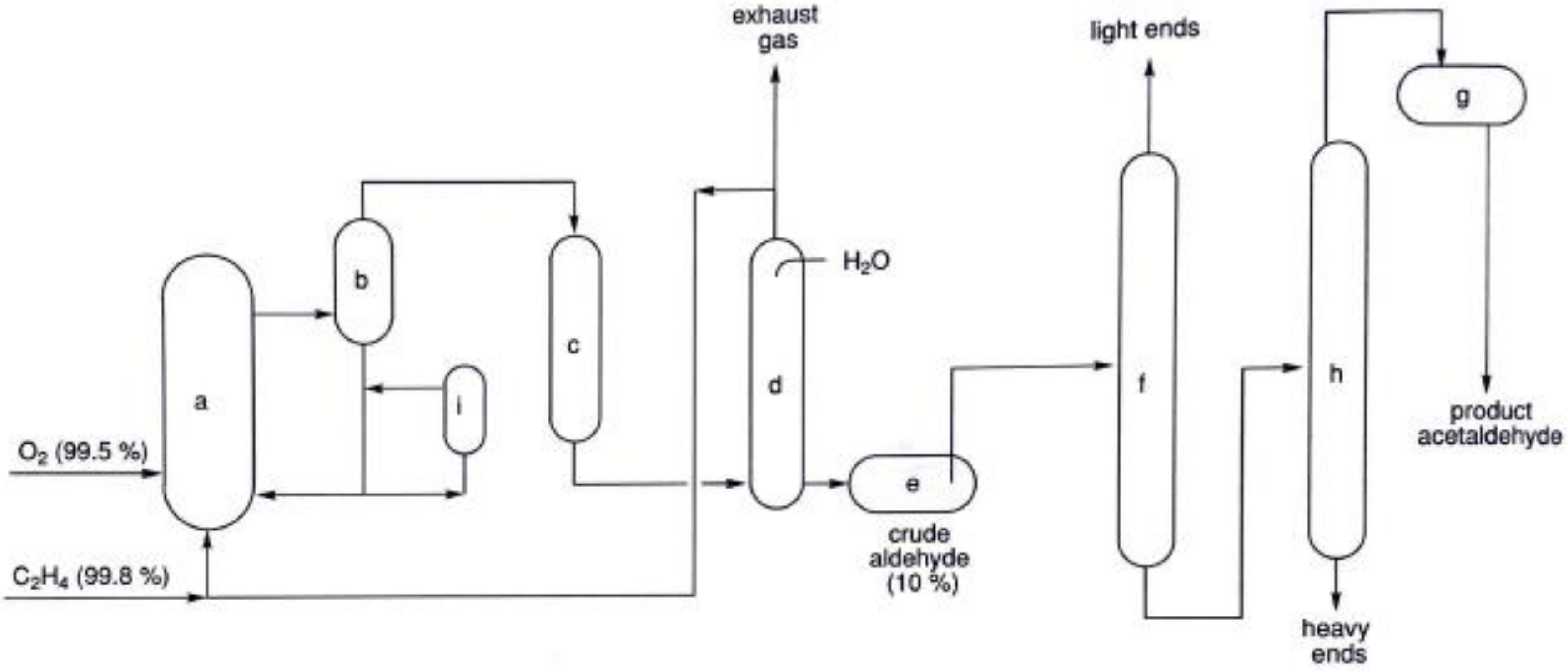
## Two stage process

**Total Conversion of ethene → No gas recycle**

**Raw Gas**

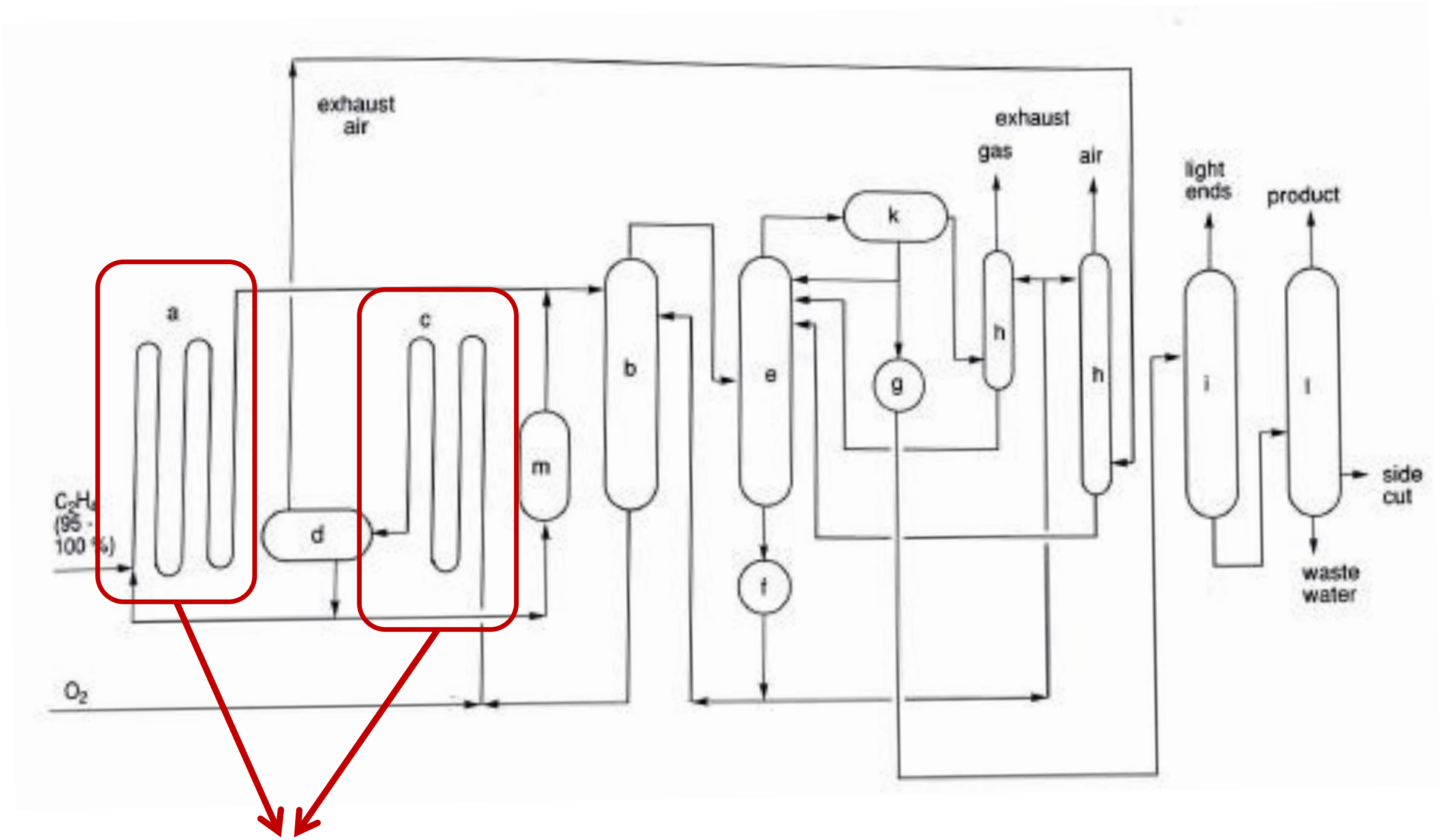
**Higher pressure + two reactors → Higher investiments**

# *Flow scheme of the Wacker process at one stage*



Reaction Conditions:
T = 100 - 130 °C; $P_{CO}$ = 4 - 10 bar.

# *Flow scheme of the Wacker process at two stages*



**reattori**

Reaction Conditions:
T = 100 - 130 °C; $P_{CO}$ = 10 bar.

# The Monsanto process

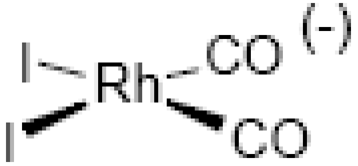
It is based on the carbonylation reaction of methanol

$$CH_3OH + CO \rightarrow CH_3COOH$$

$\Delta G$, *standard conditions, -75 kJ.mol*$^{-1}$

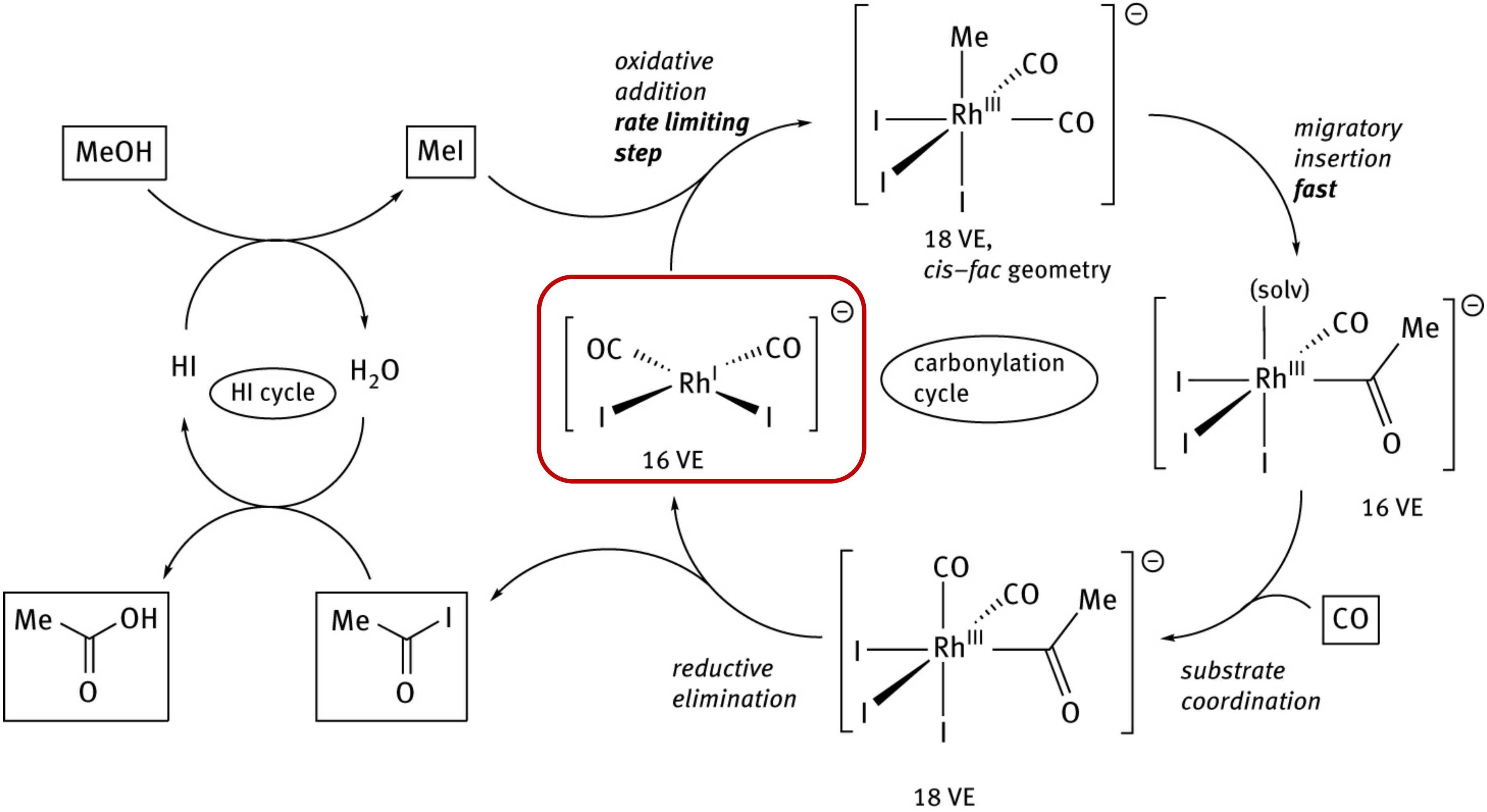
**ATOM EFFICIENCY = 100 %** **E FACTOR = 0**

in situ catalytic system:
$RhI_3$, CO e $H_2O$



$$CH_3OH + HI \leftrightarrows CH_3I + H_2O \qquad (1)$$

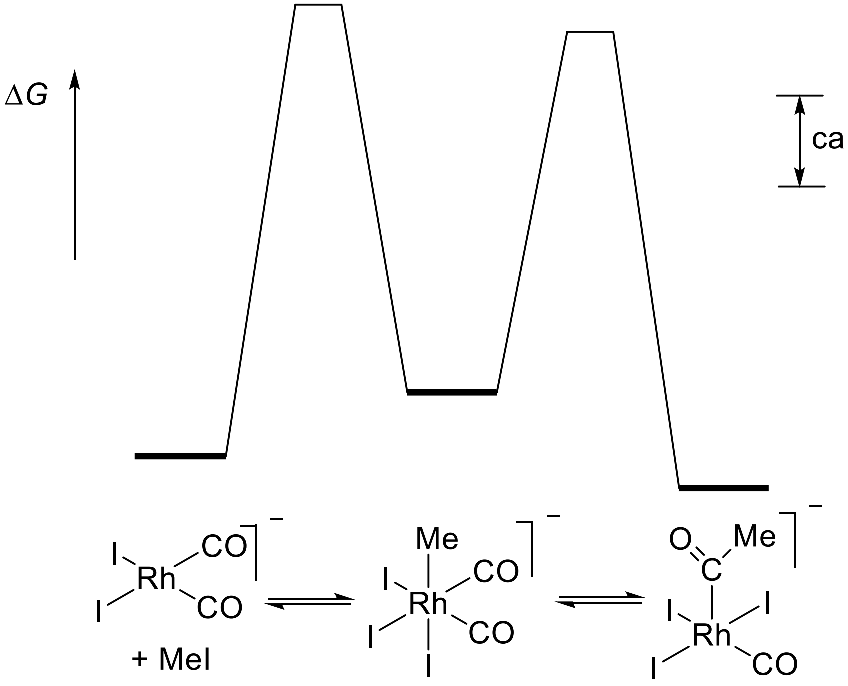
# The catalytic cycle



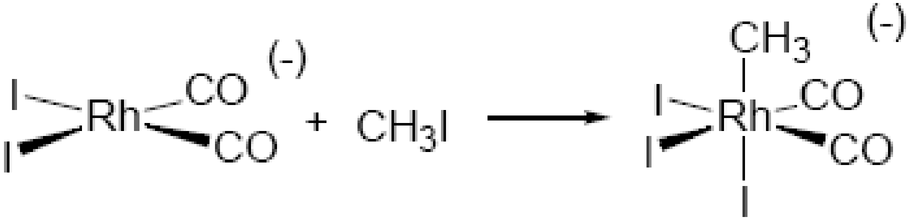
# The rate determining step of the catalytic cycle

Rate law: $v = k\,.[\mathrm{Rh\,I_2(CO)_2^{-}}].[\mathrm{CH_3I}]$

v: $\mathrm{CH_3I > CH_3Br > CH_3Cl}$



$K_1 = 4.5\cdot10^{-3}$ l/mol $K_2 = 3.2\cdot10^{3}$



$\nu_{CO}$: 2055 e 1984 $\mathrm{cm^{-1}}$

$^{13}$C NMR: two equivalent carbons with $J_{Rh}$ = 60 Hz

$^{13}$C NMR: methyl group with $J_{Rh}$ = 14 Hz at $\delta$ = -0.6 ppm

# *Side reactions*

$$CO + H_2O \rightarrow H_2 + CO_2$$

$$RhI_2(CO)_2^{-} + OH^{-} \rightarrow RhI_2(CO)(CO_2H)^{2-}$$

$$RhI_2(CO)(CO_2H)^{2-} \rightarrow RhI_2(CO)(H)^{2-} + CO_2$$

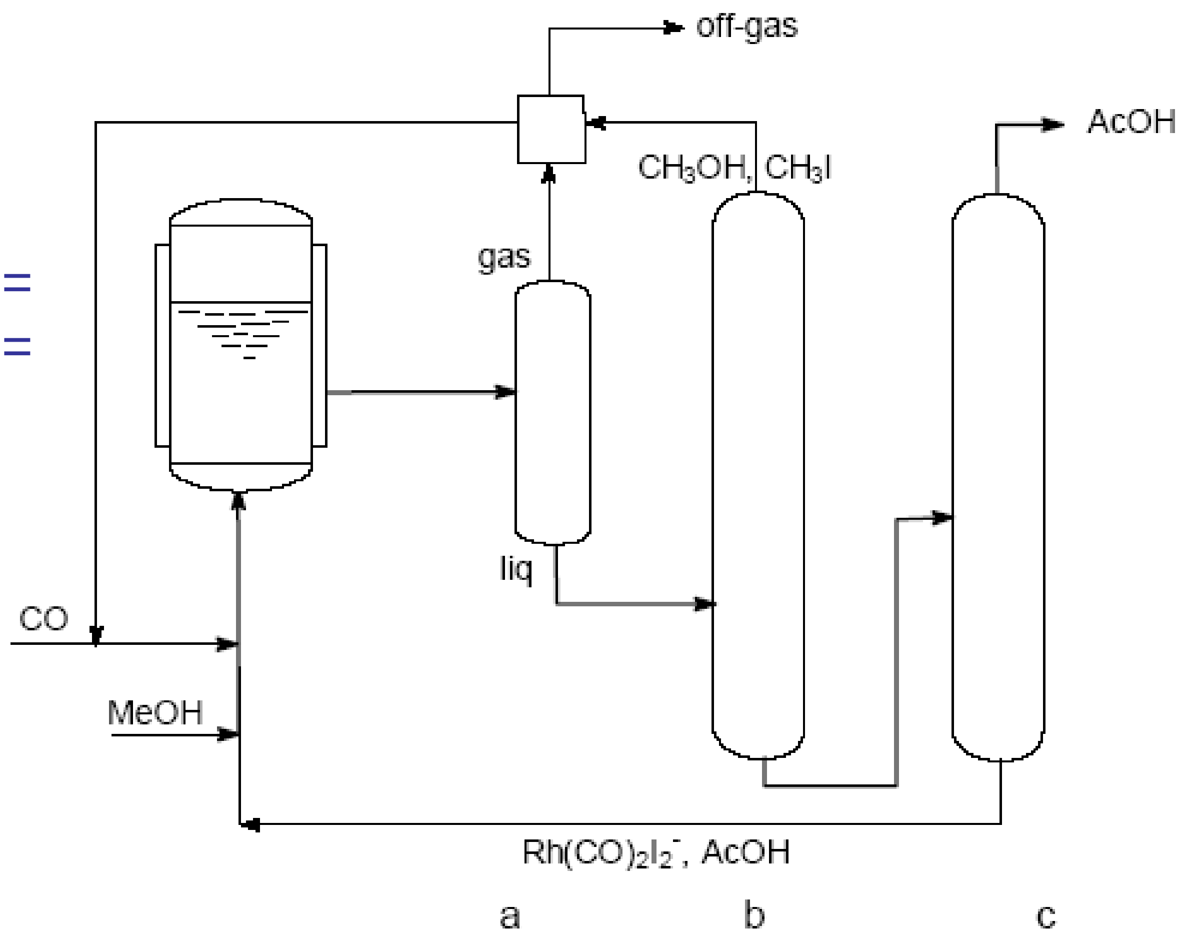
$$RhI_2(CO)(H)^{2-} + H^{+} + CO \rightarrow RhI_2(CO)_2^{-} + H_2$$

$$CH_3RhI_3(CO)_2^{-} + HI \rightarrow CH_4 + RhI_4(CO)_2^{-}$$

$$RhI_2(CO)_2^{-} + 2\,HI \leftrightarrows H_2 + RhI_4(CO)_2^{-}$$

# *Flow scheme of Monsanto process*

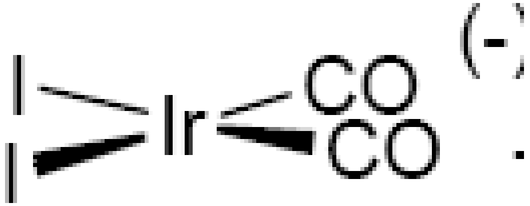
Reaction Conditions:
[Rh] = 10 mM; [$CH_3I$] = 1.5 M; T = 180 °C; $P_{CO}$ = 50 bar.



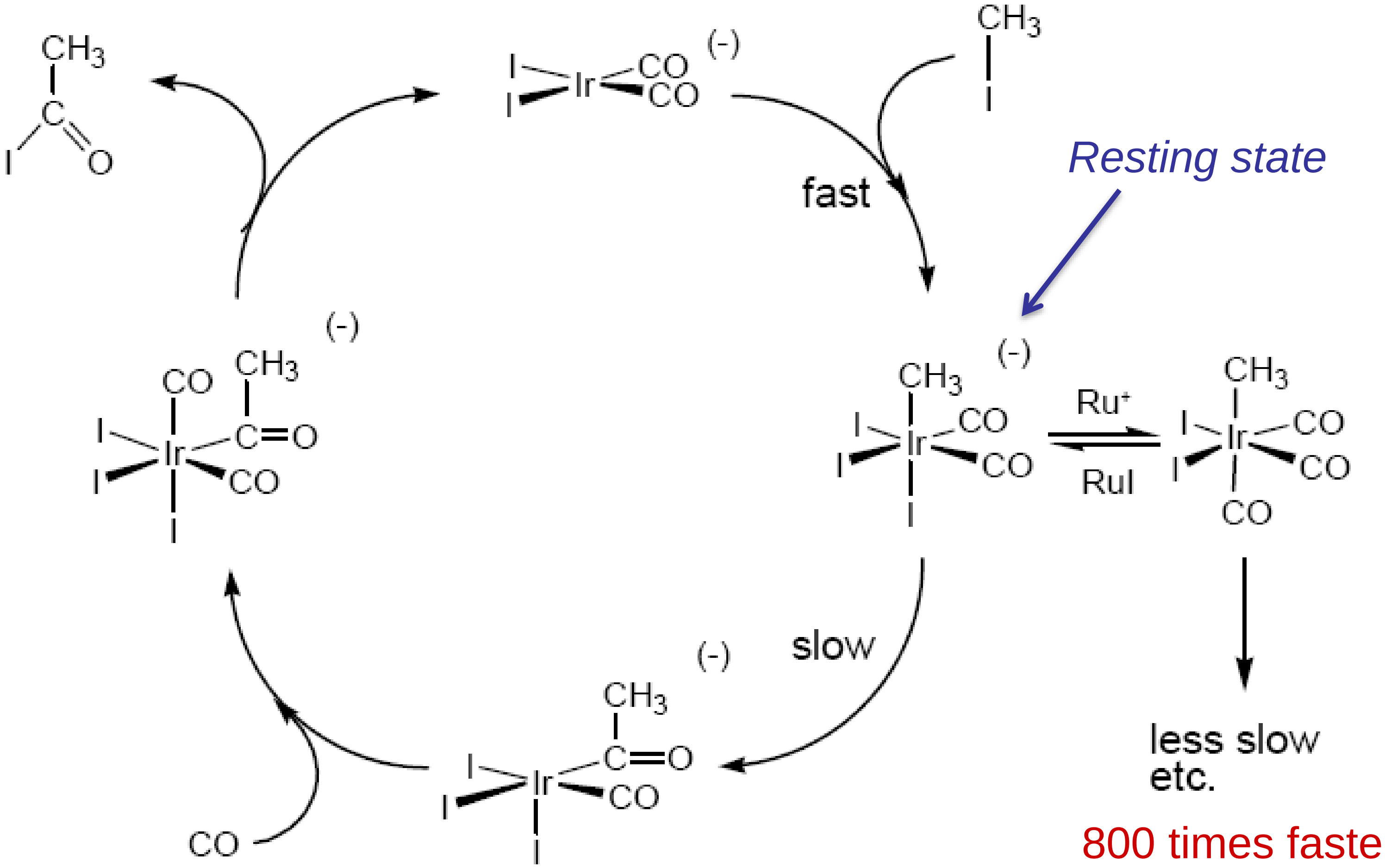
# *The CATIVA process*

It is based on the carbonylation reaction of methanol

It is based on an Iridium catalyst;
High rate at low water concentration;
High stability allowing a wide range of process condition;.
The catalyst is about 25 % faster than the Monsanto;
Thanks to lower content of sideproducts, the produced acetic acid is of better quality than that obtained by the Monsanto process.



# The catalytic cycle

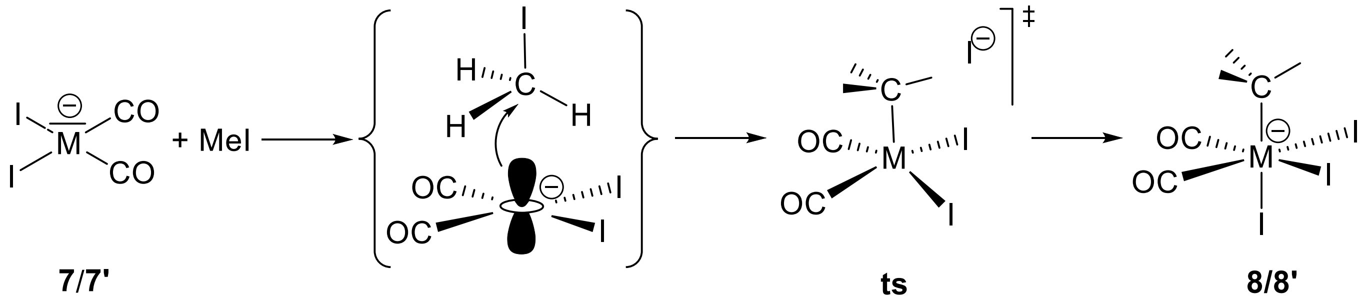


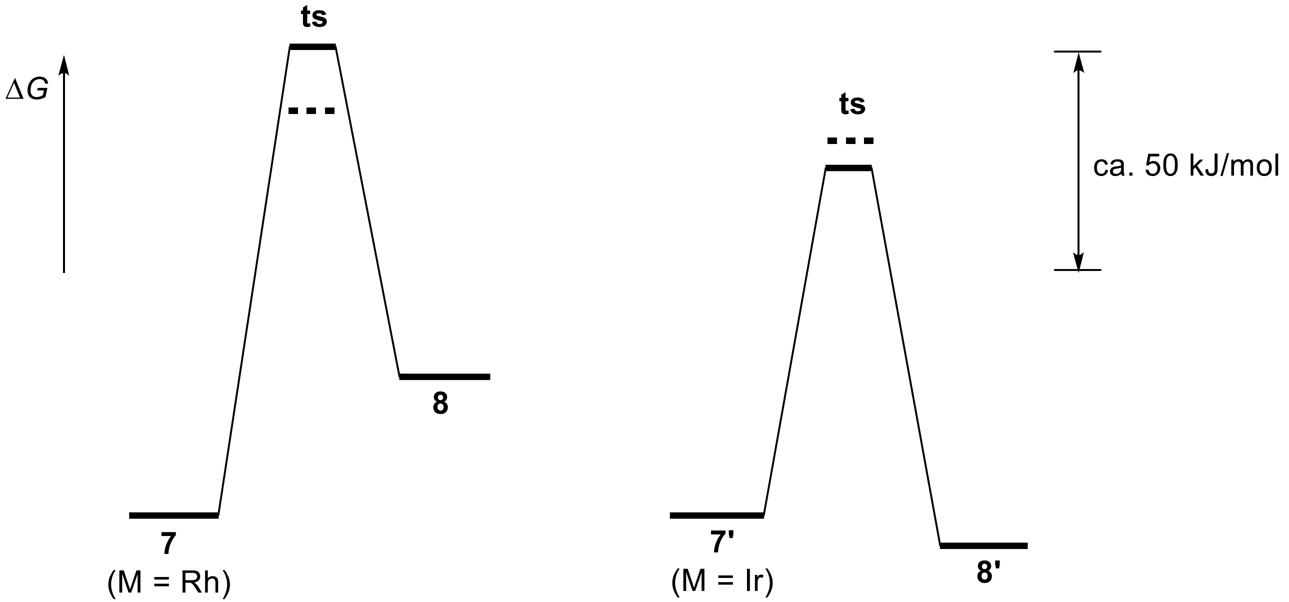
*Resting state*

800 times faster than Monsanto process

**The oxidative addition reaction is facile and it is no longer the rate determining step!**

# *The energy barriers of Monsanto and CATIVA processes*





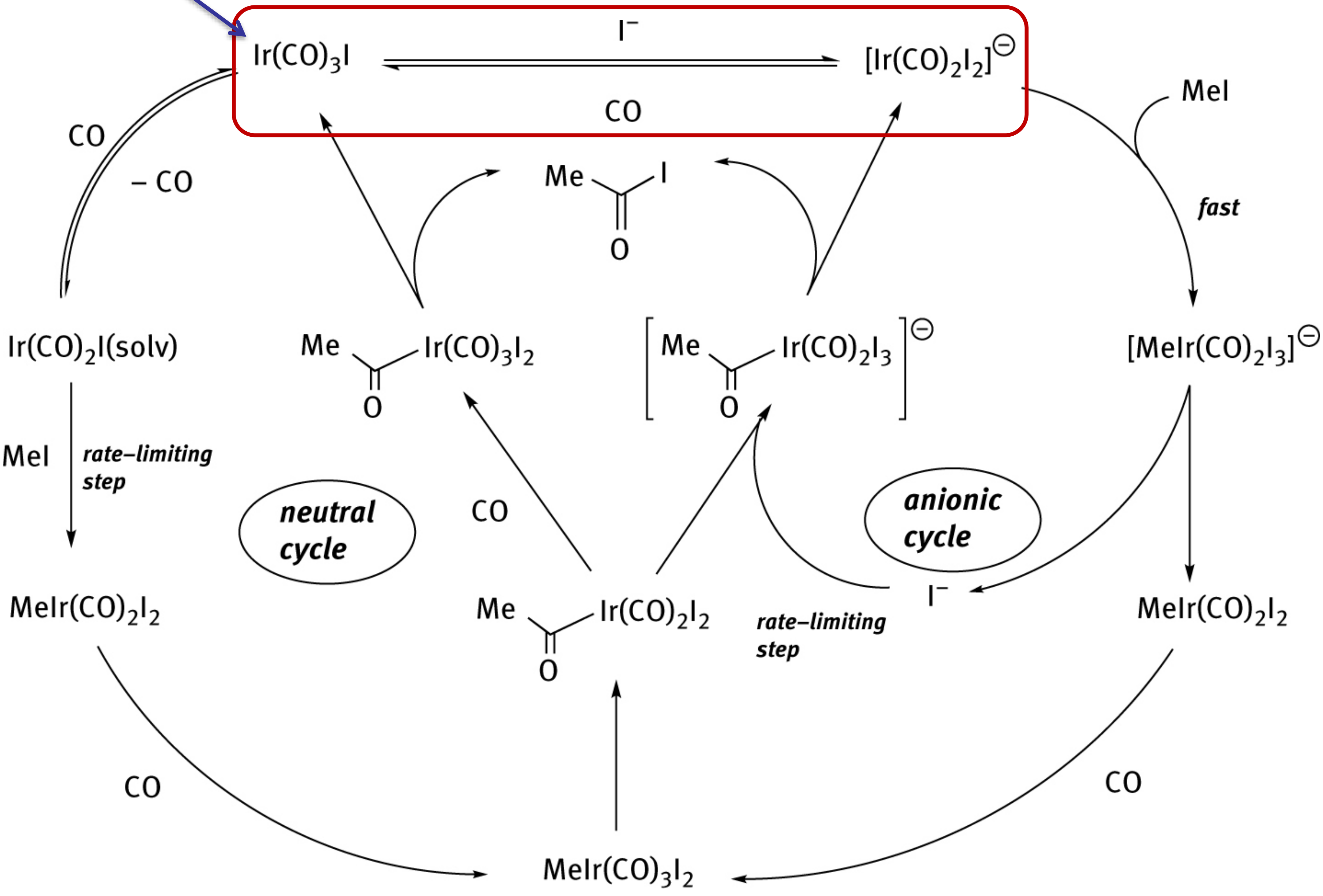
**Activation energy for the methyl migration reaction:**

Experimental values: $\Delta G^{\#}$ = 128.5 kJ/mol (Ir); $\Delta G^{\#}$ = 81.1 kJ/mol (Rh)

Calculated values: $\Delta G^{\#}$ = 116.3 kJ/mol (Ir); $\Delta G^{\#}$ = 72.2 kJ/mol (Rh)

# The catalytic cycles

Resting state



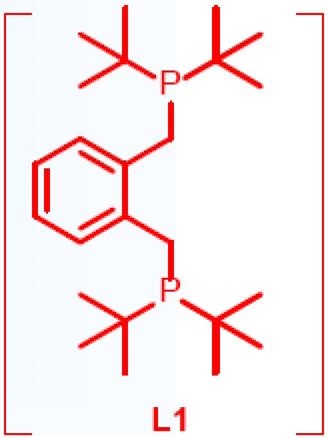
# The Lucite process

## Methoxycarbonylation of ethene

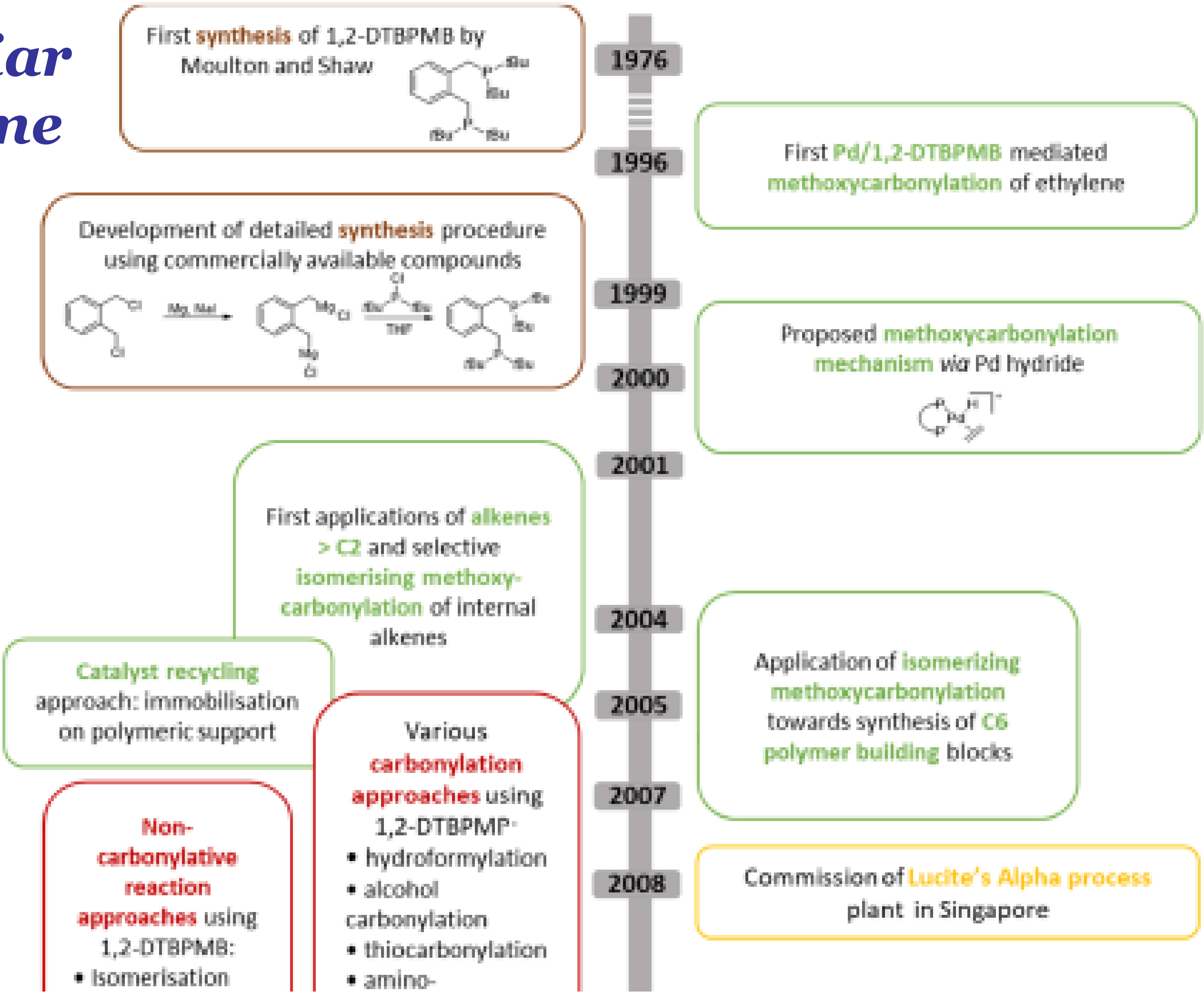
**Lucite alpha I: 120,000 t/y in Singapore**

**Lucite alpha II: 250,000 t/y in Saudi Arabia**

# *A peculiar diphosphine*



L1



First **synthesis** of 1,2-DTBPMB by Moulton and Shaw

First **Pd/1,2-DTBPMB** mediated **methoxycarbonylation** of ethylene

Development of detailed **synthesis** procedure using commercially available compounds

Proposed **methoxycarbonylation mechanism** *via* Pd hydride

First applications of **alkenes > C2** and selective **isomerising methoxy-carbonylation** of internal alkenes

**Catalyst recycling** approach: immobilisation on polymeric support

Application of **isomerizing methoxycarbonylation** towards synthesis of **C6 polymer building** blocks

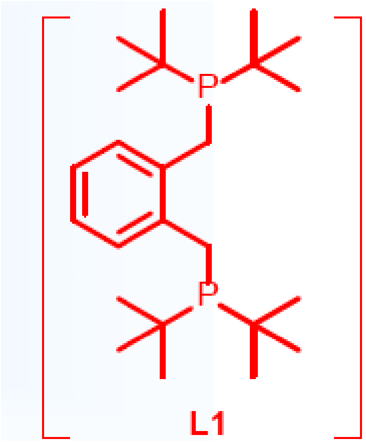
Various **carbonylation approaches** using 1,2-DTBPMP:
- hydroformylation
- alcohol carbonylation
- thiocarbonylation
- amino-

**Non-carbonylative reaction approaches** using 1,2-DTBPMB:
- Isomerisation

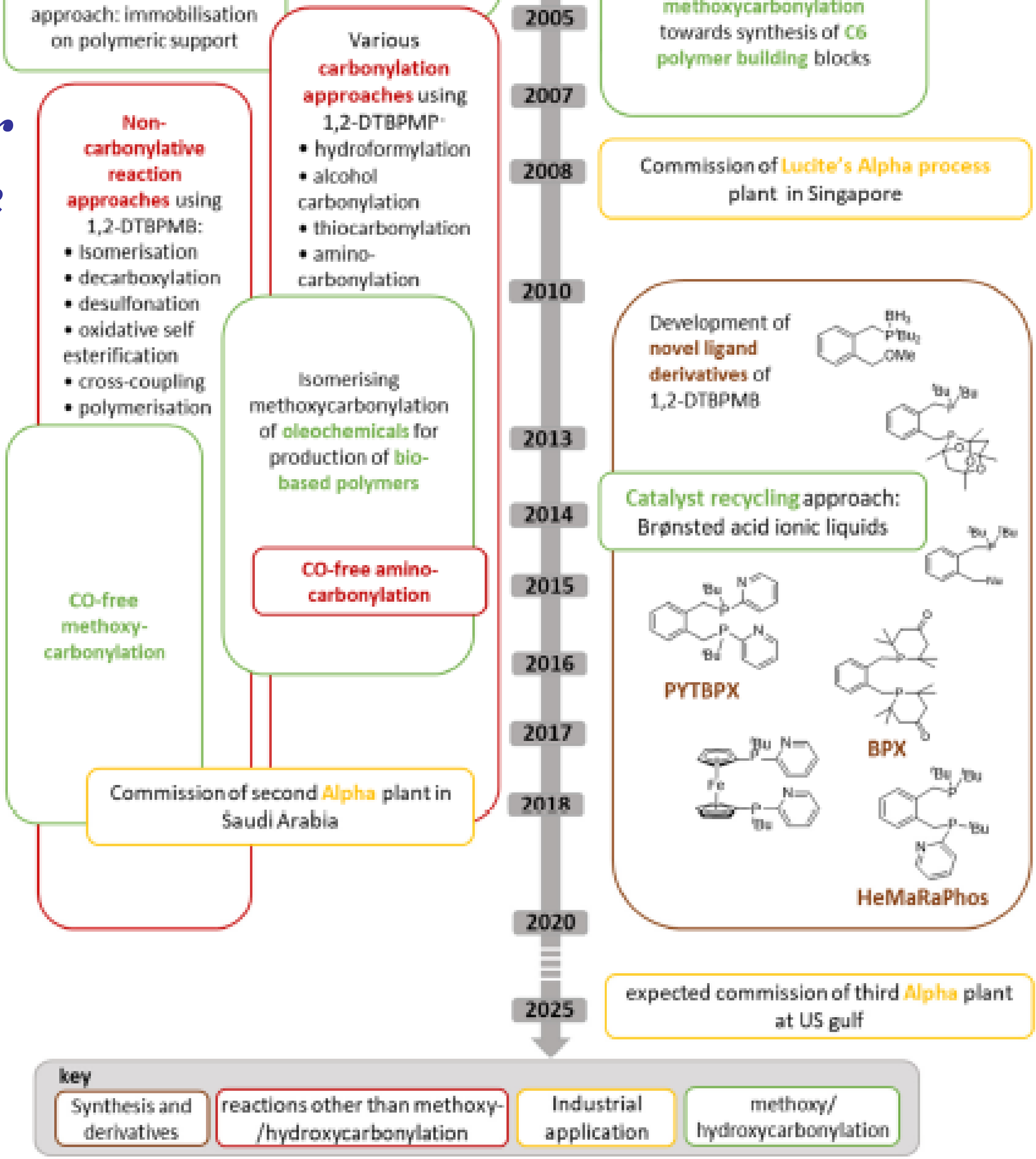
Commission of **Lucite's Alpha process** plant in Singapore

D. J. Cole-Hamilton et al *Chem. Rev.* 2021, *121*, 6610.

# *A peculiar diphosphine*



**L1**



approach: immobilisation on polymeric support

Various **carbonylation approaches** using 1,2-DTBPMP:
- hydroformylation
- alcohol carbonylation
- thiocarbonylation
- amino-carbonylation

**Non-carbonylative reaction approaches** using 1,2-DTBPMB:
- Isomerisation
- decarboxylation
- desulfonation
- oxidative self esterification
- cross-coupling
- polymerisation

Isomerising methoxycarbonylation of **oleochemicals** for production of **bio-based polymers**

**CO-free amino-carbonylation**

**CO-free methoxy-carbonylation**

Commission of second **Alpha** plant in Saudi Arabia

2005

2007

2008

2010

2013

2014

2015

2016

2017

2018

2020

2025

**methoxycarbonylation** towards synthesis of **C6 polymer building** blocks

Commission of **Lucite's Alpha process** plant in Singapore

Development of **novel ligand derivatives** of 1,2-DTBPMB

**Catalyst recycling** approach: Brønsted acid ionic liquids

expected commission of third **Alpha** plant at US gulf

**key**

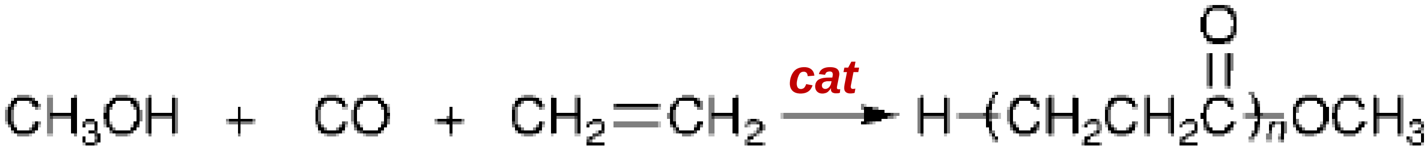
Synthesis and derivatives

reactions other than methoxy-/hydroxycarbonylation

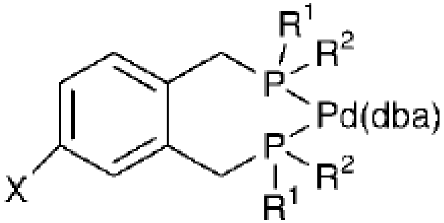
Industrial application

methoxy/ hydroxycarbonylation

# *Methoxycarbonylation of ethene*



$$CH_3OH + CO + CH_2{=}CH_2 \xrightarrow{cat} H\!-\!(CH_2CH_2\overset{\displaystyle O}{\overset{\|}{C}})_n OCH_3$$

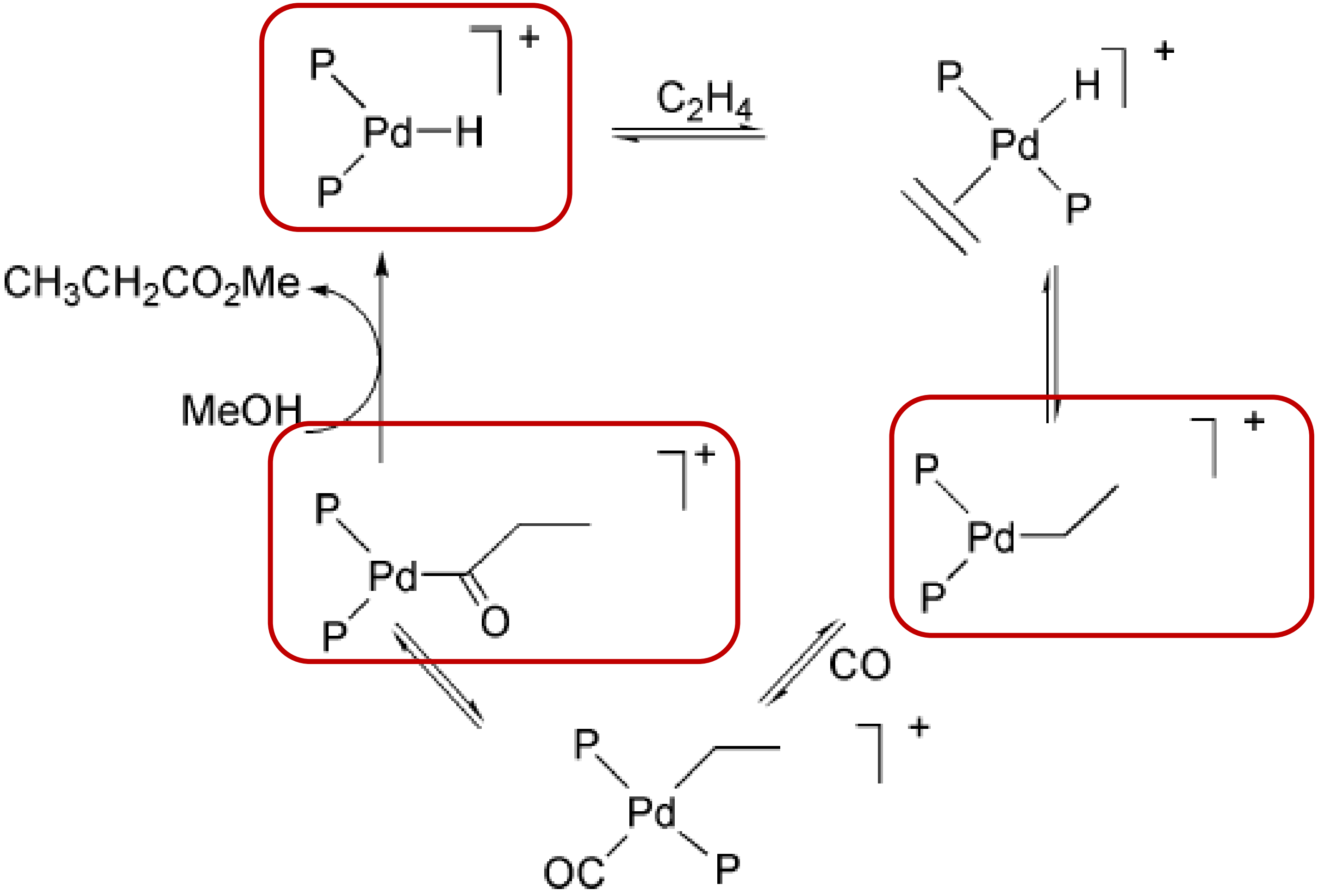


**1** $R^1 = R^2 = Bu^t$, X = H
**2** $R^1 = R^2 = Bu^t$, X = $NO_2$
**3** $R^1 = R^2 = Bu^t$, X = OMe
**4** $R^1 = R^2 = Pr^i$, X = H
**5** $R^1 = R^2 = Cy$, X = H
**6** $R^1 = R^2 = Ph$, X = H
**7** $R^1 = Bu^t$, $R^2 = Cy$, X = H

| Complex | Activity[b] | Selectivity[c] (%) | P–Pd–P bite angle[d] (°) |
|---|---|---|---|
| **1** | 12 000 | 99.9 | 103.9 |
| **2** | 11 500 | 99.9 | |
| **3** | 11 800 | 99.9 | |
| **4** | 200 | 20 | 104.3 |
| **5** | 200 | 25 | 103.9 |
| **6** | 400 | 20 | 104.6 |
| **7** | 500 | 30 | |

Reaction Conditions:
$n_{Pd}$ = 0.01 mmol; $CH_3OH$ V = 300 mL; T = 80 °C;
$P_{tot}$ = 10 bar; $n_{CF3SO3H}$ = 0.2 mmol; t = 3 h.
[b] $mol_{Et}/mol_{Pd}\cdot h$.

# *The mechanism*



# *Flow scheme of Lucite process*